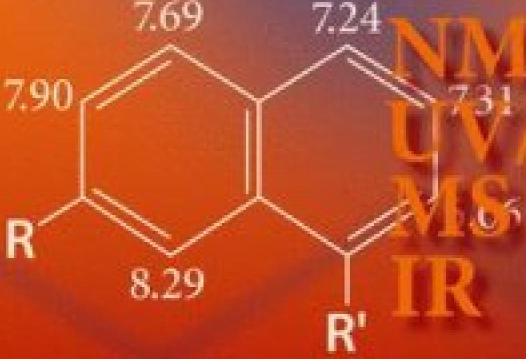

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UV/Vis
MS
IR

Ernö Pretsch
Philippe Bühlmann
Martin Badertscher

Structure Determination of Organic Compounds

Tables of Spectral Data

4th, Completely Revised Edition

Springer

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Ernö Pretsch · Philippe Bühlmann ·
Martin Badertscher

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Tables of Spectral Data

Fourth, Revised and Enlarged Edition

Springer

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Preface

The ongoing success of the earlier versions of this book motivated us to prepare a new edition. While modern techniques of nuclear magnetic resonance spectroscopy and mass spectrometry have changed the ways of data acquisition and greatly extended the capabilities of these methods, the basic parameters, such as chemical shifts, coupling constants, and fragmentation pathways remain the same. However, since the amount and quality of available data has considerably increased over the years, we decided to prepare a significantly revised manuscript. It follows the same basic concepts, i.e., it provides a representative, albeit limited set of reference data for the interpretation of ^{13}C NMR, ^{1}H NMR, IR, mass, and UV/Vis spectra. We also added a new chapter with reference data for ^{19}F and ^{31}P NMR spectroscopy and, in the chapter on infrared spectroscopy, we newly refer to important Raman bands.

Since operating systems of computers become outdated much faster than printed media, we decided against providing a compact disk with this new edition. The limited versions of the NMR spectra estimation programs can be downloaded from the home page of the developing company (www.upstream.ch/support/book_downloads.html).

We thank numerous colleagues who helped us in many different ways to complete the manuscript. We are particularly indebted to Dr. Dorothée Wegmann for her expertise with which she eliminated many errors and inconsistencies of the earlier versions. Special thanks are due to Prof. Wolfgang Robien for providing us with reference data from his outstanding ^{13}C NMR database, CSEARCH. Another high-quality source of information was the Spectral Database System of the National Institute of Advanced Industrial Science and Technology (http://riodb01.ibase.aist.go.jp/sdbs/), Tsukuba, Ibaraki (Japan).

In spite of great efforts and many checks to eliminate errors, it is likely that some mistakes or inconsistencies remain. We would like to encourage our readers to contact us with comments and suggestions under one of the following addresses: Prof. Ernö Pretsch, Institute of Biogeochemistry and Pollutant Dynamics, ETH Zürich, CH-8092 Zürich, Switzerland, e-mail: pretsche@ethz.ch, Prof. Philippe Bühlmann, Department of Chemistry, University of Minnesota, 207 Pleasant St. SE, Minneapolis, MN 55455, USA, e-mail: buhlmann@umn.edu, or Dr. Martin Badertscher, Laboratory of Organic Chemistry, ETH Zürich, CH-8093 Zürich, Switzerland, e-mail: badertscher@org.chem.ethz.ch.

Zürich and Minneapolis, November 2008

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1 Introduction

1.1 Scope and Organization

The present data collection is intended to serve as an aid in the interpretation of molecular spectra for the elucidation and confirmation of the structure of organic compounds. It consists of reference data, spectra, and empirical correlations from ^{1}H, ^{13}C, ^{19}F, and ^{31}P nuclear magnetic resonance (NMR), infrared (IR), mass, and ultraviolet–visible (UV/Vis) spectroscopy. It is to be viewed as a supplement to textbooks and specific reference works dealing with these spectroscopic techniques. The use of this book to interpret spectra only requires the knowledge of basic principles of the techniques, but its content is structured in a way that it will serve as a reference book also to specialists.

Chapters 2 and 3 contain Summary Tables and Combined Tables of the most relevant spectral characteristics of structural elements. While Chapter 2 is organized according to the different spectroscopic methods, Chapter 3 for each class of structural elements supplies spectroscopic information obtained with various techniques. These two chapters should assist users less familiar with spectra interpretation to identify the classes of structural elements present in samples of their interest. The four chapters with data from ^{13}C NMR, ^{1}H NMR, IR spectroscopy, and mass spectrometry are ordered in the same manner by compound types. These cover the various carbon skeletons (alkyl, alkenyl, alkynyl, alicyclic, aromatic, and heteroaromatic), the most important substituents (halogen, single-bonded oxygen, nitrogen, sulfur, and carbonyl), and some specific compound classes (miscellaneous compounds and natural products). Finally, a spectra collection of common solvents, auxiliary compounds (such as matrix materials and references), and commonly found impurities is provided with each method. Not only the strictly analogous order of the data but also the optical marks on the edge of the pages help fast cross-referencing between the various spectroscopic techniques. Because their data sets are less comprehensive, the chapters on ^{19}F and ^{31}P NMR and UV/Vis are organized somewhat differently. Although currently UV/Vis spectroscopy is only marginally relevant to structure elucidation, its importance might increase by the advent of high-throughput analyses. Also, the reference data presented in the UV/Vis chapter are useful in connection with optical sensors and the widely applied UV/Vis detectors in chromatography and electrophoresis.

Since a great part of the tabulated data either comes from our own measurements or is based on a large body of literature data, comprehensive references to published sources are not included. Whenever possible, the data refer to conventional modes and conditions of measurement. For example, unless the solvent is indicated, the NMR chemical shifts were normally determined with deuterochloroform. Likewise, the IR spectra were measured using solvents of low polarity, such as chloroform or

carbon disulfide. Mass spectral data were recorded with electron impact ionization at 70 eV.

While retaining the basic structure of the previous editions, numerous reference entries have been updated and new entries have been added. Altogether, about 20% of the data is new. The chapter on ^{19}F and ^{31}P NMR is entirely new, and the section on IR spectroscopy now includes references to important Raman bands.

1.2 Abbreviations and Symbols

al	aliphatic
alk	alkyl
alken	alkenyl
ar	aromatic
as	asymmetric
ax	axial
comb	combination vibration
d	doublet
δ	IR: deformation vibration
	NMR: chemical shift
DFTMP	1,1-difluoro-1-(trimethylsilyl)methylphosphonic acid
DMSO	dimethyl sulfoxide
eq	equatorial
ε	molar absorptivity
frag	fragment
γ	skeletal vibration
gem	geminal
hal	halogen
ip	in plane vibration
J	coupling constant
liq	liquid
$M^{+\cdot}$	molecular radical ion
m/z	mass to charge ratio
$\tilde{\nu}$	wavenumber
oop	out of plane vibration
sh	shoulder
st	stretching vibration
sy	symmetric
TFA	trifluoroacetic acid
THF	tetrahydrofuran
TMS	tetramethylsilane
vic	vicinal

2 Summary Tables

2.1 General Tables

2.1.1 Calculation of the Number of Double Bond Equivalents from the Molecular Formula

General Equation

$$\text{double bond equivalents} = 1 + \tfrac{1}{2} \sum_i n_i (v_i - 2)$$

n_i: number of atoms of element i in molecular formula
v_i: formal valence of element i

Short Cut

For compounds containing only C, H, O, N, S, and halogens, the following steps permit a quick and simple calculation of the number of double bond equivalents:

1. O and divalent S are deleted from the molecular formula
2. Halogens are replaced by hydrogen
3. Trivalent N is replaced by CH
4. The resulting hydrocarbon, C_nH_x, is compared with the saturated hydrocarbon, C_nH_{2n+2}. Each double bond equivalent reduces the number of hydrogen atoms by 2:

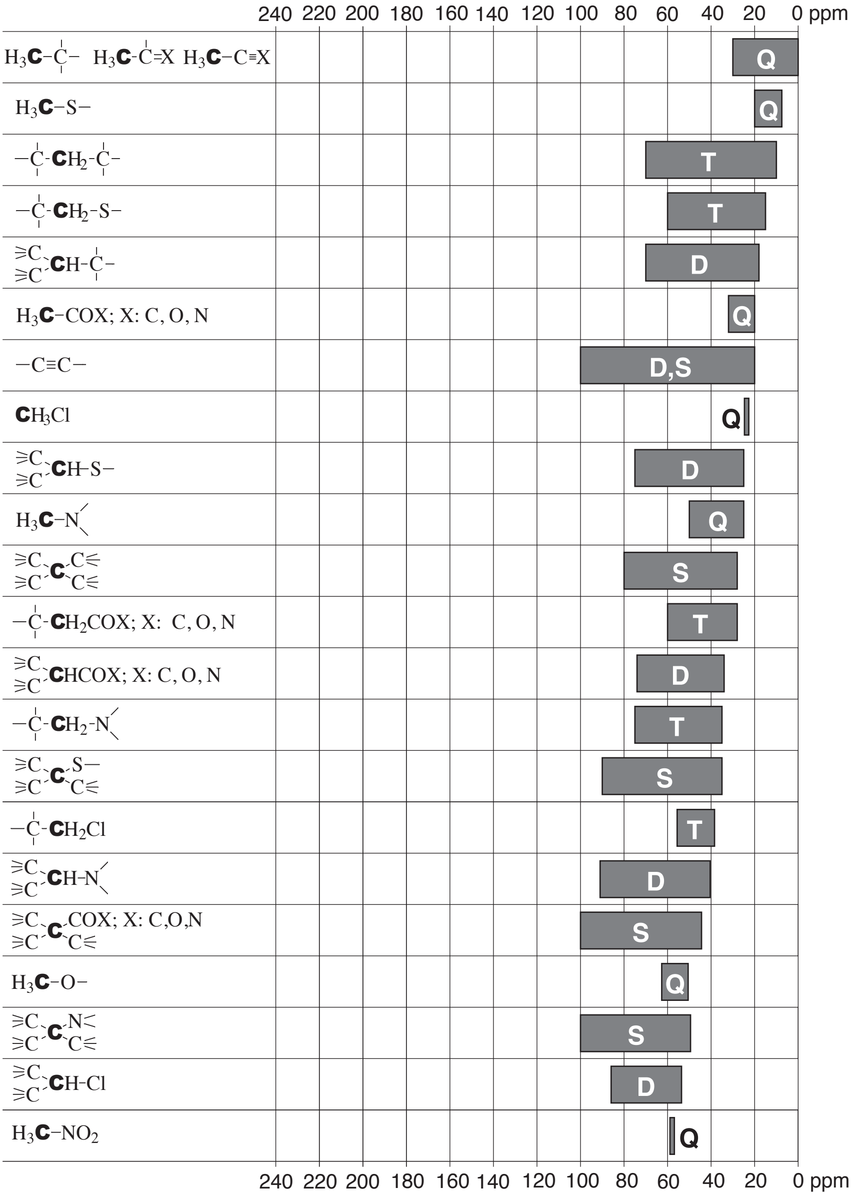
$$\text{double bond equivalents} = \tfrac{1}{2} (2\,n + 2 - x)$$

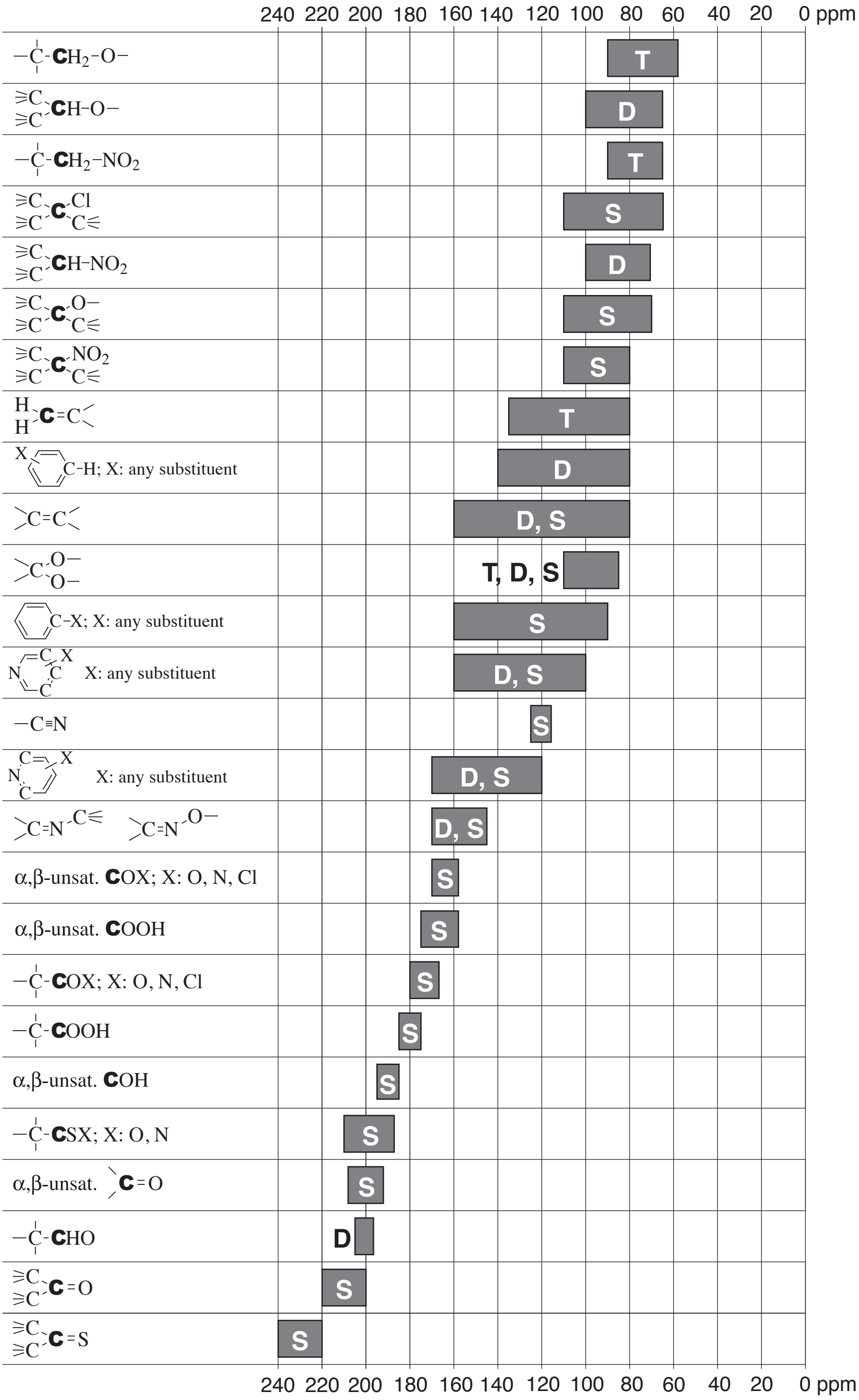
2.1.2 Properties of Selected Nuclei

Isotope	Natural abundance [%]	Spin quantum number, I	Frequency [MHz] at 2.35 Tesla	Relative sensitivity of nucleus	Relative sensitivity at natural abundance	Electric quadrupole moment [$e \times 10^{-24}$ cm^2]
^{1}H	99.985	1/2	100.0	1	1	
^{2}H	0.015	1	15.4	$9.6{\times}10^{-3}$	$1.5{\times}10^{-6}$	$2.8{\times}10^{-3}$
^{3}H	0.000	1/2	106.7	1.2	0	
^{10}B	19.58	3	10.7	$2.0{\times}10^{-2}$	$3.9{\times}10^{-3}$	$7.4{\times}10^{-2}$
^{11}B	80.42	3/2	32.1	$1.6{\times}10^{-1}$	$1.3{\times}10^{-1}$	$3.6{\times}10^{-2}$
^{13}C	1.108	1/2	25.1	$1.6{\times}10^{-2}$	$1.8{\times}10^{-4}$	
^{14}N	99.635	1	7.3	$1.0{\times}10^{-3}$	$1.0{\times}10^{-3}$	$1.9{\times}10^{-2}$
^{15}N	0.365	1/2	10.1	$1.0{\times}10^{-3}$	$3.8{\times}10^{-6}$	
^{17}O	0.037	5/2	13.6	$2.9{\times}10^{-2}$	$1.1{\times}10^{-5}$	$-2.6{\times}10^{-2}$
^{19}F	100.000	1/2	94.1	$8.3{\times}10^{-1}$	$8.3{\times}10^{-1}$	
^{31}P	100.000	1/2	40.5	$6.6{\times}10^{-2}$	$6.6{\times}10^{-2}$	
^{33}S	0.76	3/2	7.6	$2.3{\times}10^{-3}$	$1.7{\times}10^{-5}$	$-6.4{\times}10^{-2}$
^{117}Sn	7.61	1/2	35.6	$4.5{\times}10^{-2}$	$3.4{\times}10^{-3}$	
^{119}Sn	8.58	1/2	37.3	$5.2{\times}10^{-2}$	$4.4{\times}10^{-3}$	
^{195}Pt	33.8	1/2	21.5	$9.9{\times}10^{-3}$	$3.4{\times}10^{-3}$	
^{199}Hg	16.84	1/2	17.8	$5.7{\times}10^{-3}$	$9.5{\times}10^{-4}$	
^{207}Pb	22.6	1/2	20.9	$9.2{\times}10^{-3}$	$2.1{\times}10^{-4}$	

2.2 ^{13}C NMR Spectroscopy

Summary of the Regions of Chemical Shifts, δ (in ppm), for Carbon Atoms in Various Chemical Environments (carbon atoms are specified as follows: Q for CH_3, T for CH_2, D for CH, and S for C)




240 220 200 180 160 140 120 100 80 60 40 20 0 ppm
T
D
T
S
D
S
S
T
C–H; X: any substituent
D
D, S
T, D, S
C–X; X: any substituent
S
X: any substituent
D, S
S
X: any substituent
D, S
D, S
α,β-unsat. COX; X: O, N, Cl
S
α,β-unsat. COOH
S
S
S
α,β-unsat. COH
S
S
α,β-unsat.
S
D
S
S
240 220 200 180 160 140 120 100 80 60 40 20 0 ppm

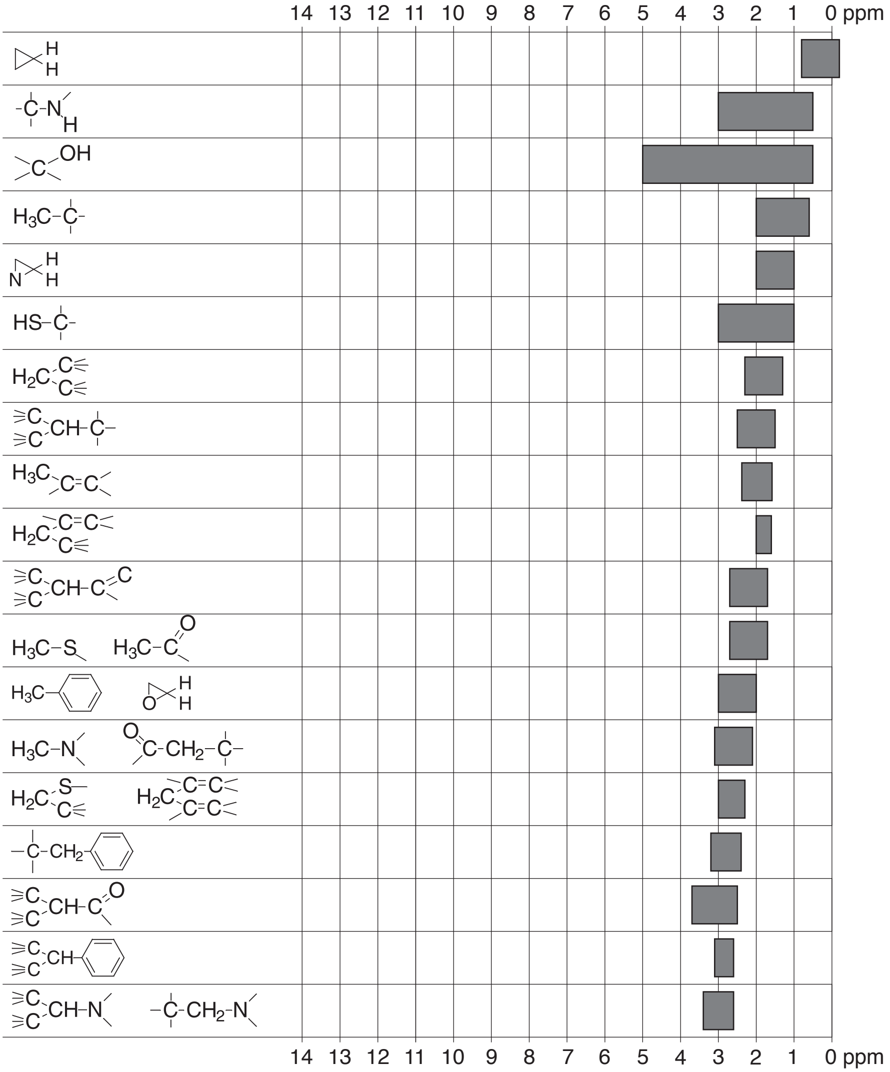
^{13}C Chemical Shifts of Carbonyl Groups (δ in ppm)

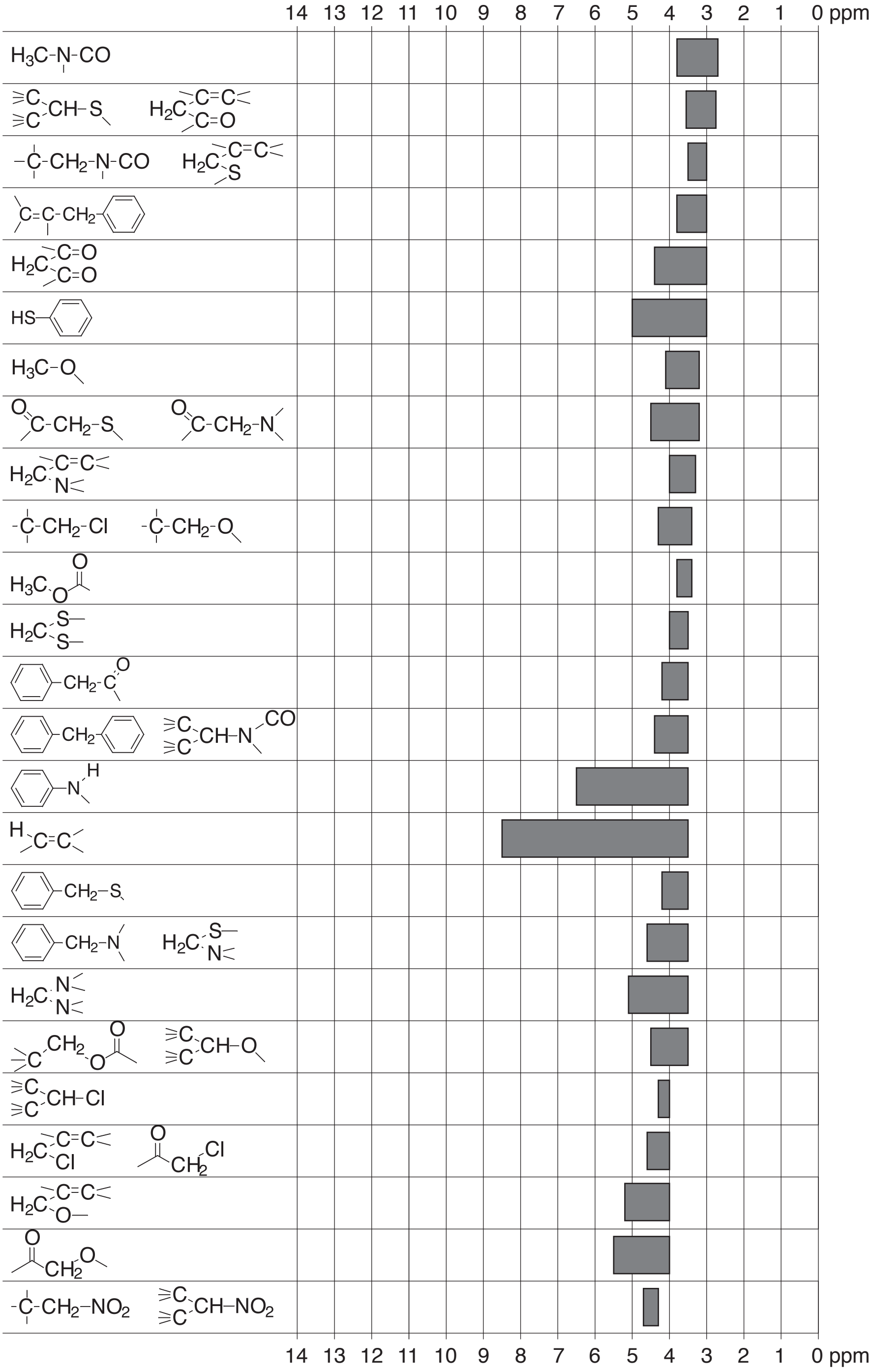
R	R–CHO	R–$COCH_3$	R–COOH	R–COO^-
–H	197.0	200.5	166.3	171.3
–CH_3	200.5	206.7	176.9	182.6
–CH_2CH_3	202.7	207.6	180.4	185.1
–$CH(CH_3)_2$	204.6	211.8	184.1	
–$C(CH_3)_3$	205.6	213.5	185.9	188.6
–*n*-C_8H_{17}	202.6	207.9	180.7	183.1
–CH_2Cl	193.3	200.1	173.7	175.9
–$CHCl_2$		193.6	170.4	171.8
–CCl_3	176.9	186.3	167.1	167.6
–cyclohexyl	204.7	209.4	182.1	185.4
–CH=CH_2	194.4	197.5	171.7	174.5
–C≡CH	176.8	183.6	156.5	
–phenyl	192.0	196.9	172.6	177.6

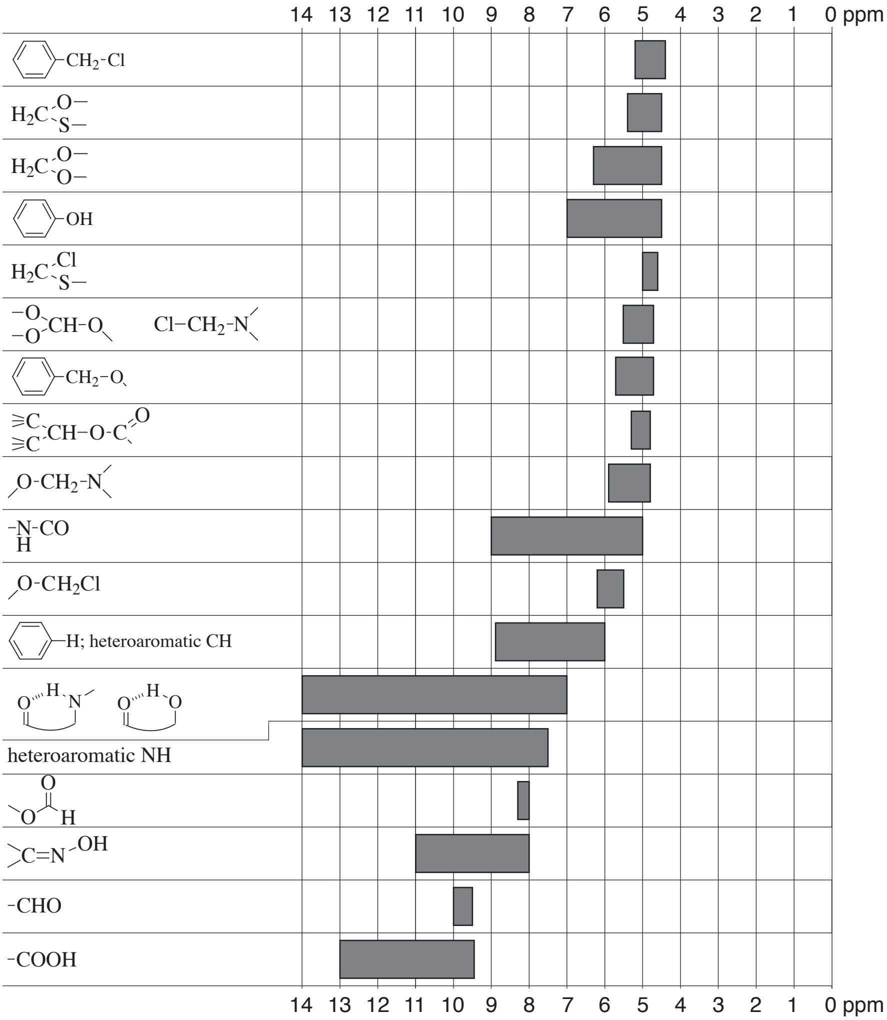
R	R–CHO	R–$COCH_3$	R–COOH	R–COO^-
–H	161.6	167.6	158.5	
–CH_3	171.3	173.4	167.4	170.4
–CH_2CH_3	173.3	177.2	170.3	174.7
–$CH(CH_3)_2$	177.4		172.8	178.0
–$C(CH_3)_3$	178.8	180.9	173.9	180.3
–*n*-C_8H_{17}	174.4	176.3	169.4	173.8
–CH_2Cl	167.8	168.3	162.1	167.7
–$CHCl_2$	165.1		157.6	165.5
–CCl_3	162.5		154.1	
–cyclohexyl	175.3	177.3		176.3
–CH=CH_2	166.5	168.3		165.6
–C≡CH	153.4			
–phenyl	166.8	169.7	162.8	168.0

2.3 ^{1}H NMR Spectroscopy

Summary of the Regions of Chemical Shifts, δ (in ppm), for Hydrogen Atoms in Various Chemical Environments

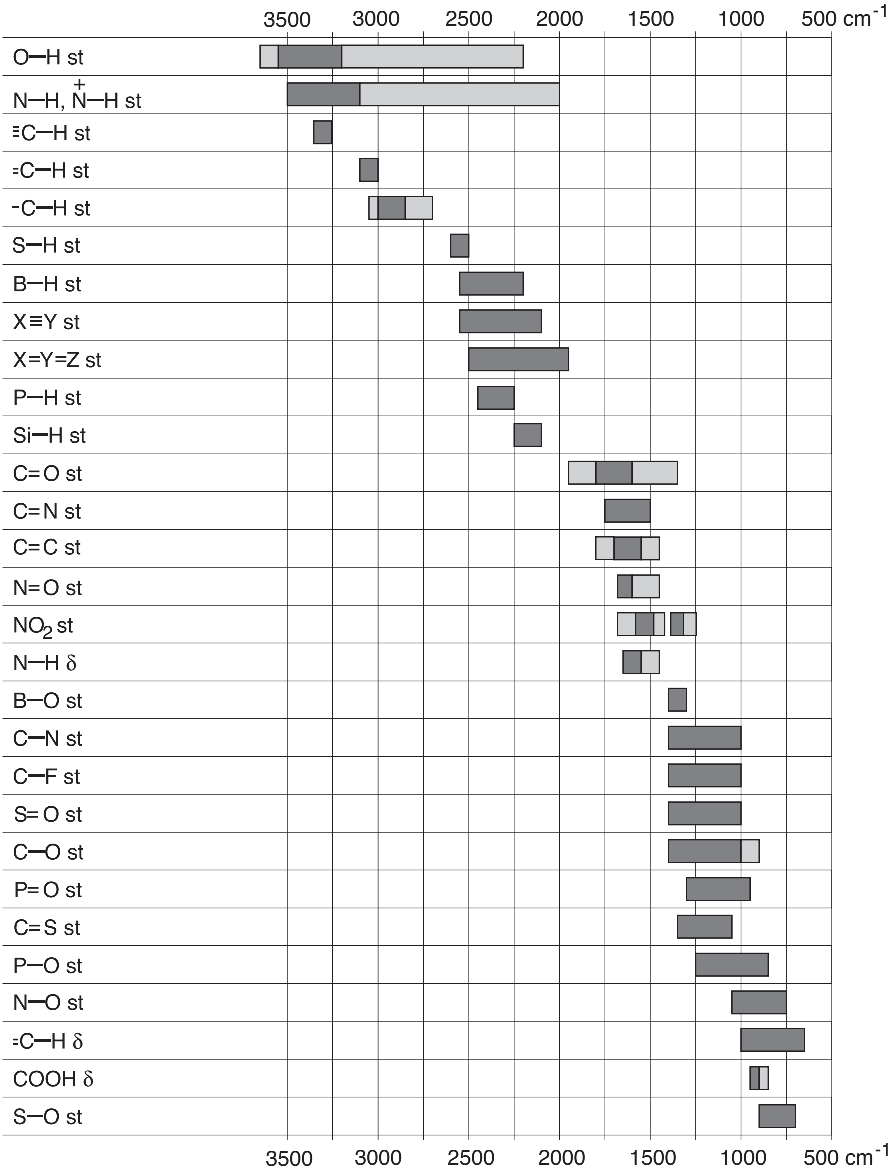



14 13 12 11 10 9 8 7 6 5 4 3 2 1 0 ppm

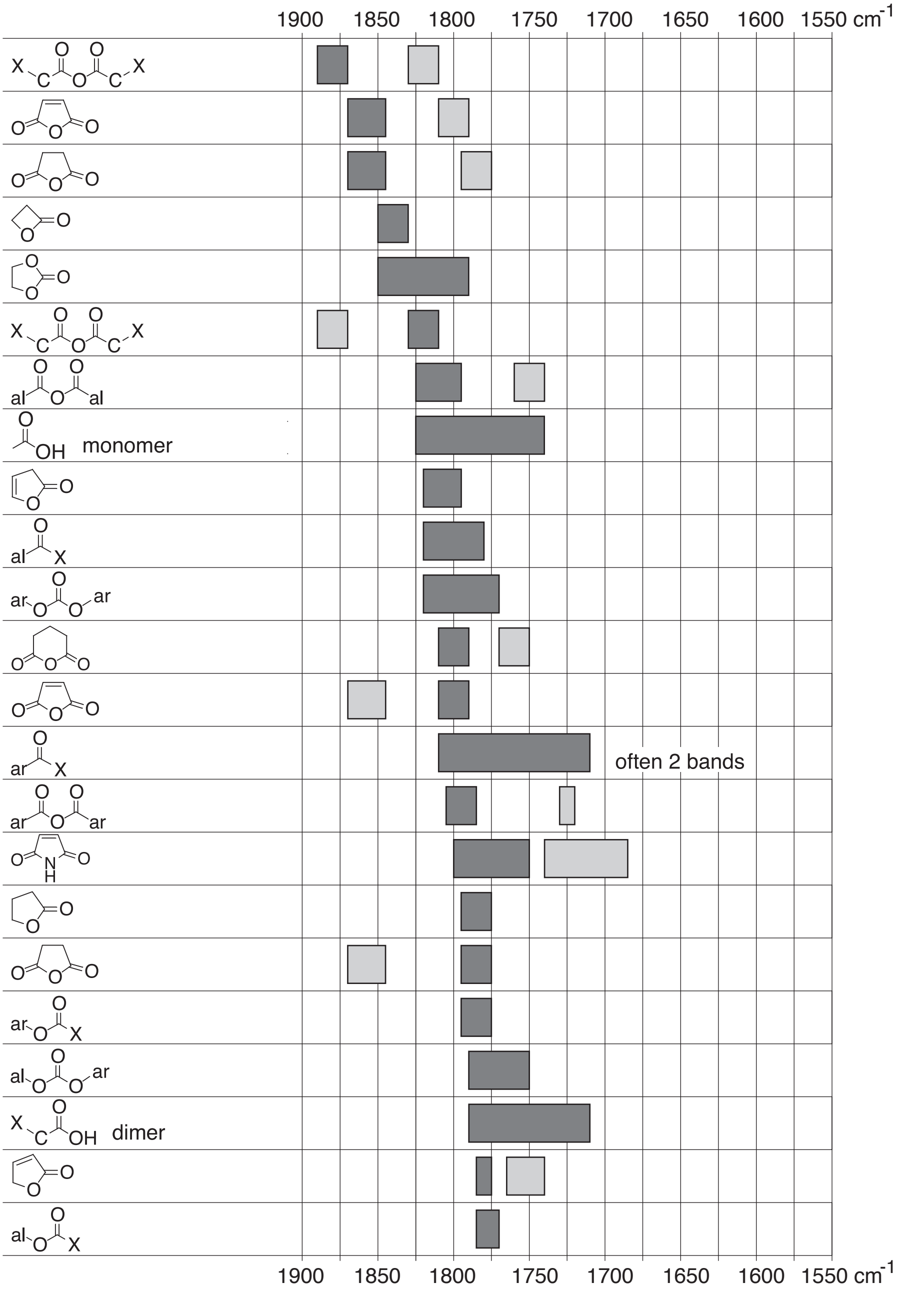

14 13 12 11 10 9 8 7 6 5 4 3 2 1 0 ppm
-CH2-Cl
H2C-O- / -S-
H2C-O- / -O-
-OH
H2C-Cl / -S-
-O / -O CH-O
Cl-CH2-N
-CH2-O
CH-O-C=O
O-CH2-N
-N-CO / H
O-CH2Cl
-H; heteroaromatic CH
heteroaromatic NH
C=N-OH
-CHO
-COOH
14 13 12 11 10 9 8 7 6 5 4 3 2 1 0 ppm

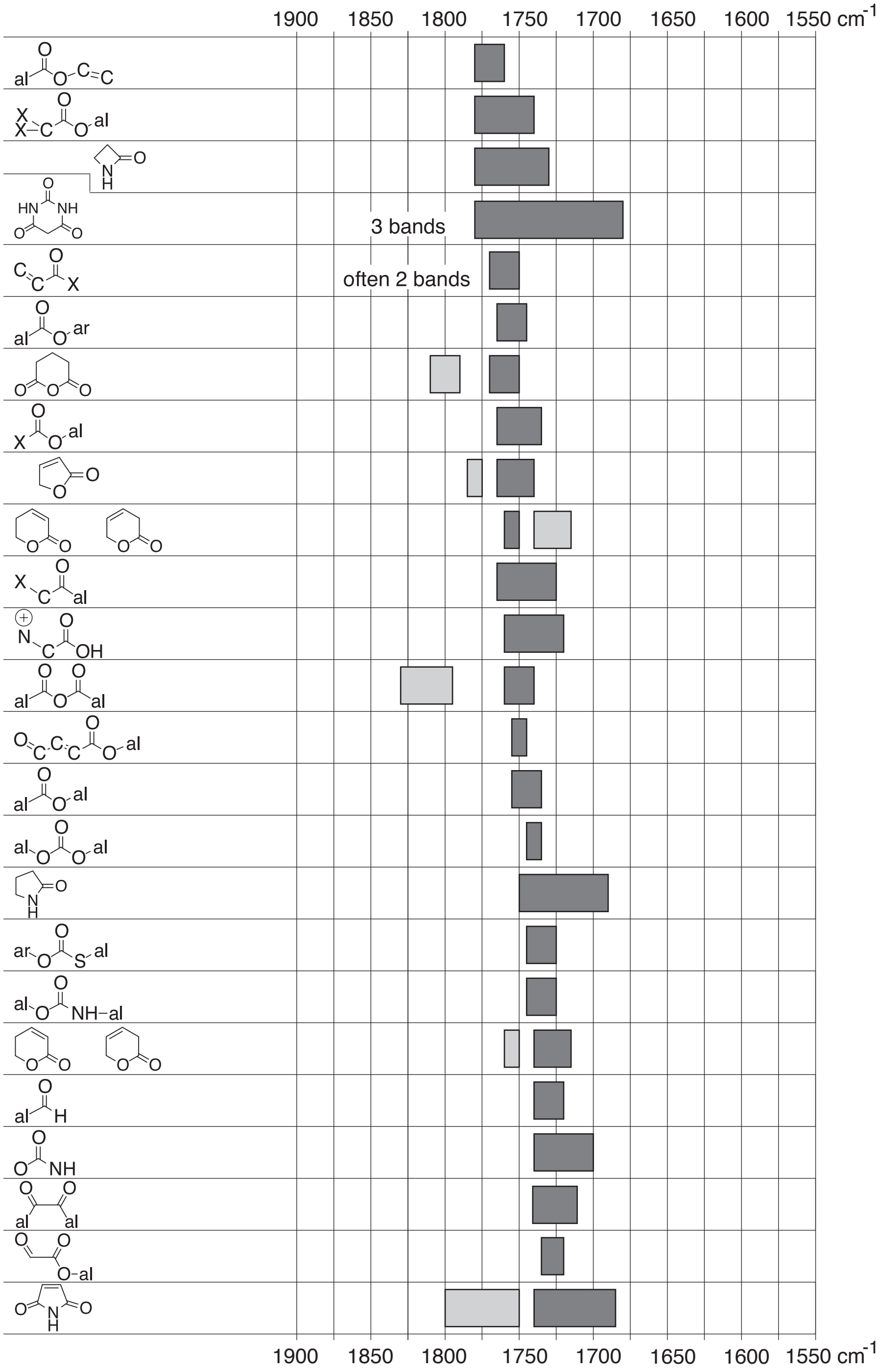
2.4 IR Spectroscopy

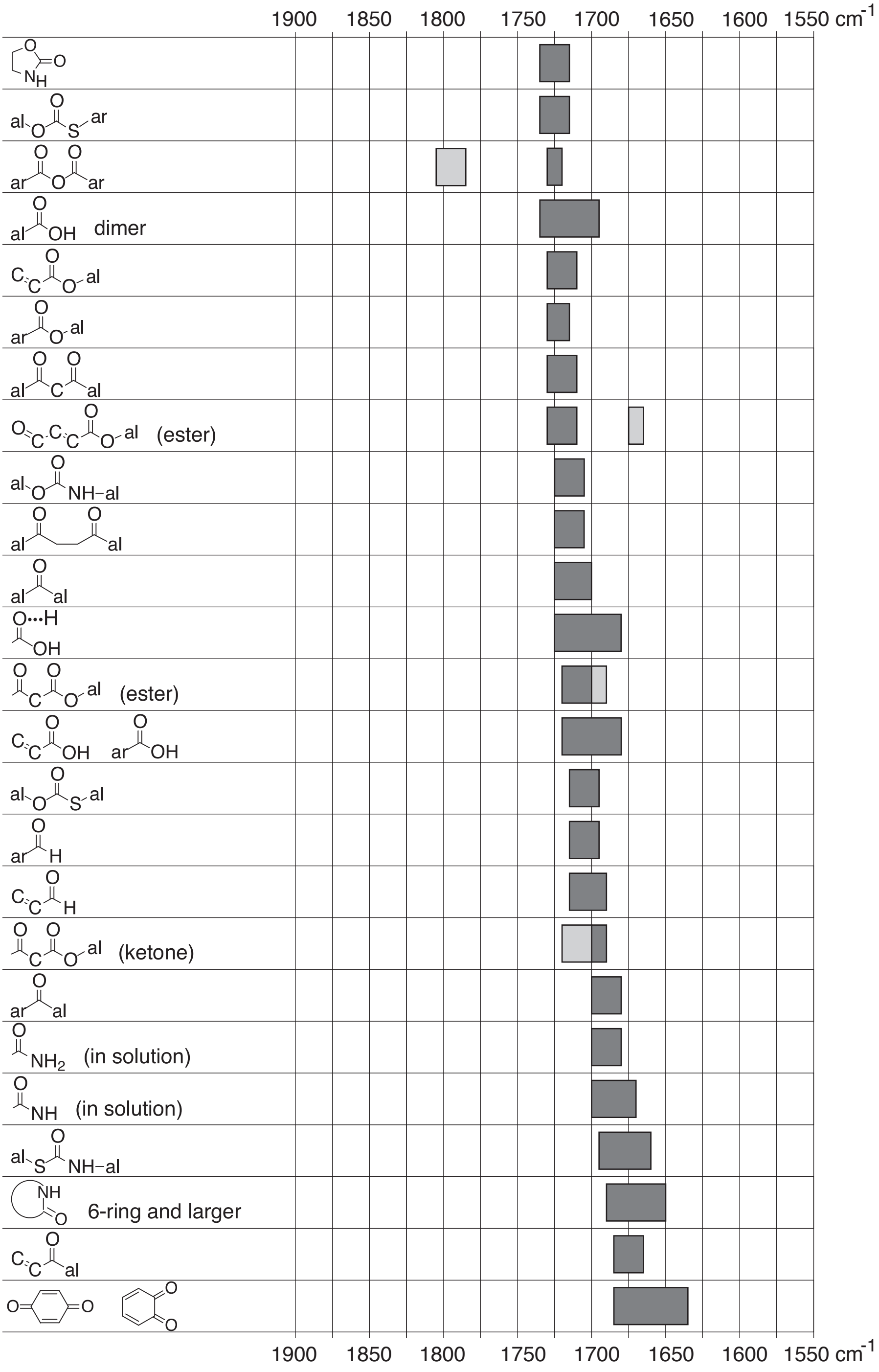
Summary of the Most Important IR Absorption Bands ($\tilde{\nu}$ in cm^{-1})

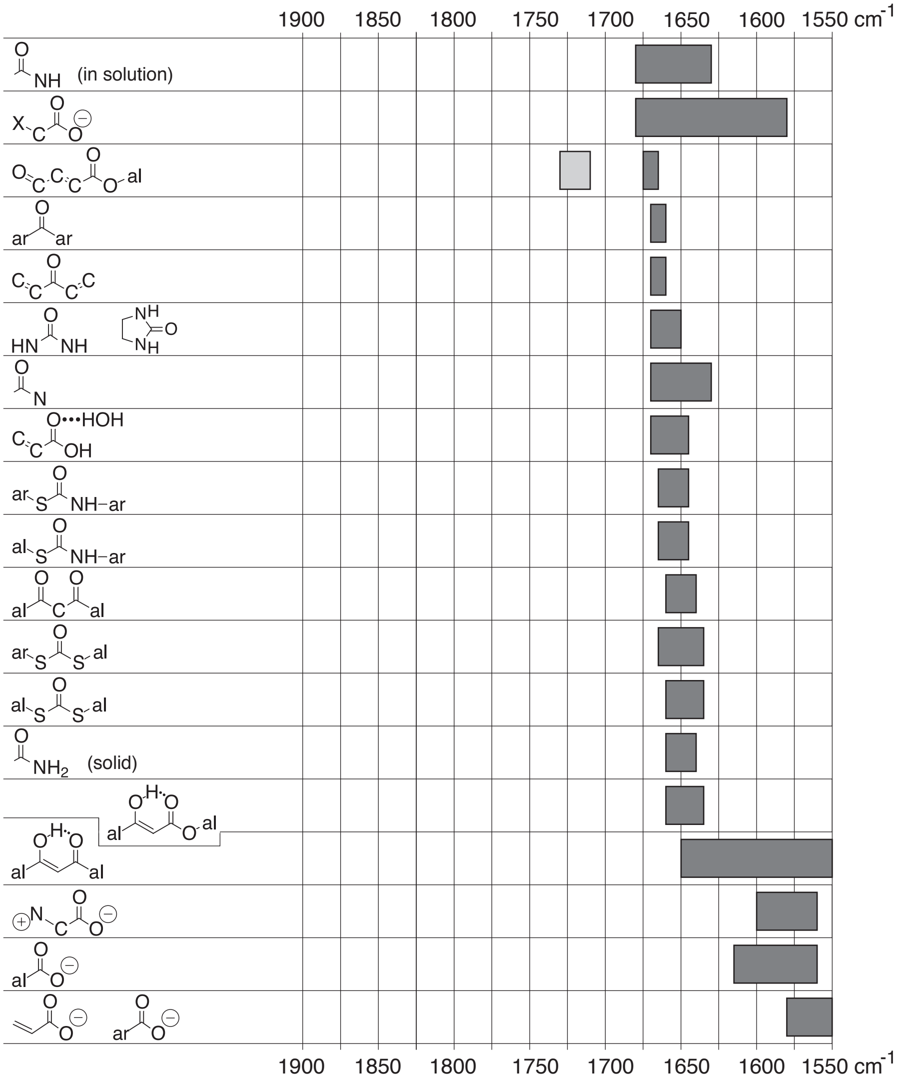


Summary of IR Absorption Bands of Carbonyl Groups ($\tilde{\nu}$ in cm^{-1})




1900 1850 1800 1750 1700 1650 1600 1550 cm-1
3 bands
often 2 bands
1900 1850 1800 1750 1700 1650 1600 1550 cm-1


1900
1850
1800
1750
1700
1650
1600
1550 cm-1
dimer
(ester)
(ester)
(ketone)
(in solution)
(in solution)
6-ring and larger


1900 1850 1800 1750 1700 1650 1600 1550 cm-1
(in solution)
(solid)

2.5 Mass Spectrometry

2.5.1 Average Masses of Naturally Occurring Elements with Masses and Representative Relative Abundances of Isotopes [1–3]

Element Isotope	Mass	Abundance	**Element** Isotope	Mass	Abundance
H	1.00794[a,b]	(in water)	**F**	18.998403	
^{1}H	1.007825	100[c]	^{19}F	18.998403	100
^{2}H	2.014102	0.0115			
			Ne	20.1797[a]	(in air)
He	4.002602[a]	(in air)	^{20}Ne	19.992440	100[c]
^{3}He	3.016029	0.000134	^{21}Ne	20.993847	0.38
^{4}He	4.002603	100	^{22}Ne	21.991385	10.22
Li	6.941[a]		**Na**	22.989769	
^{6}Li	6.015123	8.21[d]	^{23}Na	22.989769	100
^{7}Li	7.016005	100			
			Mg	24.3050	
Be	9.012182		^{24}Mg	23.985042	100
^{9}Be	9.012182	100	^{25}Mg	24.985837	12.66
			^{26}Mg	25.982593	13.94
B	10.811[a]				
^{10}B	10.012937	24.8[c]	**Al**	26.981538	
^{11}B	11.009305	100	^{27}Al	26.981538	100
C	12.0107[a]		**Si**	28.0855[a]	
^{12}C	12.000000	100	^{28}Si	27.976927	100
^{13}C	13.003355	1.08	^{29}Si	28.976495	5.080
			^{30}Si	29.973770	3.353
N	14.0067[a]				
^{14}N	14.003074	100	**P**	30.973762	
^{15}N	15.000109	0.365	^{31}P	30.973762	100
O	15.9994[a]		**S**	32.065[a]	
^{16}O	15.994915	100	^{32}S	31.972071	100[c]
^{17}O	16.999132	0.038	^{33}S	32.971459	0.79
^{18}O	17.999161	0.205	^{34}S	33.967867	4.47
			^{36}S	35.967081	0.01

Element Isotope	Mass	Abundance	**Element** Isotope	Mass	Abundance
Cl	35.453		**Cr**	51.9961	
^{35}Cl	34.968853	100[c]	^{50}Cr	49.946044	5.186
^{37}Cl	36.965903	32.0	^{52}Cr	51.940508	100
			^{53}Cr	52.940649	11.339
Ar	39.948[a]	(in air)	^{54}Cr	53.938880	2.823
^{36}Ar	35.967545	0.3379			
^{38}Ar	37.962732	0.0635	**Mn**	54.938045	
^{40}Ar	39.962383	100	^{55}Mn	54.938045	100
K	39.0983		**Fe**	55.845	
^{39}K	38.963707	100	^{54}Fe	53.939611	6.370
^{40}K	39.963998	0.0125	^{56}Fe	55.934938	100
^{41}K	40.961826	7.2167	^{57}Fe	56.935394	2.309
			^{58}Fe	57.933276	0.307
Ca	40.078				
^{40}Ca	39.962591	100	**Co**	58.933195	
^{42}Ca	41.958618	0.667	^{59}Co	58.933195	100
^{43}Ca	42.958767	0.139			
^{44}Ca	43.955482	2.152	**Ni**	58.6934	
^{46}Ca	45.953693	0.004	^{58}Ni	57.935343	100
^{48}Ca	47.952534	0.193	^{60}Ni	59.930786	38.5198
			^{61}Ni	60.931056	1.6744
Sc	44.955912		^{62}Ni	61.928345	5.3388
^{45}Sc	44.955912	100	^{64}Ni	63.927966	1.3596
Ti	47.867		**Cu**	63.546[a]	
^{46}Ti	45.952632	11.19	^{63}Cu	62.929598	100
^{47}Ti	46.951763	10.09	^{65}Cu	64.927790	44.61
^{48}Ti	47.947946	100			
^{49}Ti	48.947870	7.34	**Zn**	65.409	
^{50}Ti	49.944791	7.03	^{64}Zn	63.929142	100
			^{66}Zn	65.926033	57.96
V	50.9415		^{67}Zn	66.927127	8.49
^{50}V	49.947159	0.251	^{68}Zn	67.924844	39.41
^{51}V	50.943960	100	^{70}Zn	69.925319	1.31

Element Isotope	Mass	Abundance	**Element** Isotope	Mass	Abundance
Ga	69.723		**Rb**	85.4678	
^{69}Ga	68.925574	100[c]	^{85}Rb	84.911790	100
^{71}Ga	70.924701	66.36	^{87}Rb	86.909181	38.56
Ge	72.64		**Sr**	87.62[a]	
^{70}Ge	69.924247	55.50	^{84}Sr	83.913425	0.68
^{72}Ge	71.922076	74.37	^{86}Sr	85.909260	11.94
^{73}Ge	72.923459	21.13	^{87}Sr	86.908877	8.48
^{74}Ge	73.921178	100	^{88}Sr	87.905612	100
^{76}Ge	75.921403	21.32			
			Y	88.905848	
As	74.921597		^{89}Y	88.905848	100
^{75}As	74.921597	100			
			Zr	91.224	
Se	78.96		^{90}Zr	89.904704	100
^{74}Se	73.922476	1.79	^{91}Zr	90.905646	21.81
^{76}Se	75.919214	18.89	^{92}Zr	91.905041	33.33
^{77}Se	76.919914	15.38	^{94}Zr	93.906315	33.78
^{78}Se	77.917309	47.91	^{96}Zr	95.908273	5.44
^{80}Se	79.916521	100			
^{82}Se	81.916699	17.60	**Nb**	92.906378	
			^{93}Nb	92.906378	100
Br	79.904				
^{79}Br	78.918337	100	**Mo**	95.94	
^{81}Br	80.916291	97.28	^{92}Mo	91.906811	61.06
			^{94}Mo	93.905088	38.16
Kr	83.798	(in air)	^{95}Mo	94.905842	65.72
^{78}Kr	77.920382	0.623[c]	^{96}Mo	95.904680	68.95
^{80}Kr	79.916379	4.011	^{97}Mo	96.906022	39.52
^{82}Kr	81.913484	20.343	^{98}Mo	97.905408	100
^{83}Kr	82.914136	20.180	^{100}Mo	99.907477	39.98
^{84}Kr	83.911507	100			
^{86}Kr	85.910611	30.321			

Element Isotope	Mass	Abundance	Element Isotope	Mass	Abundance
Ru	101.07		**In**	114.818	
^{96}Ru	95.907598	17.56	^{113}In	112.904058	4.48
^{98}Ru	97.905287	5.93	^{115}In	114.903878	100
^{99}Ru	98.905939	40.44			
^{100}Ru	99.904220	39.94	**Sn**	118.710	
^{101}Ru	100.905582	54.07	^{112}Sn	111.904818	2.98
^{102}Ru	101.904349	100	^{114}Sn	113.902779	2.03
^{104}Ru	103.905433	59.02	^{115}Sn	114.903342	1.04
			^{116}Sn	115.901741	44.63
Rh	102.905504		^{117}Sn	116.902952	23.57
^{103}Rh	102.905504	100	^{118}Sn	117.901603	74.34
			^{119}Sn	118.903309	26.37
Pd	106.42		^{120}Sn	119.902195	100
^{102}Pd	101.905609	3.73	^{122}Sn	121.903439	14.21
^{104}Pd	103.904036	40.76	^{124}Sn	123.905274	17.77
^{105}Pd	104.905085	81.71			
^{106}Pd	105.903486	100	**Sb**	121.760	
^{108}Pd	107.903892	96.82	^{121}Sb	120.903816	100
^{110}Pd	109.905153	42.88	^{123}Sb	122.904214	74.79
Ag	107.8682		**Te**	127.60	
^{107}Ag	106.905097	100	^{120}Te	119.904020	0.26
^{109}Ag	108.904752	92.90	^{122}Te	121.903044	7.48
			^{123}Te	122.904270	2.61
Cd	112.411		^{124}Te	123.902818	13.91
^{106}Cd	105.906459	4.35	^{125}Te	124.904431	20.75
^{108}Cd	107.904184	3.10	^{126}Te	125.903312	55.28
^{110}Cd	109.903002	43.47	^{128}Te	127.904463	93.13
^{111}Cd	110.904178	44.55	^{130}Te	129.906224	100
^{112}Cd	111.902758	83.99			
^{113}Cd	112.904402	42.53	**I**	126.904473	
^{114}Cd	113.903359	100	^{127}I	126.904473	100
^{116}Cd	115.904756	26.07			

Element Isotope	Mass	Abundance
Xe	131.293	
^{124}Xe	123.905893	0.354[c]
^{126}Xe	125.904274	0.330
^{128}Xe	127.903531	7.099
^{129}Xe	128.904779	98.112
^{130}Xe	129.903508	15.129
^{131}Xe	130.905082	78.906
^{132}Xe	131.904154	100
^{134}Xe	133.905395	38.782
^{136}Xe	135.907219	32.916
Cs	132.905452	
^{133}Cs	132.905452	100
Ba	137.327	
^{130}Ba	129.906321	0.148
^{132}Ba	131.905061	0.141
^{134}Ba	133.904508	3.371
^{135}Ba	134.905689	9.194
^{136}Ba	135.904576	10.954
^{137}Ba	136.905827	15.666
^{138}Ba	137.905247	100
La	138.90547	
^{138}La	137.907112	0.090
^{139}La	138.906353	100
Ce	140.116	
^{136}Ce	135.907172	0.209
^{138}Ce	137.905991	0.284
^{140}Ce	139.905439	100
^{142}Ce	141.909244	12.565
Pr	140.907653	
^{141}Pr	140.907653	100

Element Isotope	Mass	Abundance
Nd	144.242	
^{142}Nd	141.907723	100
^{143}Nd	142.909815	44.9
^{144}Nd	143.910087	87.5
^{145}Nd	144.912574	30.5
^{146}Nd	145.913117	63.2
^{148}Nd	147.916893	21.0
^{150}Nd	149.920891	20.6
Sm	150.36	
^{144}Sm	143.911999	11.48
^{147}Sm	146.914898	56.04
^{148}Sm	147.914823	42.02
^{149}Sm	148.917185	51.66
^{150}Sm	149.917276	27.59
^{152}Sm	151.919732	100
^{154}Sm	153.922209	85.05
Eu	151.964	
^{151}Eu	150.919850	91.61
^{153}Eu	152.921230	100
Gd	157.25	
^{152}Gd	151.919791	0.81
^{154}Gd	153.920866	8.78
^{155}Gd	154.922622	59.58
^{156}Gd	155.922123	82.41
^{157}Gd	156.923960	63.00
^{158}Gd	157.924104	100
^{160}Gd	159.927054	88.00
Tb	158.925347	
^{159}Tb	158.925347	100

Element Isotope	Mass	Abundance	**Element** Isotope	Mass	Abundance
Dy	162.500		**Hf**	178.49	
^{156}Dy	155.924283	0.20	^{174}Hf	173.940046	0.46
^{158}Dy	157.924409	0.34	^{176}Hf	175.941409	14.99
^{160}Dy	159.925198	8.24	^{177}Hf	176.943221	53.02
^{161}Dy	160.926933	66.84	^{178}Hf	177.943699	77.77
^{162}Dy	161.926798	90.15	^{179}Hf	178.944816	38.83
^{163}Dy	162.928731	88.10	^{180}Hf	179.946550	100
^{164}Dy	163.929175	100			
			Ta	180.94788	
Ho	164.930322		^{180}Ta	179.947465	0.012
^{165}Ho	164.930322	100	^{181}Ta	180.947996	100
Er	167.259		**W**	183.84	
^{162}Er	161.928778	0.41	^{180}W	179.946704	0.39
^{164}Er	163.929200	4.78	^{182}W	181.948204	86.49
^{166}Er	165.930293	100	^{183}W	182.950223	46.70
^{167}Er	166.932048	68.26	^{184}W	183.950931	100.0
^{168}Er	167.932370	80.52	^{186}W	185.954364	92.79
^{170}Er	169.935464	44.50			
			Re	186.207	
Tm	168.934213		^{185}Re	184.952955	59.74
^{169}Tm	168.934213	100	^{187}Re	186.955753	100
Yb	173.04		**Os**	190.23	
^{168}Yb	167.933897	0.41	^{184}Os	183.952489	0.05
^{170}Yb	169.934762	9.55	^{186}Os	185.953838	3.90
^{171}Yb	170.936326	44.86	^{187}Os	186.955751	4.81
^{172}Yb	171.936382	68.58	^{188}Os	187.955838	32.47
^{173}Yb	172.938211	50.68	^{189}Os	188.958148	39.60
^{174}Yb	173.938862	100	^{190}Os	189.958447	64.39
^{176}Yb	175.942572	40.09	^{192}Os	191.961481	100
Lu	174.967		**Ir**	192.217	
^{175}Lu	174.940772	100	^{191}Ir	190.960594	59.49
^{176}Lu	175.942686	2.66	^{193}Ir	192.962926	100.0

Element Isotope	Mass	Abundance	**Element** Isotope	Mass	Abundance
Pt	195.084		**Tl**	204.3833	
^{190}Pt	189.959932	0.041	^{203}Tl	202.972344	41.88
^{192}Pt	191.961038	2.311	^{205}Tl	204.974428	100
^{194}Pt	193.962680	97.443			
^{195}Pt	194.964791	100	**Pb**	207.2[a]	
^{196}Pt	195.964952	74.610	^{204}Pb	203.973044	2.7
^{198}Pt	197.967893	21.172	^{206}Pb	205.974465	46.0
			^{207}Pb	206.975897	42.2
Au	196.966569		^{208}Pb	207.976653	100
^{197}Au	196.966569	100			
			Bi	208.980399	
Hg	200.59		^{209}Bi	208.980399	100
^{196}Hg	195.965833	0.50			
^{198}Hg	197.966769	33.39	**Th**	232.038055	
^{199}Hg	198.968280	56.50	^{232}Th	232.038055	100
^{200}Hg	199.968326	77.36			
^{201}Hg	200.970302	44.14	**U**	238.02891	
^{202}Hg	201.970643	100	^{234}U	234.040952	0.0054[e]
^{204}Hg	203.973494	23.01	^{235}U	235.043930	0.7257
			^{238}U	238.050788	100

[a] Natural variations in the isotopic composition of terrestrial materials do not allow to give a more precise value.

[b] The mole ratio of ^{2}H in hydrogen from gas cylinders was reported to be as low as 0.000032.

[c] Commercially available materials may have substantially different isotopic compositions if they were subjected to undisclosed or inadvertent isotopic fractionation.

[d] Materials depleted in ^{6}Li are commercial sources of laboratory shelf reagents and are known to have ^{6}Li abundances in the range of 2.0007–7.672 atom percent, with natural materials at the higher end of this range. Average atomic masses vary between 6.939 and 6.996; if a more accurate value is required, it must be determined for the specific material.

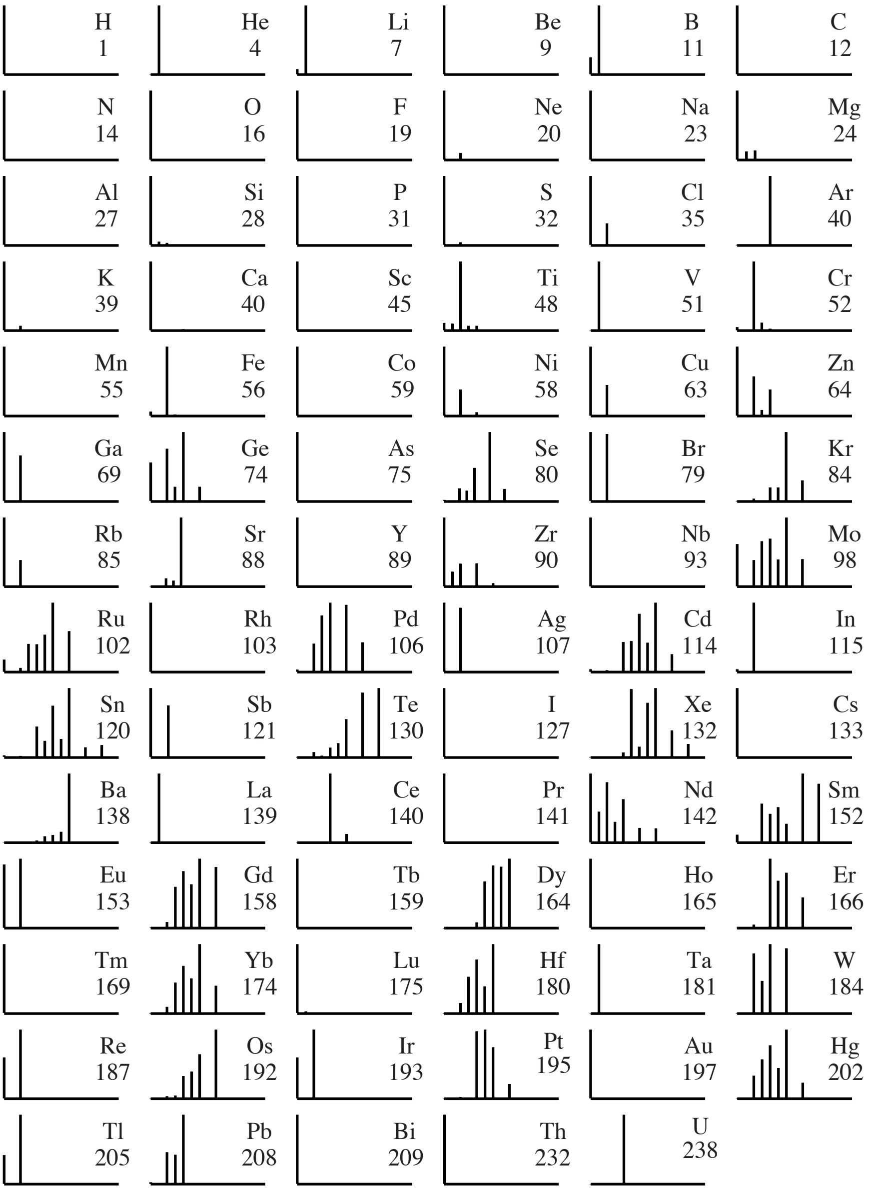
[e] Materials depleted in ^{235}U are commercial sources of laboratory shelf reagents.

2.5.2 Ranges of Natural Isotope Abundances of Selected Elements [3]

Element Isotope	Range [atom %]	**Element** Isotope	Range [atom %]	**Element** Isotope	Range [atom %]
H		**Si**		**Sr**	
^{1}H	99.9816–99.9974	^{28}Si	92.205–92.241	^{84}Sr	0.55–0.58
^{2}H	0.0026–0.0184	^{29}Si	4.678–4.692	^{86}Sr	9.75–9.99
		^{30}Si	3.082–3.102	^{87}Sr	6.94–7.14
He				^{88}Sr	82.29–82.75
^{3}He	4.6×10^{-8}–0.0041	**S**			
^{4}He	99.9959–100	^{32}S	94.454–95.281	**Ce**	
		^{33}S	0.730–0.793	^{136}Ce	0.185–0.186
Li		^{34}S	3.976–4.734	^{138}Ce	0.251–0.254
^{6}Li	7.225–7.714	^{36}S	0.013–0.019	^{140}Ce	88.446–88.449
^{7}Li	92.275–92.786			^{142}Ce	11.114–11.114
		Cl			
B		^{35}Cl	75.644–75.923	**Nd**	
^{10}B	18.929–20.386	^{37}Cl	24.077–24.356	^{142}Nd	26.80–27.30
^{11}B	79.614– 81.071			^{143}Nd	12.12–12.32
		Ca		^{144}Nd	23.79–23.97
C		^{40}Ca	96.933–96.947	^{145}Nd	8.23–8.35
^{12}C	98.853–99.037	^{42}Ca	0.646–0.648	^{146}Nd	17.06–17.35
^{13}C	0.963–1.147	^{43}Ca	0.135–0.135	^{148}Nd	5.66–5.78
		^{44}Ca	2.082–2.092	^{150}Nd	5.53–5.69
N		^{46}Ca	0.004–0.004		
^{14}N	99.579–99.654	^{48}Ca	0.186–0.188	**Pb**	
^{15}N	0.346–0.421			^{204}Pb	1.04–1.65
		V		^{206}Pb	20.84–27.48
O		^{50}V	0.2487–0.2502	^{207}Pb	17.62–23.65
^{16}O	99.738–99.776	^{51}V	99.7498–99.7513	^{208}Pb	51.28–56.21
^{17}O	0.037–0.040				
^{18}O	0.188–0.222	**Cu**		**U**	
		^{63}Cu	68.983–69.338	^{234}U	0.0050–0.0059
Ne		^{65}Cu	30.662–31.017	^{235}U	0.7198–0.7207
^{20}Ne	88.47–90.51			^{238}U	99.2739–99.2752
^{21}Ne	0.27–1.71				
^{22}Ne	9.20– 9.96				

2.5.3 Isotope Patterns of Naturally Occurring Elements

The mass of the most abundant isotope is given under the symbol of the element. The lightest isotope is shown at the left end of the x axis.



2.5.4 Calculation of Isotope Distributions

The characteristic abundance patterns resulting from the combination of more than one polyisotopic element can be calculated from the relative abundances of the different isotopes. The following polynomial expression gives the isotope distribution of a polyisotopic molecule:

$$\{p_{i1}\,A^0 + p_{i2}\,A^{(m_{i2}-m_{i1})} + p_{i3}\,A^{(m_{i3}-m_{i1})} + \ \ldots\}^{n_i} \times$$
$$\{p_{j1}\,A^0 + p_{j2}\,A^{(m_{j2}-m_{j1})} + p_{j3}\,A^{(m_{j3}-m_{j1})} + \ \ldots\}^{n_j} \times \{\ldots$$

where p_{ix} is the relative abundance of the xth isotope of element i, m_{ix} is the mass of the xth isotope of the element i, and the exponent n_i stands for the number of atoms of the element i in the molecule. The expansion of this polynomial expression after inserting the p_{ix} and m_{ix} values for all the isotopes 1, 2, 3, … of the elements i, j, … of a given molecule yields an expression that represents the isotope distribution:

$$w_0\,A^0 + w_r\,A^r + w_s\,A^s + w_t\,A^t + \ldots$$

where the values of w_0, w_r, w_s, w_t, … are the relative abundances of $M^{+\cdot}$, $[M+r]^{+\cdot}$, $[M+s]^{+\cdot}$, $[M+t]^{+\cdot}$,… , respectively. The use of $A^{(m_{ix}-m_{i1})}$ allows to determine the values of *r, s, t*,… simply by expanding the general polynomial. A numerical value for A, which has no intrinsic meaning, is never needed.

For example, for CBr_2Cl_2, the above equation gives rise to the following expression:

$$\{p_{^{12}C}\,A^0 + p_{^{13}C}\,A^{(m_{^{13}C}-m_{^{12}C})}\} \times$$
$$\{p_{^{79}Br}\,A^0 + p_{^{81}Br}\,A^{(m_{^{81}Br}-m_{^{79}Br})}\}^2 \times$$
$$\{p_{^{35}Cl}\,A^0 + p_{^{37}Cl}\,A^{(m_{^{37}Cl}-m_{^{35}Cl})}\}^2$$

For sufficient resolution, $(m_{ix} - m_{i1})$ and $(m_{jx} - m_{j1})$ differ from one another. This results in very complex isotope patterns even for very small molecules. Thus, owing to the occurrence of ^{12}C, ^{13}C, ^{79}Br, ^{81}Br, ^{35}Cl, and ^{37}Cl, there are 18 signals for CBr_2Cl_2. However, the limited resolution of many real life experiments can make many pairs of $(m_{ix} - m_{i1})$ and $(m_{jx} - m_{j1})$ indistinguishable within experimental error, thereby reducing the number of separate peaks. For example, at unit resolution, one obtains $(m_{^{13}C} - m_{^{12}C}) = 1$ and $(m_{^{81}Br} - m_{^{79}Br}) = (m_{^{37}Cl} - m_{^{35}Cl}) = 2$. Consequently, the expression for CBr_2Cl_2 becomes:

$$\{p_{^{12}C}A^0 + p_{^{13}C}A^1\} \times \{p_{^{79}Br}A^0 + p_{^{81}Br}A^2\}^2 \times \{p_{^{35}Cl}A^0 + p_{^{37}Cl}A^2\}^2 =$$

$$\{p_{^{12}C}p^2_{^{79}Br}p^2_{^{35}Cl}\}A^0 +$$

$$\{p_{^{13}C}p^2_{^{79}Br}p^2_{^{35}Cl}\}A^1 +$$

$$\{p_{^{12}C}p_{^{79}Br}p_{^{81}Br}p^2_{^{35}Cl} + p_{^{12}C}p^2_{^{79}Br}p_{^{35}Cl}p_{^{37}Cl}\}A^2 +$$

$$\{p_{^{13}C}p_{^{79}Br}p_{^{81}Br}p^2_{^{35}Cl} + p_{^{13}C}p^2_{^{79}Br}p_{^{35}Cl}p_{^{37}Cl}\}A^3 +$$

$$\{p_{^{12}C}p^2_{^{81}Br}p^2_{^{35}Cl} + 4p_{^{12}C}p_{^{79}Br}p_{^{81}Br}p_{^{35}Cl}p_{^{37}Cl} + p_{^{12}C}p^2_{^{79}Br}p^2_{^{37}Cl}\}A^4 +$$

$$\{p_{^{13}C}p^2_{^{81}Br}p^2_{^{35}Cl} + 4p_{^{13}C}p_{^{79}Br}p_{^{81}Br}p_{^{35}Cl}p_{^{37}Cl} + p_{^{13}C}p^2_{^{79}Br}p^2_{^{37}Cl}\}A^5 +$$

$$\{p_{^{12}C}p_{^{79}Br}p_{^{81}Br}p^2_{^{37}Cl} + p_{^{12}C}p^2_{^{81}Br}p_{^{35}Cl}p_{^{37}Cl}\}A^6 +$$

$$\{p_{^{13}C}p_{^{79}Br}p_{^{81}Br}p^2_{^{37}Cl} + p_{^{13}C}p^2_{^{81}Br}p_{^{35}Cl}p_{^{37}Cl}\}A^7 +$$

$$\{p_{^{12}C}p^2_{^{81}Br}p^2_{^{37}Cl}\}A^8 +$$

$$\{p_{^{13}C}p^2_{^{81}Br}p^2_{^{37}Cl}\}A^9$$

This shows that at unit resolution, CBr_2Cl_2 gives rise to only 10 peaks ($M^{+\cdot}$, $[M+1]^{+\cdot}$, $[M+2]^{+\cdot}$, ... $[M+9]^{+\cdot}$) rather than 18 peaks, as they would be expected for very high resolution. Moreover, the contribution of isotopes of low abundance can often be neglected without sacrificing much precision. For example, the effect of ^2H on isotope patterns is usually insignificant. Also, ^{13}C is often negligible when focussing on peaks of the series $[M+2n]^{+\cdot}$, which then results in patterns that are characteristic for halogens, sulfur, and silicon. In large molecules, however, isotopes of low abundance cannot be neglected. For example, in the case of buckminster fullerene (C_{60}), not only $M^{+\cdot}$ (relative intensity, 100%) and $[M+1]^{+\cdot}$ (64.80%), but also $[M+2]^{+\cdot}$ (20.65%), $[M+3]^{+\cdot}$ (4.31%), and even $[M+4]^{+\cdot}$ (0.66%) are quite significant ions.

With the above algorithm, typical isotope patterns can be readily calculated manually by applying the general equation and neglecting isotopes of low abundance. The outlined procedure can also be easily implemented and evaluated with generic computer software that allows simple calculations. Dedicated and user-friendly programs that already contain the necessary isotope abundances and masses are available. Incidentally, because the use of the above equation for systems with 1000 or more polyisotopic atoms results in excessive calculation times, more efficient but somewhat more complicated algorithms have been developed for implementation in dedicated programs [4]. Typical isotope patterns are given on the following pages.

2.5.5 Isotopic Abundances of Various Combinations of Chlorine, Bromine, Sulfur, and Silicon

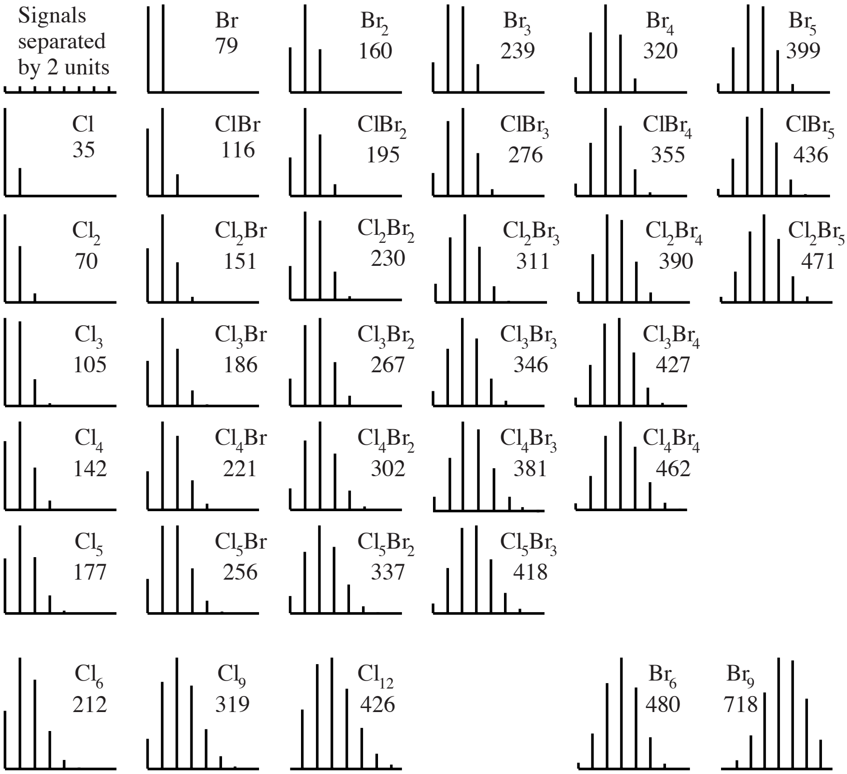
Elements	Mass	Relative abundance	Elements	Mass	Relative abundance	Elements	Mass	Relative abundance
Cl_1	35	100	Br_1	79	100	S_1	32	100
	37	31.96		81	97.28		33	0.80
							34	4.52
Cl_2	70	100	Br_2	158	51.40			
	72	63.92		160	100	S_2	64	100
	74	10.21		162	48.64		65	1.60
							66	9.05
Cl_3	105	100	Br_3	237	34.27		68	0.20
	107	95.88		239	100			
	109	30.64		241	97.28	S_3	96	100
	111	3.26		243	31.54		97	2.40
							98	13.58
Cl_4	140	78.22	Br_4	316	17.61		99	0.22
	142	100		318	68.53		100	0.61
	144	47.94		320	100			
	146	10.21		322	64.85	S_4	128	100
	148	0.82		324	15.77		129	3.20
							130	18.12
Cl_5	175	62.53	Br_5	395	10.57		131	0.43
	177	100		397	51.40		132	1.23
	179	63.92		399	100			
	181	20.43		401	97.28	S_5	160	100
	183	3.26		403	47.32		161	4.00
	185	0.21		405	9.21		162	22.66
							163	0.72
Cl_6	210	52.15	Br_6	474	5.43		164	2.05
	212	100		476	31.70		166	0.09
	214	79.90		478	77.10			
	216	34.05		480	100			
	218	8.16		482	72.96			
	220	1.04		484	28.39			
	222	0.06		486	4.60			

Elements	Mass	Relative abundance	Elements	Mass	Relative abundance	Elements	Mass	Relative abundance
Si_1	28	100	Si_2	56	100	Si_3	84	100
	29	5.08		57	10.15		85	15.23
	30	3.35		58	6.95		86	10.82
				59	0.34		87	1.03
				60	0.11		88	0.36
Cl_1Br_1	114	77.38	Cl_1Br_2	193	44.14	Cl_1Br_3	272	26.51
	116	100		195	100		274	85.85
	118	24.06		197	69.23		276	100
				199	13.35		278	48.46
							280	7.80
Cl_1Br_4	351	14.45	Cl_2Br_1	149	62.03	Cl_2Br_2	228	38.69
	353	60.84		151	100		230	100
	355	100		153	44.91		232	88.68
	357	79.42		155	6.16		234	31.09
	359	29.94					236	3.74
	361	4.14						
Cl_3Br_1	184	51.77	Cl_3Br_2	263	32.07	Cl_4Br_1	219	44.42
	186	100		265	93.14		221	100
	188	64.15		267	100		223	82.47
	190	17.12		269	49.27		225	32.28
	192	1.64		271	11.34		227	6.11
				273	0.99		229	0.45
Cl_1S_1	67	100	Cl_1S_2	99	100	Cl_2S_1	102	100
	68	0.80		100	1.60		103	0.80
	69	36.48		101	41.01		104	68.44
	70	0.26		102	0.58		105	0.51
	71	1.44		103	3.10		106	13.10
							108	0.46

Elements	Mass	Relative abundance	Elements	Mass	Relative abundance	Elements	Mass	Relative abundance
Cl_1Si_1	63	100	Cl_2Si_1	98	100	Cl_3Si_1	133	100
	64	5.08		99	5.08		134	5.08
	65	35.31		100	67.27		135	99.23
	66	1.62		101	3.25		136	4.87
	67	1.07		102	12.35		137	33.85
				103	0.52		138	1.56
				104	0.34		139	4.29

2.5.6 Isotope Patterns of Combinations of Cl and Br

The signals are separated by 2 mass units. The mass for the most abundant signal is shown under the symbol of the element. The combination of the lightest isotopes is given on the left side of the x axis. See Chapter 2.5.5 for exact abundances of many of these combinations.



2.5.7 Indicators of the Presence of Heteroatoms

In low-resolution mass spectra, one often observes characteristic isotope patterns, specific masses of fragment ions, and characteristic mass differences (Δm) between the molecular ion ($M^{+\cdot}$) and fragment ions ($frag^{+}$) or between fragment ions. High resolution mass spectra can be used to confirm the elemental composition provided that the resolution is sufficient to discriminate alternative compositions. Moreover, tandem mass spectrometry (also called MS/MS) may be used to identify characteristic losses of heteroatoms from parent or fragment ions:

Indication of O:	Δm 17 from $M^{+\cdot}$, in N-free compounds
	Δm 18 from $M^{+\cdot}$
	Δm 18 from $frag^{+}$, particularly in aliphatic compounds
	Δm 28, 29 from $M^{+\cdot}$ for aromatic compounds
	Δm 28 from $frag^{+}$ for aromatic compounds
	m/z 15, relatively abundant
	m/z 19
	m/z 31, 45, 59, 73, … + $(14)_n$
	m/z 32, 46, 60, 74, … + $(14)_n$
	m/z 33, 47, 61, 75, … + $(14)_n$ for 2 × O, in absence of S
	m/z 69 for aromatic compounds meta-disubstituted by O
Indication of N:	$M^{+\cdot}$ odd-numbered (indicates odd number of N in $M^{+\cdot}$)
	Large number of even-numbered fragment ions
	Δm 17 from $M^{+\cdot}$ or $frag^{+}$, in O-free compounds
	Δm 27 from $M^{+\cdot}$ or $frag^{+}$, for aromatic compounds or nitriles
	Δm 30, 46 for nitro compounds
	m/z 30, 44, 58, 72, … + $(14)_n$ for aliphatic compounds
Indication of S:	Isotope peak $[M+2]^{+\cdot}$ ≥ 5% of $M^{+\cdot}$
	Δm 33, 34, 47, 48, 64, 65 from $M^{+\cdot}$
	Δm 34, 48, 64 from $frag^{+}$
	m/z 33, 34, 35
	m/z 45 in O-free compounds
	m/z 47, 61, 75, 89, … + $(14)_n$ unless compound with 2 × O
	m/z 48, 64 for *S*-oxides
Indication of F:	Δm 19, 20, 50 from $M^{+\cdot}$
	Δm 20 from $frag^{+}$
	m/z 20
	m/z 57 without m/z 55 in aromatics

Indication of Cl: Isotope peak $[M+2]^{+\cdot} \geq 33\%$ of $M^{+\cdot}$
Δm 35, 36 from $M^{+\cdot}$
Δm 36 from $frag^{+}$
m/z 35/37, 36/38, 49/51

Indication of Br: Isotope peak $[M+2]^{+\cdot} \geq 98\%$ of $M^{+\cdot}$
Δm 79, 80 from $M^{+\cdot}$
Δm 80 from $frag^{+}$
m/z 79/81, 80/82

Indication of I: Isotope peak $[M+1]^{+}$ of very low abundance at relatively high mass
Δm 127 from $M^{+\cdot}$
Δm 127, 128 from $frag^{+}$
m/z 127, 128, 254

Indication of P: m/z 47 in compounds without S or $2 \times O$
m/z 99 without isotope peak at m/z 100 in alkyl phosphates

2.5.8 Rules for Determining the Relative Molecular Weight (M_r)

The molecular ion ($M^{+\cdot}$) is defined as the ion that comprises the most abundant isotopes of the elements in the molecule. Interestingly, the lightest isotopes of most elements frequently occurring in organic compounds and their common salts (H, C, N, O, F, Si, P, S, Cl, As, Br, I, Na, Mg, Al, K, Ca, Rb, Cs) are also the most abundant ones. Notable exceptions are B, Li, Se, Sr, and Ba.

$M^{+\cdot}$ is always accompanied by isotope peaks. Their relative abundance depends on the number and kind of the elements present and their natural isotopic distribution. The abundance of $[M+1]^{+\cdot}$ indicates the maximum number of carbon atoms (C_{max}) according to the following relationship:

$$C_{max} = 100 \times \text{intensity}([M+1]^{+\cdot}) \,/\, \{1.1 \times \text{intensity}(M^{+\cdot})\}$$

$[M+2]^{+\cdot}$ and higher masses indicate the number and kind of elements that have a relatively abundant heavier isotope (such as S, Si, Cl, Br). Note that, in analogy to the calculation of C_{max}, the ratio of the intensities of $[M+2]^{+\cdot}$ and $M^{+\cdot}$ for a compound with *n* silicon, *o* sulfur, *p* chlorine, or *q* bromine atoms can be approximated with quite high accuracy from $n \times 3.35\%$, $o \times 4.52\%$, $p \times 31.96\%$, or $q \times 97.28\%$, respectively (see also Chapters 2.5.4 to 2.5.6).

The mass of $M^{+\cdot}$ is always an even number if the molecule contains only elements for which the atomic mass and valence are both even- (C, O, S, Si) or both odd-numbered (H, P, F, Cl, Br, I). In the presence of other elements (e.g., ^{14}N) and isotope labels (e.g., ^{13}C, ^{2}H), $M^{+\cdot}$ becomes an odd number if they are present in an odd number.

The molecular ion can only form fragment ions of masses that differ from that of $M^{+\cdot}$ by chemically logical values (Δm). In this context, chemically illogical differences are $\Delta m = 3$ (in the absence of $\Delta m = 1$) to $\Delta m = 14$, $\Delta m = 21$ (in the absence of $\Delta m = 1$) to $\Delta m = 24$, $\Delta m = 37, 38$, and all Δm less than the mass of an element of characteristic isotope pattern in cases where the same isotope pattern is not retained in the fragment ion.

$M^{+\cdot}$ must contain all elements (and the maximum number of each) that are shown to be present in the fragment ions.

If ionization is performed by electron impact, $M^{+\cdot}$ is the ion with the lowest appearance potential.

If a pure sample flows into the ion source through a molecular leak, $M^{+\cdot}$ exhibits the same effusion rate as can be determined from the fragment ions. The abundance of $M^{+\cdot}$ is proportional to the sample pressure in the ion source.

For polar compounds, $[M+H]^{+}$ is often observed in mass spectra obtained not only with fast atom bombardment and atmospheric pressure chemical ionization but also with electron impact ionization. In this latter case, the abundance of $[M+H]^{+}$ changes in proportion to the square of the sample pressure in the ion source.

In the absence of a signal for $M^{+\cdot}$, the relative molecular weight must have a value that shows a logical and reasonable mass difference, Δm, to all the observed fragment ions.

2.5.9 Homologous Mass Series as Indications of Structural Type

Certain sequences of intensity maxima in the lower mass range and the masses of unique signals are often characteristic of a particular compound type. The intensity distribution of such ion series is in general smooth. Therefore, abrupt changes (maxima and minima) are of structural significance. The ion or ion series most indicative of a particular compound type is set in italics.

Mass values, m/z	Elemental composition	Compound types
12 + 14n	C_nH_{2n-2}	alkenes, monocycloalkanes, alkynes, dienes, cycloalkenes, polycyclic alicyclics, cyclic alcohols
13 + 14n	C_nH_{2n-1}	alkanes, alkenes, *monocycloalkanes*, alkynes, dienes, cycloalkenes, polycyclic alicyclics, alcohols, alkyl ethers, cyclic alcohols, cycloalkanones, aliphatic acids, esters, lactones, thiols, sulfides, glycols, glycol ethers, alkyl chlorides
	$C_nH_{2n-3}O$	cycloalkanones
14 + 14n	C_nH_{2n}	alkanes, alkenes, monocycloalkanes, polycyclic alicyclics, alcohols, alkyl ethers, thiols, sulfides, alkyl chlorides
	$C_nH_{2n-2}O$	cycloalkanones
15 + 14n	C_nH_{2n+1}	*alkanes*, alkenes, monocycloalkanes, alkynes, dienes, cycloalkenes, polycyclic alicyclics, alkanones, alkanals, glycols, glycol ethers, alkyl chlorides, acid chlorides
	$C_nH_{2n-1}O$	alkanones, alkanals, *cyclic alcohols*, acid chlorides
16 + 14n	$C_nH_{2n}O$	*alkanones, alkanals*
	$C_nH_{2n+2}N$	*alkyl amines, aliphatic amides*
	$C_nH_{2n}NO$	aliphatic amides
17 + 14n	$C_nH_{2n+1}O$	*alcohols, alkyl ethers,* aliphatic acids, esters, lactones, glycols, glycol ethers
	$C_nH_{2n-1}O_2$	aliphatic acids, esters, lactones
18 + 14n	$C_nH_{2n}O_2$	*aliphatic acids, esters, lactones*
19 + 14n	$C_nH_{2n+3}O$	alcohols, alkyl ethers
	$C_nH_{2n+1}O_2$	aliphatic acids, esters, lactones
	$C_nH_{2n+1}O_2$	*glycols, glycol ethers*
	$C_nH_{2n+1}S$	*thiols, sulfides*
20 + 14n	$C_8H_8 + C_nH_{2n}$	alkylbenzenes
	$C_nH_{2n+2}O_2$	glycols, glycol ethers
	$C_nH_{2n+2}S$	thiols, sulfides

Mass values m/z	Elemental composition	Compound types
21 + 14n	$C_7H_7 + C_nH_{2n}$	alkylbenzenes
	C_7H_5O	aryl ketones
	$C_nH_{2n}Cl$	*alkyl chlorides*
	$C_nH_{2n}COCl$	acid chlorides
22 + 14n	$C_6H_6N + C_nH_{2n}$	alkylanilines
	C_nH_{2n-6}	polycyclic alicyclics
23 + 14n	C_nH_{2n-5}	polycyclic alicyclics
24 + 14n	C_nH_{2n-4}	polycyclic alicyclics
25 + 14n	C_nH_{2n-3}	*alkynes, dienes, cycloalkenes*, polycyclic alicyclics
39, 52±1, 64±1, 76±2, 91±1	$C_nH_{n\pm1}$	alkylbenzenes, aromatic hydrocarbons, phenols, aryl ethers, aryl ketones

2.5.10 Mass Correlation Table

Note: As long as it makes sense chemically, CH_2, CH_4, CH_3O, and O_2 in the formulae of the second column may be replaced by N, O, P, and S, respectively.

Mass	Ion	Product ion (and neutral particle lost)		Substructure or compound type
1		$[M+1]^+$, $[M-1]^-$		particularly in FAB spectra, in which M±1 occurs even for moderately basic and acidic compounds, but intensive $M^{+\cdot}$ without M±1 is unusual
7	$Li^{+\cdot}$	$[M+7]^+$		in FAB spectra in the presence of Li^+ (with isotope signal for ^6Li)
		$[M-7]^-$		in FAB spectra of organic Li^+ salts
12	$C^{+\cdot}$			
13	CH^+			
14	$CH_2^{+\cdot}$, N^+, N_2^{++}, CO^{++}			
15	CH_3^+	$[M-15]^{+\cdot}$	(CH_3)	nonspecific; *abundant:* methyl, *N*-ethylamines
16	$O^{+\cdot}$, NH_2^+, O_2^{++}	$[M-16]^{+\cdot}$	(CH_4)	methyl (rare)
			(O)	nitro compounds, sulfones, epoxides, *N*-oxides
			(NH_2)	primary amines
17	OH^+, $NH_3^{+\cdot}$	$[M-17]^{+\cdot}$	(OH)	acids (especially aromatic acids), hydroxylamines, *N*-oxides, nitro compounds, sulfoxides, tertiary alcohols
			(NH_3)	primary amines
18	$H_2O^{+\cdot}$, NH_4^+	$[M-18]^{+\cdot}$	(H_2O)	nonspecific; *abundant:* alcohols, some acids, aldehydes, ketones, lactones, cyclic ethers **O indicator**
19	H_3O^+, F^+	$[M-19]^+$	(F)	fluoro compounds **F indicator**
20	$HF^{+\cdot}$, Ar^{++}, CH_2CN^{++}	$[M-20]^{+\cdot}$	(HF)	fluoro compounds **F indicator**
21	$C_2H_2O^{++}$			
22	CO_2^{++}			

Mass	Ion	Product ion (and neutral particle lost)	Substructure or compound type
23	Na^+	$[M+23]^+$	in FAB spectra in the presence of Na^+; sometimes strong even if Na^+ is only an impurity
		$[M-23]^-$	in FAB spectra of organic Na^+ salts
24	$C_2^{+\cdot}$		
25	C_2H^+	$[M-25]^+$ (C_2H)	terminal acetylenyl
26	$C_2H_2^{+\cdot}$, CN^+	$[M-26]^{+\cdot}$ (C_2H_2)	aromatics
		(CN)	nitriles
27	$C_2H_3^+$, $HCN^{+\cdot}$	$[M-27]^+$ (C_2H_3)	terminal vinyl, some ethyl esters and *N*-ethylamides, ethyl phosphates
		$[M-27]^{+\cdot}$ (HCN)	aromatic N, nitriles
28	$C_2H_4^{+\cdot}$, $CO^{+\cdot}$, $N_2^{+\cdot}$, $HCNH^+$	$[M-28]^{+\cdot}$ (C_2H_4)	nonspecific; *abundant:* cyclohexenes, ethyl esters, propyl ketones, propyl-substituted aromatics
		(CO)	aromatic O, quinones, lactones, lactams, unsaturated cyclic ketones, allyl aldehydes
		(N_2)	diazo compounds; air (intensity 3.7 times larger than for $O_2^{+\cdot}$, m/z 32)
29	$C_2H_5^+$, CHO^+	$[M-29]^+$ (C_2H_5)	nonspecific; *abundant:* ethyl
		(CHO)	phenols, furans, aldehydes
30	$CH_2O^{+\cdot}$, $CH_2NH_2^+$, NO^+, $C_2H_6^{+\cdot}$, $BF^{+\cdot}$, $N_2H_2^{+\cdot}$	$[M-30]^{+\cdot}$ (C_2H_6)	ethylalkanes, polymethyl compounds
		(CH_2O)	cyclic ethers, lactones, primary alcohols
		(NO)	nitro and nitroso compounds **N indicator**
31	CH_3O^+, $CH_3NH_2^{+\cdot}$, CF^+, $N_2H_3^+$	$[M-31]^+$ (CH_3O)	methyl esters, methyl ethers, primary alcohols **O indicator**
		(CH_3NH_2)	*N*-methylamines
		(N_2H_3)	hydrazides
32	$O_2^{+\cdot}$, $CH_3OH^{+\cdot}$, $N_2H_4^{+\cdot}$, $S^{+\cdot}$	$[M-32]^{+\cdot}$ (O_2)	cyclic peroxides; air (intensity 3.7 times smaller than for $N_2^{+\cdot}$, m/z 28)
		(CH_3OH)	methyl esters, methyl ethers
		(S)	sulfides (with ^{34}S isotope signal) **O indicator**

Mass	Ion	Product ion (and neutral particle lost)	Substructure or compound type
33	$CH_3OH_2^+$, SH^+, CH_2F^+	$[M-33]^+$ (SH)	nonspecific (with isotope signal for ^{34}S) **S indicator**
		($CH_3 + H_2O$)	nonspecific **O indicator**
		(CH_2F)	fluoromethyl
34	$SH_2^{+\cdot}$	$[M-34]^{+\cdot}$ (SH_2)	nonspecific (with ^{34}S isotope signal) **S indicator**
		(OH + OH)	nitro compounds
35	SH_3^+, Cl^+	$[M-35]^+$ (Cl)	chloro compounds (with ^{37}Cl isotope signal)
		($OH + H_2O$)	nitro compounds **2 × O indicator**
36	$HCl^{+\cdot}$, C_3^+	$[M-36]^{+\cdot}$ (HCl)	chloro compounds
		($H_2O + H_2O$)	**2 × O indicator**
37	C_3H^+		
	$^{37}Cl^+$		chloro compounds (with isotope signal for ^{35}Cl)
38	$C_3H_2^{+\cdot}$		
39	$C_3H_3^+$	$[M-39]^+$ (C_3H_3)	aromatics
	K^+	$[M+39]^+$	in FAB spectra often strong even if K^+ is only an impurity (with isotope signal for ^{41}K)
		$[M-39]^-$	in FAB spectra of organic K^+ salts
40	$C_3H_4^{+\cdot}$, $Ar^{+\cdot}$, CH_2CN^+	$[M-40]^{+\cdot}$ (CH_2CN)	cyanomethyl
41	$C_3H_5^+$, $CH_3CN^{+\cdot}$	$[M-41]^+$ (C_3H_5)	alicyclics (*especially* polyalicyclics), alkenes
		(CH_3CN)	2-methyl-*N*-aromatics, *N*-methyl-anilines
42	$C_3H_6^{+\cdot}$, $C_2H_2O^{+\cdot}$, CON^+, $C_2H_4N^+$	$[M-42]^{+\cdot}$ (C_3H_6)	nonspecific; *abundant:* propyl esters, butyl ketones, butylaromatics, methylcyclohexenes
		(C_2H_2O)	acetates (*especially* enol acetates), acetamides, cyclohexenones, α,β-unsaturated ketones
43	$C_3H_7^+$, $C_2H_3O^+$, $CONH^{+\cdot}$	$[M-43]^+$ (C_3H_7)	nonspecific; *abundant:* propyl, alicyclics, cycloalkanones, cycloalkylamines, cycloalkanols, butylaromatics
		(CH_3CO)	methyl ketones, acetates, aromatic methyl ethers

Mass	Ion	Product ion (and neutral particle lost)	Substructure or compound type
44	$CO_2^{+\cdot}$, $C_2H_6N^+$, $C_2H_4O^{+\cdot}$, $CS^{+\cdot}$, $C_3H_8^{+\cdot}$, $CH_4Si^{+\cdot}$	$[M-44]^{+\cdot}$ (C_3H_8)	propylalkanes
		(C_2H_6N)	*N,N*-dimethylamines, *N*-ethylamines
		(C_2H_4O)	cycloalkanols, cyclic ethers, ethylene ketals, aliphatic aldehydes (McLafferty rearrangement)
		(CO_2)	anhydrides, lactones, carboxylic acids
45	$C_2H_5O^+$, $C_2H_7N^{+\cdot}$, CHS^+ (with isotope signal for ^{34}S)	$[M-45]^+$ (C_2H_5O)	ethyl esters, ethyl ethers, lactones, ethyl sulfonates, ethyl sulfones
		(CHO_2)	carboxylic acids
		(C_2H_7N)	*N,N*-dimethylamines, *N*-ethylamines
			O indicator
			S indicator
46	$C_2H_5OH^{+\cdot}$, NO_2^+	$[M-46]^{+\cdot}$ (C_2H_6O)	ethyl esters, ethyl ethers, ethyl sulfonates
		($H_2O + C_2H_4$)	primary alcohols
		($H_2O + CO$)	carboxylic acids
		(NO_2)	nitro compounds
47	CH_3S^+, CCl^+, $C_2H_5OH_2^+$, $CH(OH)_2^+$, PO^+	$[M-47]^+$ (CH_3S)	methyl sulfides (with isotope signal for ^{34}S)
			2 × O indicator
			S indicator
			P indicator
48	$CH_3SH^{+\cdot}$, $SO^{+\cdot}$, $CHCl^{+\cdot}$	$[M-48]^{+\cdot}$ (CH_4S)	methyl sulfides
		(SO)	sulfoxides, sulfones, sulfonates (with isotope signal for ^{34}S)
49	CH_2Cl^+, $CH_3SH_2^+$ (with isotope signal for ^{34}S)	$[M-49]^+$ (CH_2Cl)	chloromethyl (with ^{37}Cl isotope signal)
50	$C_4H_2^{+\cdot}$, $CH_3Cl^{+\cdot}$, $CF_2^{+\cdot}$	$[M-50]^{+\cdot}$ (CF_2)	trifluoromethylaromatics, perfluoro-alicyclics
51	$C_4H_3^+$, CHF_2^+		

Mass	Ion	Product ion (and neutral particle lost)		Substructure or compound type
52	$C_4H_4^{+\cdot}$			
53	$C_4H_5^{+}$			
54	$C_4H_6^{+\cdot}$,	$[M-54]^{+\cdot}$	(C_4H_6)	cyclohexenes
	$C_2H_4CN^{+}$		(C_2H_4CN)	cyanoethyl
55	$C_4H_7^{+}$,	$[M-55]^{+}$	(C_4H_7)	nonspecific; *abundant:* alicyclics,
	$C_3H_3O^{+}$			butyl esters, *N*-butylamides
56	$C_4H_8^{+\cdot}$,	$[M-56]^{+\cdot}$	(C_4H_8)	butyl esters, *N*-butylamides, pentyl ketones, cyclohexenes, tetralins, pentylaromatics
	$C_3H_4O^{+\cdot}$		(C_3H_4O)	methylcyclohexenones, β-tetralones
57	$C_4H_9^{+}$,	$[M-57]^{+}$	(C_4H_9)	nonspecific
	$C_3H_5O^{+}$,		(C_3H_5O)	ethyl ketones
	$C_3H_2F^{+}$			
58	$C_3H_8N^{+}$,	$[M-58]^{+\cdot}$	(C_4H_{10})	alkanes
	$C_3H_6O^{+\cdot}$		(C_3H_6O)	α-methylalkanals, methyl ketones, isopropylidene glycols **N indicator** **O indicator**
59	$C_3H_7O^{+}$,	$[M-59]^{+}$	(C_3H_7O)	propyl esters, propyl ethers
	$C_2H_5NO^{+\cdot}$		$(C_2H_3O_2)$	methyl esters
			(C_3H_9N)	amines, amides **O indicator**
60	$C_2H_4O_2^{+\cdot}$,	$[M-60]^{+\cdot}$	(C_3H_8O)	propyl esters, propyl ethers
	$CH_2NO_2^{+}$,		$(C_2H_4O_2)$	acetates
	$C_2H_6NO^{+}$,		$(CH_3OH + CO)$	methyl esters **O indicator**
	$C_2H_4S^{+\cdot}$		(C_2H_4S)	
61	$C_2H_5O_2^{+}$,	$[M-61]^{+}$	$(C_2H_5O_2)$	glycols, ethylene ketals **2 × O indicator**
	$C_2H_5S^{+}$		(C_2H_5S)	ethyl sulfides (with ^{34}S isotope signal) **S indicator**
62	$C_2H_6O_2^{+\cdot}$,	$[M-62]^{+\cdot}$	$(C_2H_6O_2)$	methoxymethyl ethers, ethylene glycols, ethylene ketals
	$C_2H_3Cl^{+\cdot}$			
	$C_2H_6S^{+\cdot}$		(C_2H_6S)	ethyl sulfides (with ^{34}S isotope signal)
63	$C_5H_3^{+}$,	$[M-63]^{+}$	(C_2H_4Cl)	chloroethyl
	$C_2H_4Cl^{+}$,		$(CO + Cl)$	carboxylic acid chlorides (with ^{37}Cl isotope signal)
	$COCl^{+}$			
64	$C_5H_4^{+\cdot}$,	$[M-64]^{+\cdot}$	(SO_2)	sulfones, sulfonates
	$SO_2^{+\cdot}$, $S_2^{+\cdot}$		(S_2)	disulfides (with ^{34}S isotope signal)

Mass	Ion	Product ion (and neutral particle lost)	Substructure or compound type
65	$C_5H_5^+$, $H_2PO_2^+$	$[M\text{-}65]^+$ (S_2H) (SO_2H)	disulfides
66	$C_5H_6^{+\cdot}$	$[M\text{-}66]^{+\cdot}$ (C_5H_6)	cyclopentenes
	$S_2H_2^{+\cdot}$		disulfides (with ^{34}S isotope signal)
67	$C_5H_7^+$, $C_4H_3O^+$	$[M\text{-}67]^+$ (C_4H_3O)	furyl ketones
68	$C_5H_8^{+\cdot}$, $C_4H_4O^{+\cdot}$, $C_3H_6CN^+$	$[M\text{-}68]^{+\cdot}$ (C_5H_8) (C_4H_4O)	cyclohexenes, tetralins cyclohexenones, β-tetralones
69	$C_5H_9^+$, $C_4H_5O^+$, $C_3HO_2^+$	$[M\text{-}69]^+$ (C_5H_9)	alicyclics, alkenes
	CF_3^+	(CF_3)	trifluoromethyl
70	$C_5H_{10}^{+\cdot}$		alkanes, alkenes, alicyclics
	$C_4H_6O^{+\cdot}$		cycloalkanones
	$C_4H_8N^+$		pyrrolidines
71	$C_5H_{11}^+$		alkanes, larger alkyl groups
	$C_4H_7O^+$		alkanones, alkanals, tetrahydrofurans
72	$C_4H_8O^{+\cdot}$		alkanones, alkanals **O indicator**
	$C_4H_{10}N^+$		aliphatic amines **N indicator**
	$C_6^{+\cdot}$		perhalogenated benzenes
73	$C_4H_9O^+$		alcohols, ethers, esters **O indicator**
	$C_3H_5O_2^+$		carboxylic acids, esters, lactones
	$C_3H_9Si^+$		trimethylsilyl compounds
74	$C_4H_{10}O^{+\cdot}$		ethers
	$C_3H_6O_2^{+\cdot}$		methyl esters of carboxylic acids, α-methyl carboxylic acids
75	$C_3H_7O_2^+$		methyl acetals, glycols **2 × O indicator**
	$C_3H_7S^+$		sulfides, thiols (with ^{34}S isotope signal) **S indicator**
	$C_2H_7SiO^+$		trimethylsilyloxyl compounds
76	$C_6H_4^{+\cdot}$		aromatics
77	$C_6H_5^+$		aromatics
	$C_3H_6Cl^+$		chloro compounds (with ^{37}Cl isotope signal)

Mass	Ion	Compound type	
78	$C_6H_6^{+\cdot}$	aromatics	
	$C_5H_4N^+$	pyridines	
	$C_3H_7Cl^{+\cdot}$	chloro compounds (with ^{37}Cl isotope signal)	
79	$C_6H_7^+$	aromatics with H-containing substituents	
	$C_5H_5N^{+\cdot}$	pyridines, pyrroles	
	Br^+	bromo compounds (with ^{81}Br isotope signal)	
80	$C_6H_8^{+\cdot}$	cyclohexenes, polycyclic alicyclics	
	$C_5H_4O^{+\cdot}$	cyclopentenones	
	$HBr^{+\cdot}$	bromo compounds (with ^{81}Br isotope signal)	
	$C_5H_6N^+$	pyrroles, pyridines	
81	$C_6H_9^+$	cyclohexanes, cyclohexenyls, dienes	
	$C_5H_5O^+$	furans, pyrans	
	$^{81}Br^+$	bromo compounds (with ^{79}Br isotope signal)	
82	$C_6H_{10}^{+\cdot}$	cyclohexanes	
	$C_5H_6O^{+\cdot}$	cyclopentenones, dihydropyrans	
	$C_5H_8N^+$	tetrahydropyridines	
	$C_4H_6N_2^{+\cdot}$	pyrazoles, imidazoles	
	$CCl_2^{+\cdot}$	chloro compounds (with isotope signals at m/z 84 and 86)	
83	$C_6H_{11}^+$	alkenes, alicyclics, monosubstituted alkanes	
	$C_5H_7O^+$	cycloalkanones	
84	$C_5H_{10}N^+$	piperidines, *N*-methylpyrrolidines	
85	$C_6H_{13}^+$	alkanes	
	$C_5H_9O^+$	alkanones, alkanals, tetrahydropyrans, fatty acid derivatives	
	$CClF_2^+$	chlorofluoroalkanes (with ^{37}Cl isotope signal)	
86	$C_5H_{10}O^{+\cdot}$	alkanones, alkanals	
	$C_5H_{12}N^+$	aliphatic amines	**N indicator**
87	$C_5H_{11}O^+$	alcohols, ethers, esters	**O indicator**
	$C_4H_7O_2^+$	esters, carboxylic acids	
88	$C_4H_8O_2^{+\cdot}$	ethyl esters of carboxylic acids, α-methyl-methyl esters, α-C_2-carboxylic acids	
89	$C_4H_9O_2^+$	diols, glycol ethers	**2 × O indicator**
	$C_4H_9S^+$	sulfides (with ^{34}S isotope signal)	
90	$C_7H_6^{+\cdot}$	disubstituted aromatics	
91	$C_7H_7^+$	aromatics	
	$C_4H_8Cl^+$	alkyl chlorides (with ^{37}Cl isotope signal)	

Mass	Ion	Compound type
92	$C_7H_8^{+\cdot}$	alkylbenzenes
	$C_6H_6N^+$	alkylpyridines
93	$C_6H_5O^+$	phenols, phenol derivatives
	$C_6H_7N^{+\cdot}$	anilines
	CH_2Br^+	bromo compounds (with ^{81}Br isotope signal)
94	$C_6H_6O^{+\cdot}$	phenol esters, phenol ethers
	$C_5H_4NO^+$	pyrryl ketones, pyridone derivatives
95	$C_5H_3O_2^+$	furyl ketones
96	$C_7H_{12}^{+\cdot}$	alicyclics
97	$C_7H_{13}^+$	alicyclics, alkenes
	$C_6H_9O^+$	cycloalkanones
	$C_5H_5S^+$	alkylthiophenes (with ^{34}S isotope signal)
98	$C_6H_{12}N^+$	*N*-alkylpiperidines
99	$C_7H_{15}^+$	alkanes
	$C_6H_{11}O^+$	alkanones
	$C_5H_7O_2^+$	ethylene ketals
	$H_4PO_4^+$	alkyl phosphates
104	$C_8H_8^{+\cdot}$	tetralin derivatives, phenylethyl derivatives
	$C_7H_4O^{+\cdot}$	disubstituted α-ketobenzenes
105	$C_8H_9^+$	alkylaromatics
	$C_7H_5O^+$	benzoyl derivatives
	$C_6H_5N_2^+$	diazophenyl derivatives
106	$C_7H_8N^+$	alkylanilines
111	$C_5H_3OS^+$	thiophenoyl derivatives (with ^{34}S isotope signal)
115	$C_9H_7^+$	aromatics
	$C_6H_{11}O_2^+$	esters
	$C_5H_7O_3^+$	diesters
119	$C_9H_{11}^+$	alkylaromatics
	$C_8H_7O^+$	tolyl ketones
	$C_2F_5^+$	perfluoroethyl derivatives
	$C_7H_5NO^{+\cdot}$	phenyl carbamates
120	$C_7H_4O_2^{+\cdot}$	γ-benzopyrones, salicylic acid derivatives
	$C_8H_{10}N^+$	pyridines, anilines
121	$C_8H_9O^+$ and $C_7H_5O_2^+$	hydroxybenzene derivatives

Mass	Ion	Compound type
127	$C_{10}H_7^+$	naphthalenes
	$C_6H_7O_3^+$	unsaturated diesters
	$C_6H_6NCl^{+\cdot}$	chlorinated *N*-aromatics (with ^{37}Cl isotope signal)
	I^+	iodo compounds
128	$C_{10}H_8^{+\cdot}$	naphthalenes
	$C_6H_5OCl^{+\cdot}$	chlorinated hydroxybenzene derivatives (with ^{37}Cl isotope signal)
	$HI^{+\cdot}$	iodo compounds
130	$C_9H_8N^+$	quinolines, indoles
	$C_9H_6O^{+\cdot}$	naphthoquinones
131	$C_{10}H_{11}^+$	tetralins
	$C_5H_7S_2^+$	thioethylene ketals (with ^{34}S isotope signal)
	$C_3F_5^+$	perfluoroalkyl derivatives
135	$C_4H_8Br^+$	alkyl bromides (with ^{81}Br isotope signal at m/z 137)
141	$C_{11}H_9^+$	naphthalenes
142	$C_{10}H_8N^+$	quinolines
149	$C_8H_5O_3^+$	phthalates
152	$C_{12}H_8^{+\cdot}$	diphenyl aromatics
165	$C_{13}H_9^+$	diphenylmethane derivatives (fluorenyl cation)
167	$C_8H_7O_4^+$	phthalates
205	$C_{12}H_{13}O_3^+$	phthalates
223	$C_{12}H_{15}O_4^+$	phthalates

2.5.11 References

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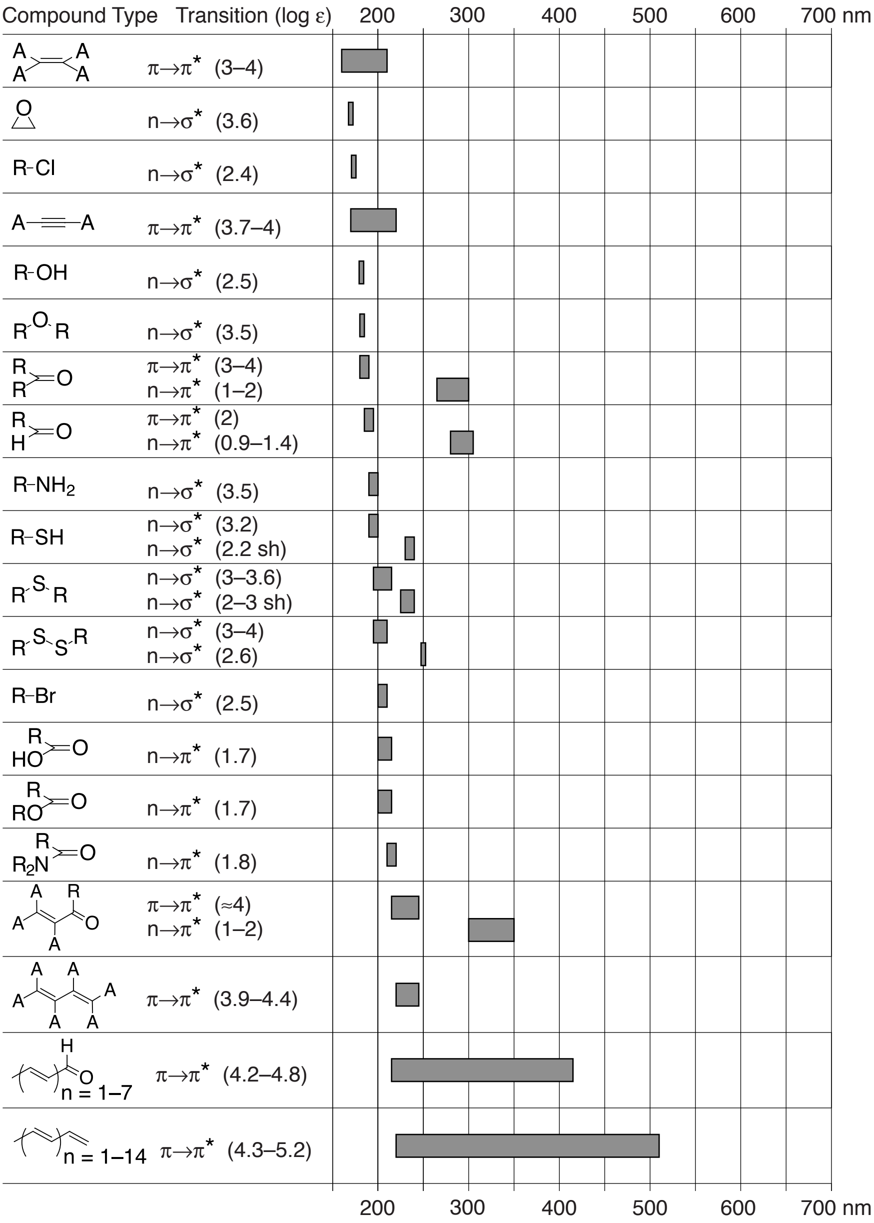
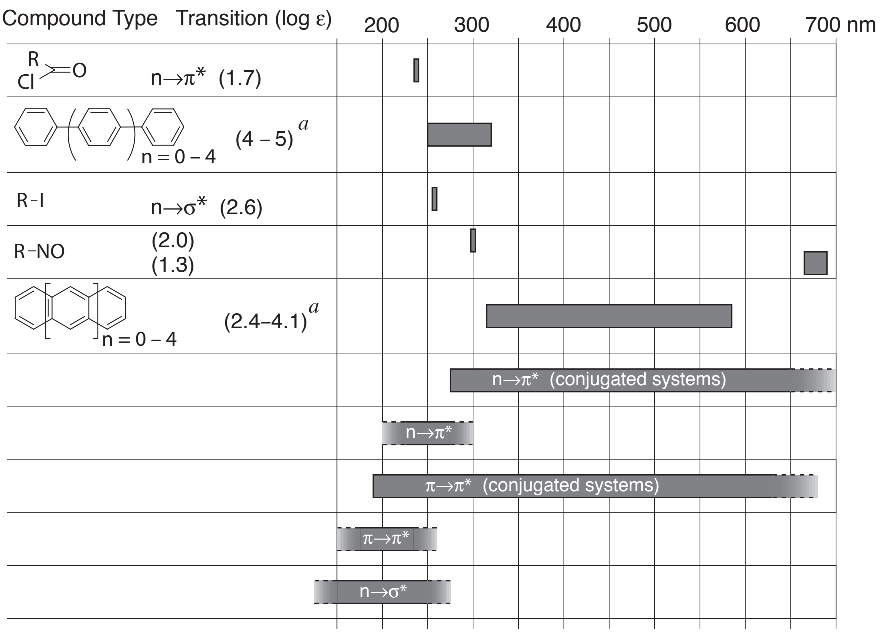
2.6 UV/Vis Spectroscopy

UV/Vis Absorption Bands of Various Compound Types (A: alkyl or H; R: alkyl; sh: shoulder)





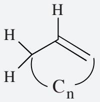
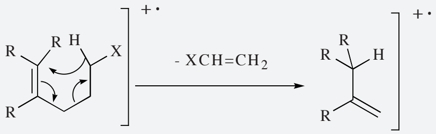
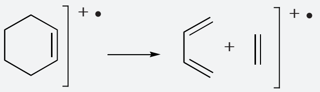
[a] longest wavelength absorption maximum

3 Combination Tables

3.1 Alkanes, Cycloalkanes

	Assignment	Range	Comments
^{13}C	CH_3 CH_2 CH C	5–35 ppm 5–45 ppm 25–60 ppm 30–60 ppm	CH_3, CH_2, CH, and C can be differentiated by multipulse experiments (DEPT, APT), off-resonance decoupling, 2D CH correlation spectra, or based on relaxation times Lower shift values in three-membered rings
^{1}H	CH_3 CH_2 CH	0.8–1.2 ppm 1.1–1.8 ppm 1.1–1.8 ppm	Lower shift values in three-membered rings
IR	CH st	3000–2840 cm^{-1}	Higher frequency in three-membered rings
	CH_3 δ as	≈1460 cm^{-1}	
	CH_2 δ	≈1460 cm^{-1}	
	CH_3 δ sy	≈1380 cm^{-1}	Doublet for *geminal* methyl groups
	CH_2 γ	770–720 cm^{-1}	In C–$(CH_2)_n$–C with n ≥ 4 at ca. 720 cm^{-1}
MS	Molecular ion	m/z 14n + 2	Weak in *n*-alkanes Very weak in isoalkanes
	Fragments		*n*-Alkanes: local maxima at 14n + 1, intensity variations: smooth, minimum at $[M-15]^+$ Isoalkanes: local maxima at 14n + 1, intensity distribution: irregular (relative maxima due to fragmentation at branching points with charge retention at the most highly substituted C)
	Rearrangements	 m/z 14n m/z 14n - 2	*n*-Alkanes: unspecific Isoalkanes: elimination of alkenes Monocycloalkanes: elimination of alkanes
UV			No absorption above 200 nm

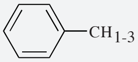
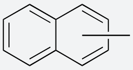
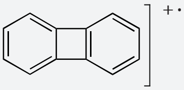
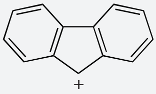
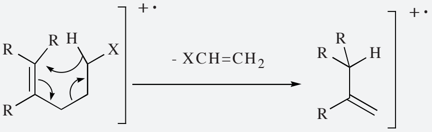
3.2 Alkenes, Cycloalkenes

	Assignment	Range	Comments
^{13}C	C=C C–(C=C)	100–150 ppm 10–60 ppm	Considerable differences between *Z* and *E*:
^{1}H	H–(C=C)	4.5–6.5 ppm	Coupling constants, $\|J_{gem}\|$ 0–3 Hz J_{cis} 5–12 Hz J_{trans} 12–18 Hz
	CH_3–(C=C) CH_2–(C=C)	≈1.7 ppm ≈2.0 ppm	Coupling constants, $^3J_{CH_3-CH=C} \approx 7$ Hz $^3J_{CH_2-CH=C} \approx 7$ Hz In rings, $\|J\|$ smaller:  n = 2, $^3J \approx 0.5$ Hz n = 3, $^3J \approx 1.5$ Hz n = 4, $^3J \approx 4.0$ Hz Long-range coupling, $^4J_{HC-C=CH}$ 0–2 Hz
IR	H–C(=C) st C=C st H–C(=C) δ oop CH_2–(C=C) δ	3100–3000 cm^{-1} 1690–1635 cm^{-1} 1000–675 cm^{-1} 1440 cm^{-1}	 Of variable intensity
MS	Molecular ion	m/z 14n m/z 14n - 2	Alkenes: moderate intensity Monocycloalkenes: medium intensity
	Fragments	m/z 14n - 1 m/z 14n - 3	Local maxima for alkenes Local maxima for monocyclic alkenes Usually, double bonds cannot be localized
	Rearrange- ments		*n*-Alkenes: unspecific except for:  Cyclohexenes: retro-Diels–Alder reaction: 
UV	C=C π→π* $(C=C)_2$ π→π*	< 210 nm (log ε 3–4) 215–280 nm (log ε 3.5–4.5)	Isolated double bonds; for highly substituted double bonds often absorption tail

3.3 Alkynes

	Assignment	Range	Comments
^{13}C	C≡C	65–85 ppm	Coupling constant $^{2}J_{HC\equiv^{13}C}$ ≈50 Hz; often leading to unexpected signs of signals in DEPT spectra and unexpected signals in 2D heteronuclear correlation spectra
	C–(C≡C)	0–30 ppm	
^{1}H	H–(C≡C)	1.5–3.0 ppm	Coupling constants, $\lvert{}^{4}J_{CH-C\equiv CH}\rvert$ ≈3 Hz $\lvert{}^{5}J_{CH-C\equiv C-CH}\rvert$ ≈3 Hz
	CH_3–(C≡C)	≈1.8 ppm	
	CH_2–(C≡C)	≈2.2 ppm	
	CH–(C≡C)	≈2.6 ppm	
IR	H–C(≡C) st	3340–3250 cm^{-1}	Sharp, intensive
	C≡C st	2260–2100 cm^{-1}	Sometimes very week
MS	Molecular ion		Weak, in the case of 1-alkynes up to C_7 often absent
	Fragments		$[M-1]^+$ often significant
	Rearrange-ments		Extensive rearrangements, not very character-istic
UV	C≡C π→π*	< 210 nm (log ε 3.7–4.0)	Isolated double bonds; for highly substituted double bonds often absorption tail

3.4 Aromatic Hydrocarbons

	Assignment	Range	Comments
^{13}C	ar C ar CH al C–C ar	120–150 ppm 110–130 ppm 10–60 ppm	Same ranges for polycyclic aromatic hydrocarbons
^{1}H	H–C ar	6.5–7.5 ppm	In polycyclic aromatic hydrocarbons up to ≈9 ppm Coupling constants, $^{3}J_{ortho}$ ≈7 Hz $^{4}J_{meta}$ ≈2 Hz $^{5}J_{para}$ <1 Hz
	CH_3–C ar CH_2–C ar CH–C ar	≈2.3 ppm ≈2.6 ppm ≈2.9 ppm	Often line broadening due to long-range coupling with aromatic protons
IR	ar C–H st	3080–3030 cm^{-1}	Often multiple bands, weak
	comb	2000–1650 cm^{-1}	Very weak
	ar C–C st	≈1600 cm^{-1} ≈1500 cm^{-1} ≈1450 cm^{-1}	Of variable intensity, sometimes not all bands observable
	ar C–H δ oop	960–650 cm^{-1}	Strong, frequently multiple bands
MS	Molecular ion		Strong, often base peak
	Fragments	m/z 39, 50–53, 63–65, 75–78, $[M-26]^{+\cdot}$, $[M-39]^{+}$	Often doubly charged fragment ions
		Benzylic cleavage	 m/z 91 (90, 92)
		Other typical fragments	 m/z 127
			 m/z 152
			 m/z 165
	Rearrangements		

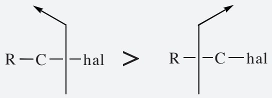
	Assignment	Range	Comments
UV		≈200–210 nm (log ε ≈4) ≈260 nm (log ε ≈2.4)	In benzene and alkylbenzenes

3.5 Heteroaromatic Compounds

	Assignment	Range	Comments
^{13}C	ar C–X	120–160 ppm	
	ar C–C	100–150 ppm	
^{1}H	H–C ar	6–9 ppm	Coupling constants in 6-membered rings similar to those in aromatic hydrocarbons; smaller in 5-membered rings
	H–N ar	7–14 ppm	Strongly solvent dependent, generally broad
IR	ar C–H st	3100–3000 cm^{-1}	Often multiple bands, weak
	ar N–H st	3500–2800 cm^{-1}	
	ar C–C st	≈1600 cm^{-1} ≈1500 cm^{-1} ≈1450 cm^{-1}	Often split, sometimes not all bands observable
	ar C–H δ oop	1000–650 cm^{-1}	Often strong, frequently multiple bands
MS	Molecular ion		Strong, often base peak
	Fragments	m/z 39, 50–53, 63–65, 75–78, $[M-26]^{+\bullet}$, $[M-39]^{+}$	Often doubly charged fragment ions
		m/z 45 $[CHS]^{+}$	*S*-Heteroaromatics
		Benzyl-analogous cleavage	
	Rearrangements		Loss of HCN (Δm 27, *N*-heteroaromatics) Loss of CO (Δm 28, *O*-heteroaromatics) Loss of CS (Δm 44, *S*-heteroaromatics)
UV			cf. UV/Vis Reference Spectra, Chapter 8.5.3

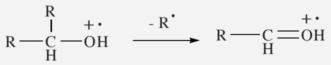
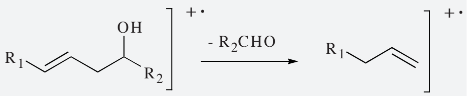
3.6 Halogen Compounds

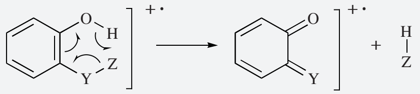
	Assignment	Range	Comments
13C	al C–F	70–100 ppm	CF_3: ≈115 ppm
	(C)=C–F	125–175 ppm	Coupling, with ^{19}F (isotope abundance, 100%; I = 1/2): $\|^1J_{CF}\|$ 100–300 Hz
	C=(C–F)	65–115 ppm	$\|^2J_{CF}\|$ 10–40 Hz
	ar C–F	140–165 ppm	$\|^3J_{CF}\|$ 5–10 Hz
	ar C–(C–F)	105–135 ppm	$\|^4J_{CF}\|$ 0–5 Hz
	al C–Cl	30–60 ppm	
	(C)=C–Cl	100–150 ppm	
	C=(C–Cl)	100–155 ppm	
	ar C–Cl	120–150 ppm	
	ar C–(C–Cl)	125–135 ppm	
	al C–Br	10–45 ppm	
	(C)=C–Br	90–140 ppm	
	C=(C–Br)	90–140 ppm	
	ar C–Br	110–140 ppm	
	ar C–(C–Br)	125–135 ppm	
	al C–I	-20 to +30 ppm	
	(C)=C–I	60–110 ppm	
	C=(C–I)	120–150 ppm	
	ar C–I	85–115 ppm	
	ar C–(C–I)	125–145 ppm	
1H	CH_2–F	≈4.3 ppm	Coupling, with ^{19}F (isotope abundance, 100%; I = 1/2): $\|^2J_{HF}\|$ 40–80 Hz; $\|^3J_{HF}\|$ 0–50 Hz; $\|^4J_{CF}\|$ 0–5 Hz
	CH_2–Cl	≈3.5 ppm	
	CH_2–Br	≈3.4 ppm	
	CH_2–I	≈3.1 ppm	
	H–CX=C	5.5–8.0 ppm	Similar shifts for all halogens
	H–C=CF	4.0–6.0 ppm	
	H–C=CCl	4.5–6.5 ppm	
	H–C=CBr	5.0–7.0 ppm	
	H–C=CI	5.5–7.5 ppm	
	H–phenyl–hal	7.0–7.6 ppm	Shielding by F in *ortho* and *para* positions; small effects for Cl and Br; deshielding by I in *ortho*, and shielding in *meta* position
IR	C–F st	1400–1000 cm^{-1}	Strong
	C–Cl st	850–600 cm^{-1}	Strong
	C–Br st	700–500 cm^{-1}	Strong
	C–I st	650–450 cm^{-1}	Strong

	Assignment	Range	Comments
MS	Molecular ion		Often weak for saturated aliphatic halogen compounds, often absent from spectra of aliphatic polyhalogenated compounds Characteristic isotope pattern for Cl and Br
	Fragments	m/z 69, 50–53	CF_3 Upon fragmentation of the C–hal bond, the positive charge preferably remains on the alkyl side, and on the halogen side upon fragmentation of the neighboring bond: 
	Rearrange-ments	$[M-20]^{+\bullet}$ $[M-50]^{+\bullet}$ or $[frag-50]^{+}$ $[M-36]^{+\bullet}$	HF elimination CF_2 elimination HCl elimination
UV	hal n → π*	≤280 nm (log ε ≈2.5)	For C–I; for C–Br and C–Cl in general only absorption tail, for C–F no absorption

3.7 Oxygen Compounds

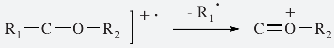
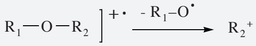
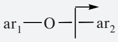
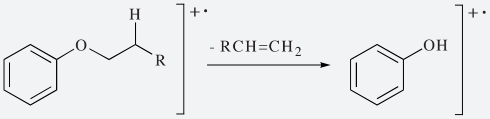
3.7.1 Alcohols and Phenols

	Assignment	Range	Comments
^{13}C	al C–OH	50–80 ppm	Shift with respect to C–H ≈50 ppm
	al C–(C–OH)	10–60 ppm	Hardly any shift with respect to C–(C–CH_3)
	al C–(C–C–OH)	10–60 ppm	Shift with respect to C–(C–C–CH_3) ≈-5 ppm
	ar C–OH	140–155 ppm	Shift with respect to C–H ≈+25 ppm
	ar C–(C–OH)	100–130 ppm	Shift with respect to C–(C–H): *ortho* ≈-13 ppm, *meta* ≈+1 ppm, *para* ≈-8 ppm
^{1}H	HO–C al HO–C ar	0.5–6 ppm 4–12 ppm	Often broad; position and shape strongly depend on experimental conditions
	CH_2–(OH) CH–(OH)	3.5–4.0 ppm 3.8–4.2 ppm	
	ar CH–(C–OH)	6.5–7.0 ppm	Shift with respect to CH–(C–H): *ortho* ≈-0.6 ppm, *meta* ≈-0.1 ppm, *para* ≈-0.5 ppm
IR	O–H st	3650–3200 cm^{-1}	Position and shape depend on the degree of association. Often different bands for H-bonded and free OH
	C–O(H) st	1260–970 cm^{-1}	Strong
MS	Molecular ion		Aliphatic: weak, often missing in the case of primary and highly branched alcohols; in this case, peaks at highest mass are often due to $[M-18]^{+\bullet}$ or $[M-15]^{+}$ Aromatic: strong
	Fragments	Aliphatic: m/z 31, 45, 59, … $[M-33]^{+}$	Primary: m/z 31 > m/z 45 ≈ m/z 59 Secondary, tertiary: local maxima due to α-cleavage: 
		Aromatic: $[M-28]^{+\bullet}$ (CO) $[M-29]^{+}$ (CHO)	CO and CHO elimination also from fragments. H_2O elimination ($[M-18]^{+\bullet}$) only with alkyl substituent in *ortho* position
	Rearrangements	Aliphatic: $[M-18]^{+\bullet}$ $[M-46]^{+\bullet}$	Elimination of H_2O from $M^{+\bullet}$ followed by alkene elimination; elimination of H_2O from products of α-cleavage
		Unsaturated	Vinylcarbinols: spectra similar to those of ketones Allyl alcohols: specific aldehyde elimination: 

	Assignment	Range	Comments
MS		Aromatic:	*Ortho* effect with appropriate substituents:  with Y–Z as –CO–OR, C–hal, –O–R, and similar
UV	Aliphatic		No absorption above 200 nm
	Aromatic	≈200–210 nm (log ε ≈3.8) ≈270 nm (log ε ≈2.4)	In alkaline solution, shift to longer wavelength and increase in intensity due to deprotonation

3.7.2 Ethers

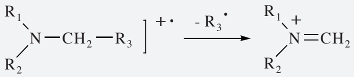
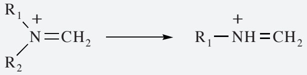
	Assignment	Range	Comments
^{13}C	al C–O	50–90 ppm	Oxiranes: outside the normal range
	al C–(C–O)	10–60 ppm	Hardly any shift with respect to C–(C–CH_3)
	al C–(C–C–O)	10–60 ppm	Shift with respect to C–(C–C–CH_3) ≈-5 ppm
	O–C–O	85–110 ppm	
	(C)=C–O	115–165 ppm	Shift with respect to (C)=C–C ≈+15 ppm
	C=(C–O)	70–120 ppm	Shift with respect to C=(C–C) ≈-30 ppm
	ar C–O	140–155 ppm	Shift with respect to ar C–H ≈+25 ppm
	ar C–(C–O)	100–130 ppm	Shift with respect to ar C–(C–H): *ortho* ≈-15 ppm; *meta* ≈+1 ppm; *para* ≈-8 ppm
^{1}H	CH_3–O	3.3–4.0 ppm	Singlet
	CH_2–O	3.4–4.2 ppm	
	O–CH_2–O	4.5–6.0 ppm	
	CH–O	3.5–4.3 ppm	
	$CH(O)_3$	≈ 5–6 ppm	
	H–C(O)=C	5.7–7.5 ppm	Shift with respect to H–C(H)=C ≈+1.2 ppm
	H–C=C–O	3.5–5.0 ppm	Shift with respect to H–C(=C–H) ≈-1 ppm
	ar CH–C–O	6.6–7.6 ppm	
IR	H–C(–O) st	2880–2815 cm^{-1}	For CH_3–O and CH_2–O; similar range for corresponding amines
	H–$CH(O)_2$ st	2880–2750 cm^{-1}	Two bands
	C–O–C st	1310–1000 cm^{-1}	Strong, sometimes two bands

	Assignment	Range	Comments
MS	Molecular ion		Aliphatic: weak, tendency to protonate Aromatic: strong
	Fragments	Aliphatic: m/z 31, 45, 59, … $[M\text{-}33]^+$	Base peak of aliphatic ethers generally due to fragmentation of the bond next to the ether bond:  or due to heterolytic cleavage of the C–O bond (especially for polyethers): 
		Alkyl aryl ethers	Preferential loss of the alkyl chain
		Diaryl ethers	Preferential loss of CO (Δm 28) from $M^{+\bullet}$ and/or $[M\text{-}H]^+$ as well as: 
	Rearrange-ments	Aliphatic: $[M\text{-}18]^{+\bullet}$ $[M\text{-}46]^{+\bullet}$	Elimination of water or alcohol
		Aromatic	Ethyl and higher alkyl ethers: alkene elimination to the phenol: 
UV	Aliphatic		No absorption above 200 nm
	Aromatic		Shift to higher wavelength and increase in intensity due to the ether group

3.8 Nitrogen Compounds

3.8.1 Amines

	Assignment	Range	Comments
^{13}C	al C–N	25–70 ppm	Shift with respect to C–H ≈+20 ppm
	al C–(C–N)	10–60 ppm	Shift with respect to C–(C–CH_3) ≈+2 ppm
	al C–(C–C–N)	10–60 ppm	Shift with respect to C–(C–C–CH_3) ≈-2 ppm
	(C)=C–N	120–170 ppm	Shift with respect to (C)=C–C ≈+20 ppm
	C=(C–N)	75–125 ppm	Shift with respect to C=(C–C) ≈-25 ppm
	ar C–N	130–150 ppm	Shift with respect to C–H ≈+20 ppm
	ar C–(C–N)	100–130 ppm	Shift with respect to C–(C–H): *ortho* ≈-15 ppm, *meta* ≈+1 ppm, *para* ≈-10 ppm
^1H	HN–C al	0.5–4.0 ppm	
	HN–C ar	2.5–5.0 ppm	
	HN^+–C al or ar	6.0–9.0 ppm	Often broad
	CH_3–N	2.3–3.1 ppm	Singlet
	CH_2–N	2.5–3.5 ppm	
	CH–N	3.0–3.7 ppm	
	CH–N^+	3.2–4.0 ppm	
	ar CH–C–N	6.0–7.5 ppm	Shift with respect to CH–(C–H): *ortho* ≈-0.8 ppm, *meta* ≈-0.2 ppm, *para* ≈-0.7 ppm
	ar CH–C–N^+	7.5–8.0 ppm	Shift with respect to CH–(C–H): *ortho* ≈+0.7 ppm, *meta* ≈+0.4 ppm, *para* ≈+0.3 ppm
IR	N–H st	3500–3200 cm^{-1}	Position and shape depend on the degree of association. Often different bands for H-bonded and free NH. For NH_2, always at least two bands
	N^+–H st	3000–2000 cm^{-1}	Broad, similar to COOH but more structured
	N–H δ	1650–1550 cm^{-1}	Weak or absent
	N^+–H δ	1600–1460 cm^{-1}	Often weak
	H–C(–N) st	2850–2750 cm^{-1}	For CH_3–N and CH_2–N in amines; similar range for corresponding ethers

	Assignment	Range	Comments
MS	Molecular ion		Odd nominal mass number for odd number of N atoms Aliphatic: weak, tendency to protonate, $[M+H]^+$ is often important Aromatic: strong, no tendency to protonate
	Fragments	Aliphatic: m/z 30, 44, 58, …	Base peak of aliphatic amines generally due to fragmentation of the bond next to the amine bond: 
	Rearrange-ments		Elimination of alkenes following amine cleavage: 
UV	Aliphatic		No absorption above 200 nm
	Aromatic		In acidic solutions, shift to lower wavelength and decrease in intensity

3.8.2 Nitro Compounds

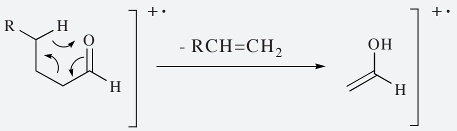
	Assignment	Range	Comments
^{13}C	al $C–NO_2$	55–110 ppm	Shift with respect to C–H ≈+50 ppm
	al $C–(C–NO_2)$	10–50 ppm	Shift with respect to C–(C–C) ≈-6 ppm
	al $C–(C–CNO_2)$	10–60 ppm	Shift with respect to C–(C–C–C) ≈-2 ppm
	ar $C–NO_2$	130–150 ppm	Shift with respect to C–H ≈+20 ppm
	ar $C–(C–NO_2)$	120–140 ppm	Shift with respect to C–(C–H): *ortho* ≈-5 ppm, *meta* ≈+1 ppm, *para* ≈+6 ppm
^{1}H	al $CH–NO_2$	4.2–4.6 ppm	
	ar $CH–C–NO_2$	7.5–8.5 ppm	Shift with respect to CH–(C–H): *ortho* ≈+1 ppm, *meta* ≈+0.3 ppm, *para* ≈+0.4 ppm
IR	NO_2 st as	1660–1490 cm^{-1}	Strong to very strong
	NO_2 st sy	1390–1260 cm^{-1}	Strong to very strong
MS	Molecular ion		Odd nominal mass number for odd number of N atoms Aliphatic: weak or absent Aromatic: strong
	Fragments	$[M-16]^{+\cdot}$, $[M-46]^+$	
	Rearrange-ments	m/z 30, $[M-17]^+$, $[M-30]^+$, $[M-47]^{+\cdot}$	
UV	Aliphatic	≈275 nm (log ε <2)	
	Aromatic	≈350 nm (log ε ≈2)	

3.9 Thiols and Sulfides

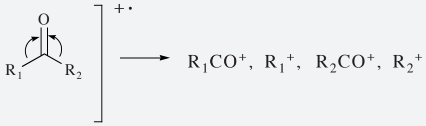
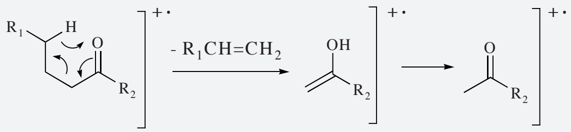
	Assignment	Range	Comments
^{13}C	al C–S ar C–S	5–60 ppm 120–140 ppm	No significant shift with respect to C–C
^{1}H	HS–C al HS–C ar al CH–S ar CH–S	1.0–2.0 ppm 2.0–4.0 ppm 2.0–3.2 ppm 7.0–7.5 ppm	*Vicinal* coupling constant, J, 5–9 Hz
IR	S–H st	2600–2540 cm^{-1}	Frequently weak
MS	Molecular ion		^{34}S-isotope peak at $[M+2]^{+\bullet}$ ≈4.5% Aliphatic: intensity higher than for corresponding alcohols and ethers
	Fragments	m/z 47, 61, 75, …	Sulfide cleavage: $[R_1{-}S{-}CH_2{-}R_2]^{+\bullet} \xrightarrow{-R_2^{\bullet}} R_1{-}\overset{+}{S}{=}CH_2$
	Rearrangements	m/z 34, 35, 48 $[M-33]^{+}$ $[M-34]^{+\bullet}$	 Alkene elimination after sulfide cleavage
UV	Aliphatic	<225 nm (log ε 3–4) 220–250 nm (log ε 2–3)	

3.10 Carbonyl Compounds

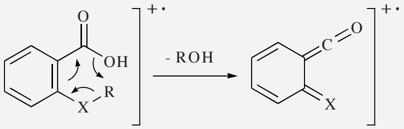
3.10.1 Aldehydes

	Assignment	Range	Comments
^{13}C	CHO	190–205 ppm	Coupling constant $^{1}J_{CH}$ 172 Hz
	al C–(CHO)	30–70 ppm	Coupling constant $\vert ^{2}J_{CH}\vert$ 20–50 Hz
	al C–(C–CHO)	5–50 ppm	Shift with respect to C–(C–CH_3) ≈-10 ppm
	(C)=C–(CHO)	110–160 ppm	
	C=(C–CHO)	110–160 ppm	
	ar C–(CHO)	120–150 ppm	
^{1}H	H–(C=O)	9.0–10.5 ppm	
	al CH–(CHO)	2.0–2.5 ppm	$^{3}J_{HH}$ 0–3 Hz
	(CH)=CH(CHO)	5.5–7.0 ppm	$^{3}J_{HH}$ ≈8 Hz
	CH=(CH–CHO)	5.5–7.0 ppm	
	ar CH–(C–CHO)	7.2–8.0 ppm	Shift with respect to CH–(C–H): *ortho* ≈+0.6 ppm, *meta* ≈+0.2 ppm, *para* ≈+0.3 ppm
IR	comb	2900–2700 cm^{-1}	Two weak bands
	C=O	1765–1645 cm^{-1}	Aliphatic: ≈1730 cm^{-1} Conjugated: ≈1690 cm^{-1}
MS	Molecular ion		Aliphatic: moderate Aromatic: strong
	Fragments	$[M-1]^{+}$	For aliphatic aldehydes, only significant up to C_7
		$[M-29]^{+}$	
	Rearrange-ments	m/z 44 $[M-44]^{+\bullet}$	Aliphatic aldehydes 
UV	n → π*	270–310 nm (log ε ≈1)	Saturated aldehydes
		≥207 nm (log ε ≈4)	α,β-Unsaturated aldehydes
		≥250 nm (log ε >3)	Aromatic aldehydes

3.10.2 Ketones

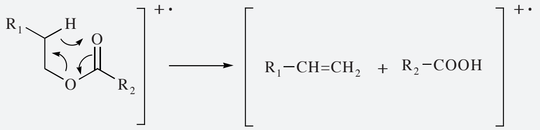
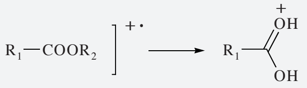
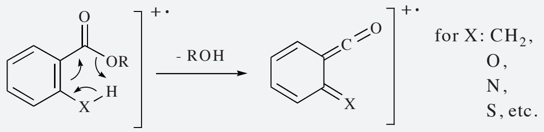
	Assignment	Range	Comments
^{13}C	C=O	195–220 ppm	
	al C–(C=O)	25–70 ppm	
	al C–(C–C=O)	5–50 ppm	Shift with respect to C–(C–CH_3) ≈-6 ppm
	(C)=C–(C=O)	105–160 ppm	
	C=(C–C=O)	105–160 ppm	
	ar C–(C=O)	120–150 ppm	
^{1}H	al CH–(C=O)	2.0–3.6 ppm	al CH–C(=O)–C al 2.0–2.6 ppm al CH–C(=O)–C ar 2.5–3.6 ppm
	CH=CH–(C=O)	5.5–7.0 ppm	
	ar CH–(C–C=O)	7.2–8.0 ppm	Shift with respect to CH–(C–H): *ortho* ≈+0.6 ppm *meta* ≈+0.1 ppm *para* ≈+0.2 ppm
IR	C=O st	1775–1650 cm^{-1}	Aliphatic: ≈1715 cm^{-1} Cyclic: ring size ≥6: ≈1715 cm^{-1} ring size <6: ≥1750 cm^{-1} Conjugated: ≈1690–1665 cm^{-1}
MS	Molecular ion		Aliphatic: moderate Aromatic: strong
	Fragments		Ketone cleavages: 
	Rearrange- ments	m/z 44 $[M-44]^{+\cdot}$	Aliphatic ketones
			
UV	$\pi \rightarrow \pi^*$	<200 nm (log ε 3–4)	Saturated ketones
	$n \rightarrow \pi^*$	250–300 nm (log ε 1–2)	Saturated ketones
		≥215 nm (log ε ≈4)	α,β-Unsaturated ketones
		≥245 nm (log ε >3)	Aromatic ketones

3.10.3 Carboxylic Acids

	Assignment	Range	Comments
13C	COOH	170–185 ppm	For COO^-, shift with respect to COOH: 0 to +8 ppm
	al C–(COOH)	25–70 ppm	
	al C–(C–COOH)	5–50 ppm	Shift with respect to C–(C–CH_3) ≈-6 ppm
	(C)=C–(COOH)	105–160 ppm	
	C=(C–COOH)	105–160 ppm	
	ar C–(COOH)	120–150 ppm	
1H	COOH	10.0–13.0 ppm	Position and shape strongly depend on experimental conditions
	al CH–(COOH)	2.0–2.6 ppm	
	CH=CH–(COOH)	5.2–7.5 ppm	
	ar CH–(C–COOH)	7.2–8.0 ppm	Shift with respect to CH–(C–H): *ortho* ≈+0.8 ppm, *meta* ≈+0.2 ppm, *para* ≈+0.3 ppm
IR	COO–H st	3550–2500 cm^{-1}	Broad
	C=O st	1800–1650 cm^{-1}	Aliphatic: ≈1715 cm^{-1} Conjugated: ≈1695 cm^{-1} For COO^-, two bands: 1580 and 1420 cm^{-1}
	COO–H δ oop	≈920 cm^{-1}	For dimers
MS	Molecular ion		Aliphatic: moderate, strong for long chains, tendency to protonate Aromatic: strong
	Fragments	$[M-17]^+$ $[M-45]^+$	Strong for aromatic acids
	Rearrange-ments	m/z 60, 61 $[M-18]^{+\cdot}$	Aliphatic acids Aromatic acids *Ortho* effect with aromatic acids:  for X: CH_2, O, N, S, etc.
UV	n → π*	<220 nm (log ε 1–2) ≥193 nm (log ε ≈4) ≥230 nm (log ε >3)	Saturated acids α,β-Unsaturated acids Aromatic acids

3.10.4 Esters and Lactones

	Assignment	Range	Comments
13C	COOR	165–180 ppm	Shift with respect to COOH -5 to -10 ppm
	al C–(COOR)	20–70 ppm	
	al C–(OCOR)	50–100 ppm	Shift with respect to C–(OH) +2 to +10 ppm
	(C)=C–(COOR)	105–160 ppm	
	C=(C–COOR)	105–160 ppm	
	(C)=C–(OCOR)	100–150 ppm	
	C=(C–OCOR)	80–130 ppm	
	ar C–(COOR)	120–150 ppm	
	ar C–(OCOR)	130–160 ppm	
	ar C=(C–OCOR)	105–130 ppm	
1H	al CH–COOR	2.0–2.5 ppm	CH_3COOR ≈2.0 ppm CH_2COOR ≈2.3 ppm CHCOOR ≈2.5 ppm
	al CH–OCOR	3.5–5.3 ppm	CH_3OCOR ≈3.5–3.9 ppm CH_2COOR ≈4.0–4.5 ppm CHCOOR ≈4.8–5.3 ppm
	CH=CH–COOR	5.5–8.0 ppm	Shift with respect to CH=CH–H: *gem* ≈+0.8 ppm, *cis* ≈+1.1 ppm *trans* ≈+0.5 ppm
	CH=CH–OCOR	6.0–8.0 ppm	Shift with respect to CH=CH–H: *gem* ≈+2.1 ppm, *cis* ≈-0.4 ppm *trans* ≈-0.6 ppm
	ar CH–C–COOR	7.0–8.0 ppm	Shift with respect to CH–(C–H): *ortho* ≈+0.7 ppm, *meta* ≈+0.1 ppm, *para* ≈+0.2 ppm
	ar CH–C–OCOR	6.8–7.5 ppm	Shift with respect to CH–(C–H): *ortho* ≈-0.2 ppm, *meta* ≈0 ppm, *para* ≈-0.1 ppm
IR	C=O st	1745–1730 cm^{-1}	Strong; range for aliphatic esters Higher wavenumbers for hal–C–COOR, COO–C=C, COO–C ar, and for small ring lactones Lower wavenumbers for C=C–COOR and ar C–COOR
	C–O st	1330–1050 cm^{-1}	Mostly two bands, at least one of them strong For COO^-, two bands: 1580 and 1420 cm^{-1}

	Assignment	Range	Comments
MS	Molecular ion		Aliphatic esters: weak, tendency to protonate Aliphatic lactones: medium to weak, tendency to protonate Aromatic esters and lactones: strong
	Fragments	$[M - RO]^+$ $[M - ROCO]^+$	Esters Esters Lactones: loss of α-substituents (attached to ether carbon), decarbonylation, for aromatic lactones also double decarbonylation
	Rearrangements		Alkene elimination from the alcohol moiety:  Elimination of the alcohol side chain with double H transfer (for alcohols with $C_{n>2}$):  Alcohol elimination from *ortho*-substituted aromatic esters: 
		$[M-18]^{+\cdot}$	Lactones
UV	n → π*	<220 nm (log ε 1–2) ≥193 nm (log ε ≈4) ≥230 nm (log ε >3)	Aliphatic esters α,β-Unsaturated esters Aromatic esters

3.10.5 Amides and Lactams

	Assignment	Range	Comments
^{13}C	$CONR_2$	165–180 ppm	
	al C–($CONR_2$)	20–70 ppm	
	al C–(C–$CONR_2$)	5–50 ppm	Shift with respect to C–(C–CH_3) ≈-6 ppm
	al C–(NCOR)	25–80 ppm	Shift with respect to C–(NH) ≈-1 to -2 ppm
	C=C–($CONR_2$)	105–160 ppm	
	ar C–($CONR_2$)	120–150 ppm	
	ar C–(NCOR)	110–150 ppm	
^{1}H	CONH	5–10 ppm	Frequently broad to very broad; splitting due to H–N–C–H coupling often recognizable only in the CH signal
	al CH–$CONR_2$	2.0–2.5 ppm	
	al CH–NCOR	2.7–4.8 ppm	CH_3NCOR ≈2.7–3.0 ppm CH_2NCOR ≈3.1–3.5 ppm CHNCOR ≈3.8–4.8 ppm
	CH=CH–$CONR_2$	5.2–7.5 ppm	Shift with respect to CH=CH–(H): *gem* ≈+1.4 ppm, *cis* ≈+1.0 ppm *trans* ≈+0.5 ppm
	C=CH–NCOR CH=C–NCOR	6.0–8.0 ppm 4.5–6.0 ppm	Shift with respect to CH=CH–(H): *gem* ≈+2.1 ppm, *cis* ≈-0.6 ppm *trans* ≈-0.7 ppm
	ar CH–C($CONR_2$)	7.5–8.5 ppm	Shift with respect to CH–C–(H): *ortho* ≈+0.6 ppm, *meta* ≈+0.1 ppm, *para* ≈+0.2 ppm
	ar CH–C(NCOR)	6.8–7.5 ppm	Shift with respect to CH–C–(H): *ortho* ≈0 ppm, *meta* ≈0 ppm, *para* ≈-0.2 ppm
IR	N–H st	3500–3100 cm^{-1}	Position and shape depend on the extent of association, often different bands for H-bonded and free NH, always at least two bands for NH_2
	C=O st (amide I)	1700–1650 cm^{-1}	Strong; range for amides as well as for δ- and larger lactams, higher wavenumbers for β- and γ-lactams
	N–H δ and N–C=O st sy (amide II)	1630–1510 cm^{-1}	Often strong, missing in the case of tertiary amides and lactams

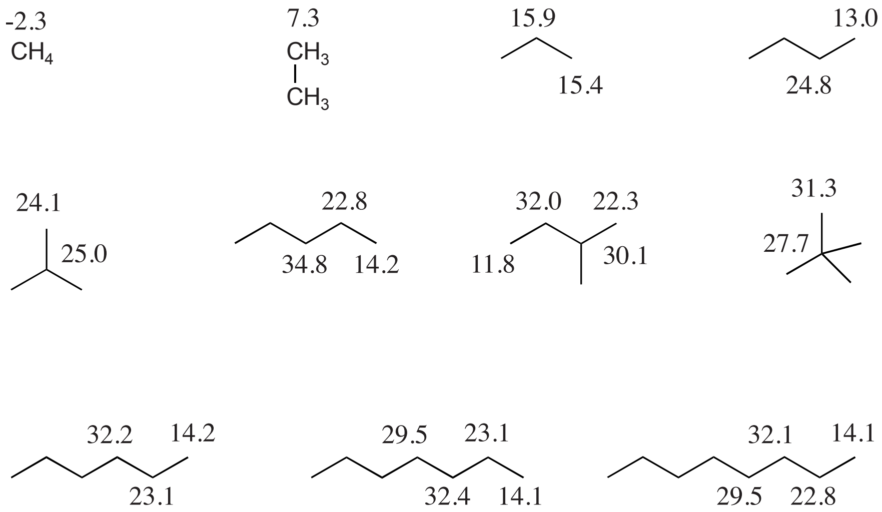
	Assignment	Range	Comments
MS	Molecular ion		Aliphatic amides: moderate, tendency to protonate Aromatic amides: strong
	Fragments		Amides: cleavage on both sides of the carbonyl group followed by loss of CO; large number of fragments of even mass Lactams: loss of α-substituent, loss of CO
	Rearrange-ments		Amides: elimination of the amine moiety, elimination of alkene from the amine or acid moiety in analogy to esters
		$[M\text{-}18]^{+\bullet}$	Lactams
UV	n → π*	<220 nm (log ε 1–2)	Aliphatic amides and lactams

4 ^{13}C NMR Spectroscopy

4.1 Alkanes

4.1.1 Chemical Shifts

^{13}C Chemical Shifts (δ in ppm)



^{13}C Chemical Shifts of Methyl Groups (δ in ppm)

	Substituent R	δ_{CH_3-R}
	–H	-2.3
C	$-CH_3$	7.3
	$-CH_2CH_3$	15.4
	$-CH(CH_3)_2$	24.1
	$-C(CH_3)_3$	31.3
	$-(CH_2)_6CH_3$	14.1
	$-CH_2$–phenyl	15.7
	$-CH_2F$	15.8
	$-CH_2Cl$	18.7
	$-CH_2Br$	19.1
	$-CH_2I$	20.4
	$-CHCl_2$	31.6
	$-CHBr_2$	31.8
	$-CCl_3$	46.3
	$-CBr_3$	49.4
	$-CH_2OH$	18.2
	$-CH_2OCH_3$	14.7
	$-CH_2OCH_2CH_3$	15.4
	$-CH_2OCH{=}CH_2$	14.6
	$-CH_2O$–phenyl	14.9
	$-CH_2OCOCH_3$	14.4
	$-CH_2NH_2$	19.0
	$-CH_2NHCH_3$	14.3
	$-CH_2N(CH_3)_2$	12.8
	$-CH_2NO_2$	12.3
	$-CH_2SH$	19.7
	$-CH_2S(O)_2CH_3$	6.7
	$-CH_2S(O)_2OH$	8.0
	$-CH_2CHO$	5.2
	$-CH_2COCH_3$	7.0
	$-CH_2COOH$	9.6
	–cyclopentyl	20.5
	–cyclohexyl	23.1
	$-CH{=}CH_2$	18.7
	$-C{\equiv}CH$	3.7
	–phenyl	21.4
	–1-naphthyl	19.1
	–2-naphthyl	21.5

	Substituent R	δ_{CH_3-R}
C	–2-pyridyl	24.2
	–3-pyridyl	18.0
	–4-pyridyl	20.6
	–2-furyl	13.7
	–2-thienyl	14.7
	–2-pyrrolyl	11.8
	–2-indolyl	13.4
	–3-indolyl	9.8
	–4-indolyl	21.6
	–5-indolyl	21.5
	–6-indolyl	21.7
	–7-indolyl	16.6
X	–F	71.6
	–Cl	25.6
	–Br	9.6
	–I	-24.0
O	–OH	50.2
	$-OCH_3$	60.9
	$-OCH_2CH_3$	57.6
	$-OCH(CH_3)_2$	54.9
	$-OC(CH_3)_3$	49.4
	$-OCH_2CH{=}CH_2$	57.4
	–O–cyclohexyl	55.1
	$-OCH{=}CH_2$	52.5
	–O–phenyl	54.8
	$-OCOCH_3$	51.5
	–OCO–cyclohexyl	51.2
	$-OCOCH{=}CH_2$	51.5
	–OCO–phenyl	51.8
	$-OCOOCH_3$	54.9
	$-OS(O)_2$–4-tolyl	56.3
	$-OS(O)_2OCH_3$	59.1
	$-OP(OCH_3)_2$	48.8
N	$-NH_2$	28.3
	$-NH_3^+$	26.5
	$-NHCH_3$	38.2
	–NH–cyclohexyl	33.5
	–NH–phenyl	30.2

	Substituent R	$\delta_{CH_3\text{–}R}$
N	$–N(CH_3)_2$	47.5
	–*N*-pyrrolidinyl	42.7
	–*N*-piperidinyl	47.7
	$–N(CH_3)$phenyl	39.9
	–*N*-pyrrolyl	35.9
	–*N*-imidazolyl	32.2
	–*N*-pyrazolyl	38.4
	–*N*-indolyl	32.1
	$–NHCOCH_3$	26.1
	$–N(CH_3)CHO$	31.5, 36.5
	$–N(CH_3)COCH_3$	35.0, 38.0
	$–N(CH_3)P[N(CH_3)_2]_2$	33.9
	$–NO_2$	61.2
	–C≡N	1.7
	–NC	26.8
	–NCS	29.1
S	–SH	6.5
	$–SCH_3$	19.3
	$–S–n\text{-}C_8H_{17}$	15.5
	–S–phenyl	15.6
	$–SSCH_3$	22.0
	$–S(O)CH_3$	40.1
	$–S(O)_2CH_3$	42.6
	$–S(O)_2CH_2CH_3$	39.3
	$–S(O)_2Cl$	52.6
	$–S(O)_2OH$	39.6
	$–S(O)_2ONa$	41.1
O=C	–CHO	31.2
	$–COCH_3$	30.7
	$–COCH_2CH_3$	27.5
	$–COCCl_3$	21.1
	$–COCH{=}CH_2$	25.7
	–CO–cyclohexyl	27.6
	–CO–phenyl	25.7
	–COOH	21.7
	$–COO^-$	24.4
	$–COOCH_3$	20.6
	$–COOCOCH_3$	21.8
	$–CONH_2$	22.3
	$–CON(CH_3)_2$	21.5
	–COSH	32.6
	$–COSCH_3$	30.2
	$–COCOCH_3$	23.2
	–COCl	33.6
	–COBr	39.1
	$–COSi(CH_3)_3$	35.7
M	–Li	-16.6
	$–B(CH_3)_2$	14.8
	$–B^-(CH_3)_3\ Li^+$	6.2
	$–Si(CH_3)_2CH{=}CH_2$	-2.0
	$–SiCl_3$	9.8
	$–Ge(CH_3)_3$	-3.6
	$–Sn(CH_3)_3$	-9.3
	$–Pb(CH_3)_3$	-4.2
	$–P(CH_3)(n\text{-}C_4H_9)$	14.4
	$–P^+(CH_3)_3\ I^-$	10.7
	$–As(CH_3)_2$	11.2
	$–As^+(CH_3)_3\ I^-$	8.4
	$–In(CH_3)_2$	-6.3

^{13}C Chemical Shifts of Monosubstituted Alkanes (δ in ppm)

	Substituent	Methyl	Ethyl		1-Propyl		
		$-CH_3$	$-CH_2$	$-CH_3$	$-CH_2$	$-CH_2$	$-CH_3$
	–H	-2.3	7.3	7.3	15.4	15.9	15.4
C	$-CH=CH_2$	18.7	27.4	13.4	36.2	22.4	13.6
	–C≡CH	3.7	12.3	13.8	20.6	22.2	13.4
	–phenyl	21.4	29.1	15.8	38.3	24.8	13.8
X	–F	71.6	80.1	15.8	85.2	23.6	9.2
	–Cl	25.6	39.9	18.9	46.8	26.3	11.6
	–Br	9.6	27.6	19.4	35.6	26.4	13.0
	–I	-24.0	-1.6	20.6	9.1	27.0	15.3
O	–OH	50.2	57.8	18.2	64.2	25.9	10.3
	$-OCH_3$	60.9	67.7	14.7	74.5	23.2	10.5
	$-OCH_2CH_3$	57.6	66.0	15.4	72.5	23.2	10.7
	$-OCH(CH_3)_2$	54.9					
	$-OC(CH_3)_3$	49.4	56.8	16.4			
	–O–phenyl	54.8	63.2	14.9	69.4	22.8	10.6
	$-OCOCH_3$	51.5	60.4	14.4	66.2	22.4	10.5
	–OCO–phenyl	51.8	60.8	14.4	66.4	22.2	10.5
	$-OS(O)_2$–4-tolyl	56.3	66.9	14.7	72.2	22.3	10.0
N	$-NH_2$	28.3	36.9	19.0	44.6	27.4	11.5
	$-NHCH_3$	38.2	45.9	14.3	54.0	23.2	12.5
	$-N(CH_3)_2$	47.6	53.6	12.8	61.8	20.6	11.9
	$-NHCOCH_3$	26.1	34.4	14.6	40.7	22.5	11.1
	$-NO_2$	61.2	70.8	12.3	77.4	21.2	10.8
	–C≡N	1.7	10.8	10.6	19.3	19.0	13.3
	–NC	26.8	36.4	15.3	43.4	22.9	11.0
S	–SH	6.5	19.1	19.7	26.4	27.6	12.6
	$-SCH_3$	19.3					
	$-SSCH_3$	22.0	31.8	14.7			
	$-S(O)CH_3$	40.1					
	$-S(O)_2CH_3$	42.6	48.2	6.7	56.3	16.3	13.0
	$-S(O)_2Cl$	52.6	60.2	9.1	67.1	18.4	12.1
	$-S(O)_2OH$	39.6	46.7	8.0	53.7	18.8	13.7
C=O	–CHO	31.3	36.7	5.2	45.7	15.7	13.3
	$-COCH_3$	30.7	35.2	7.0	45.2	17.5	13.5
	–CO–phenyl	25.7	31.7	8.3	40.4	17.7	13.8
	–COOH	21.7	28.5	9.6	36.2	18.7	13.7
	$-COOCH_3$	20.6	27.2	9.2	35.6	18.9	13.8
	$-CONH_2$	22.3	29.0	9.7			
	–COCl	33.6	41.0	9.3	48.9	18.8	13.0

^{13}C Chemical Shifts of Monosubstituted Alkanes (δ in ppm, contd.)

		2-Propyl		*tert*-Butyl	
		–CH	$–CH_3$	–C	$–CH_3$
	–H	15.9	15.4	25.0	24.1
C	$–CH{=}CH_2$	32.3	22.1	33.8	29.4
	–C≡CH	20.3	22.8	27.4	31.1
	–phenyl	34.3	24.0	34.6	31.4
X	–F	87.3	22.6	93.5	28.3
	–Cl	53.7	27.3	66.7	34.6
	–Br	44.8	28.5	62.1	36.4
	–I	20.9	31.2	43.0	40.4
O	–OH	64.0	25.3	68.9	31.2
	$–OCH_3$	72.6	21.4	72.7	27.0
	$–OCH_2CH_3$			72.6	27.7
	$–OCH(CH_3)_2$	68.5	23.0	73.0	28.5
	$–OC(CH_3)_3$	63.5	25.2	76.3	33.8
	–O–phenyl	69.3	22.0		
	$–OCOCH_3$	67.5	21.9	79.9	28.1
	–OCO–phenyl	68.2	21.9	80.7	28.2
N	$–NH_2$	43.0	26.5	47.2	32.9
	$–NHCH_3$	50.5	22.5	50.4	28.2
	$–N(CH_3)_2$	55.5	18.7	53.6	25.4
	$–NHCOCH_3$	40.5	22.3	49.9	28.6
	$–NO_2$	78.8	20.8	85.2	26.9
	–C≡N	19.8	19.9	28.1	28.5
	–NC	45.5	23.4	54.0	30.7
S	–SH	29.9	27.4	41.1	35.0
	$–SCH_2CH_3$	34.4	23.4		
	$–S(O)_2CH_3$	53.5	15.2	57.6	22.7
	$–S(O)_2Cl$	67.6	17.1	74.2	24.5
	$–S(O)_2OH$	52.9	16.8	55.9	25.0
O=C	–CHO	41.1	15.5	42.4	23.4
	$–COCH_3$	41.6	18.2	44.3	26.5
	–CO–phenyl	35.2	19.1	43.5	27.9
	–COOH	34.1	18.8	38.7	27.1
	$–COOCH_3$	34.1	19.1	38.7	27.3
	$–CONH_2$	34.9	19.5	38.6	27.6
	–COCl	46.5	19.0	49.4	27.1

^{13}C Chemical Shifts of 1-Substituted *n*-Octanes (δ in ppm)

	Substituent	1 $-CH_2$	2 $-CH_2$	3 $-CH_2$	4 $-CH_2$	5 $-CH_2$	6 $-CH_2$	7 $-CH_2$	8 $-CH_3$
	–H	14.1	22.8	32.1	29.5	29.5	32.1	22.8	14.1
C	$-CH{=}CH_2$	34.5	~29.6	~29.6	~29.6	~29.6	32.2	23.0	13.9
	–phenyl	36.2	31.7	~29.6	~29.6	~29.6	32.1	22.8	14.1
X	–F	84.2	30.6	25.3	29.3	29.3	31.9	22.7	14.1
	–Cl	45.1	32.8	27.0	29.0	29.2	31.9	22.8	14.1
	–Br	33.8	33.0	28.3	28.8	29.2	31.8	22.7	14.1
	–I	6.9	33.7	30.6	28.6	29.1	31.8	22.6	14.1
O	–OH	63.1	32.9	25.9	29.5	29.4	31.9	22.8	14.1
	–O–*n*-C_8H_{17}	71.1	30.0	26.3	29.6	29.4	32.0	22.8	14.1
	–O–phenyl	68.0	26.2	29.3	29.4	29.4	31.9	22.7	14.1
	–OCO–*n*-propyl	64.4	28.8	26.1	29.3	29.3	31.9	22.8	14.1
	–OCO–phenyl	65.1	28.8	26.1	29.3	29.3	31.9	22.7	14.1
	–ONO	68.3	29.2	26.0	29.3	29.3	31.9	22.7	14.0
N	$-NH_2$	42.4	34.1	27.0	29.6	29.4	31.9	22.7	14.1
	$-N(CH_3)_2$	60.1	29.5*	≈27.9*	≈27.7*	29.7*	32.0	22.8	14.4
	$-N^+(CH_3)_3\ Cl^-$	66.6	26.2	23.2	29.1*	29.0*	31.6	22.5	14.0
	$-NO_2$	75.8	26.2	27.9	≈29.6	≈29.6	31.4	22.6	14.0
	–C≡N	17.2	25.5	≈29.9	≈29.9	≈29.9	31.8	22.7	14.0
S	–SH	24.7	34.2	28.5	29.2	29.1	31.9	22.7	14.1
	$-SCH_3$	34.5	29.0	29.4	29.4	29.4	31.9	22.8	14.1
	–S(O)–*n*-C_8H_{17}	52.6	≈29.1	≈29.1	≈29.1	≈29.1	31.8	22.7	14.1
C=O	–CHO	44.0	22.2	≈29.3	≈29.3	≈29.3	31.9	22.7	14.1
	$-COCH_3$	43.7	24.1	≈29.5	≈29.5	≈29.5	32.0	22.8	14.1
	–CO–phenyl	38.6	24.4	29.5	29.5	29.5	31.9	22.7	14.0
	–COOH	34.2	24.8	≈29.3	≈29.3	≈29.3	31.9	22.7	14.1
	$-COOCH_3$	34.2	25.1	29.3	29.3	29.3	31.9	22.8	14.1
	$-CONH_2$	35.5	25.4	29.1	29.1	29.1	31.6	22.3	14.0
	–COCl	47.2	25.1	28.5	29.1	29.1	31.8	22.7	14.1
Si	$-Si(OCH_3)_3$	9.2	22.7	33.2	29.3	29.3	32.0	22.7	14.1

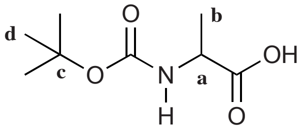
* Assignment uncertain

Estimation of ^{13}C Chemical Shifts of Aliphatic Compounds (δ in ppm)

The chemical shifts of *sp*^{3}-hybridized carbon atoms can be estimated with the help of an additivity rule using the shift value of methane (-2.3 ppm) and increments (Z) for substituents in α, β, γ, and δ position (see next pages). Some substituents occupy two positions. Thus, the quaternary carbon atom **c** in the example given below is in δ position relative to the carbon atom **a** since the *sp*^{3}-hybridized oxygen of the β-COO group occupies the γ position. This simple linear model needs corrections in case of strong branching of the observed C atom and/or its neighbors (steric corrections, S). Substituents for which such corrections are necessary are those with varying branching, i.e., a varying number of directly bonded H atoms. They are marked with an asterisk (*) in the Table of Increments (next page). Further correction terms are needed if γ substituents are in a sterically fixed position (conformational corrections, K).

The chemical shifts estimated with this additivity rule, in general, differ by less than ca. 4 ppm from the experimental values. Larger discrepancies may be expected for highly branched systems (particularly for quaternary carbon atoms). For carbon atoms bearing several halogen, oxygen, and/or other strongly deshielding substituents, additional correction terms are needed [1]. Without such corrections, deviations can be so large as to render the rule useless.

Example: Estimation of chemical shifts for *N*-(*tert*-butoxycarbonyl)alanine



a	base value	-2.3
	1 α-C	9.1
	1 α-COOH	20.1
	1 α-NH	28.3
	1 β-COO	2.0
	1 δ-C	0.3
	1 S(tert,2)	-3.7
	estimated	53.8
	exp	49.0

b	base value	-2.3
	1 α-C	9.1
	1 β-COOH	2.0
	1 β-NH	11.3
	1 γ-COO	-2.8
	1 S(prim,3)	-1.1
	estimated	16.2
	exp	17.3

c	base value	-2.3
	3 α-C	27.3
	1 α-OCO	56.5
	1 γ-NH	-5.1
	1 δ-C	0.3
	3 S(quat,1)	-4.5
	estimated	72.2
	exp	78.1

d	base value	-2.3
	1 α-C	9.1
	2 β-C	18.8
	1 β-OCO	6.5
	1 δ-NH	0.0
	1 S(prim,4)	-3.4
	estimated	28.7
	exp	28.1

Estimation of ^{13}C Chemical Shifts of Aliphatic Compounds (δ in ppm)

$$\delta = -2.3 + \sum_i Z_i + \sum_j S_j + \sum_k K_k$$

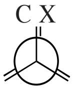
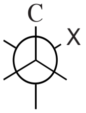
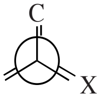
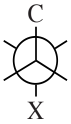
	Substituent	Increment Z_i for substituents in position α	β	γ	δ
	–H	0.0	0.0	0.0	0.0
C	–C*≤	9.1	9.4	-2.5	0.3
	–C*=C<	19.5	6.9	-2.1	0.4
	–C≡C–	4.4	5.6	-3.4	-0.6
	–phenyl	22.1	9.3	-2.6	0.3
X	–F	70.1	7.8	-6.8	0.0
	–Cl	31.0	10.0	-5.1	-0.5
	–Br	18.9	11.0	-3.8	-0.7
	–I	–7.2	10.9	-1.5	-0.9
O	–O–*	49.0	10.1	-6.2	0.3
	–OCO–	56.5	6.5	-6.0	0.0
	–ONO	54.3	6.1	-6.5	-0.5
N	–N*<	28.3	11.3	-5.1	0.0
	–$N^{+,*}$≤	30.7	5.4	-7.2	-1.4
	–NH_3^+	26.0	7.5	-4.6	0.0
	–NO_2	61.6	3.1	-4.6	-1.0
	–C≡N	3.1	2.4	-3.3	-0.5
	–NC	31.5	7.6	-3.0	0.0
S	–S*–	10.6	11.4	-3.6	-0.4
	–SCO–	17.0	6.5	-3.1	0.0
	–S*(O)–	31.1	7.0	-3.5	0.5
	–S*$(O)_2$–	30.3	7.0	-3.7	0.3
	–S$(O)_2$Cl	54.5	3.4	-3.0	0.0
	–SCN	23.0	9.7	-3.0	0.0
O=C	–CHO	29.9	-0.6	-2.7	0.0
	–CO–	22.5	3.0	-3.0	0.0
	–COOH	20.1	2.0	-2.8	0.0
	–COO^-	24.5	3.5	-2.5	0.0
	–COO–	22.6	2.0	-2.8	0.0
	–CO–N<	22.0	2.6	-3.2	-0.4
	–COCl	33.1	2.3	-3.6	0.0
	–C=NOH syn	11.7	0.6	-1.8	0.0
	–C=NOH anti	16.1	4.3	-1.5	0.0
	–CS–N<	33.1	7.7	-2.5	0.6
	–Sn	-5.2	4.0	-0.3	0.0

Steric Corrections, S_j

Observed ^{13}C center	S for number of substituents at the α atom[a]			
	1	2	3	4
primary (CH_3)	0.0	0.0	-1.1	-3.4
secondary (CH_2)	0.0	0.0	-2.5	-6.0
tertiary (CH)	0.0	-3.7	-8.5	-10.0
quaternary (C)	-1.5	-8.0	-10.0	-12.5

[a] To be applied to each of the neighboring atoms that has an unspecified number of non-hydrogen substituents (marked with an asterisk (*) in the Table of Increments, Z_i).

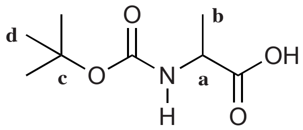
Conformational Corrections, K_k, for γ Substituents

Conformation		K
synperiplanar (eclipsed)		-4.0
synclinal (gauche)		-1.0
anticlinal		0.0
antiperiplanar (anti)		2.0
not fixed		0.0

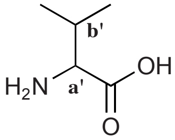
One can also use the chemical shifts of a reference compound as the base value if its structure is closely related to that assumed for the unknown. The increments corresponding to the structural elements missing in the reference compound are then added to the base value, while those of structural elements present in the reference but absent in the unknown are subtracted (see example on next page).

Example: Estimation of the chemical shifts for the carbon atoms **a** and **b** in *N*-(*tert*-butoxycarbonyl)alanine using the chemical shifts of valine as base values (**a'**, **b'**):

Target:



Reference:



a		
	base value (**a'**)	61.9
	1 β-COO	2.0
	1 δ-C	0.3
	1 S(tert,2)	-3.7
	- 2 β-C	-18.8
	- 1 S(tert,3)	8.5
	estimated	50.2
	exp	49.0

b		
	base value (**b'**)	30.3
	1 γ-COO	-2.8
	1 S(prim,3)	-1.1
	- 2 α-C	-18.2
	- 1 S(tert,3)	8.5
	estimated	16.6
	exp	17.3

4.1.2 Coupling Constants

^{13}C-^{1}H Coupling Constants

Coupling through one bond ($^{1}J_{CH}$ in Hz)

The ^{13}C-^{1}H coupling constant of 125 Hz in methane increases in the presence of electronegative substituents and can be estimated by using the following additivity rule:

$$J_{CHZ_1Z_2Z_3} = 125.0 + \sum_i Z_i$$

Substituent	Increment Z_i
–H	0.0
–CH_3	1.0
–$C(CH_3)_3$	-3.0
–CH_2Cl	3.0
–CH_2Br	3.0
–CH_2I	7.0
–$CHCl_2$	6.0
–CCl_3	9.0
–C≡C	7.0
–phenyl	1.0
–F	24.0
–Cl	27.0

Substituent	Increment Z_i
–Br	27.0
–I	26.0
–OH	18.0
–O–phenyl	18.0
–NH_2	8.0
–$NHCH_3$	7.0
–$N(CH_3)_2$	6.0
–C≡N	11.0
–$S(O)CH_3$	13.0
–CHO	2.0
–$COCH_3$	-1.0
–COOH	5.5

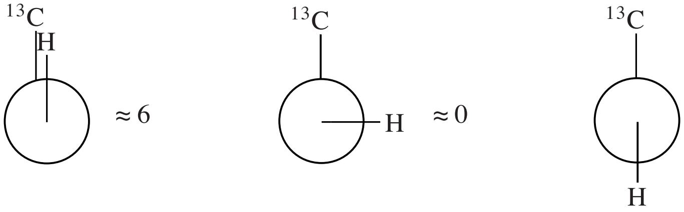
Example: Estimation of ^{13}C-^{1}H coupling constant of $CHCl_3$:
J = 125.0 + 3 × 27.0 = 206.0 Hz (exp: 209.0 Hz).

Coupling through more than one bond (|J_{CH}| in Hz)

The coupling constants can be estimated from the corresponding ^{1}H-^{1}H coupling constants [2]: $J_{CH} \approx 0.62\ J_{HH}$

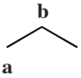
Typical values:		Examples:	
$^2J_{CH}$	1– 6	$^1H–CH_2–^{13}CH_3$	4.5
$^3J_{CH}$	0–10	$^1H–CH_2–CH_2–^{13}CH_3$	5.8

The ^{13}C-^{1}H coupling constants for coupling across three bonds depend on the dihedral angle in the same way as the vicinal ^{1}H-^{1}H coupling constants (see Chapter 5.1.2):

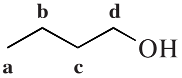


^{13}C-^{13}C Coupling Constants (|J_{CC}| in Hz)

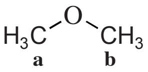
H_3C-CH_3 ^1J 34.6



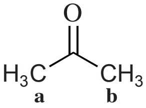
$^1J_{ab}$ 34.6



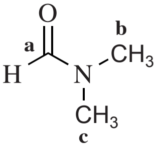
$^2J_{ac}$ 4.6
$^3J_{ad}$ 4.6
$^2J_{bd}$ <1



$^2J_{ab}$ 2.4



$^2J_{ab}$ 16.1



$^2J_{ab}$ 0.5
$^2J_{ac}$ 4.9

The ^{13}C-^{13}C coupling constants for coupling over three bonds depend on the dihedral angle in the same way as the vicinal ^{1}H-^{1}H (see Chapter 5.1.2) and ^{13}C-^{1}H coupling constants. Maximum values of ca. 4–6 Hz are observed for dihedral angles of 0° and 180° and minimal values around 0 Hz at 90°.

4.1.3 References

[1] A. Fürst, E. Pretsch, W. Robien, A comprehensive parameter set for the prediction of the ^{13}C NMR chemical shifts of *sp*^{3}-hybridized carbon atoms in organic compounds, *Anal. Chim. Acta* **1990**, *233*, 213.

[2] J.L. Marshall, Carbon-Carbon and Carbon-Proton NMR couplings, Verlag Chemie International, Deerfield Beach, FL, 1983.

4.2 Alkenes

4.2.1 Chemical Shifts

^{13}C Chemical Shifts (δ in ppm)

The ^{13}C chemical shifts of the carbons of C=C double bonds typically range from ca. 80–160 ppm; a wider range of 40–210 ppm is observed with O and N substituents. In unsaturated *acyclic hydrocarbons*, they can be predicted with high accuracy (see below). To estimate the ^{13}C chemical shifts in all other *substituted alkenes*, one can use the substituent effects listed for chemical shifts in vinyl groups. However, since no configuration-dependent parameters are available, the values thus estimated are less accurate than those for unsaturated acyclic hydrocarbons.

The ^{13}C chemical shifts of sp^3-hybridized carbon atoms in the vicinity of double bonds can be estimated using the additivity rule given in Chapter 4.1.1. The conformational correction factors, K, for γ substituents of *cis*- vs. *trans*-disubstituted alkenes differ by 6 ppm because the relative position of these substituents is fixed by the double bond.

Estimation of the ^{13}C Chemical Shifts of sp^2-Hybridized Carbon Atoms in Unsaturated Acyclic Hydrocarbons (δ in ppm)

C–C–C'=**C**–C–C–C
γ' β' α' ↑ α β γ
Base value: 123.3

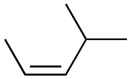
Increments for C substituents:

at C atom under consideration (**C**)		at neighboring C atom (C')	
α	10.6	α'	-7.9
β	4.9	β'	-1.8
γ	-1.5	γ'	1.5

Steric corrections:

• for each pair of *cis*-α,α'-substituents	-1.1
• for a pair of geminal α,α-substituents	-4.8
• for a pair of geminal α',α'-substituents	2.5
• if one or more β-substituents are present	2.3

Example: Estimation of chemical shifts of *cis*-4-methyl-2-pentene



a **b**

a		
	base value	123.3
	1 α-C	10.6
	1 α'-C	-7.9
	2 β'-C	-3.6
	cis-α,α'	-1.1
	estimated	121.3
	exp	121.8

b		
	base value	123.3
	1 α-C	10.6
	2 β-C	9.8
	1 α'-C	-7.9
	cis-α,α'	-1.1
	1 β-substituent	2.3
	estimated	137.0
	exp	138.8

Effect of Substituents on the ^{13}C Chemical Shifts of Vinyl Compounds (δ in ppm)

1 **2**

$R-CH=CH_2$

$$\delta_{C_i} = 123.3 + Z_i$$

	Substituent R	Z_1	Z_2
	–H	0.0	0.0
C	$-CH_3$	12.9	-7.4
	$-CH_2CH_3$	17.2	-9.8
	$-CH_2CH_2CH_3$	15.7	-8.8
	$-CH(CH_3)_2$	22.7	-12.0
	$-(CH_2)_3-$	14.6	-8.9
	$-C(CH_3)_3$	26.0	-14.8
	$-CH_2Cl$	10.2	-6.0
	$-CH_2Br$	10.9	-4.5
	$-CH_2I$	14.2	-4.0
	$-CH_2OH$	14.2	-8.4
	$-CH_2OCH_2CH_3$	12.3	-8.8
	$-CH=CH_2$	13.6	-7.0
	–C≡CH	-6.0	5.9
	–phenyl	12.5	-11.0
X	–F	24.9	-34.3
	–Cl	2.8	-6.1
	–Br	-8.6	-0.9
	–I	-38.1	7.0

	Substituent R	Z_1	Z_2
O	–OH	25.7	-35.3
	$-OCH_3$	29.4	-38.9
	$-OCH_2CH_3$	28.8	-37.1
	$-O(CH_2)_3CH_3$	28.1	-40.4
	$-OCOCH_3$	18.4	-26.7
N	$-N(CH_3)_2$	28.0*	-32.0*
	$-N^+(CH_3)_3$	19.8	-10.6
	–*N*-pyrrolidonyl	6.5	-29.2
	$-NO_2$	22.3	-0.9
	–C≡N	-15.1	14.2
	–NC	-3.9	-2.7
S	$-SCH_2CH_3$	9.0	-12.8
	$-S(O)_2CH=CH_2$	14.3	7.9
O=C	–CHO	15.3	14.5
	$-COCH_3$	13.8	4.7
	–COOH	5.0	9.8
	$-COOCH_2CH_3$	6.3	7.0
	–COCl	8.1	14.0
Si	$-Si(CH_3)_3$	16.9	6.7
	$-SiCl_3$	8.7	16.1

* Estimated values

The values listed on the preceding page can also be used to estimate the ^{13}C chemical shifts of *sp*^{2}-hybridized carbon atoms in alkenes with more than one substituent (note that the *cis*/*trans* configuration is not taken into account):

$$\delta_{C_i} = 123.3 + \sum Z_i$$

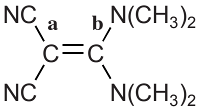
Example: Estimation of chemical shifts of 1-bromo-1-propene

a **b**
Br–CH=CH–CH$_3$

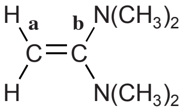
a			
	base value	123.3	
	Z_1(Br)	-8.6	
	$Z_2(CH_3)$	-7.4	
	estimated	107.3	
	exp	108.9	(*cis*)
		104.7	(*trans*)

b			
	base value	123.3	
	Z_2(Br)	-0.9	
	$Z_1(CH_3)$	12.9	
	estimated	135.3	
	exp	129.4	(*cis*)
		132.7	(*trans*)

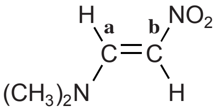
The following examples show some larger deviations between measured and estimated (in parentheses) chemical shifts. This is usually to be expected when several substituents are present that strongly interact with the π electrons of the double bond:



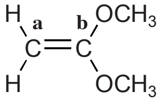
a 39.1 (29.1)
b 171.0 (207.7)



a 69.2 (59.3)
b 163.0 (179.3)

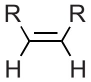
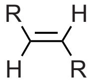


a 151.0 (150.4)
b 111.4 (113.6)



a 54.7 (45.5)
b 167.9 (182.1)

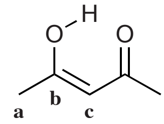
^{13}C Chemical Shifts of *cis*- and *trans*-1,2-Disubstituted Alkenes (δ in ppm)

Substituent R		
$-CH_3$	123.3	124.5
$-CH_2CH_3$	131.2	131.3
–Cl	118.1	119.9
–Br	116.4	109.4
–I	96.5	79.4
–C≡N	120.8	120.2
$-OCH_3$	130.3	135.2
–COOH	130.4	134.2
$-COOCH_3$	130.1	133.5

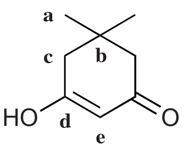
^{13}C *Chemical Shifts of Enols (δ in ppm)*

The carbon atom bonded to the enolic OH group is strongly deshielded so that its shift is close to that of a carbonyl carbon. The other carbon atom of the double bond is strongly shielded.

Enol:

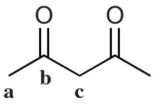


a	22.5
b	190.5
c	99.0

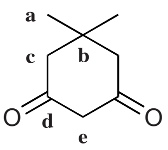


a	28.3
b	32.8
c	46.2
d	191.1
e	103.3

Ketone:

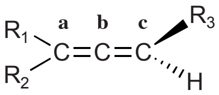


a	28.5
b	201.1
c	56.6



a	28.3
b	31.0
c	54.2
d	203.6
e	57.3

^{13}C Chemical Shifts of Allenes (δ in ppm)



R_1	R_2	R_3	**a**	**b**	**c**
–H	–H	–H	74.8	213.5	74.8
$–CH_3$	–H	–H	84.4	210.4	74.1
$–CH_3$	$–CH_3$	–H	93.4	207.3	72.1
$–CH_3$	–H	$–CH_3$	85.4	207.1	85.4
$–CH_2CH_3$	–H	–H	91.7	208.9	75.3
$–C(CH_3)_3$	$–C(CH_3)_3$	–H	119.6	207.0	75.8
$–CH{=}CH_2$	–H	–H	93.9	211.4	75.1
–C≡CH	–H	–H	74.8	217.7	77.3
–phenyl	–H	–H	94.4	210.0	78.8
–F	–H	–H	129.8	200.2	93.9
–Cl	–H	–H	88.8	207.9	84.5
–Br	–H	–H	72.7	207.6	83.8
–I	–H	–H	35.3	208.0	78.3
$–OCH_3$	–H	–H	123.1	202.0	90.3
$–N(CH_3)_2$	–H	–H	113.1	204.2	85.5
–C≡N	–H	–H	67.4	218.7	80.7
$–SCH_3$	–H	–H	90.0	206.1	81.3
–COOH	–H	–H	88.1	217.7	80.0

4.2.2 Coupling Constants

^{13}C-^{1}H Coupling Constants (|J_{CH}| in Hz)

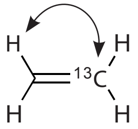
Coupling through one bond

C=C

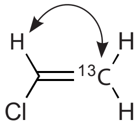
$CH_2{=}CH_2$ $^1J_{CH}$ 156.4

$CH_2{=}C{=}CH_2$ $^1J_{CH}$ 167.8

Coupling through two bonds



$^2J_{CH}$ -2.4

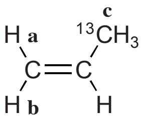


$^2J_{CH}$ 6.9

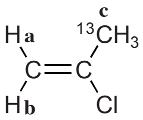
Additivity rule for the estimation of $^2J_{CH}$ of alkenes: see [2].

Coupling through three bonds

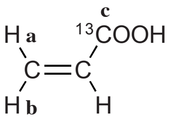
The *trans*-^{1}H–C=C–^{13}C coupling constant of alkenes is always larger than the corresponding *cis* coupling constant so that an assignment is possible if both isomers are available: see [3].



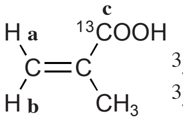
$^3J_{ac}$ 7.6
$^3J_{bc}$ 12.6



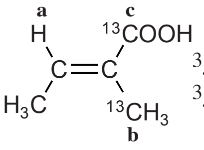
$^3J_{ac}$ 4.1
$^3J_{bc}$ 8.1



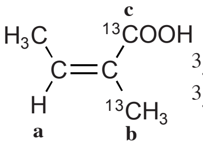
$^3J_{ac}$ 7.6
$^3J_{bc}$ 14.1



$^3J_{ac}$ 7.6
$^3J_{bc}$ 14.1



$^3J_{ab}$ 7.7
$^3J_{ac}$ 7.4



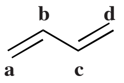
$^3J_{ab}$ 6.9
$^3J_{ac}$ 13.2

^{13}C-^{13}C Coupling Constants (|J_{CC}| in Hz)

$CH_2{=}CH_2$ $^1J_{CC}$ 67.6

a b c
$CH_2{=}CH{-}CH_3$ $^1J_{ab}$ 70.0
$^1J_{bc}$ 41.9

$CH_2{=}C{=}CH_2$ $^1J_{CC}$ 98.7



$^1J_{ab}$ 68.8 $^2J_{ac}$ <1
$^1J_{bc}$ 53.7 $^3J_{ad}$ 9.0

4.2.3 References

[1] R.H.A.M. Janssen, R.J.J.Ch. Lousberg, M.J.A. de Bie, An additivity relation for carbon-13 chemical shifts in substituted allenes, *Recl. Trav. Chim. Pays-Bas* **1981**, *100*, 85.
[2] U. Vögeli, D. Herz, W. von Philipsborn, Geminal C,H spin coupling in substituted alkenes, *Org. Magn. Reson.* **1980**, *13*, 200.
[3] U. Vögeli, W. von Philipsborn, Vicinal C,H spin coupling in substituted alkenes. *Org. Magn. Reson.* **1975**, *7*, 617.

C≡C

4.3 Alkynes

4.3.1 Chemical Shifts

^{13}C Chemical Shifts of Alkynes (δ in ppm)

$$\overset{}{R}-\overset{a}{C}\equiv\overset{b}{C}-H$$

	Substituent R	a	b
	–H	71.9	71.9
C	$-CH_3$	80.4	68.3
	$-CH_2CH_3$	85.5	67.1
	$-CH_2CH_2CH_3$	84.0	68.7
	$-CH_2CH_2CH_2CH_3$	83.0	66.0
	$-CH(CH_3)_2$	89.2	67.6
	$-C(CH_3)_3$	92.6	66.8
	–cyclohexyl	88.7	68.3
	$-CH_2OH$	83.0	73.8
	$-CH=CH_2$	82.8	80.0
	$-C\equiv C-CH_3$	68.8	64.7
	–phenyl	84.6	78.3
O	$-OCH_2CH_3$	90.9	26.5
S	$-SCH_2CH_3$	72.6	81.4
O=C	–CHO	81.8	83.1
	$-COCH_3$	81.9	78.1
	–COOH	74.0	78.6
	$-COOCH_3$	74.8	75.6

Additivity rule for estimating the chemical shifts of *sp*-hybridized carbon atoms in alkynes: see [1].

4.3.2 Coupling Constants

^{13}C-^{1}H Coupling Constants ($|J_{CH}|$ in Hz) [2]

$$\overset{a}{H}-\overset{b}{C}\equiv\overset{c}{C}-H$$

$^{1}J_{ab}$ 249.0
$^{2}J_{ac}$ 49.3 (in substituted acetylenes: 40–60)

$$\overset{a}{H}-\overset{b}{C}\equiv\overset{c}{C}-\overset{d\,e}{CH_3}$$

$^{2}J_{ac}$ 50.1 $^{3}J_{ad}$ 3.4
$^{2}J_{ce}$ -10.4 $^{3}J_{be}$ 4.7

$$\overset{a}{H_3C}-\overset{b}{C}\equiv\overset{c}{C}-CH_3$$

$^{2}J_{ab}$ -10.3 $^{3}J_{ac}$ 4.3

C≡C

With acetylenes, the results of multipulse experiments (such as DEPT, INEPT, SEFT, or APT) to determine the number of protons attached to the carbon atoms must be interpreted with care. As a consequence of the unusually large ^{13}C-^{1}H coupling constants through one and two bonds, the sign of the signals may be opposite to the expected one. For the same reasons, unexpected signals may occur in two-dimensional heteronuclear correlation spectra (HSQC, HMBC).

^{13}C-^{13}C Coupling Constants ($|^{1}J_{CC}|$ in Hz)

H−C≡C−H	$^{1}J_{CC}$	171.5	H−C^{a}≡C^{b}−C^{c}≡C−H	$^{1}J_{ab}$	190.3
				$^{1}J_{bc}$	153.4

4.3.3 References

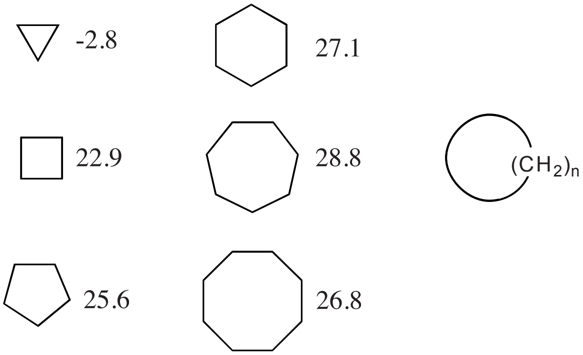
[1] W. Höbold, R. Radeglia, D. Klose, Inkrementen-Berechnung von ^{13}C-chemischen Verschiebungen in *n*-Alkinen, *J. Prakt. Chem.* **1976**, *318*, 519.

[2] K. Hayamizu, O. Yamamoto, ^{13}C,^{1}H Spin coupling constants of dimethylacetylene, *Org. Magn. Reson.* **1980**, *13*, 460.

4.4 Alicyclics

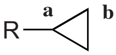
4.4.1 Chemical Shifts

Saturated Monocyclic Alicyclics (δ in ppm)



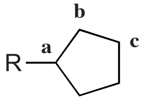
n	δ
9	26.0
10	25.1
11	26.3
12	23.8
13	26.2
14	25.2
15	27.0
20	28.0
30	29.3
40	29.4
72	29.7

^{13}C Chemical Shifts of Monosubstituted Cyclopropanes (δ in ppm)



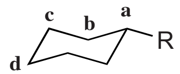
	Substituent R	a	b	other
	–H	-2.8	-2.8	
C	$–CH_3$	4.9	5.6	CH_3 19.4
	$–CH_2CH_3$	12.8	4.1	CH_2 27.8, CH_3 13.6
	$–CH_2CH_2CH_2CH_3$	10.9	4.4	1-CH_2 34.7, 2-CH_2 32.0
	$–C(CH_3)_3$	22.7	0.3	C 29.3, CH_3 28.2
	$–CH_2Cl$	13.6	5.5	CH_2 50.3
	$–CH_2OH$	12.7	2.2	CH_2 66.5
	$–CH{=}CH_2$	14.7	6.6	CH 142.4, CH_2 111.5
	–phenyl	15.3	9.2	C 143.9, CH 125.3–128.2
X	–Cl	27.3	8.9	
	–Br	14.2	9.1	
	–I	-20.1	10.4	
O	–OH	45.7	6.8	
N	$–NH_2$	24.0	7.4	
	$–NO_2$	54.3	11.7	
	–C≡N	-4.5	6.2	CN 121.5
O=C	–CHO	22.7	7.4	CO 202.1
	$–COCH_3$	20.1	9.6	CO 207.3, CH_3 29.1
	–CO–phenyl	17.1	11.5	
	–COOH	12.7	8.9	CO 181.6
	$–COOCH_3$	12.2	7.7	CO 174.7, CH_3 51.1

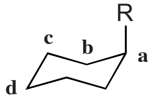
^{13}C Chemical Shifts of Monosubstituted Cyclopentanes (δ in ppm)



	Substituent R	a	b	c	other
	–H	26.0	26.0	26.0	
C	$–CH_3$	34.8	34.8	25.4	CH_3 21.4
	$–CH_2CH_3$	42.3	32.6	25.4	CH_2 29.2, CH_3 13.2
	$–CH(CH_3)_2$	47.4	30.0	24.7	CH 33.9, CH_3 21.7
	$–C(CH_3)_3$	50.3	26.5	25.1	C 32.5, CH_3 27.6
	$–CH_2OH$	41.2	28.3	24.5	CH_2 67.0
X	–F	95.5	32.8	22.5	$^1J_{CF}$ 173.5, $^2J_{CF}$ 22.1, $^3J_{CF}$ <1.5 Hz
	–Cl	62.0	37.2	23.1	
	–Br	53.5	37.9	23.3	
	–I	28.7	40.7	24.9	
O	–OH	73.7	35.4	23.4	
	$–OCH_3$	82.2	31.4	23.1	CH_3 56.0
	$–OCOCH_3$	77.7	33.8	24.9	CO 170.8, CH_3 21.7
N	$–NH_2$	53.4	36.4	24.0	
	$–NO_2$	87.0	32.6	24.8	
	–C≡N	27.0	30.5	24.2	CN 123.4
S	–SH	38.3	37.7	24.6	
O=C	–CO–phenyl	46.4	30.0	26.3	
	–COOH	43.0	29.2	25.1	CO 183.8
	$–COOCH_3$	43.7	30.0	25.8	CO 177.0, CH_3 51.4

^{13}C Chemical Shifts of Equatorially and Axially Monosubstituted Cyclohexanes (δ in ppm)



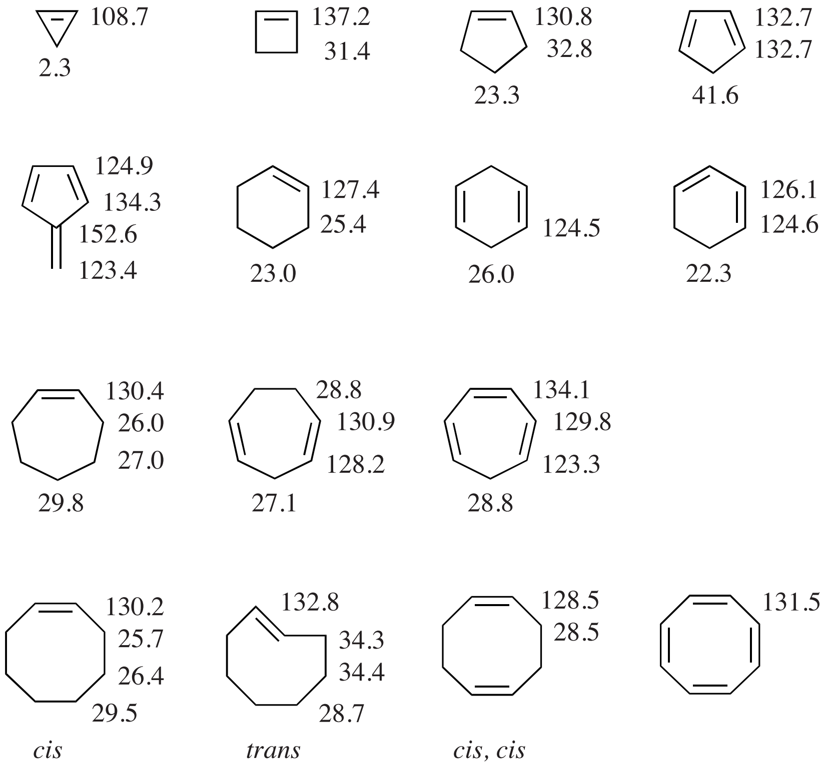


	Substituent R	a	b	c	d	a	b	c	d
	–H	27.1	27.1	27.1	27.1	27.1	27.1	27.1	27.1
C	$–CH_3$	33.2	36.0	27.1	27.0	28.4	32.4	20.6	26.9
	$–CH_2CH_3$	40.1	33.4	26.9	27.2	35.5	30.0	21.4	27.1
	$–CH_2CH_2CH_3$	40.0	33.6	26.6	26.9				
	$–CH(CH_3)_2$	44.6	30.0	26.8	27.3	41.1	30.2	21.6	27.1
	$–CH_2CH_2CH_2CH_3$	38.4	34.1	27.1	27.3				
	$–C(CH_3)_3$	48.8	28.1	27.7	27.1				
	–cyclohexyl	44.3	30.8	27.4	27.4				
	$–CH=CH_2$	42.1	32.3	26.0	27.1	37.0	30.0	21.2	27.1
	–C≡CH	28.7	32.1	25.2	24.4	28.0	30.0	21.2	25.7
	–phenyl	45.1	34.9	27.4	26.7	35.2	30.1	21.9	27.7
X	–F	91.0	32.8	23.6	25.3	88.1	30.1	19.8	25.0
	–Cl	59.8	37.4	26.1	25.4	60.1	33.9	20.4	26.0
	–Br	52.4	38.3	27.3	25.6	55.4	34.9	21.5	26.4
	–I	31.2	40.1	28.3	25.4	38.3	36.0	22.8	26.1
O	–OH	70.4	35.8	25.1	26.3	65.5	33.2	20.5	27.1
	$–OCH_3$	79.2	32.2	24.5	26.4	74.9	30.0	21.1	26.6
	$–OCOCH_3$	72.3	32.2	24.4	26.1				
	–OCO–phenyl	72.8	31.5	24.1	24.7	69.0	29.3	20.3	24.7
	$–OSi(CH_3)_3$	70.5	36.0	24.7	25.0	66.1	33.1	19.8	25.0
N	$–NH_2$	51.1	37.6	25.8	26.3	47.4	33.8	20.0	27.1
	$–NHCH_3$	58.7	32.7	25.7	26.8				
	$–N(CH_3)_2$	64.3	29.2	26.5	26.9				
	$–NH_3^+Cl^–$	51.8	32.2	24.8	25.2				
	–N=C=N–cyclohexyl	55.7	35.0	24.8	25.5				
	$–NO_2$	84.6	31.4	24.7	25.5				
	$–N_3$	59.5	31.5	24.5	24.5	56.8	29.0	20.1	25.2
	–C≡N	28.0	29.6	24.6	25.1	26.4	27.4	21.9	25.0
	–NC	51.9	33.7	24.4	25.2	50.3	30.5	20.1	25.2
	–NCS	55.3	33.9	24.5	24.8	52.8	31.3	20.4	24.8
S	–SH	38.3	38.1	26.6	25.3	35.9	33.1	19.4	25.7
C=O	–CHO	50.1	26.0	25.2	26.1	46.4	24.7	22.7	27.1
	$–COCH_3$	51.5	29.0	26.6	26.3				
	–COOH	43.7	29.6	26.2	26.6				
	$–COO^–$	47.2	30.9	26.9	26.9				
	$–COOCH_3$	43.4	29.6	26.0	26.4	39.1	27.7	24.1	26.7
	–COCl	55.4	29.7	25.5	25.9				

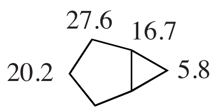
Estimation of ^{13}C Chemical Shifts of Alicyclic Compounds (δ in ppm)

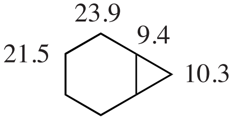
The ^{13}C chemical shift of the parent compound (e.g., 22.9 for cyclobutane, 26.0 for cyclopentane, and 27.1 ppm for cyclohexane) and the same increments as for alkanes (see Chapter 4.1) can be used to estimate the chemical shifts of *sp*^{3}-hybridized carbon atoms of alicyclic compounds. Appropriate use of the conformational correction terms, K, is especially important with axial and equatorial substituents in cyclohexanes. The additivity rule is, however, not suitable for estimating chemical shifts of substituted cyclopropanes.

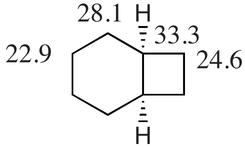
^{13}C Chemical Shifts of Unsaturated Alicyclics (δ in ppm)

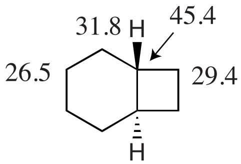


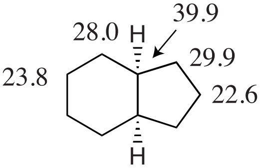
^{13}C Chemical Shifts of Condensed Alicyclics (δ in ppm)

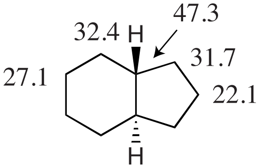


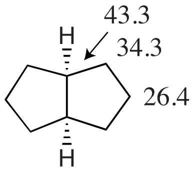


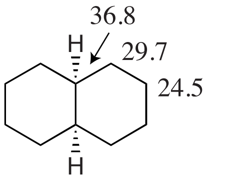


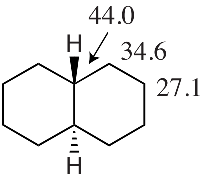


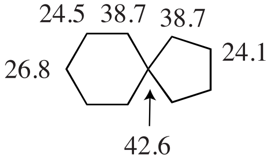


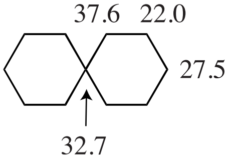


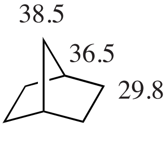


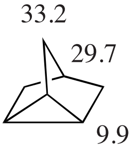


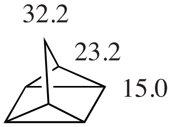


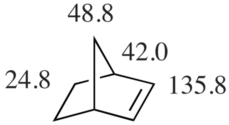


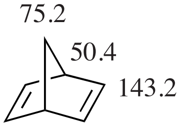


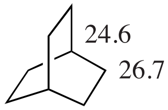


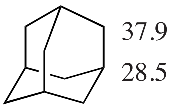


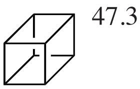


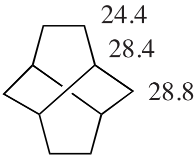


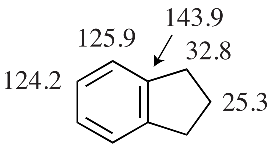


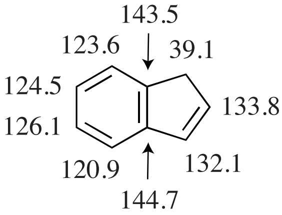


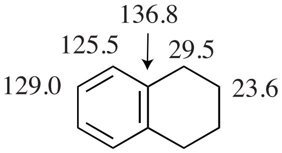








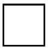
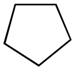
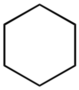




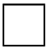
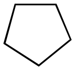
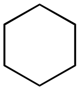
4.4.2 Coupling Constants

^{13}C-^{1}H Coupling Constants

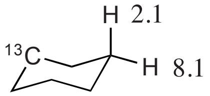
Coupling through one bond ($|^1J_{CH}|$ in Hz)

160  134  128  125

Coupling through two bonds ($|^2J_{CH}|$ in Hz)

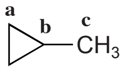
2.6  3.5  3.0  3.7

Coupling through three bonds ($|^3J_{CH}|$ in Hz)

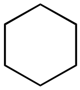


^{13}C-^{13}C Coupling Constants ($|^1J_{CC}|$ in Hz)

12.4



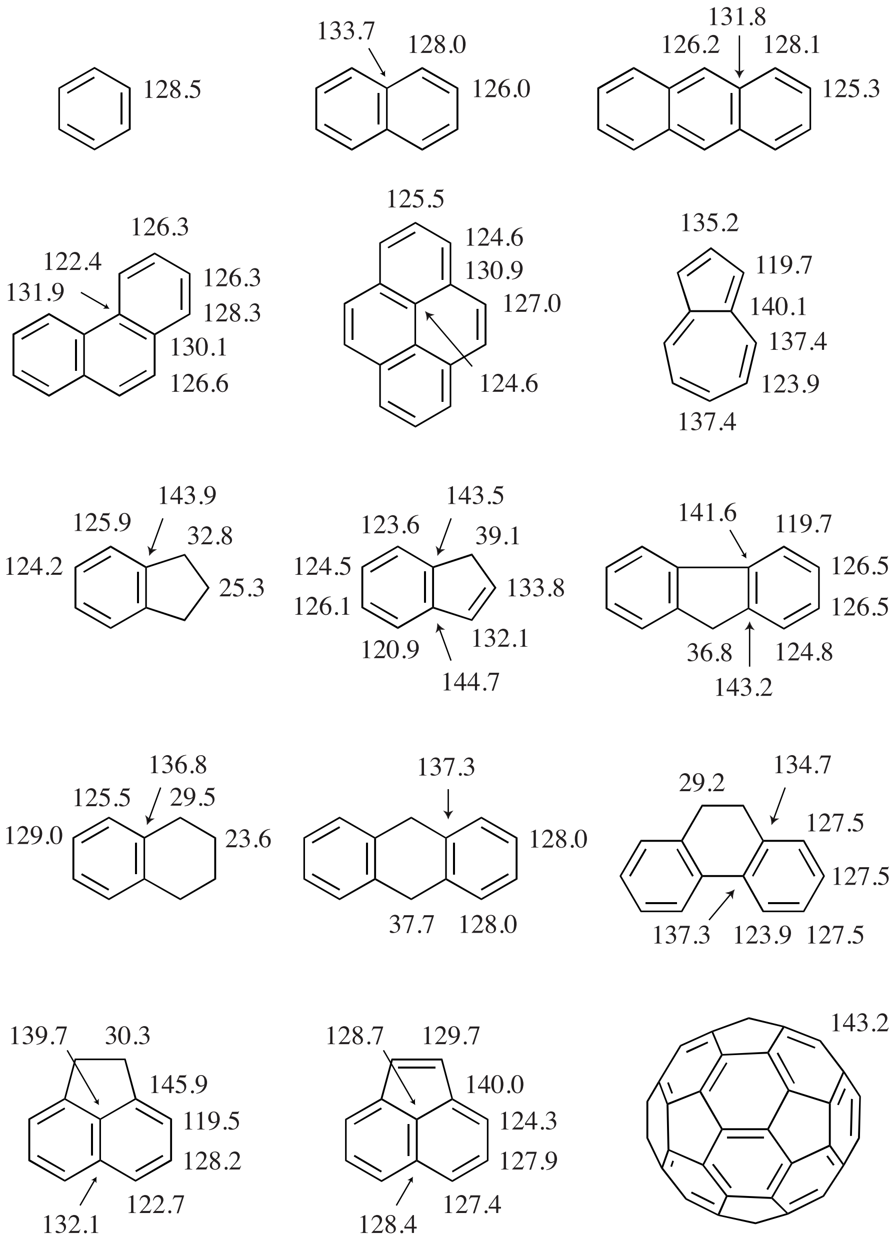
$^1J_{ab}$ 13.4
$^1J_{bc}$ 44.0

 32.7

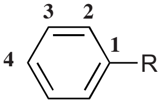
4.5 Aromatic Hydrocarbons

4.5.1 Chemical Shifts

¹³C Chemical Shifts of Aromatic Hydrocarbons (δ in ppm) [1]



Effect of Substituents on ^{13}C Chemical Shifts of Monosubstituted Benzenes (δ in ppm)



$$\delta_{C_i} = 128.5 + Z_i$$

Substituent R	Z_1	Z_2	Z_3	Z_4
$-CH_3$	9.2	0.7	-0.1	-3.0
$-CH_2CH_3$	11.7	-0.6	-0.1	-2.8
$-CH_2CH_2CH_3$	10.3	-0.2	0.1	-2.7
$-CH(CH_3)_2$	20.2	-2.2	-0.3	-2.8
$-CH_2CH_2CH_2CH_3$	10.9	-0.2	-0.2	-2.8
$-C(CH_3)_3$	18.6	-3.3	-0.4	-3.1
–cyclopropyl	15.1	-3.3	-0.6	-3.6
–cyclopentyl	17.8	-1.5	-0.4	-2.9
–cyclohexyl	16.3	-1.8	-0.3	-2.8
–1-adamantyl	22.2	-2.9	-0.5	-3.1
$-CH_2F$	8.5	-0.7	0.4	0.5
$-CF_3$	2.5	-3.2	0.3	3.3
$-CH_2Cl$	9.3	0.3	0.2	0.0
$-CHCl_2$	11.9	-2.4	0.1	1.2
$-CCl_3$	16.3	-1.7	-0.1	1.8
$-CH_2Br$	9.5	0.7	0.3	0.2
$-CH_2I$	10.5	0.0	0.0	-0.9
$-CH_2OH$	12.4	-1.2	0.2	-1.1
$-CH_2OCH_3$	8.7	-0.9	-0.1	-0.9
$-CH_2NH_2$	14.9	-1.4	-0.2	-2.0
$-CH_2NHCH_3$	12.6	-0.3	-0.3	-1.8
$-CH_2N(CH_3)_2$	7.8	0.5	-0.3	-1.5
$-CH_2NO_2$	2.2	2.2	2.2	1.2
$-CH_2CN$	1.6	0.5	-0.8	-0.7
$-CH_2SH$	12.5	-0.6	0.0	-1.6
$-CH_2SCH_3$	9.8	0.4	-0.1	-1.6
$-CH_2S(O)CH_3$	0.8	1.5	0.4	-0.2
$-CH_2S(O)_2CH_3$	-0.1	2.1	0.6	0.6
$-CH_2CHO$	7.4	1.3	0.5	-1.1
$-CH_2COCH_3$	5.8	0.8	0.1	-1.6
$-CH_2COOH$	6.5	1.4	0.4	-1.2
$-CH_2Li$	32.2	-22.0	-0.4	-24.3
$-CH{=}CH_2$	8.9	-2.3	-0.1	-0.8
$-C(CH_3){=}CH_2$	12.6	-3.1	-0.4	-1.2
$-C{\equiv}CH$	-6.2	3.6	-0.4	-0.3
–phenyl	8.1	-1.1	0.5	-1.1
–2-pyridyl	11.2	-1.4	0.5	-1.4
–4-pyridyl	9.6	-1.6	0.5	0.5

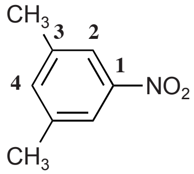
	Substituent R	Z_1	Z_2	Z_3	Z_4
X	–F	33.6	-13.0	1.6	-4.4
	–Cl	5.3	0.4	1.4	-1.9
	–Br	-5.4	3.3	2.2	-1.0
	–I	-31.2	8.9	1.6	-1.1
O	–OH	28.8	-12.8	1.4	-7.4
	–ONa	39.6	-8.2	1.9	-13.6
	$–OCH_3$	33.5	-14.4	1.0	-7.7
	$–OCH{=}CH_2$	28.2	-11.5	0.7	-5.8
	–O–phenyl	27.6	-11.2	-0.3	-6.9
	$–OCOCH_3$	22.4	-7.1	0.4	-3.2
	$–OSi(CH_3)_3$	26.8	-8.4	0.9	-7.1
	–OPO(O–phenyl)$_2$	21.9	-8.4	1.2	-3.0
	–OCN	25.0	-12.7	2.6	-1.0
N	$–NH_2$	18.2	-13.4	0.8	-10.0
	$–NHCH_3$	15.0	-16.2	0.8	-11.6
	$–N(CH_3)_2$	16.0	-15.4	0.9	-10.5
	–NH–phenyl	14.7	-10.6	0.9	-10.5
	–N(phenyl)$_2$	13.1	-7.0	0.9	-5.6
	$–NH_3^+$	0.1	-5.8	2.2	2.2
	$–NH_2^+CH(CH_3)_2$	5.5	-4.1	1.1	0.7
	$–N^+(CH_3)_3$	19.5	-7.3	2.5	2.4
	$–N(O)(CH_3)_2$	26.2	-8.4	0.8	0.6
	$–NHCOCH_3$	9.7	-8.1	0.2	-4.4
	–NHOH	21.5	-13.1	-2.2	-5.3
	$–NHNH_2$	22.8	-16.5	0.5	-9.6
	–N=CH–phenyl	24.7	-6.5	1.3	-1.5
	$–N{=}NCH_3$	22.2	-6.2	0.5	-3.0
	–NO	37.4	-7.6	0.8	7.1
	$–NO_2$	19.9	-4.9	0.9	6.1
	–C≡N	-16.0	3.5	0.7	4.3
	–NC	-1.8	-2.2	1.4	0.9
	–NCO	5.1	-3.7	1.1	-2.8
	–NCS	3.0	-2.7	1.3	-1.0
	$–N^+{\equiv}N$	-12.7	6.0	5.7	16.0
S	–SH	4.0	0.7	0.3	-3.2
	$–SCH_3$	10.0	-1.9	0.2	-3.6
	$–SC(CH_3)_3$	4.5	9.0	-0.3	0.0
	$–S(CH_3)_2^+$	-1.0	3.1	2.2	6.3
	$–SCH{=}CH_2$	5.8	2.0	0.2	-1.8
	–S–phenyl	7.3	2.5	0.6	-1.5
	–S–S–phenyl	7.5	-1.3	0.8	-1.1
	$–S(O)CH_3$	17.6	-5.0	1.1	2.4
	$–S(O)_2CH_3$	12.3	-1.4	0.8	5.1
	$–S(O)_2OH$	15.0	-2.2	1.3	3.8
	$–S(O)_2OCH_3$	6.4	-0.6	1.5	5.9

	Substituent R	Z_1	Z_2	Z_3	Z_4
S	$-S(O)_2F$	4.6	0.0	1.5	7.5
	$-S(O)_2Cl$	15.6	-1.7	1.2	6.8
	$-S(O)_2NH_2$	10.8	-3.0	0.3	3.2
	–SCN	-3.7	2.5	2.2	2.2
O=C	–CHO	8.2	1.2	0.5	5.8
	$-COCH_3$	8.9	0.1	-0.1	4.4
	$-COCF_3$	-5.6	1.8	0.7	6.7
	–COC≡CH	7.4	1.0	0.0	5.9
	–CO–phenyl	9.3	1.6	-0.3	3.7
	–COOH	2.1	1.6	-0.1	5.2
	–COONa	9.7	4.6	2.2	4.6
	$-COOCH_3$	2.0	1.2	-0.1	4.3
	$-CONH_2$	5.0	-1.2	0.1	3.4
	$-CON(CH_3)_2$	6.0	-1.5	-0.2	1.0
	–COCl	4.7	2.7	0.3	6.6
	–COSH	6.2	-0.6	0.2	5.4
	$-CH{=}NCH_3$	8.8	0.5	0.1	2.3
	–CS–phenyl	18.7	1.0	-0.6	2.4
P	$-P(CH_3)_2$	13.6	1.6	-0.6	-1.0
	$-P(\text{phenyl})_2$	8.9	5.2	0.0	0.1
	$-P^+(\text{phenyl})_2CH_3$	-9.7	5.2	2.0	6.7
	$-PO(CH_3)_2$	2.5	1.1	0.1	3.0
	$-PO(\text{phenyl})_2$	5.8	3.9	-0.1	3.0
	$-PO(OH)_2$	-1.9	3.6	1.5	5.6
	$-PO(OCH_2CH_3)_2$	1.6	3.6	-0.2	3.4
	$-PS(CH_3)_2$	6.7	2.0	0.2	2.9
	$-PS(OCH_2CH_3)_2$	6.1	2.8	-0.4	3.4
M	–Li	-43.2	-12.7	2.4	3.1
	–MgBr	-35.8	-11.4	2.7	4.0
	$-SiH_3$	-0.5	7.3	-0.4	1.3
	$-SiH_2CH_3$	4.8	6.3	-0.5	1.0
	$-Si(CH_3)_3$	11.6	4.9	-0.7	0.4
	$-Si(\text{phenyl})_3$	5.8	7.9	-0.6	1.1
	$-SiCl_3$	3.0	4.6	0.1	4.2
	$-Ge(CH_3)_3$	13.7	4.5	-0.5	-0.2
	$-Sn(CH_3)_3$	13.2	7.2	-0.4	-0.4
	$-Pb(CH_3)_3$	20.1	8.0	-0.1	-1.0
	$-AsH_2$	1.7	7.9	0.8	0.0
	$-As(\text{phenyl})_2$	11.1	5.0	0.1	-0.1
	$-As(O)(OH)_2$	3.8	1.6	0.8	4.5
	$-SeCH{=}CH_2$	0.7	4.7	0.4	-1.4
	–SeCN	-5.3	5.1	2.9	2.1
	$-Sb(\text{phenyl})_2$	9.8	7.7	0.3	0.0
	–Hg–phenyl	41.6	9.3	-0.9	-1.6
	–HgCl	22.5	8.0	-0.6	-0.9

Estimation of ^{13}C Chemical Shifts of Multiply Substituted Benzenes and Naphthalenes (δ in ppm)

The ^{13}C chemical shifts of multiply substituted benzenes and naphthalenes (see next pages) can be estimated using the substituent effects in the corresponding monosubstituted hydrocarbons.

Example: Estimation of the chemical shifts for 3,5-dimethylnitrobenzene



C-1	base value	128.5
	$Z_1(NO_2)$	19.9
	$2\ Z_3(CH_3)$	-0.2
	estimated	148.2
	exp	148.5

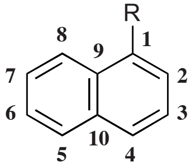
C-2	base value	128.5
	$Z_2(NO_2)$	-4.9
	$Z_2(CH_3)$	0.7
	$Z_4(CH_3)$	-3.0
	estimated	121.3
	exp	121.7

C-3	base value	128.5
	$Z_1(CH_3)$	9.2
	$Z_3(CH_3)$	-0.1
	$Z_3(NO_2)$	0.9
	estimated	138.5
	exp	139.6

C-4	base value	128.5
	$2\ Z_2(CH_3)$	1.4
	$Z_4(NO_2)$	6.1
	estimated	136.0
	exp	136.2

Larger discrepancies between estimated and experimental values are to be expected if the substituents are *ortho* to each other or if strongly electron-donating and electron-accepting groups occur simultaneously.

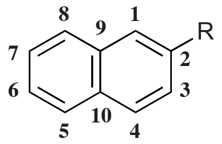
Effect of Substituents in Position 1 on ^{13}C Chemical Shifts of Monosubstituted Naphthalenes (δ in ppm)



for R: H $\delta_{C_1} = 128.0$
$\delta_{C_2} = 125.9$
$\delta_{C_9} = 133.6$

	Substituent R	C-1	C-2	C-3	C-4	C-5	C-6	C-7	C-8	C-9	C-10
C	$-CH_3$	6.0	0.5	0.6	-1.8	0.3	-0.7	-0.5	-4.1	-1.1	-0.2
	$-C(CH_3)_3$	17.9	-2.8	-0.9	-0.6	1.6	-1.4	-1.4	-1.2	-1.6	2.2
	$-CH_2Br$	4.0	1.1	-0.9	1.3	0.5	-0.1	0.3	-4.6	-2.8	0.1
	$-CH_2OH$	8.2	-0.9	-0.6	0.1	0.5	-0.3	0.1	-4.5	-2.6	0.0
	$-CF_3$	-1.9	-1.3	-1.8	5.0	1.0	0.8	2.0	-3.4	1.0	-3.9
X	–F	31.5	-16.1	0.1	-3.8	0.1	1.4	0.7	-7.1	-9.3	2.1
	–Cl	3.9	0.2	-0.2	-0.9	0.2	3.1	0.8	-3.6	-2.8	1.0
	–Br	-5.4	3.6	-0.2	-0.5	-0.1	0.4	1.0	-1.3	-2.0	0.6
	–I	-28.4	12.3	1.7	1.7	1.4	1.6	2.6	4.4	1.3	1.3
O	–OH	23.5	-17.2	-0.1	-7.3	-0.4	0.5	0.3	-6.6	-9.3	1.0
	$-OCH_3$	27.3	-22.3	-0.2	-7.9	-0.7	0.3	-0.9	-6.1	-8.1	0.8
	$-OCOCH_3$	18.6	-7.9	-0.6	-2.1	0.0	0.4	0.4	-6.9	-6.9	0.9
N	$-NH_2$	14.0	-16.5	0.3	-9.3	0.3	-0.3	-1.3	-7.3	-10.2	0.6
	$-N(CH_3)_2$	23.7	-11.2	0.6	-4.6	1.0	0.4	-0.3	-3.2	-3.9	2.1
	$-NH_3^+$	-3.8	-4.6	-0.9	3.4	1.4	2.1	2.8	-9.0	-7.4	1.2
	$-NHCOCH_3$	5.7	-4.4	-0.5	-3.0	0.0	-0.1	-0.3	-5.3	-5.9	0.1
	$-NO_2$	18.5	-2.1	-2.0	6.5	0.5	1.3	3.4	-5.1	-8.7	0.6
	–C≡N	-19.2	5.1	-2.4	3.8	-0.7	0.2	1.2	-4.5	-2.8	-2.2
O=C	–CHO	2.9	10.8	-1.4	6.7	0.2	0.6	2.7	-3.5	-3.6	-0.3
	$-COCH_3$	6.9	2.9	-1.7	4.9	0.3	0.4	2.0	-2.0	-3.5	0.2
	–COOH	-1.5	3.6	-2.4	4.3	-0.6	-0.9	0.6	-3.2	-3.2	-0.8
	$-COOCH_3$	-0.9	4.5	-1.2	5.4	0.7	0.5	1.9	-1.8	-1.9	0.5
	$-CON(CH_3)_2$	6.8	-2.1	-0.8	0.9	0.4	0.4	1.0	0.1	-4.1	-0.2
	–COCl	1.2	10.6	-0.5	9.3	1.9	2.1	4.5	-2.1	-2.1	1.0
	$-Si(CH_3)_3$	9.8	5.1	-0.4	1.7	1.2	-0.8	-0.7	0.1	3.8	0.2

Effect of Substituents in Position 2 on ^{13}C Chemical Shifts of Monosubstituted Naphthalenes (δ in ppm)



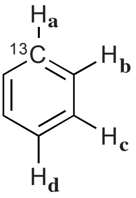
for R: H $\delta_{C_1} = 128.0$
$\delta_{C_2} = 125.9$
$\delta_{C_9} = 133.6$

	Substituent R	C-1	C-2	C-3	C-4	C-5	C-6	C-7	C-8	C-9	C-10
C	$-CH_3$	-1.3	9.3	2.0	-0.8	-0.5	-1.1	-0.2	-0.6	-0.1	-2.0
	$-C(CH_3)_3$	-3.3	22.5	-3.0	-0.4	0.0	-0.7	-0.2	-0.6	0.4	-1.3
	$-CH_2Br$	-1.7	9.0	1.9	-0.4	-0.5	0.7	0.3	0.6	-0.6	-0.7
	$-CH_2OH$	-2.7	12.3	-4.4	-0.1	-0.4*	-0.2*	0.1*	-0.2*	-0.3	-0.8
	$-CF_3$	-2.0	1.9	-4.2	1.1*	0.1*	2.4*	1.5	1.1	-1.1	1.3
X	–F	-17.0	34.9	-9.6	2.4	0.0	-0.7	1.1	-0.6	0.7	-3.0
	–Cl	-1.4	5.7	0.8	1.5	-0.2	0.2	1.1	-1.1	0.7	-1.9
	–Br	1.8	-6.2	3.1	1.5	-0.3	0.2	0.8	-1.1	-2.0	0.7
	–I	9.2	-34.1	9.0	2.3	0.5	1.3	1.5	-0.6	2.1	-0.8
O	–OH	-18.6	27.3	-8.3	1.8	-0.3	-2.4	0.5	-1.7	0.9	-4.7
	$-OCH_3$	-22.2	31.8	-7.1	1.5	-0.3	-2.2	0.5	-1.2	1.0	-4.3
	$-OCOCH_3$	-9.5	22.5	-4.8	1.3	-0.4	-0.3	0.6	-0.4	0.1	-2.2
N	$-NH_2$	-20.6	16.7	-8.9	-0.2	-1.6	-4.8	-0.9	-3.5	-0.1	-7.0
	$-N(CH_3)_2$	-21.1	23.6	-8.8	1.2	0.0	-3.4	0.7	-1.1	2.4	-5.9
	$-NH_3^+$	-5.9	-0.3	-6.5	3.2	0.2	2.3	2.0	0.2	0.1	-0.3
	$-NHCOCH_3$	-11.0	9.6	-5.7	0.6	-0.4*	-0.9	1.6*	-1.6	0.2	-3.0
	$-NO_2$	-3.4	20.0	-6.7	1.7	0.1	4.0	2.2	2.1	-1.1	2.4
	–C≡N	5.8	-16.7	0.1	1.0	-0.2	3.0	1.6	0.2	-1.6	0.7
O=C	–CHO	6.2	7.9	-3.6	0.8	-0.3	2.9	0.9	1.8	2.4	-1.4
	$-COCH_3$	1.9	8.3	-2.2	0.2	-0.4	2.3	0.7	1.4	1.8	-1.3
	–COOH	2.7	2.4	-0.6	0.2	-0.3	2.4	0.9	1.3	-1.3	1.5
	$-COOCH_3$	3.0	1.8	-0.5	0.2	-0.1	2.4	0.9	1.4	-1.0	1.9
	–COCl	2.5	9.1	-0.7	0.2*	-0.4	2.2*	0.8	1.2		-1.4
	$-Si(CH_3)_3$	5.8	11.9	3.9	-1.0	0.1	0.3	-0.2	0.1	-0.5	0.2

* Assignment uncertain

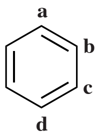
4.5.2 Coupling Constants

^{13}C-^{1}H Coupling Constants (|J| in Hz)

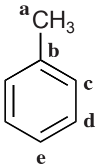


	In benzene:	In derivatives:
$^{1}J_{^{13}CH_a}$	159.0	
$^{2}J_{^{13}CH_b}$	1.0	1–4
$^{3}J_{^{13}CH_c}$	7.6	7–10
$^{4}J_{^{13}CH_d}$	-1.3	

^{13}C-^{13}C Coupling Constants (|J| in Hz)



$^{1}J_{ab}$	57.0
$^{2}J_{ac}$	2.5
$^{3}J_{ad}$	10.0



$^{1}J_{ab}$	44.2
$^{2}J_{ac}$	3.1
$^{3}J_{ad}$	3.8
$^{4}J_{ae}$	0.9

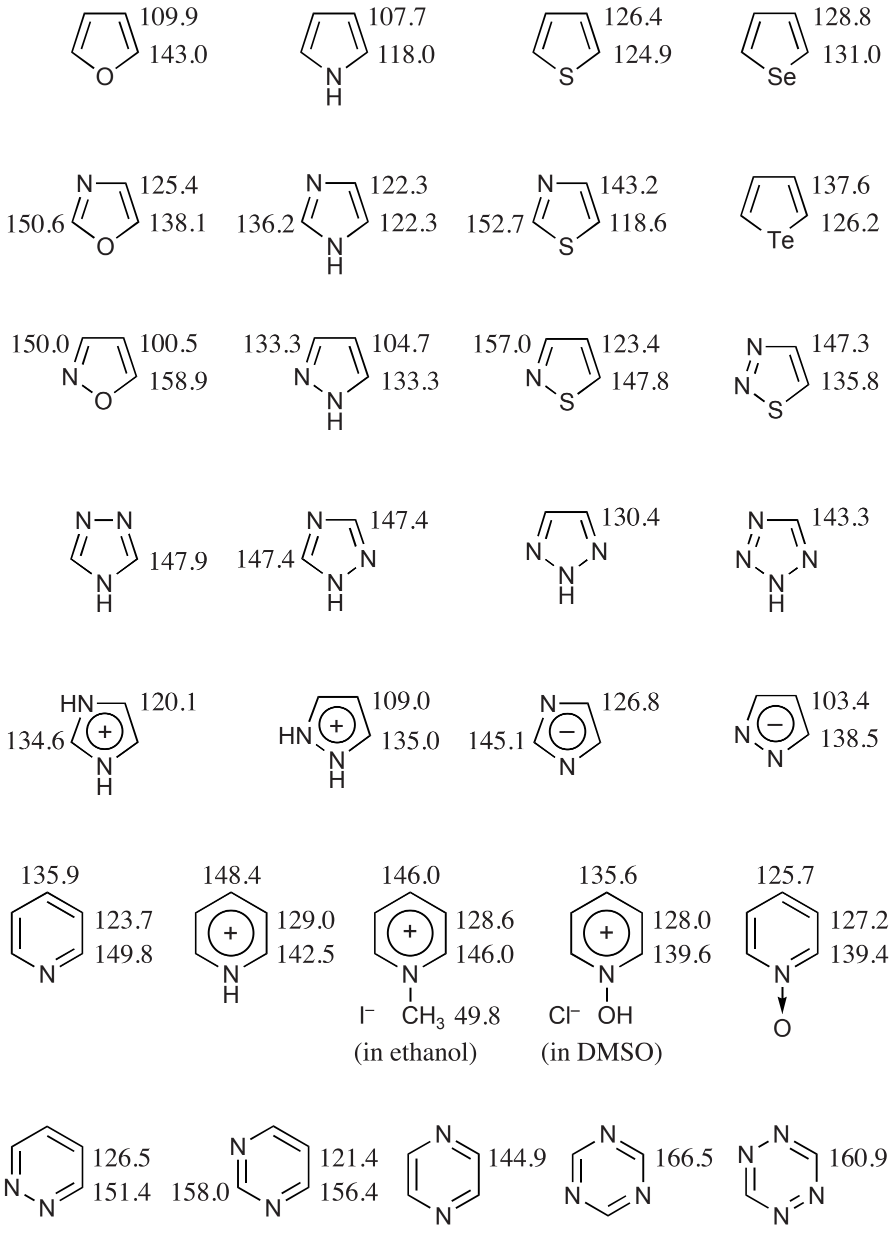
4.5.3 References

[1] P.E. Hansen, ^{13}C NMR of polycyclic aromatic hydrocarbons. A review, *Org. Magn. Reson.* **1979**, *12*, 109.

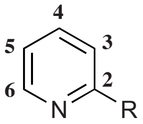
4.6 Heteroaromatic Compounds

4.6.1 Chemical Shifts

^{13}C Chemical Shifts of Monocyclic Heteroaromatics (δ in ppm)



Effect of Substituents in Position 2 on ¹³C Chemical Shifts of Monosubstituted Pyridines (δ in ppm)

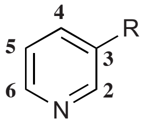


for R: H $\delta_{C_{2,6}} = 149.8$
$\delta_{C_{3,5}} = 123.7$
$\delta_{C_4} = 135.9$

	Substituent R	C-2	C-3	C-4	C-5	C-6
C	$-CH_3$	8.6	-0.5	0.3	-3.0	-0.7
	$-CH_2CH_3$	13.7	-1.7	0.4	-2.8	-0.6
	$-CH{=}CH_2$	5.9	-1.3	1.1	-2.5	-0.3
	–phenyl	7.7	-1.6	0.8	-3.2	0.2
X	–F	13.9	-14.0	5.4	-2.5	-2.0
	–Cl	1.8	0.8	2.8	-1.4	0.0
	–Br	-7.5	4.6	2.6	-1.1	0.5
	–I	-31.6	11.3	1.7	-0.8	1.0
O	–OH*	15.5	-3.6	-1.1	-17.0	-8.2
	$-OCH_3$	14.3	-12.7	2.6	-7.1	-2.9
	–O–phenyl	13.9	-12.2	3.5	-5.3	-2.0
	$-OCOCH_3$	7.6	-7.3	3.4	-1.8	-1.6
N	$-NH_2$	8.4	-15.1	1.8	-9.7	-1.6
	$-NHCH_3$	10.9	-16.2	1.5	-11.3	-1.3
	$-N(CH_3)_2$	9.6	-17.9	1.2	-12.3	-1.9
	$-NHCOCH_3$	1.4	-9.8	2.6	-3.9	-2.1
	$-NO_2$	6.9	-5.7	3.9	5.4	-0.8
	–C≡N	-15.8	4.8	1.1	3.2	1.4
S	–SH	30.4	10.7	2.1	-10.6	-12.1
	$-SCH_3$	10.2	-4.6	0.0	-2.2	-0.5
	$-S(O)CH_3$	16.2	-4.4	2.2	0.9	-0.2
	$-S(O)_2CH_3$	8.5	-2.6	2.4	3.7	0.3
O=C	–CHO	3.0	-2.0	1.2	4.2	0.4
	$-COCH_3$	3.8	-2.1	0.9	3.4	-0.8
	–COOH	-3.7	0.0	2.5	4.2	-1.7
	$-COOCH_3$	-1.7	1.5	1.1	3.3	0.0
	$-CONH_2$	-0.3	-1.2	1.4	2.8	-1.5
M	$-Si(CH_3)_3$	18.6	5.0	-2.0	-1.1	0.3
	$-Sn(CH_3)_3$	23.3	7.6	-2.7	-1.7	0.6
	$-Pb(CH_3)_3$	33.4	9.2	-2.6	-2.3	1.1

* Keto form (2-pyridone)

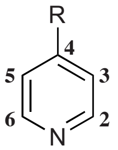
Effect of Substituents in Position 3 on ^{13}C Chemical Shifts of Monosubstituted Pyridines (δ in ppm)



for R: H $\delta_{C2,6} = 149.8$
$\delta_{C3,5} = 123.7$
$\delta_{C4} = 135.9$

	Substituent R	C-2	C-3	C-4	C-5	C-6
C	$–CH_3$	1.3	8.9	0.0	-0.9	-2.3
	$–CH_2CH_3$	-0.4	15.4	-0.8	-0.5	-2.7
	–phenyl	-1.4	12.8	-1.8	-0.3	-1.3
X	–F	-11.5	36.1	-13.2	0.8	-3.9
	–Cl	-0.3	8.1	-0.4	0.6	-1.4
	–Br	2.1	-2.7	2.7	1.1	-0.9
	–I	7.1	-28.5	8.9	2.3	0.3
O	–OH	-10.7	31.3	-12.4	1.2	-8.6
	$–OCH_3$	-12.5	31.5	-15.9	0.1	-8.4
	$–OCOCH_3$	-6.5	23.4	-7.0	-0.1	-3.2
N	$–NH_2$	-11.9	21.4	-14.4	0.8	-10.8
	$–NHCH_3$	-13.6	23.1	-18.2	0.6	-11.9
	$–N(CH_3)_2$	-14.0	23.3	-17.1	0.1	-11.6
	–C≡N	3.6	-13.8	4.2	0.5	4.2
S	–SH	-12.8	26.1	-11.3	7.3	-2.8
	$–SCH_3$	-13.6	24.6	-11.7	10.6	-3.0
O=C	–CHO	2.4	7.8	-0.2	0.5	5.4
	$–COCH_3$	3.5	8.5	-0.7	-0.2	0.0
	–COOH	-6.4	13.0	11.1	4.3	-6.0
	$–COOCH_3$	-0.6	1.0	-0.5	-1.8	1.8
	$–CONH_2$	2.7	5.9	1.1	1.2	-1.5
M	$–Si(CH_3)_3$	2.7	9.1	3.0	-2.3	-1.2
	$–Ge(CH_3)_3$	3.9	12.8	4.2	-0.4	-0.1
	$–Sn(CH_3)_3$	5.9	13.0	7.1	0.1	-0.3
	$–Sn(n\text{-}C_4H_9)_3$	6.6	12.6	7.7	0.0	-0.4
	$–Pb(n\text{-}C_4H_9)_3$	7.1	21.7	8.5	0.9	-1.8

Effect of Substituents in Position 4 on ^{13}C Chemical Shifts of Monosubstituted Pyridines (δ in ppm)



for R: H $\delta_{C2,6} = 149.8$

$\delta_{C3,5} = 123.7$

$\delta_{C4} = 135.9$

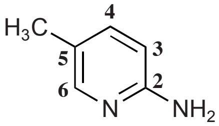
	Substituent R	C-2	C-3	C-4
C	$-CH_3$	0.5	0.7	10.6
	$-CH_2CH_3$	-0.1	-0.5	16.8
	$-CH(CH_3)_2$	0.4	-1.9	21.2
	$-C(CH_3)_3$	0.9	-2.6	23.9
	$-CH{=}CH_2$	0.3	-3.0	8.4
	–phenyl	0.4	-2.2	12.2
X	–F	2.7	-11.9	32.8
	–Br	3.0	3.3	-3.2
	–I	0.2	9.1	-30.8
O	–OH*	-9.8	-6.2	45.4
	$-OCH_3$	0.9	-13.9	29.0
	$-OCOCH_3$	1.7	-6.7	23.9
N	$-NH_2$	0.7	-13.8	19.3
	$-NHCH_3$	0.5	-15.9	19.8
	$-N(CH_3)_2$	0.6	-16.3	19.2
	$-C{\equiv}N$	2.1	2.1	-15.9
S	–SH	-16.9	5.9	54.3
	$-SCH_3$	0.1	-3.3	14.6
O=C	–CHO	1.7	-0.7	5.3
	$-COCH_3$	1.6	-2.7	6.6
	$-COOCH_3$	1.0	-0.8	1.4
	$-CONH_2$	0.4	-0.9	6.2
M	$-Si(CH_3)_3$	-2.8	2.4	11.9
	$-Ge(CH_3)_3$	-1.1	4.4	16.8
	$-Sn(CH_3)_3$	-1.1	7.3	16.2
	$-Pb(CH_3)_3$	-0.5	9.1	24.6

* Keto form (4-pyridone)

Estimation of ^{13}C Chemical Shifts of Multiply Substituted Pyridines (δ in ppm)

The ^{13}C chemical shifts in multiply substituted pyridines can be estimated using the substituent effects in the monosubstituted parent compound.

Example: Estimation of the chemical shifts for 2-amino-5-methylpyridine



C-2	base value	149.8
	$2\text{-}NH_2$	8.4
	$5\text{-}CH_3$	-2.3
	estimated	155.9
	exp	156.9

C-3	base value	123.7
	$2\text{-}NH_2$	-15.1
	$5\text{-}CH_3$	-0.9
	estimated	107.7
	exp	108.4

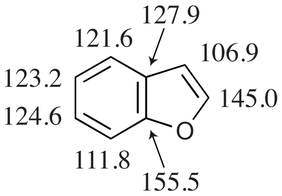
C-4	base value	135.9
	$2\text{-}NH_2$	1.8
	$5\text{-}CH_3$	0.0
	estimated	137.7
	exp	138.6

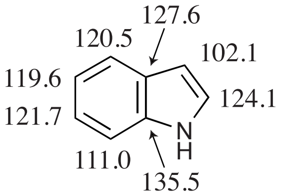
C-5	base value	123.7
	$2\text{-}NH_2$	-9.7
	$5\text{-}CH_3$	8.9
	estimated	122.9
	exp	122.5

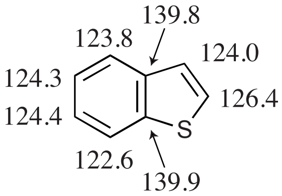
C-6	base value	149.8
	$2\text{-}NH_2$	-1.6
	$5\text{-}CH_3$	1.3
	estimated	149.5
	exp	147.6

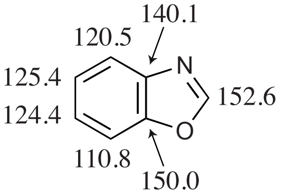
Larger discrepancies between estimated and experimental values are to be expected if the substituents are *ortho* to each other and if strongly electron-donating and -accepting groups occur simultaneously. Also, tautomerization and zwitterion formation have large effects on ^{13}C chemical shifts.

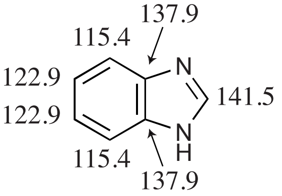
^{13}C Chemical Shifts of Condensed Heteroaromatics (δ in ppm)

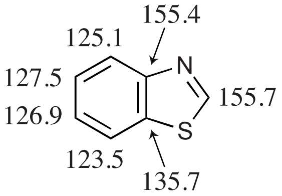


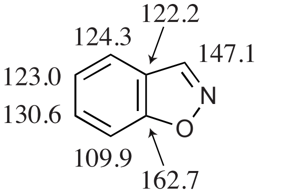


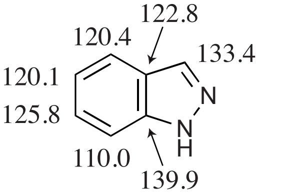


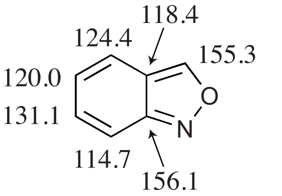


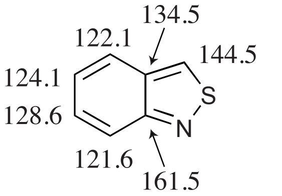


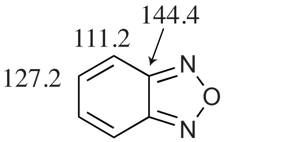


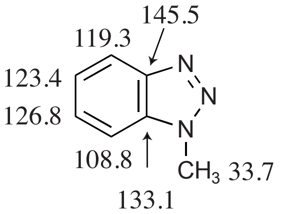


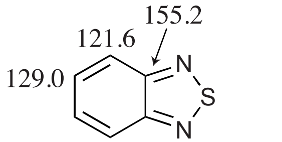


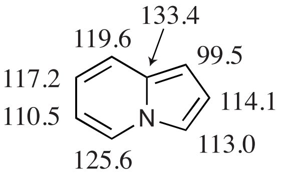


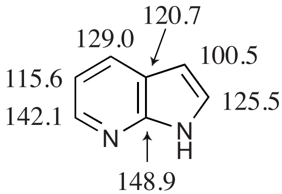


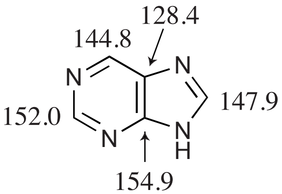


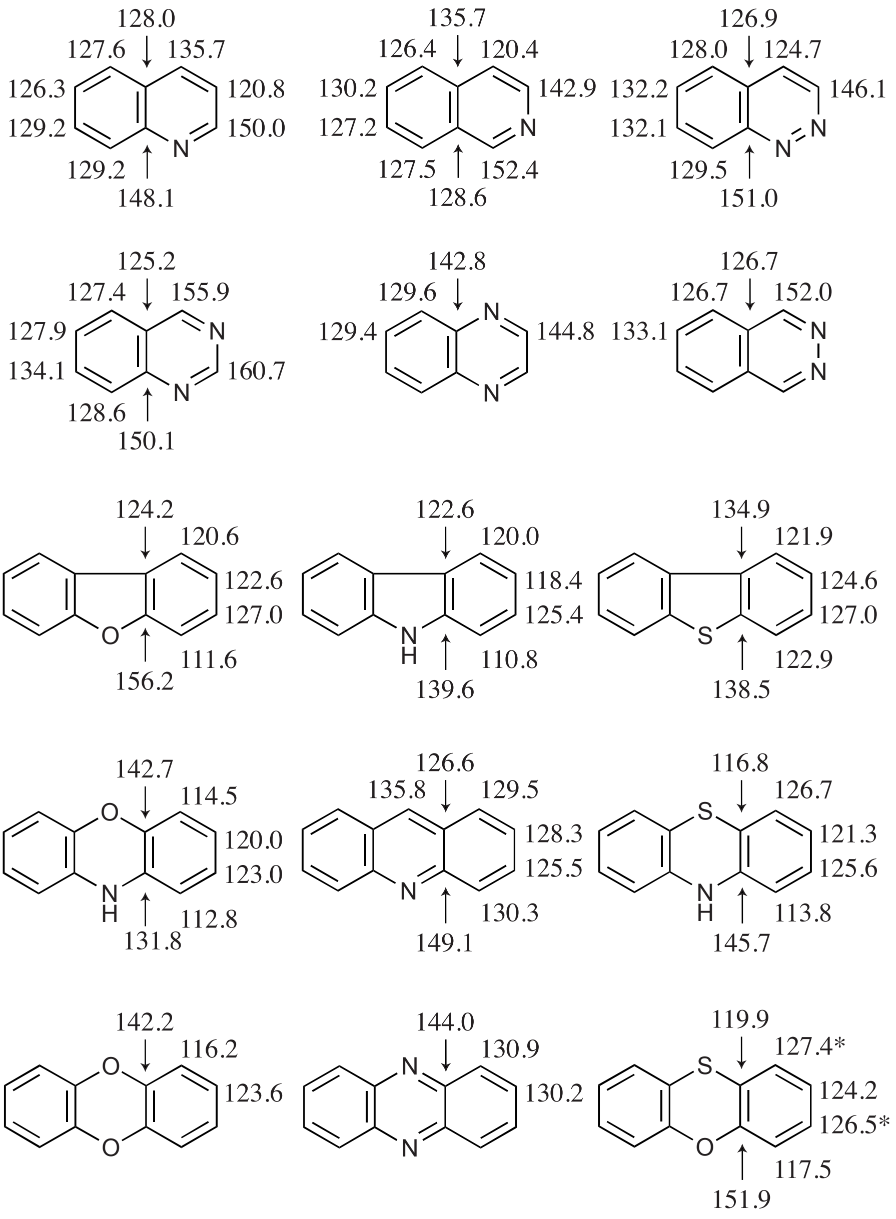








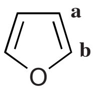




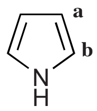
* Assignment uncertain

4.6.2 Coupling Constants

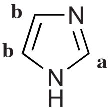
^{13}C-^{1}H Coupling Constants ($|^{1}J|$ in Hz)



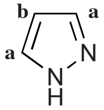
$^{1}J_{C_aH_a}$ 175
$^{1}J_{C_bH_b}$ 202



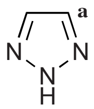
$^{1}J_{C_aH_a}$ 169
$^{1}J_{C_bH_b}$ 183



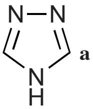
$^{1}J_{C_aH_a}$ 206
$^{1}J_{C_bH_b}$ 189



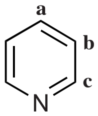
$^{1}J_{C_aH_a}$ 186
$^{1}J_{C_bH_b}$ 177



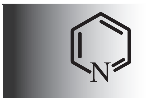
$^{1}J_{C_aH_a}$ 194



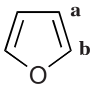
$^{1}J_{C_aH_a}$ 209



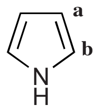
$^{1}J_{C_aH_a}$ 161
$^{1}J_{C_bH_b}$ 163
$^{1}J_{C_cH_c}$ 178



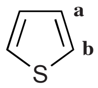
^{13}C-^{13}C Coupling Constants ($|^{1}J|$ in Hz)



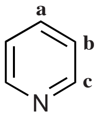
$^{1}J_{C_aC_b}$ 69.1



$^{1}J_{C_aC_b}$ 65.6



$^{1}J_{C_aC_b}$ 64.2



$^{1}J_{C_aC_b}$ 53.7
$^{1}J_{C_bC_c}$ 54.3

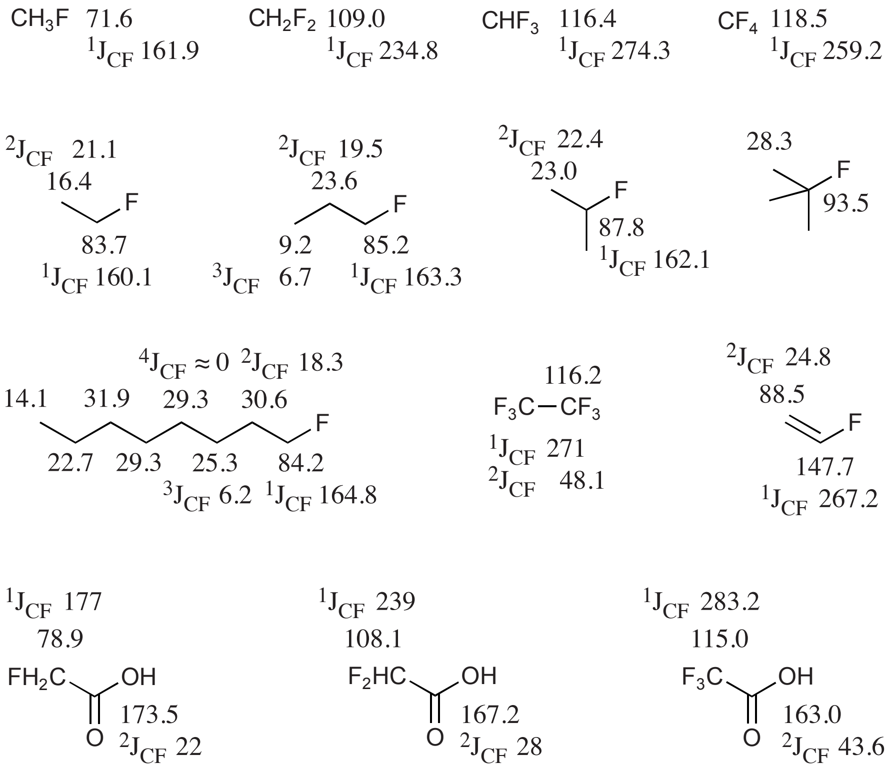
4.7 Halogen Compounds

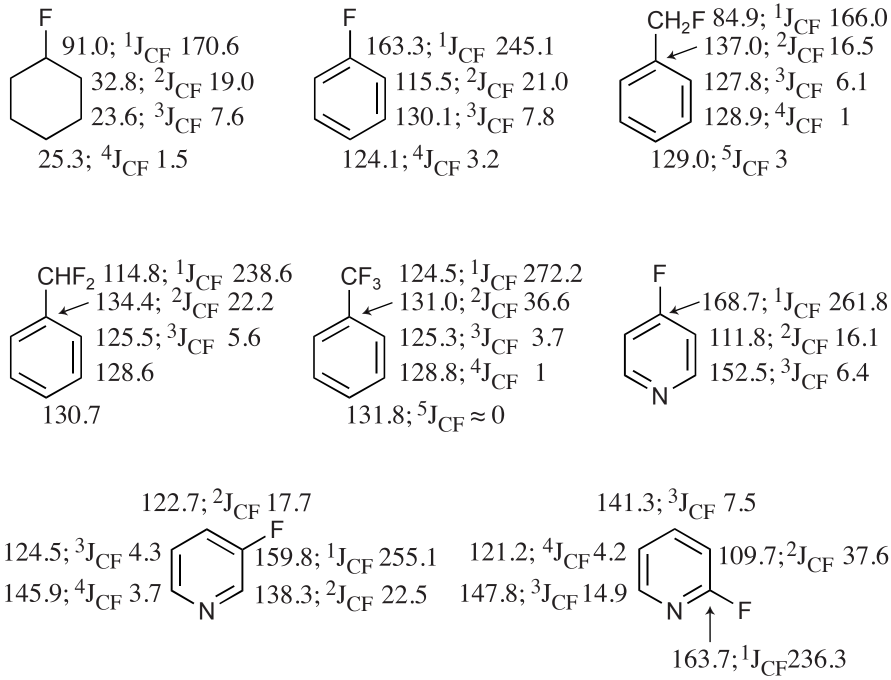
The additivity rules for estimating the ^{13}C chemical shifts of various skeletons can be applied to those haloalkanes that do not have more than one halogen atom at a given carbon atom. In all other cases, the simple linear models fail but correction terms for non-additivity are available for halomethanes and derivatives (see [1, 2]).

4.7.1 Fluoro Compounds

^{19}F (natural abundance 100%) has a spin quantum number I of 1/2. The signals of carbon atoms up to a distance of about four bonds are split by coupling to ^{19}F.

^{13}C Chemical Shifts and ^{19}F-^{13}C Coupling Constants (δ in ppm, |J| in Hz)



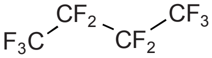


Estimation of ^{13}C Chemical Shifts of Linear Perfluoroalkanes (δ in ppm)

$$\delta = 124.8 + \sum Z_i$$

Increments Z_i for the CF_2 or CF_3 substituent in position:		
α	β	γ
-8.6	1.8	0.5

Example: Estimation of the chemical shifts in perfluorobutane



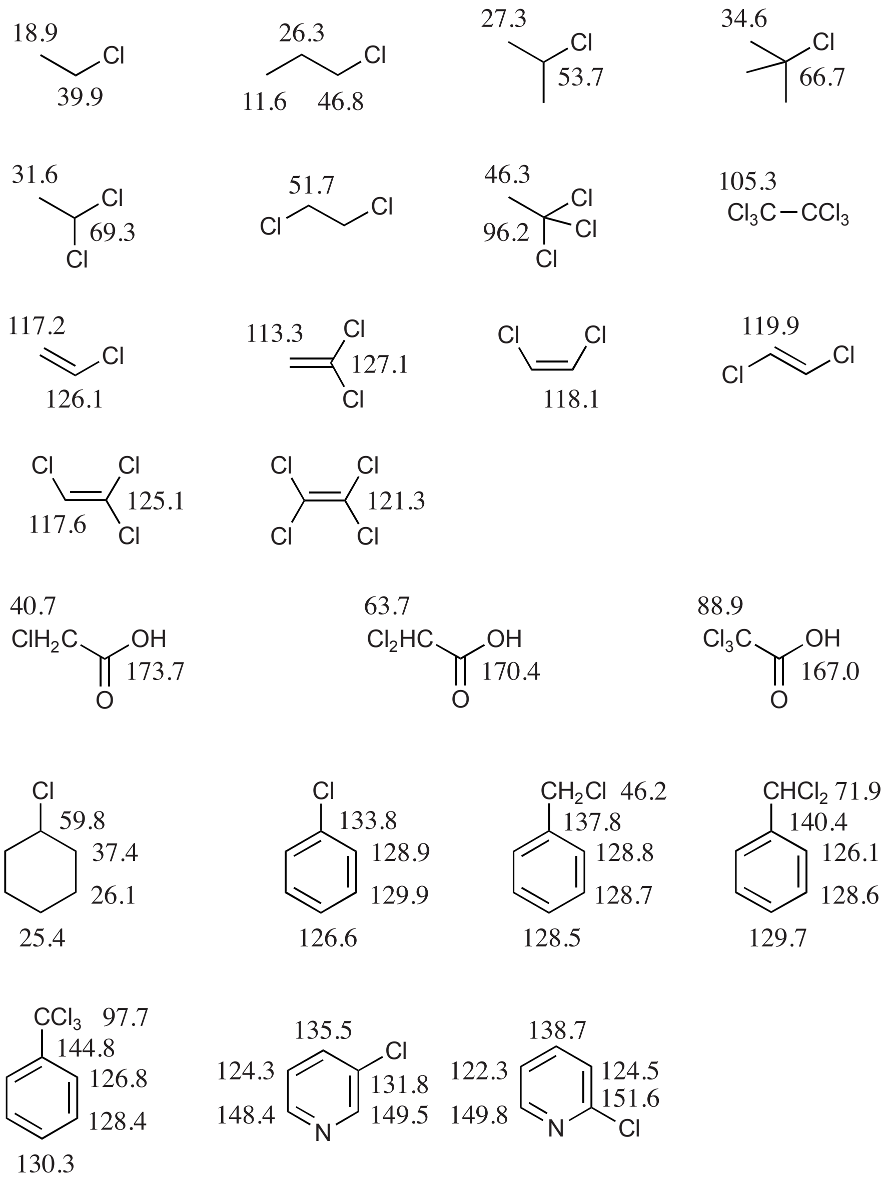
CF_3	base value	124.8
	1 α-CF_2	-8.6
	1 β-CF_2	1.8
	1 γ-CF_3	0.5
	estimated	118.5
	exp	118.5

CF_2	base value	124.8
	1 α-CF_3	-8.6
	1 α-CF_2	-8.6
	1 β-CF_3	1.8
	estimated	109.4
	exp	109.3

4.7.2 Chloro Compounds

^{13}C Chemical Shifts of Chloro Compounds (δ in ppm)

25.6	54.0	77.2	96.1
CH_3Cl	CH_2Cl_2	$CHCl_3$	CCl_4



4.7.3 Bromo Compounds

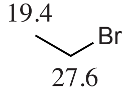
^{13}C Chemical Shifts of Bromo Compounds (δ in ppm)

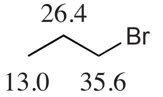
9.6
CH_3Br

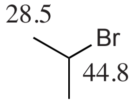
21.4
CH_2Br_2

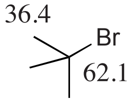
12.1
$CHBr_3$

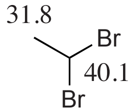
-28.7
CBr_4

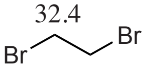


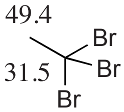




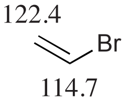


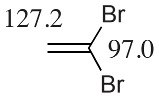


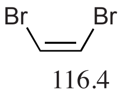


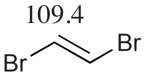


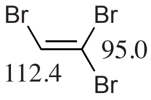
53.4
Br_3C-CBr_3

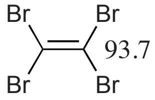


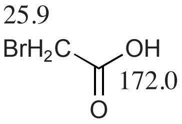


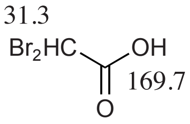


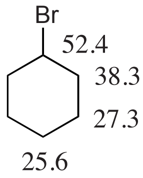


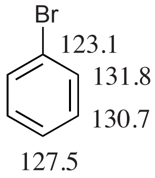


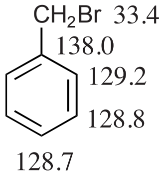


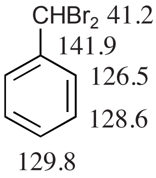


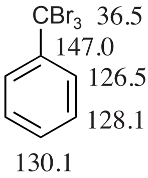


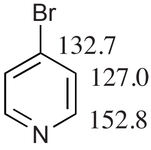


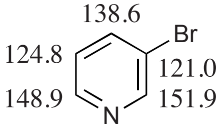


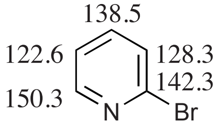












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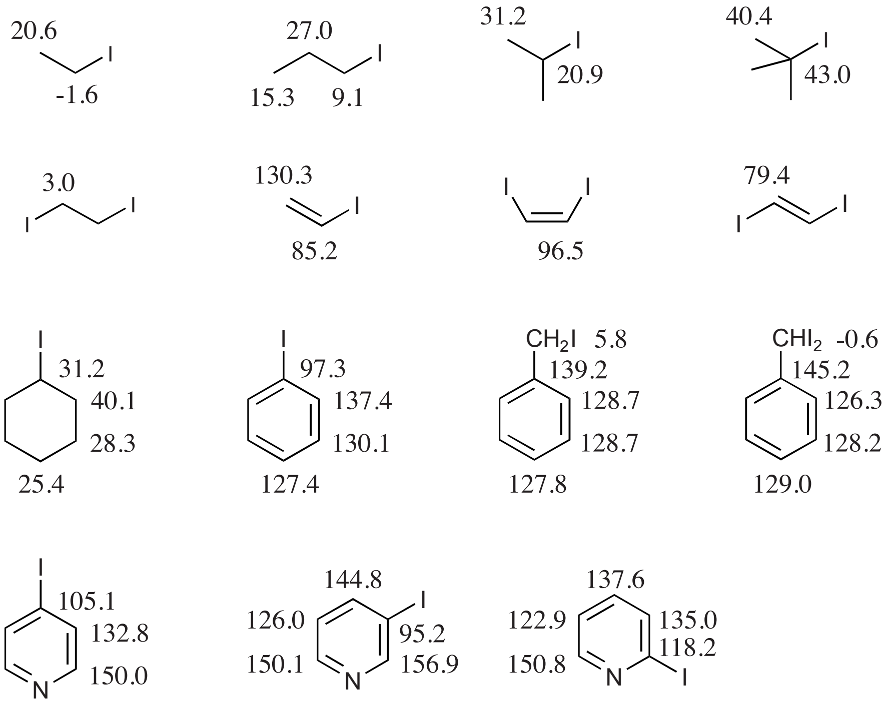
4.7.4 Iodo Compounds

^{13}C Chemical Shifts of Iodo Compounds (δ in ppm)

-24.0	-54.0	-139.9	-292.5
CH_3I	CH_2I_2	CHI_3	CI_4



4.7.5 References

[1] G.R. Somayajulu, J.R. Kennedy, T.M. Vickrey, B.J. Zwolinski, Carbon-13 chemical shifts for 70 halomethanes, *J. Magn. Reson.* **1979**, *33*, 559.

[2] A. Fürst, W. Robien, E. Pretsch, A comprehensive parameter set for the prediction of the ^{13}C NMR chemical shifts of *sp*^{3}-hybridized carbon atoms in organic compounds, *Anal. Chim. Acta* **1990**, *233*, 213.

[3] D.W. Ovenall, J.J. Chang, Carbon-13 NMR of fluorinated compounds using wide-band fluorine decoupling, *J. Magn. Reson.* **1977**, *25*, 361.

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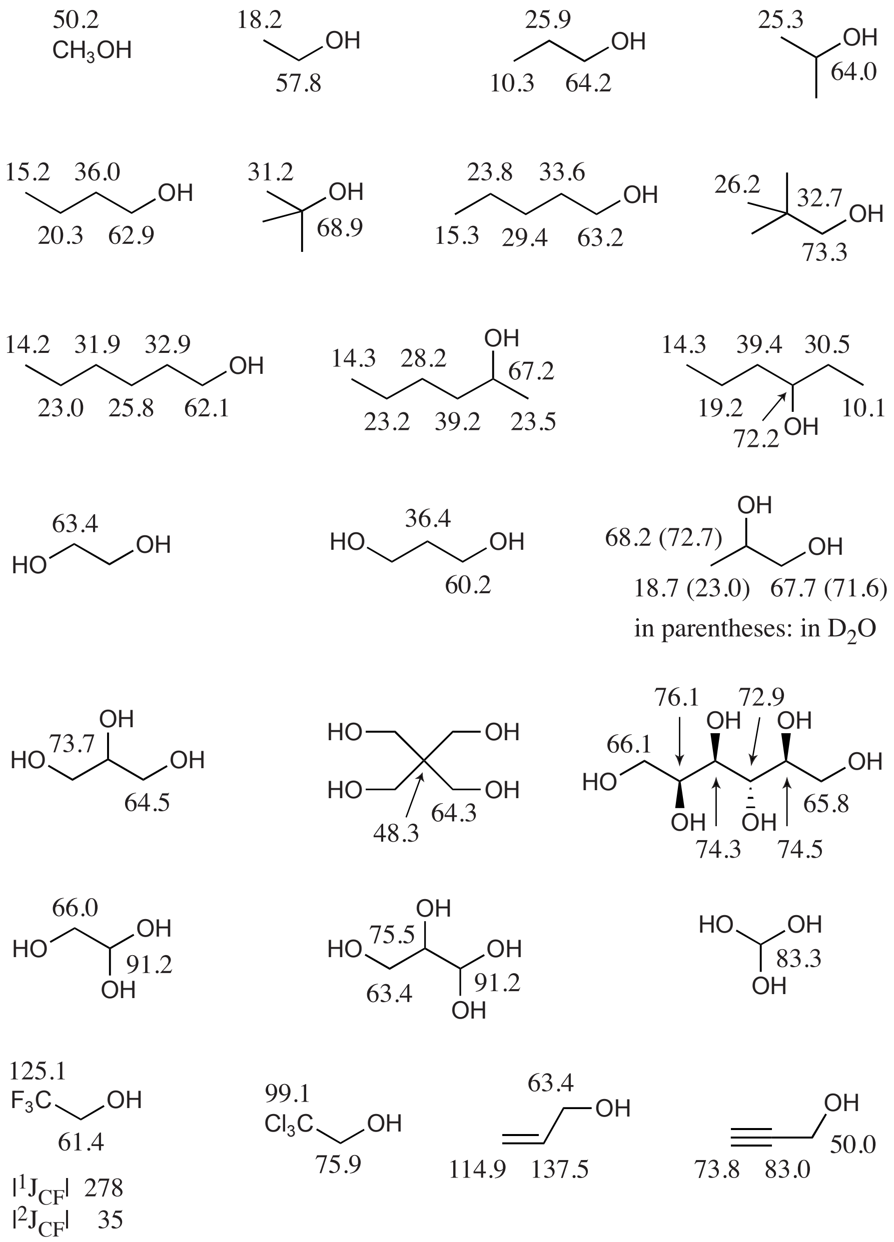
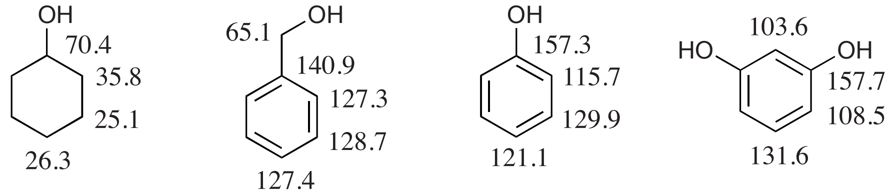
4.8 Alcohols, Ethers, and Related Compounds

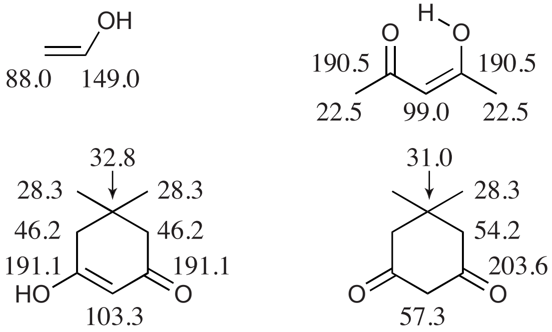
4.8.1 Alcohols

^{13}C Chemical Shifts of Alcohols (δ in ppm)



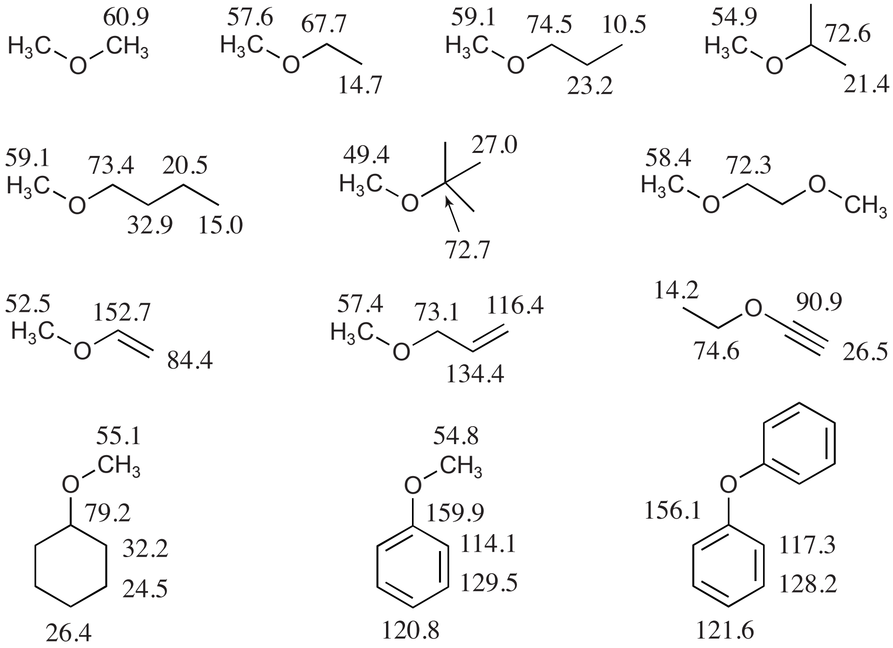


^{13}C Chemical Shifts of Enols (δ in ppm)

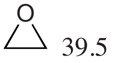


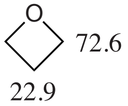
4.8.2 Ethers

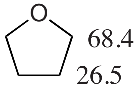
^{13}C Chemical Shifts of Ethers (δ in ppm)

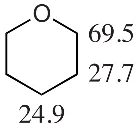


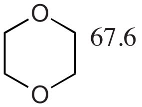
13C Chemical Shifts of Cyclic Ethers (δ in ppm)

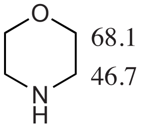


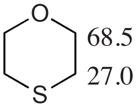


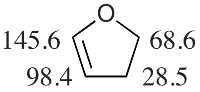


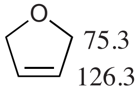


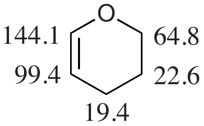


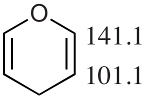




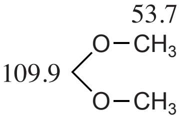


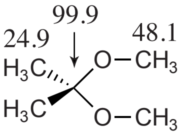


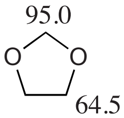


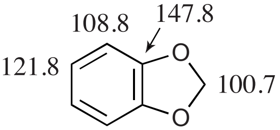


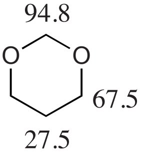
13C Chemical Shifts of Acetals, Ketals, and Ortho Esters (δ in ppm)

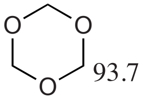


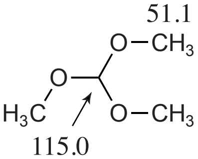


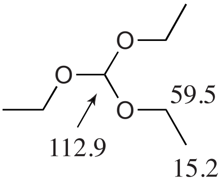


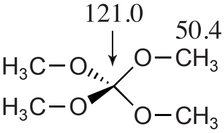


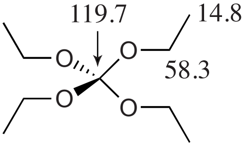










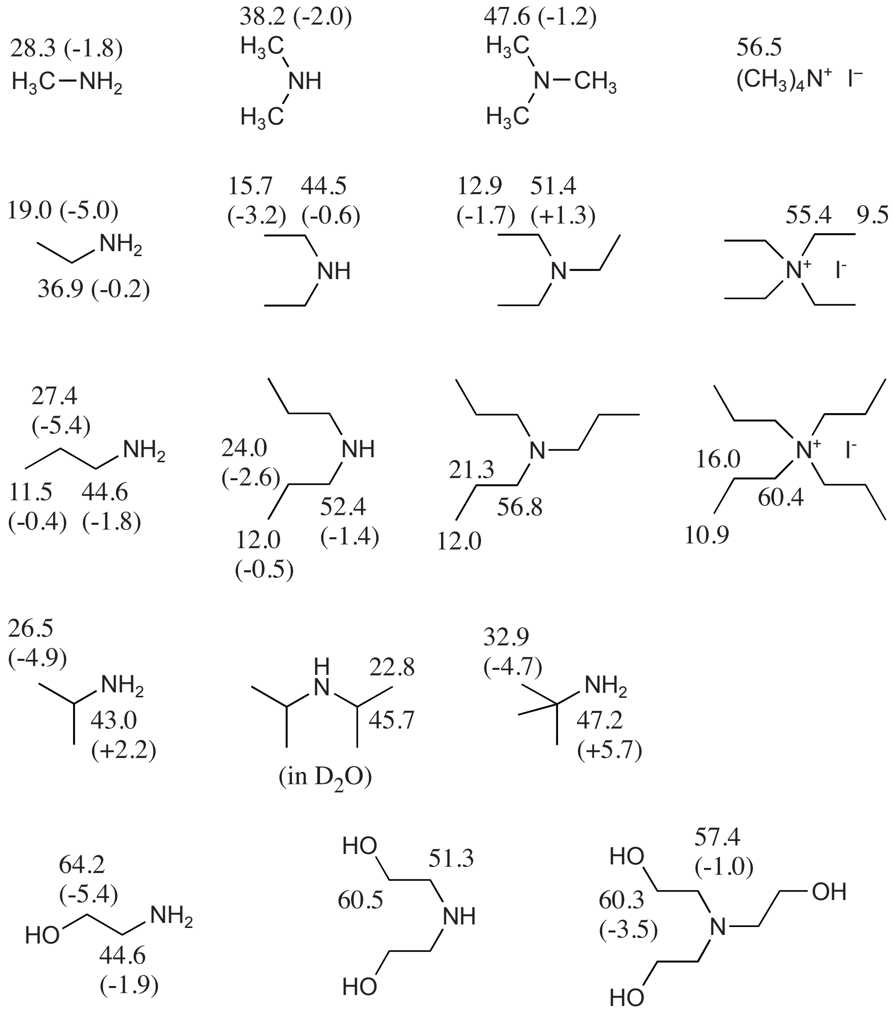


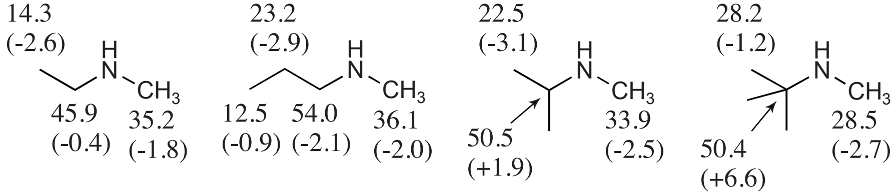
4.9 Nitrogen Compounds

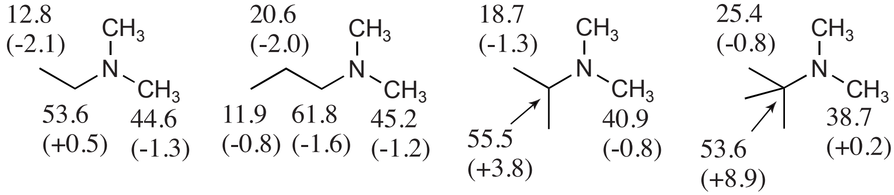
4.9.1 Amines

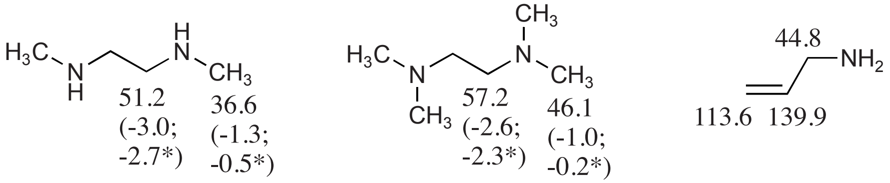
^{13}C Chemical Shifts of Amines and Ammonium Salts (δ in ppm)

The protonation of amines causes a shielding of the carbon atoms in the vicinity of the nitrogen. This shielding amounts to -2 ppm for an α carbon atom, -3 to -4 for a β carbon, and -0.5 to -1.0 ppm for a γ carbon. The most frequent exceptions occur in branched systems: Tertiary and quaternary carbon atoms in the α-position are generally deshielded by protonation of the nitrogen (Δδ = +0.5 to +9 ppm) [1]. In the following, shifts induced by protonation ($\delta_{\text{amine hydrochloride}}$ - δ_{amine}, measured in D_2O) are given in parentheses.

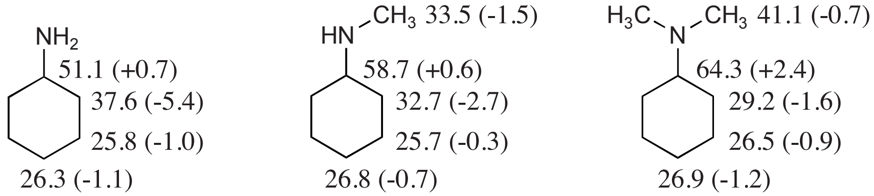


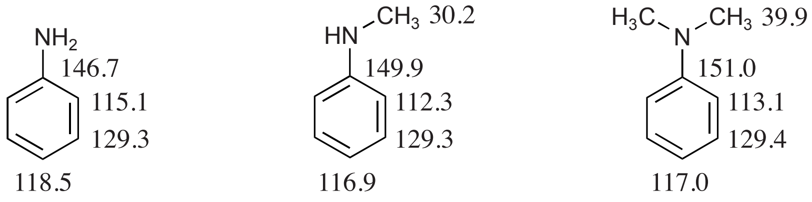


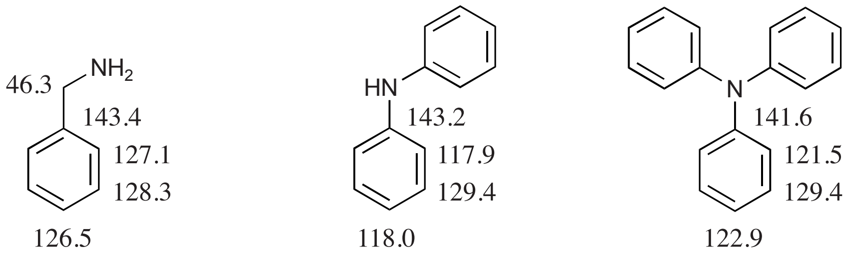




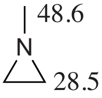
* doubly protonated form   * doubly protonated form

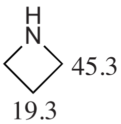


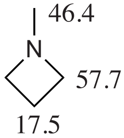


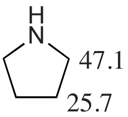


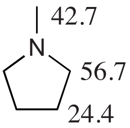
[13]C Chemical Shifts of Cyclic Amines (δ in ppm)

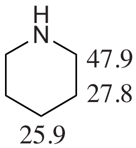


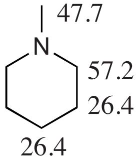


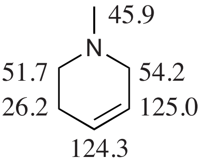


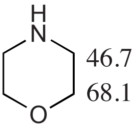


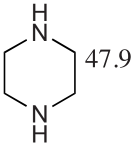


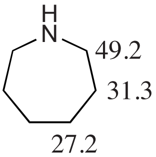








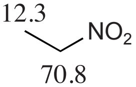


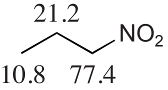


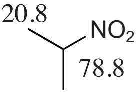
4.9.2 Nitro and Nitroso Compounds

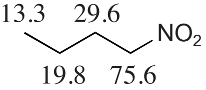
[13]C Chemical Shifts of Nitro and Nitroso Compounds (δ in ppm)

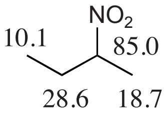
61.2
CH_3NO_2

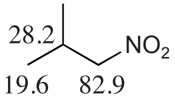


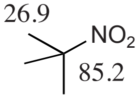


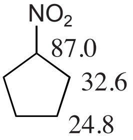


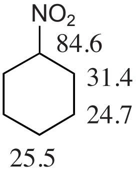


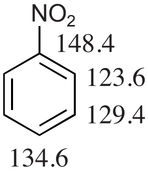


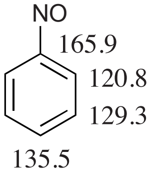






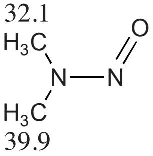


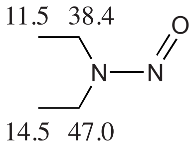


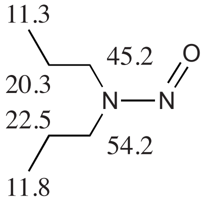


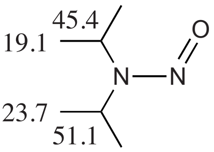
4.9.3 Nitrosamines and Nitramines

^{13}C Chemical Shifts of Nitrosamines (δ in ppm)

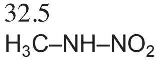


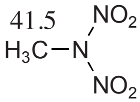


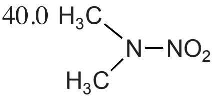




^{13}C Chemical Shifts of Nitramines (δ in ppm)

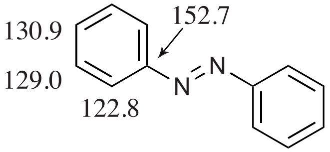


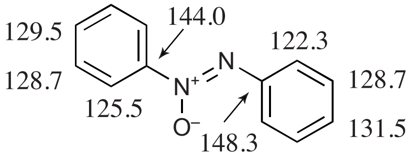




4.9.4 Azo and Azoxy Compounds

^{13}C Chemical Shifts of Azo and Azoxy Compounds (δ in ppm)

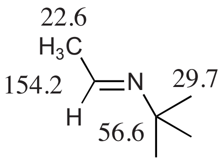


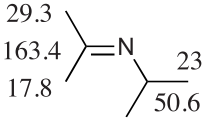


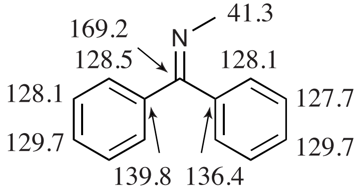
N

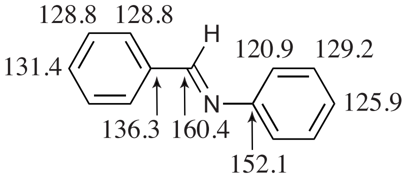
4.9.5 Imines and Oximes

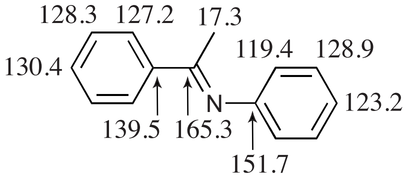
^{13}C Chemical Shifts of Imines and Oximes (δ in ppm)

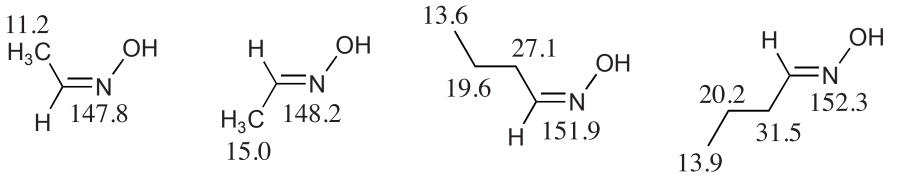


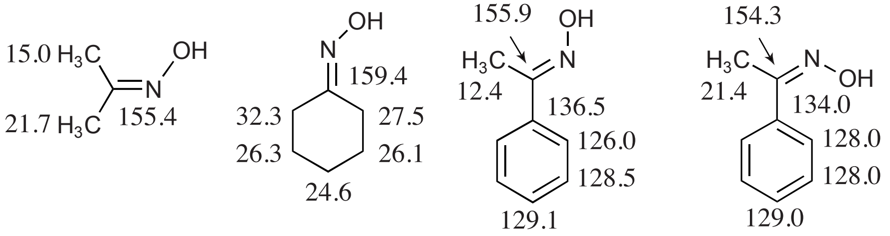






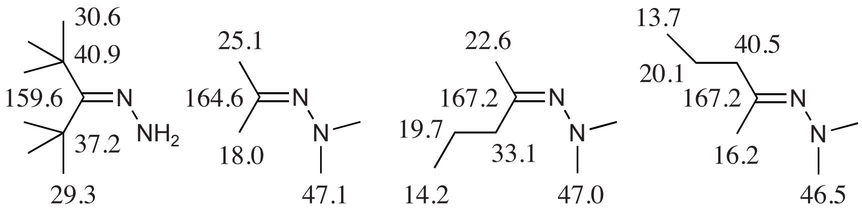


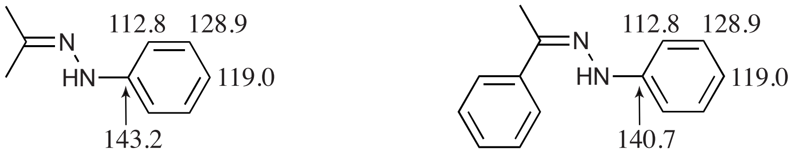


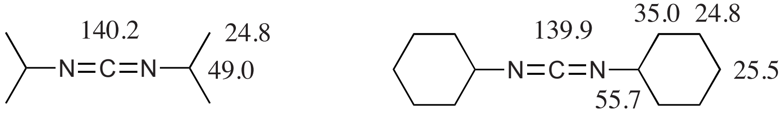


4.9.6 Hydrazones and Carbodiimides

^{13}C Chemical Shifts of Hydrazones and Carbodiimides (δ in ppm)

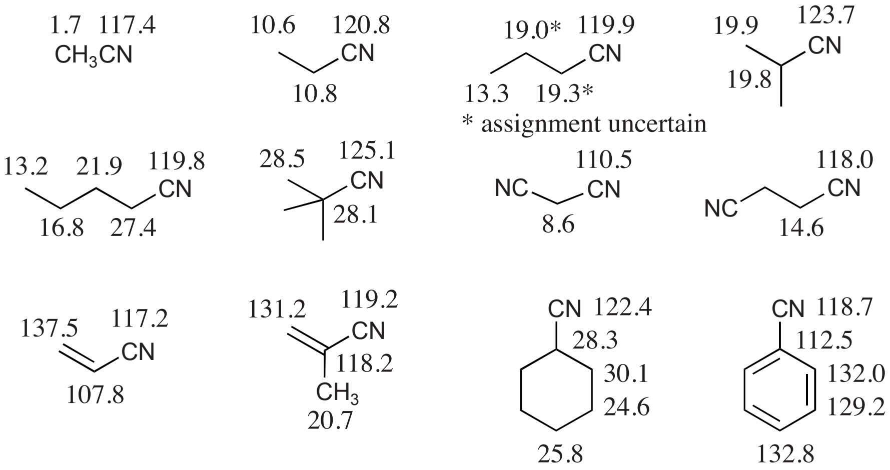






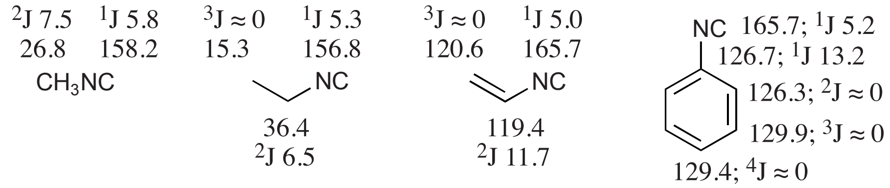
4.9.7 Nitriles and Isonitriles

¹³C Chemical Shifts of Nitriles (δ in ppm)



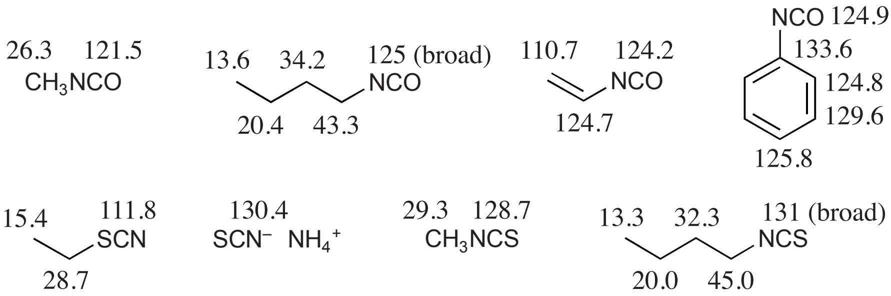
¹³C Chemical Shifts and ¹³C–¹⁴N Couplings of Isonitriles (δ in ppm, |J| in Hz)

Because of the symmetrical electron distribution around the nitrogen atom, the ^{13}C-^{14}N coupling can be observed in the ^{13}C NMR spectra of isonitriles, leading to triplets with intensities of 1:1:1 (spin quantum number of ^{14}N: I = 1, natural abundance, 99.6%).



4.9.8 Isocyanates, Thiocyanates, and Isothiocyanates

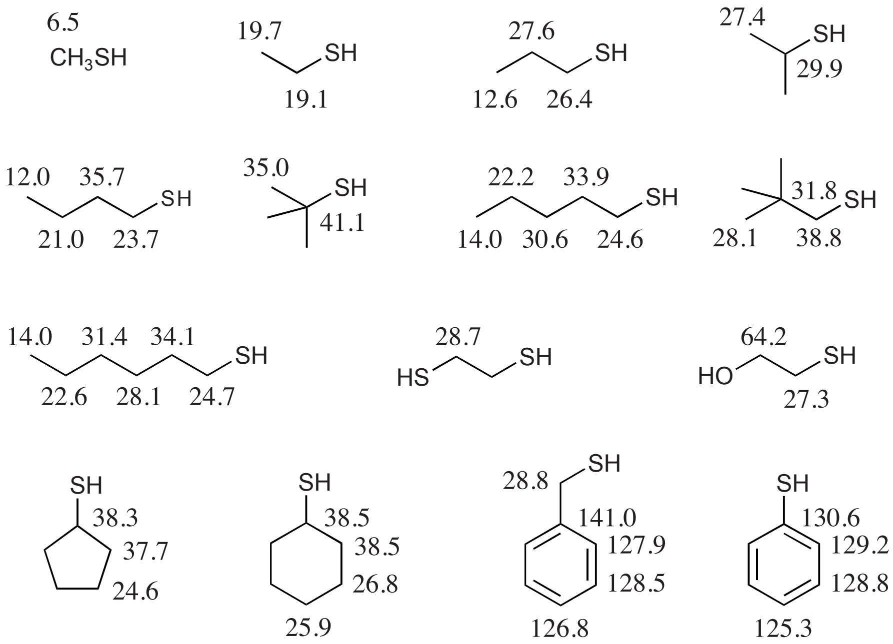
¹³C Chemical Shifts (δ in ppm)



4.10 Sulfur Compounds

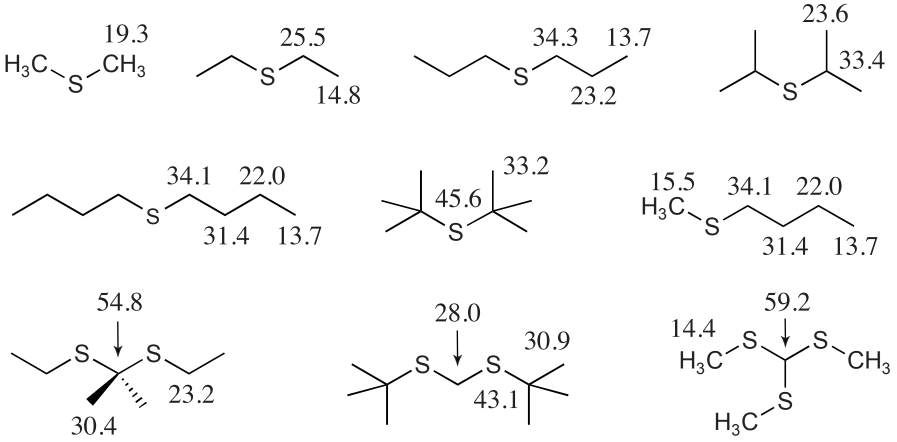
4.10.1 Thiols

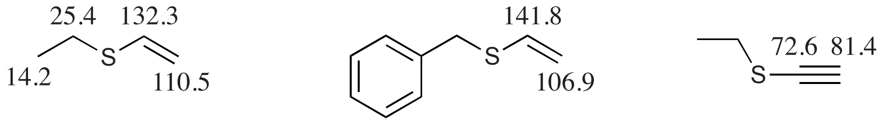
¹³C Chemical Shifts (δ in ppm)

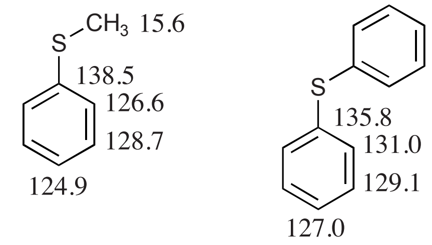


4.10.2 Sulfides

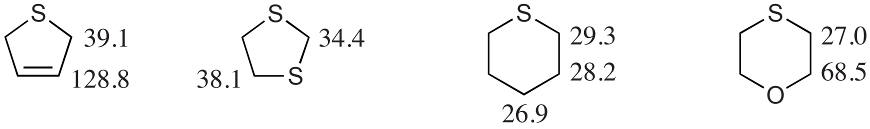
¹³C Chemical Shifts (δ in ppm)

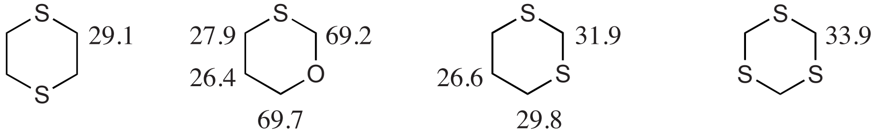






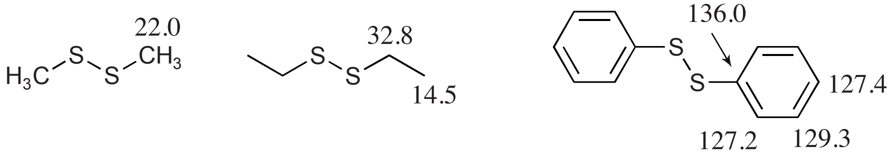
^{13}C Chemical Shifts of Cyclic Sulfides (δ in ppm)



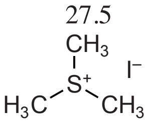


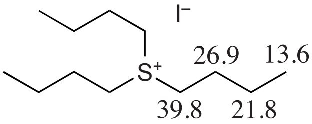
4.10.3 Disulfides and Sulfonium Salts

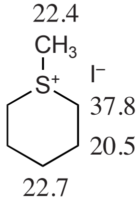
^{13}C Chemical Shifts of Disulfides (δ in ppm)

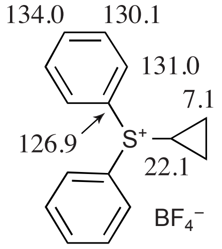


^{13}C Chemical Shifts of Sulfonium Salts (δ in ppm)



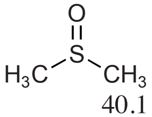


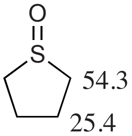


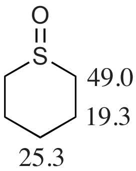


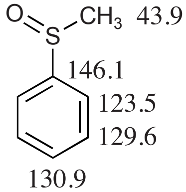
4.10.4 Sulfoxides and Sulfones

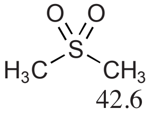
^{13}C Chemical Shifts of Sulfoxides and Sulfones (δ in ppm)

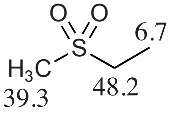


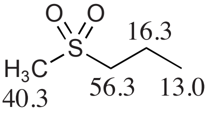


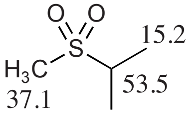


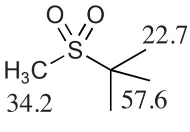


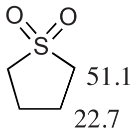


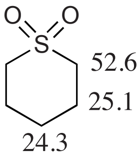


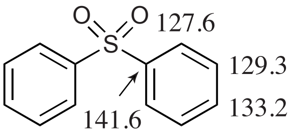






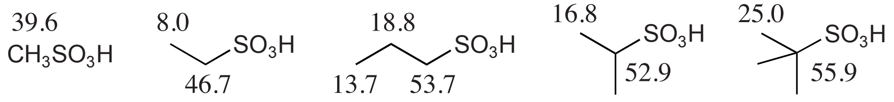


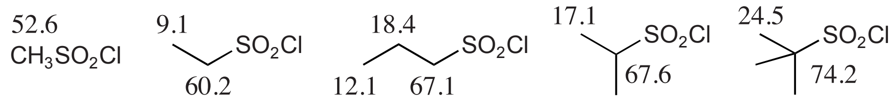


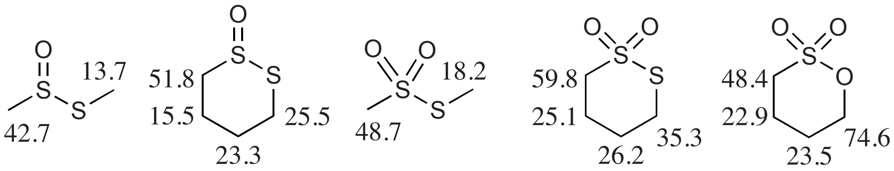


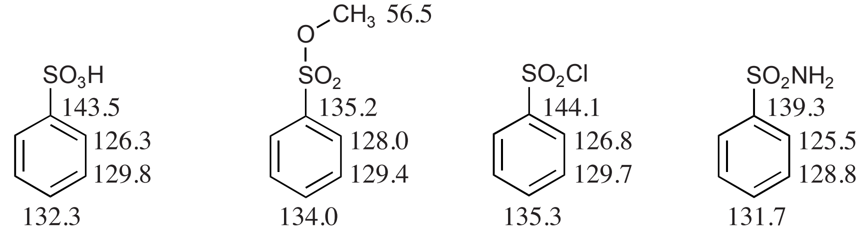
4.10.5 Sulfonic and Sulfinic Acids and Derivatives

^{13}C Chemical Shifts of Sulfonic and Sulfinic Acids and Derivatives (δ in ppm)





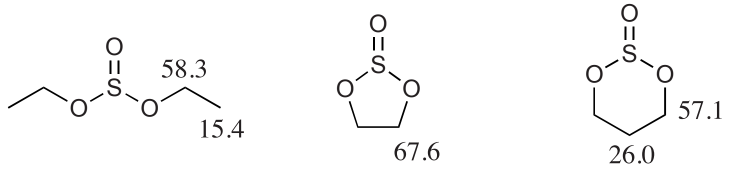


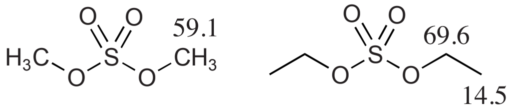


S

4.10.6 Sulfurous and Sulfuric Acid Derivatives

^{13}C Chemical Shifts of Sulfurous and Sulfuric Acid Derivatives (δ in ppm)





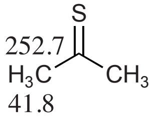
4.10.7 Sulfur-Containing Carbonyl Derivatives

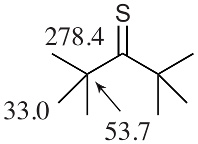
^{13}C Chemical Shifts (δ in ppm)

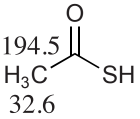
The ^{13}C chemical shifts of thiocarbonyl groups are higher by about 30 ppm than those of the corresponding carbonyl groups:

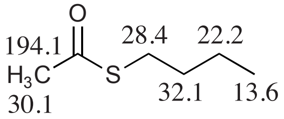
$$\delta_{C=S} \approx 1.5 \times \delta_{C=O} - 57.5$$

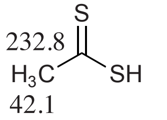
Carbonyl groups of thiocarboxylic acids and their esters are deshielded by about 20 ppm with respect to the corresponding oxygen compounds.

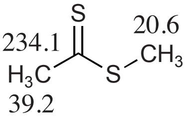


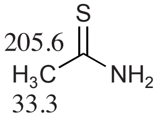


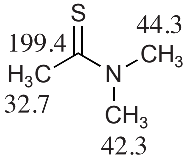


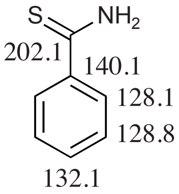


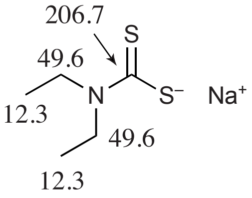




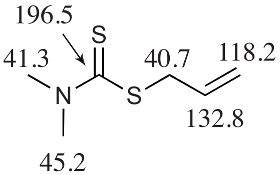


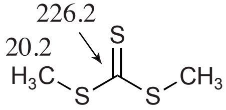


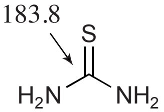


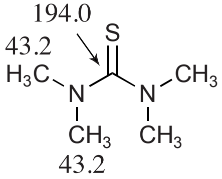


(in D_2O)









4.11 Carbonyl Compounds

4.11.1 Aldehydes

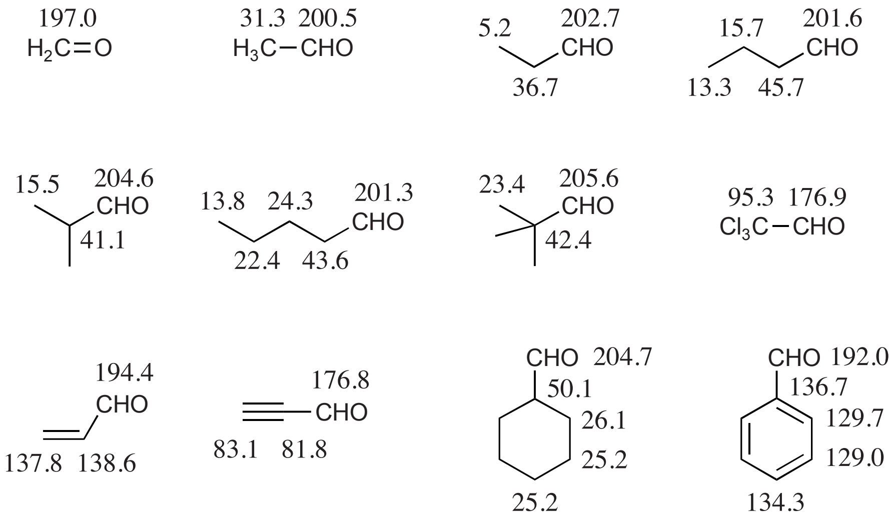
Additivity Rule for Estimating the ^{13}C Chemical Shifts of Aldehyde Carbon Atoms (δ in ppm)

$$\delta_{C=O} = 193.0 + \sum Z_i$$

$$-C_\beta-C_\alpha-CHO$$

Substituent i	Z_α	Z_β
–C≤	6.5	2.6
$-CH=CH_2$	-0.8	0.0
$-CH=CH-CH_3$	0.2	0.0
–phenyl	-1.2	0.0

^{13}C Chemical Shifts of Aldehydes (δ in ppm)



4.11.2 Ketones

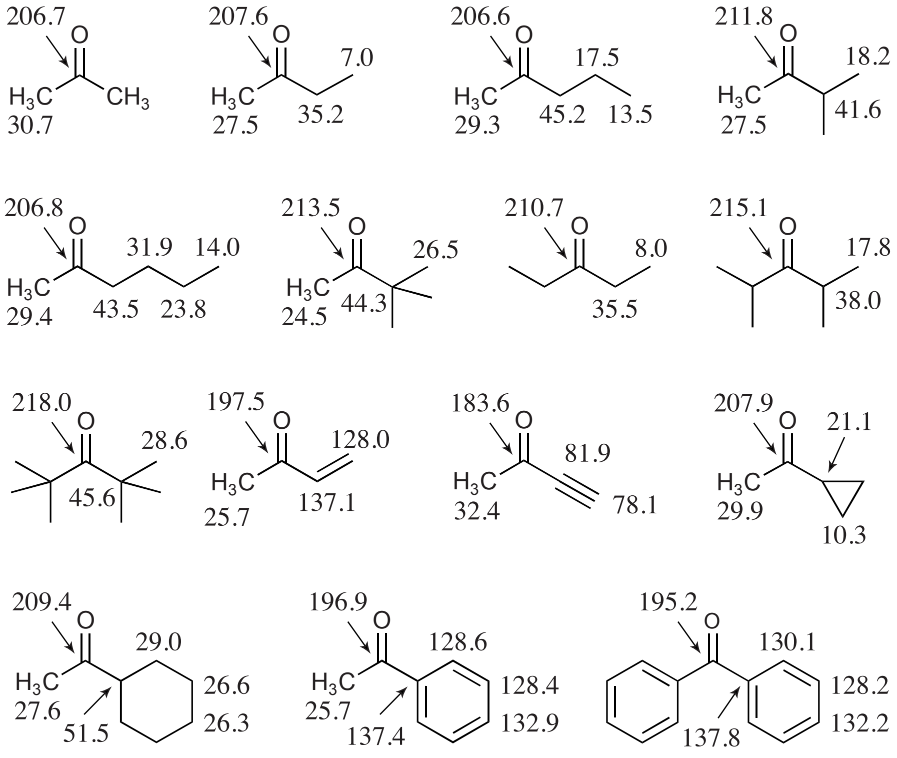
Additivity Rule for Estimating the ^{13}C Chemical Shifts of Ketone Carbon Atoms (δ in ppm)

$$\delta_{C=O} = 193.0 + \sum Z_i$$

$$-C_\beta-C_\alpha-\overset{\overset{\displaystyle O}{\|}}{C}-C_\alpha-C_\beta-$$

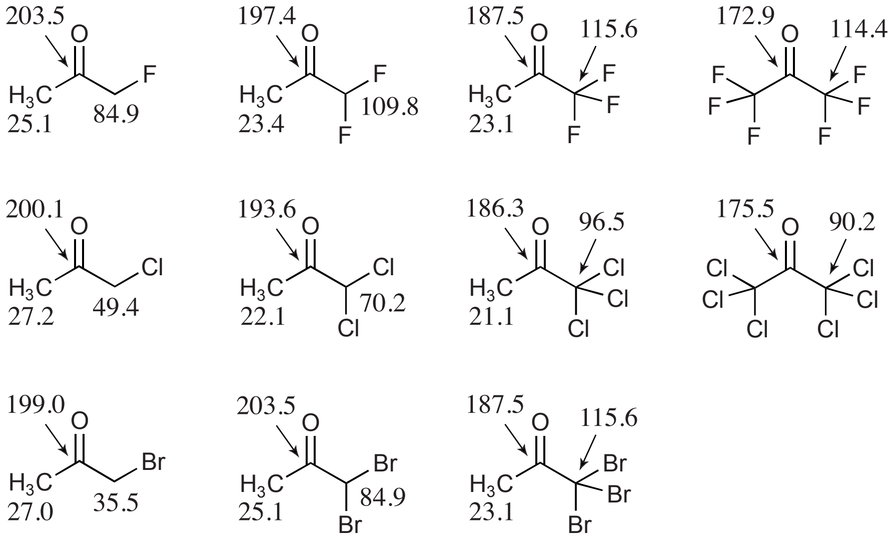
Substituent i	Z_α	Z_β
–C⋲	6.5	2.6
$-CH{=}CH_2$	-0.8	0.0
$-CH{=}CH-CH_3$	0.2	0.0
–phenyl	-1.2	0.0

^{13}C Chemical Shifts of Ketones (δ in ppm)

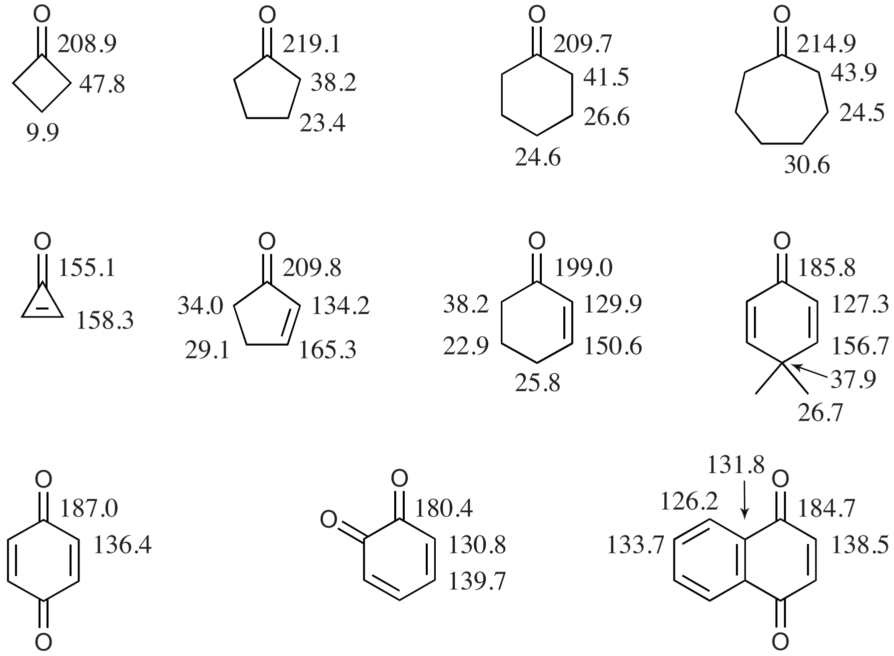


C = X

^{13}C Chemical Shifts of Halogenated Aliphatic Ketones (δ in ppm)

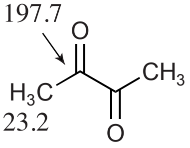


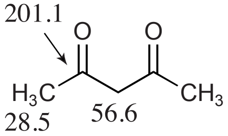
^{13}C Chemical Shifts of Cyclic Ketones and Quinones (δ in ppm)

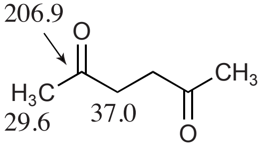


C = X

¹³C Chemical Shifts of Diketones (δ in ppm)







Enol form: see Chapter 4.8.1

4.11.3 Carboxylic Acids

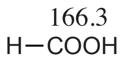
Additivity Rule for Estimating the ¹³C Chemical Shifts of Carboxyl Carbon Atoms (δ in ppm)

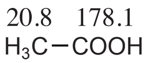
$$\delta_{C=O} = 166.0 + \sum Z_i$$

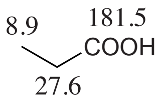
$$–C_\gamma–C_\beta–C_\alpha–COOH$$

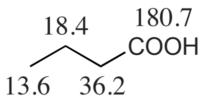
Substituent i	Z_α	Z_β	Z_γ
–C⋚	12.0	3.0	-1.0
–CH=CH_2	5.0	0.5	-1.5
–phenyl	6.0	1.0	-2.0

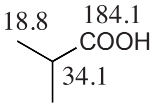
¹³C Chemical Shifts of Carboxylic Acids (δ in ppm)

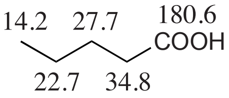


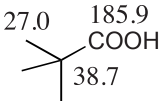


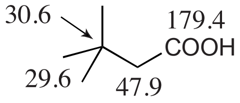


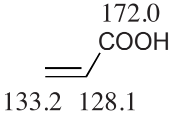


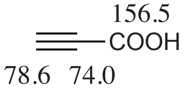


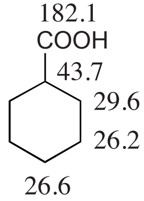


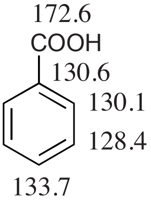








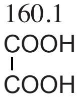
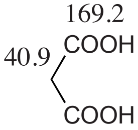
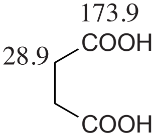
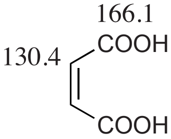
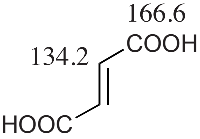




^{13}C Chemical Shifts of Halogenated Carboxylic Acids (δ in ppm)

115.0 163.0	40.7 173.7	63.7 170.4	88.9 167.1
$F_3C-COOH$	$ClH_2C-COOH$	$Cl_2HC-COOH$	$Cl_3C-COOH$

^{13}C Chemical Shifts of Dicarboxylic Acids (δ in ppm)

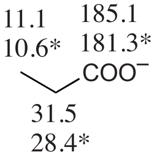






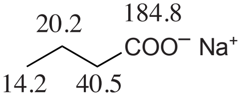
^{13}C Chemical Shifts of Carboxylate Anions (δ in ppm)

Measured in water unless indicated otherwise.

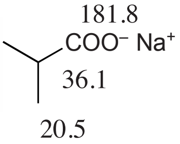
171.3
$H-COO^-$

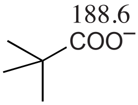
24.4 182.6
20.8* 177.6*
H_3C-COO^-
* in $CDCl_3$

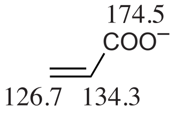

* in $CDCl_3$/DMSO

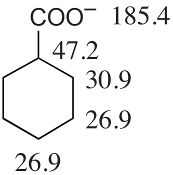

(in DMSO)

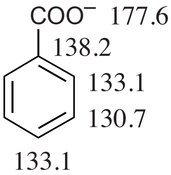
C = X


(in DMSO)









45.0 175.9	65.6 171.8	96.2 167.6
ClH_2C-COO^-	$Cl_2HC-COO^-$	Cl_3C-COO^-

4.11.4 Esters and Lactones

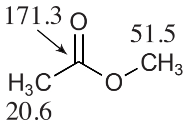
Additivity Rule for Estimating the ^{13}C Chemical Shifts of Ester Carbon Atoms (δ in ppm)

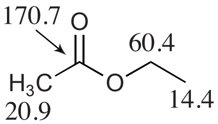
$$\delta_{C=O} = 166.0 + \sum Z_i$$

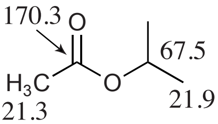
$$–C_\gamma–C_\beta–C_\alpha–COO–C_{\alpha'}–$$

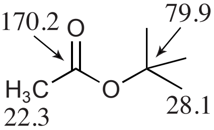
Substituent i	Z_α	Z_β	Z_γ	$Z_{\alpha'}$
–C≶	12.0	3.0	-1.0	-5.0
$–CH=CH_2$	5.0			-9.0
–phenyl	6.0	1.0		-8.0

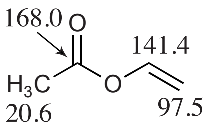
^{13}C Chemical Shifts of Acetic Acid Esters (δ in ppm)

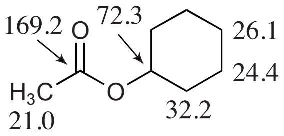


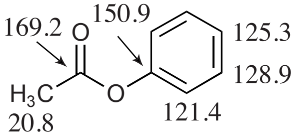




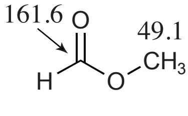


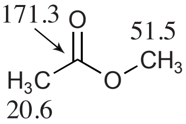


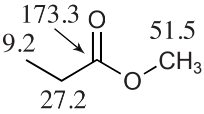


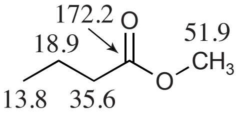


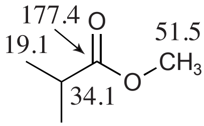
^{13}C Chemical Shifts of Methyl Esters (δ in ppm)

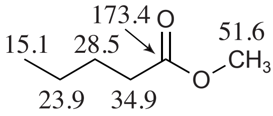


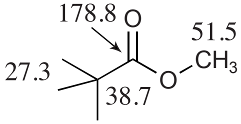


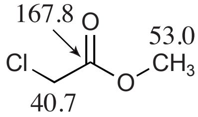


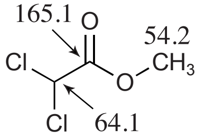


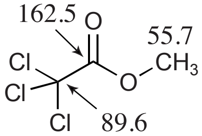


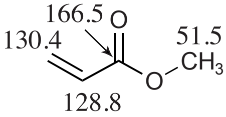


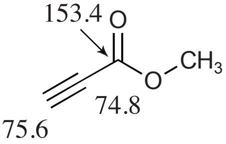


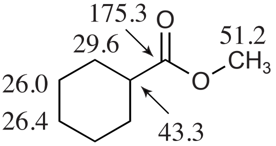


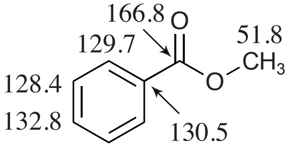


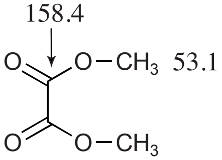


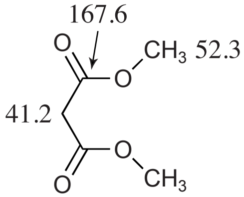


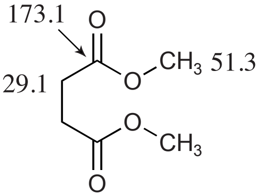


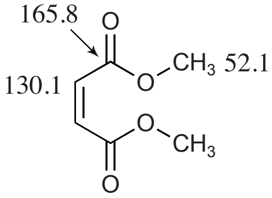


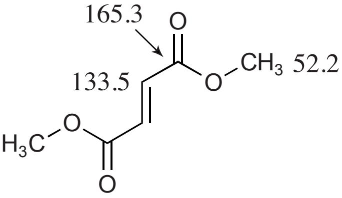


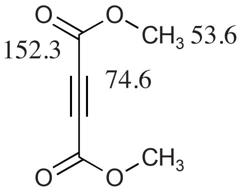






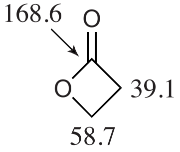


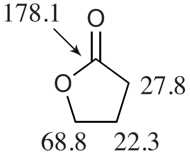


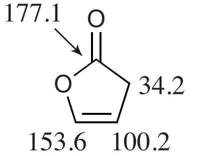


^{13}*C Chemical Shifts of Lactones (δ in ppm)*

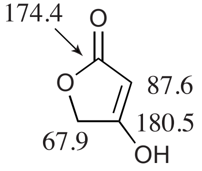
C = X

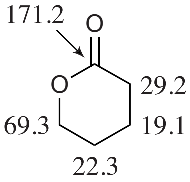


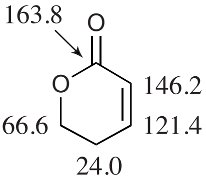


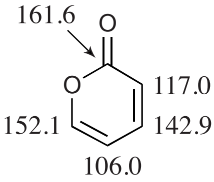


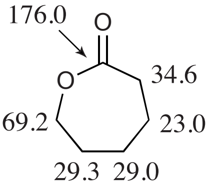
(in acetone)











4.11.5 Amides and Lactams

Additivity Rule for Estimating the ^{13}*C Chemical Shifts of Amide Carbon Atoms (δ in ppm)*

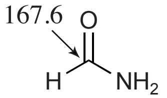
$$\delta_{C=O} = 166.0 + \sum Z_i$$

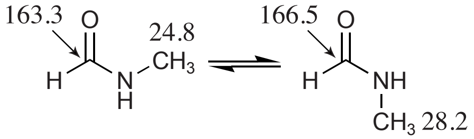
$$-C_\gamma-C_\beta-C_\alpha-CO-N\begin{matrix} C_{\alpha'}-C_{\beta'} \\ C_{\alpha'}-C_{\beta'} \end{matrix}$$

Substituent i	Z_α	Z_β	Z_γ	$Z_{\alpha'}$	$Z_{\beta'}$
–C⋜	7.7	4.5	-0.7	-1.5	-0.3
$-CH=CH_2$	3.3				
–phenyl	4.7			-4.5	

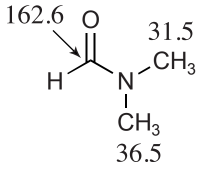
^{13}*C Chemical Shifts of Amides (δ in ppm)*

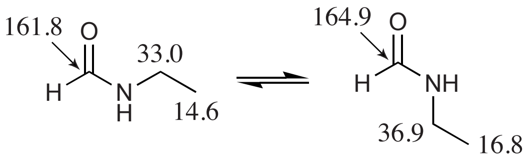
Formamides:



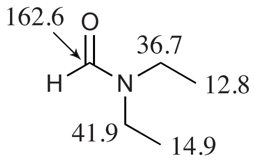


≈ 90% ≈ 10%

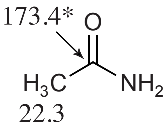




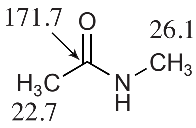
≈ 90% ≈ 10%

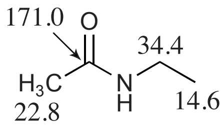


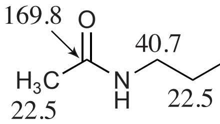
Primary and Secondary Acetamides:

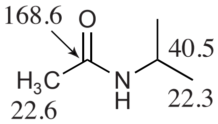


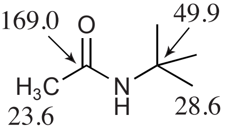
* in water: 177.0

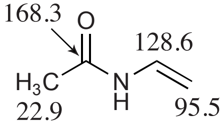


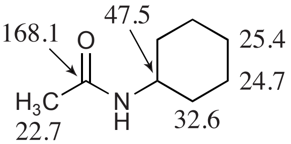


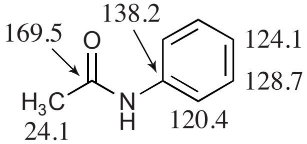




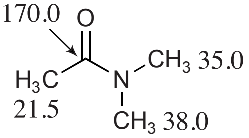


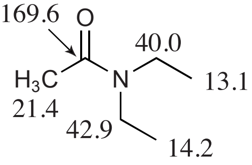


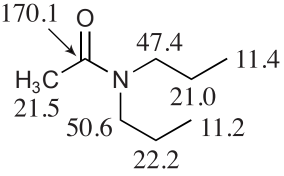


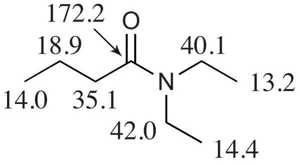


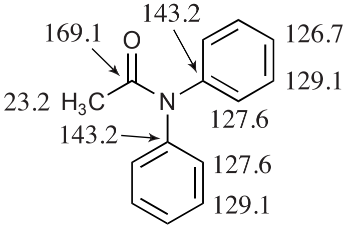
Tertiary Amides:





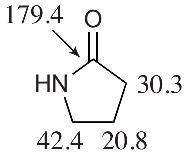


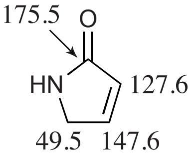


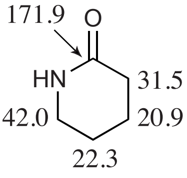


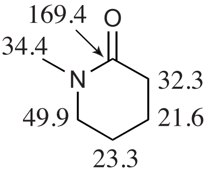
^{13}C Chemical Shifts of Lactams (δ in ppm)

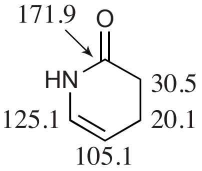
C = X

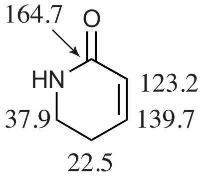


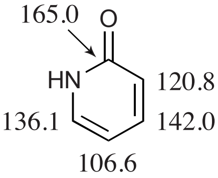


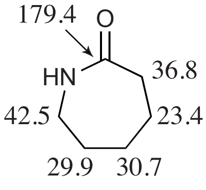






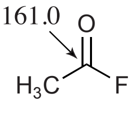


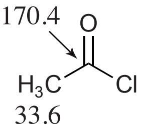


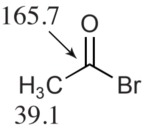


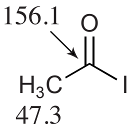
4.11.6 Miscellaneous Carbonyl Derivatives

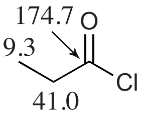
^{13}C Chemical Shifts of Carboxylic Acid Halides (δ in ppm)

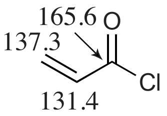


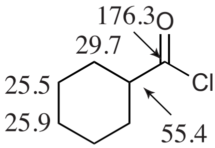


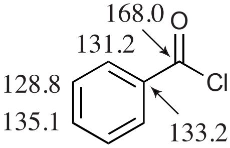




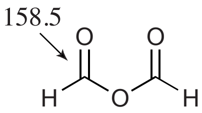


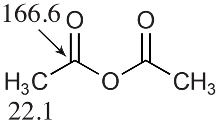


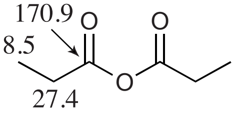


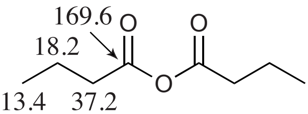


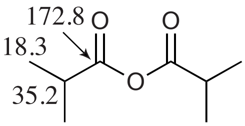
^{13}C Chemical Shifts of Carboxylic Acid Anhydrides (δ in ppm)

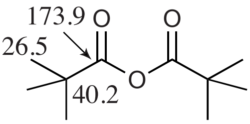


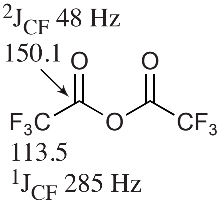


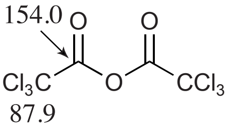


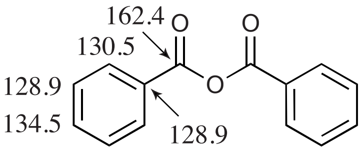


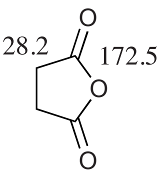


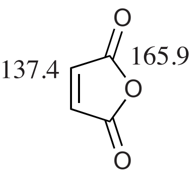


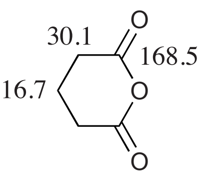


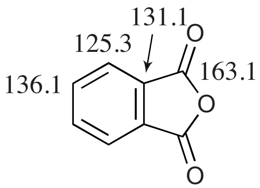




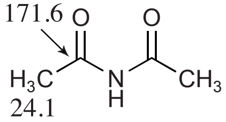


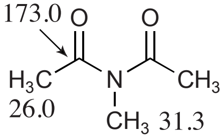


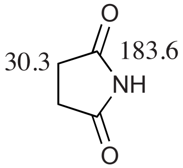


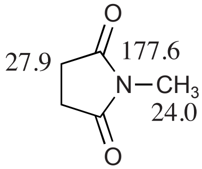


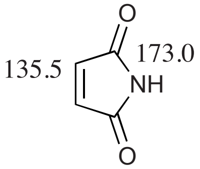
^{13}C Chemical Shifts of Carboxylic Acid Imides (δ in ppm)

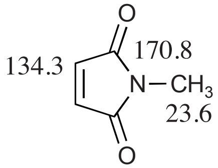


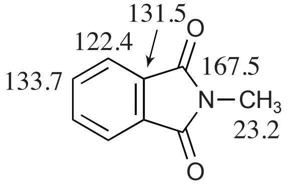




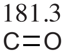


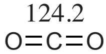


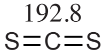


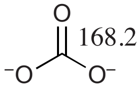


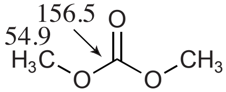
^{13}C Chemical Shifts of Carbonic Acid Derivatives (δ in ppm)

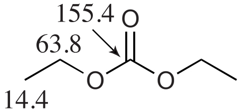


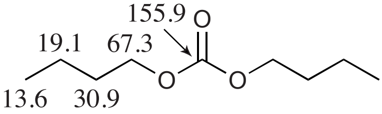


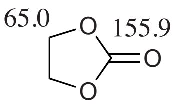


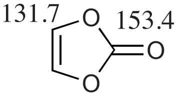


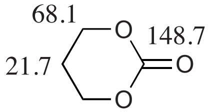


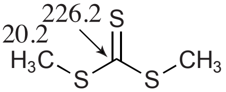


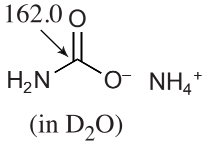


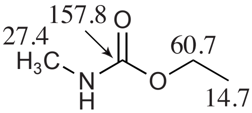


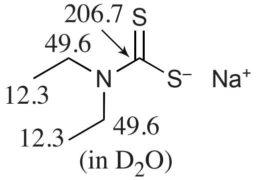




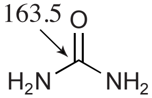


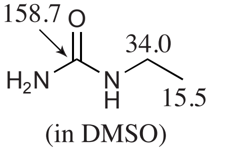


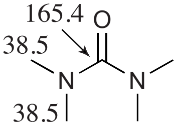


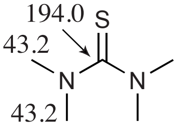


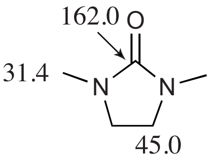
C = X

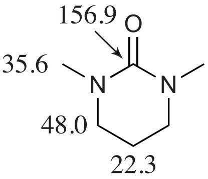


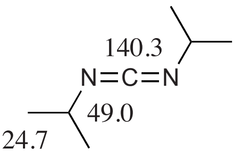








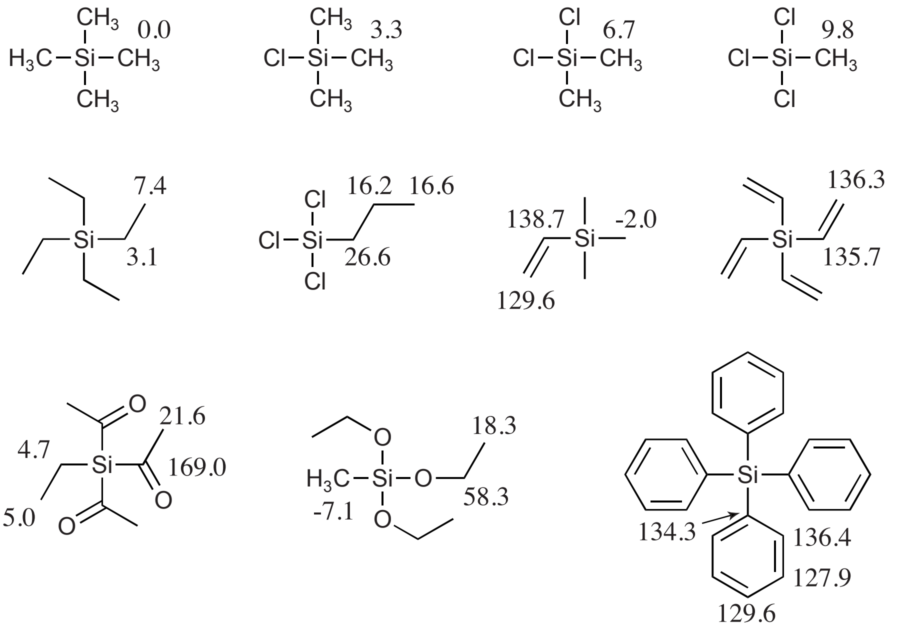




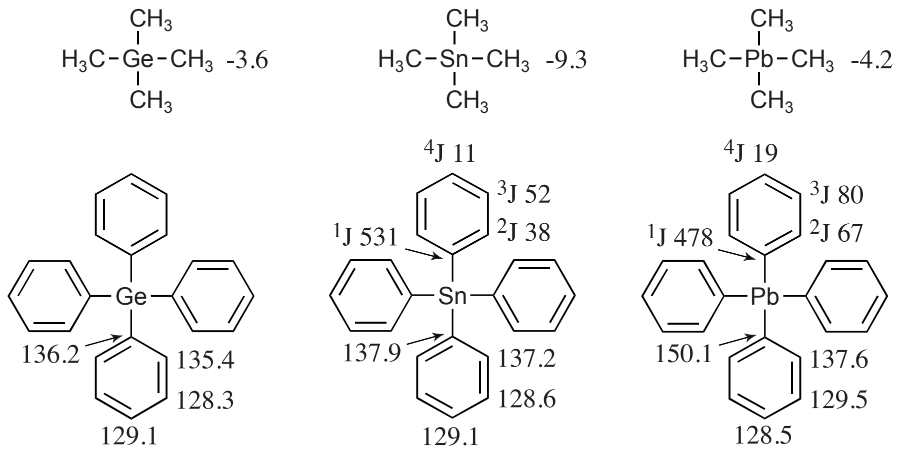
4.12 Miscellaneous Compounds

4.12.1 Compounds with Group IV Elements

^{13}C Chemical Shifts of Silicon Compounds (δ in ppm)



^{13}C Chemical Shifts and Coupling Constants of Germanium and Lead Compounds (δ in ppm, |J| in Hz)



Couplings with ^{119}Sn (8.58%, I = 1/2)

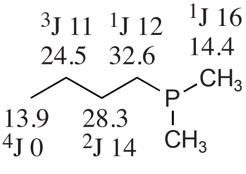
Couplings with ^{207}Pb (22.6%, I = 1/2)

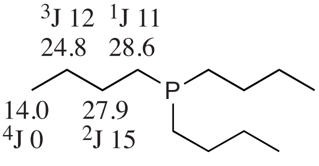
P Si

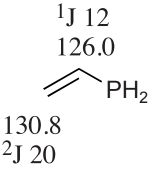
4.12.2 Phosphorus Compounds

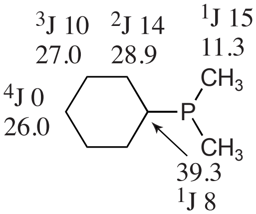
^{31}P (natural abundance, 100%) has a spin quantum number I of 1/2. Couplings to protons through up to 3–4 bonds are usually observed.

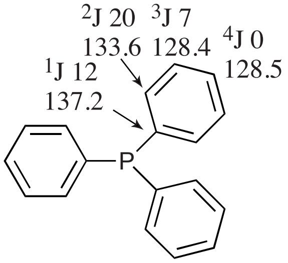
Phosphines and Phosphonium Compounds (δ in ppm, $|J_{^{31}P^{13}C}|$ in Hz)

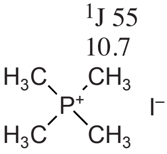


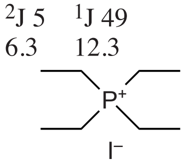


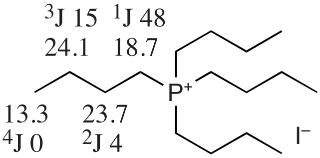


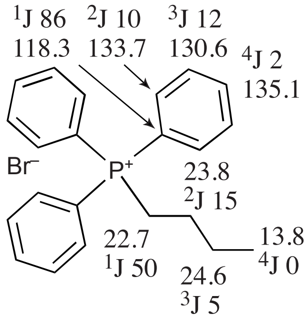


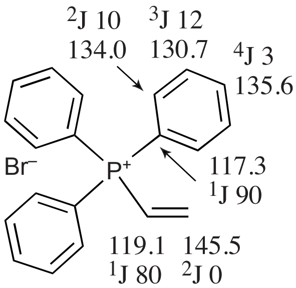


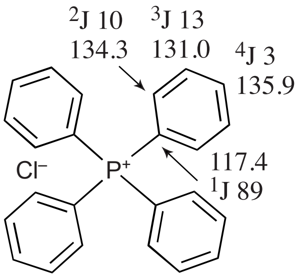




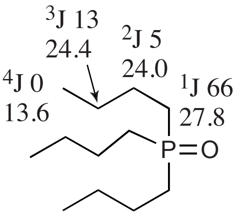


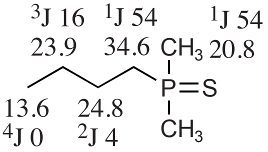


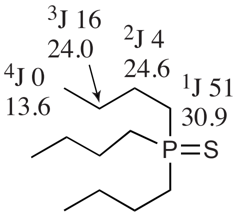


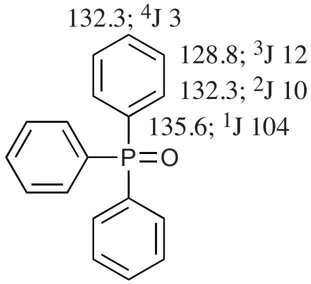


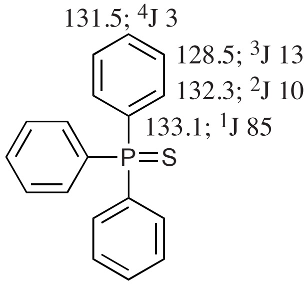
Phosphine Oxides and Sulfides (δ in ppm, $|J_{^{31}P^{13}C}|$ in Hz)



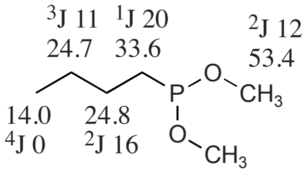


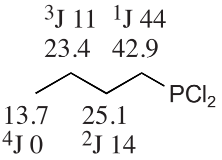


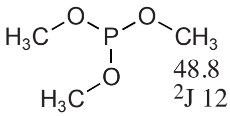


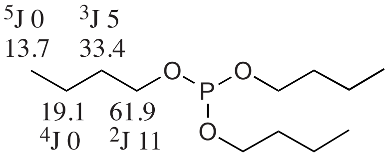


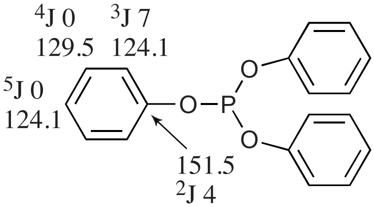
Phosphinic and Phosphorous Acid Derivatives (δ in ppm, $|J_{^{31}P^{13}C}|$ in Hz)

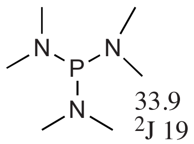


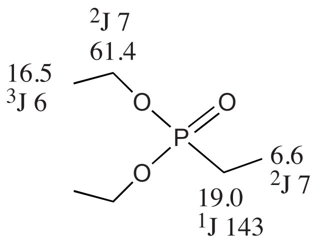


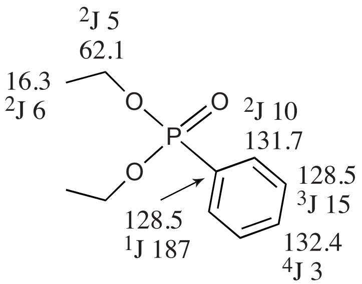




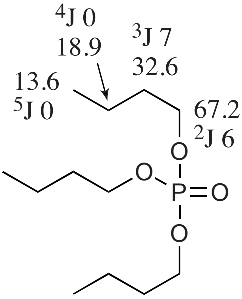


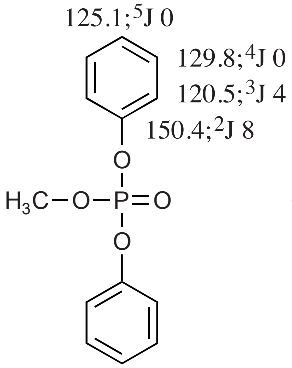


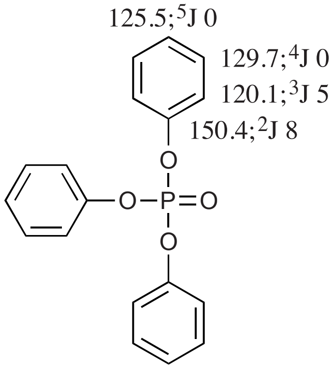




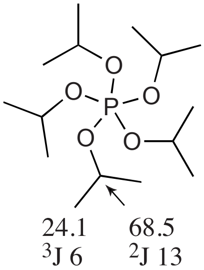
Phosphoric Acid Derivatives (δ in ppm, $|J_{^{31}P^{13}C}|$ in Hz)

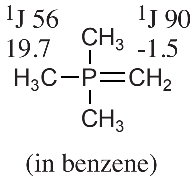


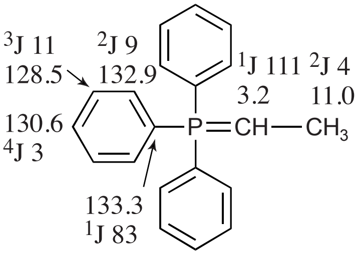




Phosphoranes and Phosphorus Ylides (δ in ppm, $|J_{^{31}P^{13}C}|$ in Hz)

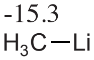


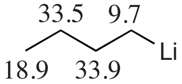


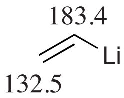


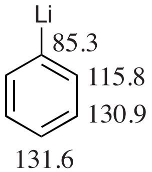
4.12.3 Miscellaneous Organometallic Compounds

Lithium Compounds (δ in ppm)

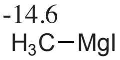


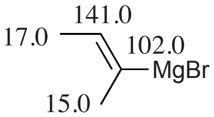


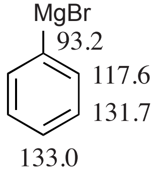




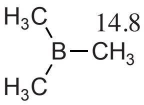
Magnesium Compounds (δ in ppm)

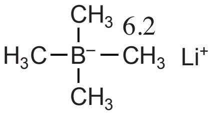


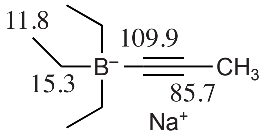


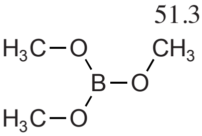


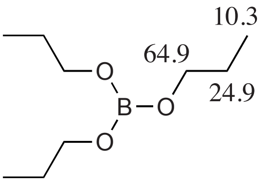
Boron Compounds (δ in ppm, |J| in Hz)

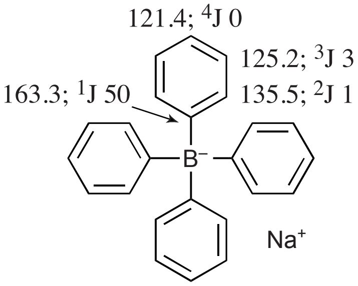






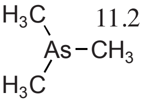


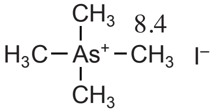


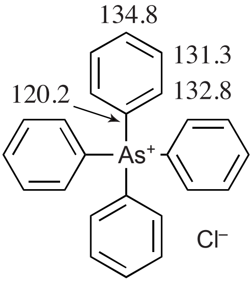


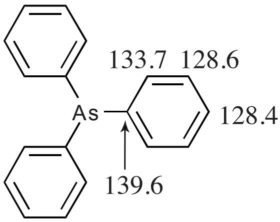
Couplings with ^{11}B
(80.4%, I = 3/2)

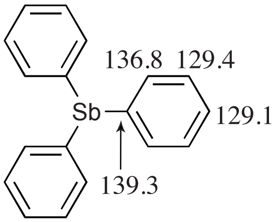
Arsenic, Antimony, and Bismuth Compounds (δ in ppm)

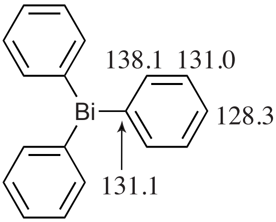






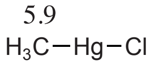


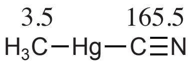


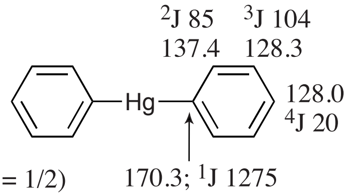


P Si

Mercury Compounds (δ in ppm, |J| in Hz)





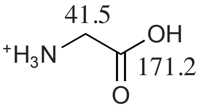


Couplings with ^{199}Hg (16.8%, I = 1/2)

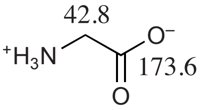
4.13 Natural Products

4.13.1 Amino Acids

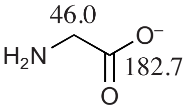
13C Chemical Shifts (δ in ppm; solvent: water)



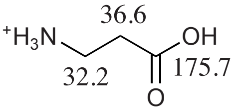
(pH 0.45)



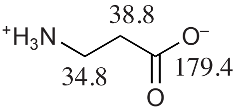
(pH 4.53)



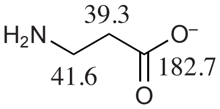
(pH 12.01)



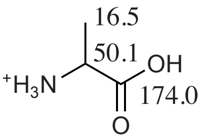
(pH 0.49)



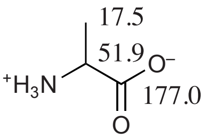
(pH 5.03)



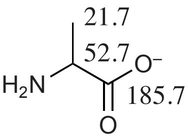
(pH 12.56)



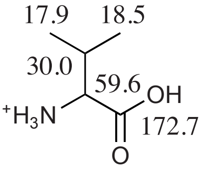
(pH 0.43)



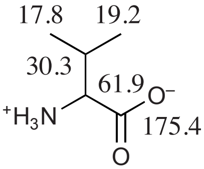
(pH 4.96)



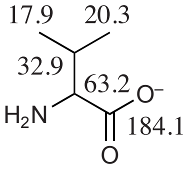
(pH 12.52)



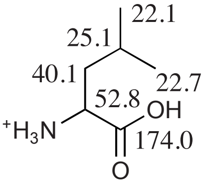
(pH 1.34)



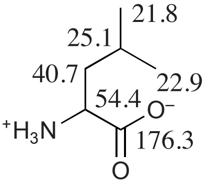
(pH 5.64)



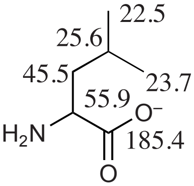
(pH 12.60)



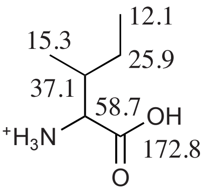
(pH 0.37)



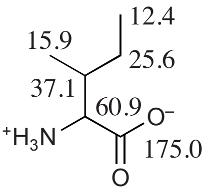
(pH 7.00)



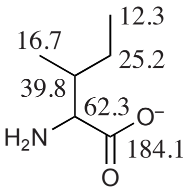
(pH 13.00)



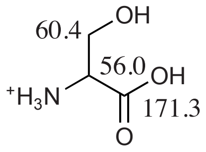
(pH 0.28)



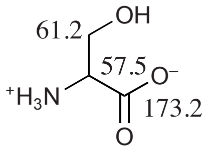
(pH 6.04)



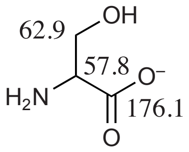
(pH 12.84)



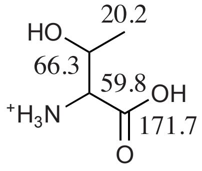
(pH 1.12)



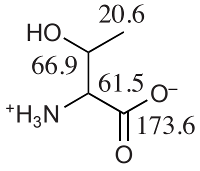
(pH 6.05)



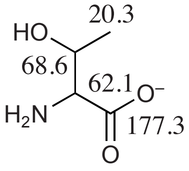
(pH 9.28)



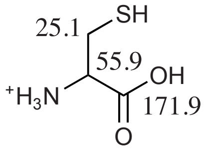
(pH 1.36)



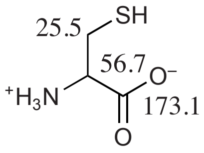
(pH 5.87)



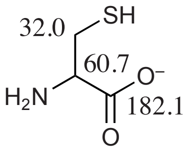
(pH 9.27)



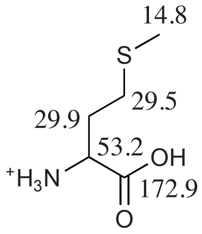
(pH 1.75)



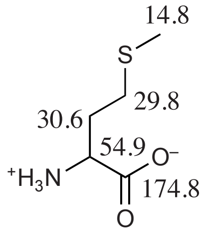
(pH 5.14)



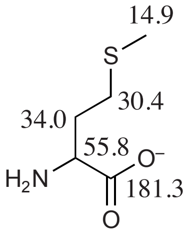
(pH 11.02)



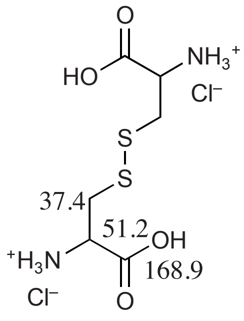
(pH 1.55)



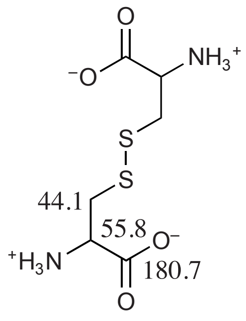
(pH 5.83)



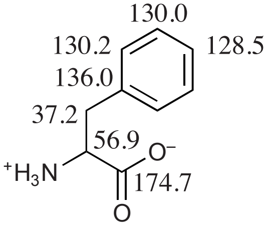
(pH 9.80)

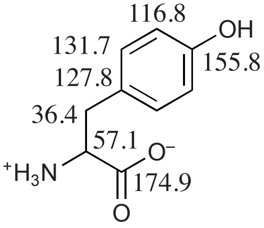


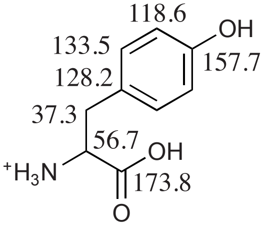
(in DMSO)



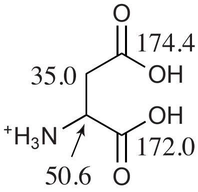
(in D_2O)



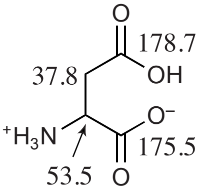




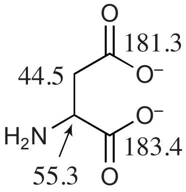
(in D_2O/DCl)



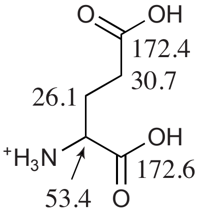
(pH 0.41)



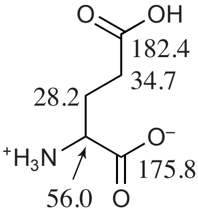
(pH 6.73)



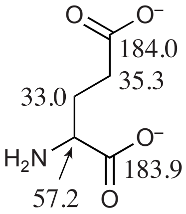
(pH 12.73)



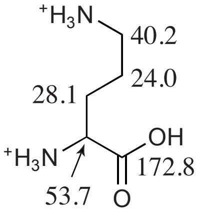
(pH 0.32)



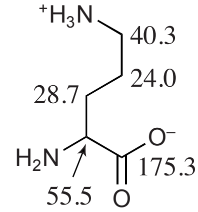
(pH 6.95)



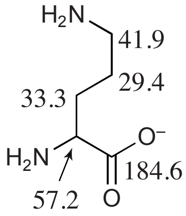
(pH 12.51)



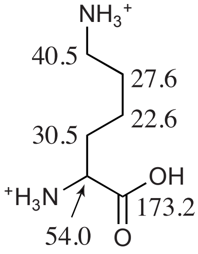
(pH 0.46)



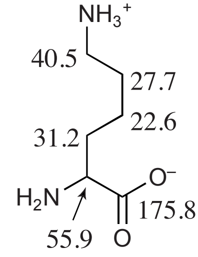
(pH 5.02)



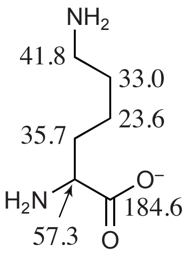
(pH 13.53)



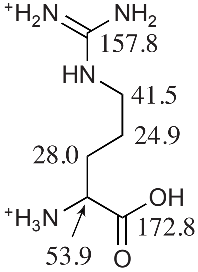
(pH 0.50)



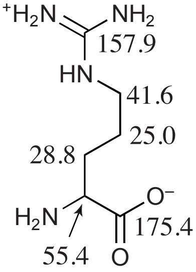
(pH 6.03)



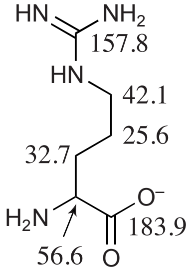
(pH 13.85)



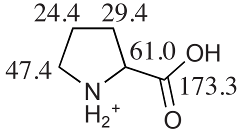
(pH 1.33)



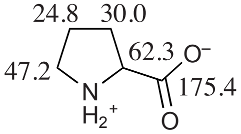
(pH 7.87)



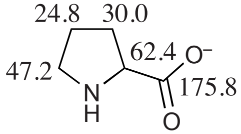
(pH 11.52)



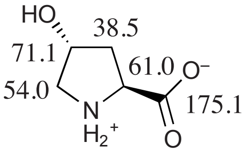
(pH 1.27)

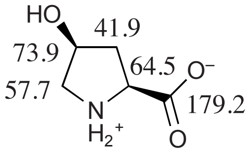


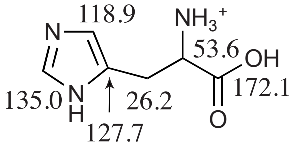
(pH 7.26)



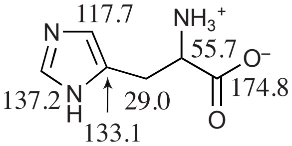
(pH 9.80)



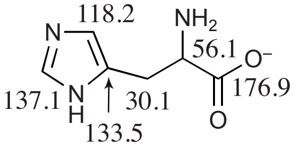




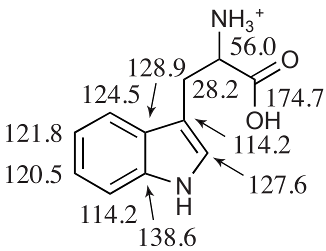
(pH 1.74)



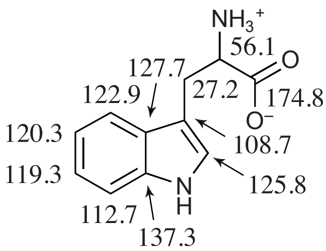
(pH 7.82)



(pH 9.21)



(in D_2O/DCl)

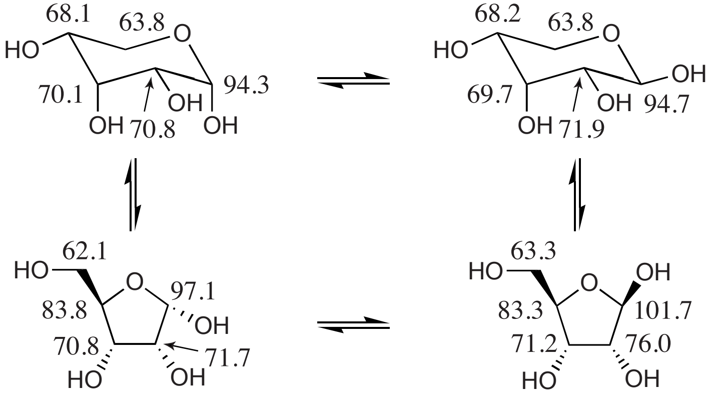


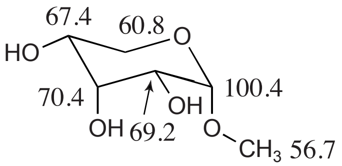
(80 °C)

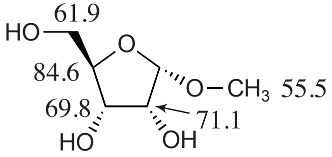
4.13.2 Carbohydrates

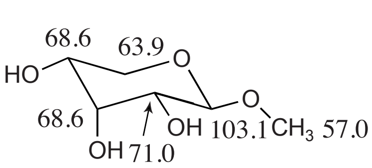
¹³C Chemical Shifts of Monosaccharides (δ in ppm)

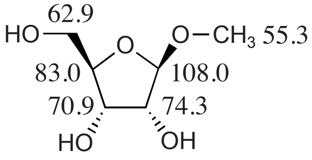
Ribose



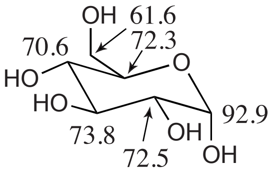


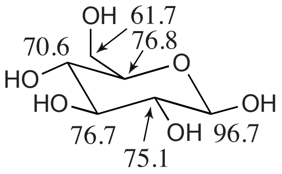


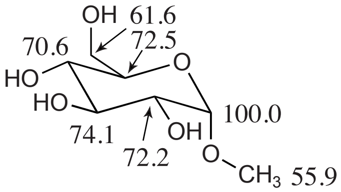


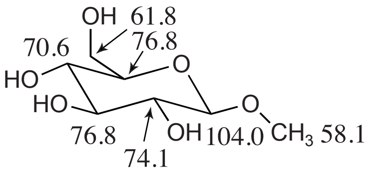


Glucose

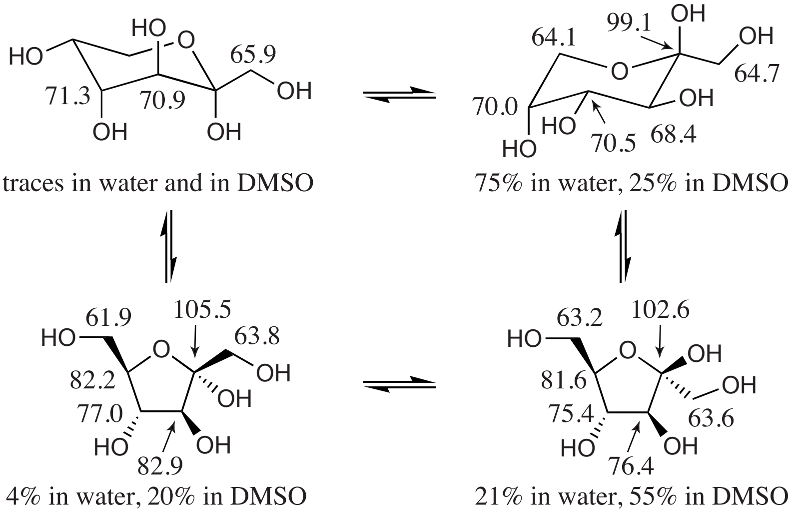








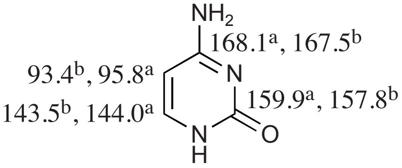
Fructose



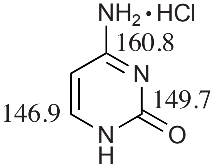
^{13}C-^{1}H Coupling Constants through One Bond ($^{1}J_{CH}$ in Hz)

4.13.3 Nucleotides and Nucleosides

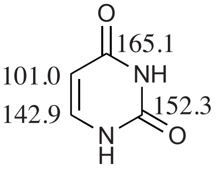
^{13}C Chemical Shifts of Nucleotides (δ in ppm)



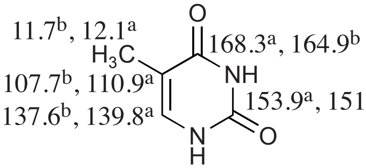
(a in D_2O, b in DMSO)



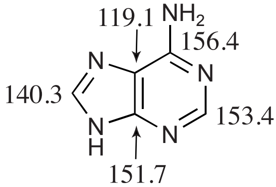
(in D_2O)



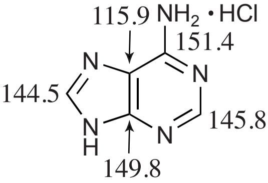
(in DMSO)



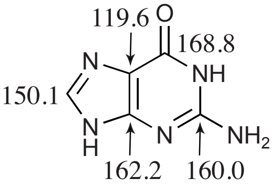
(a in D_2O, b in DMSO)



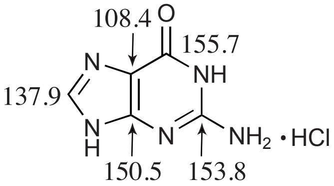
(in DMSO)



(in D_2O)

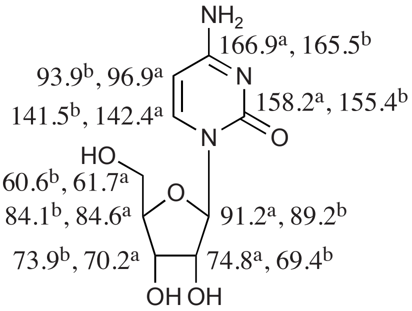


(in D_2O)

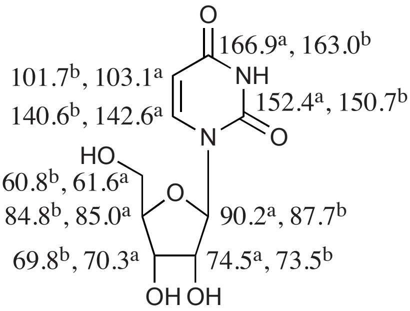


(in DMSO)

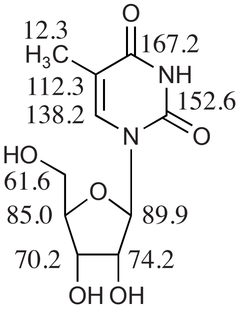
Nucleosides (δ in ppm)



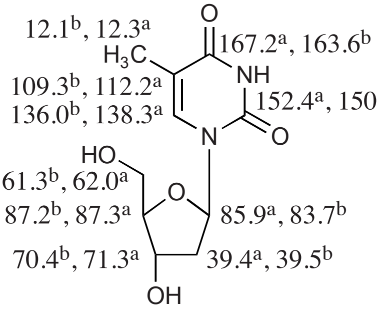
(a in D_2O, b in DMSO)



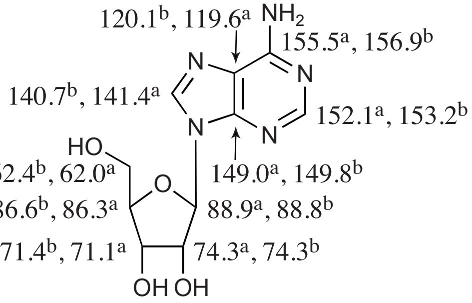
(a in D_2O, b in DMSO)



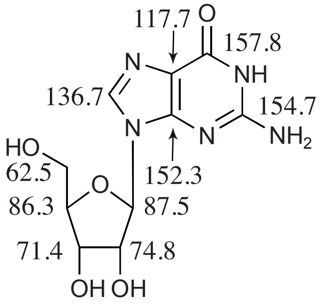
(in D_2O)



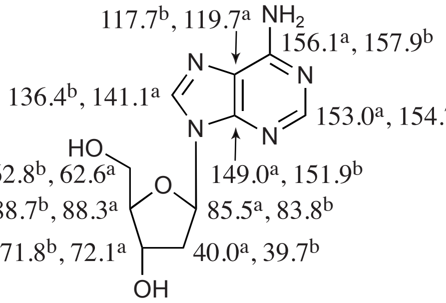
(^{a} in D_2O, ^{b} in DMSO)



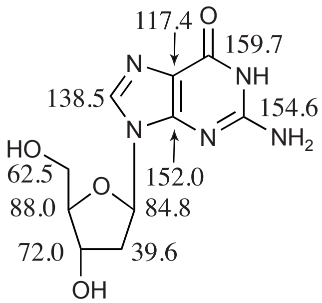
(^{a} in D_2O, ^{b} in DMSO)



(in DMSO)



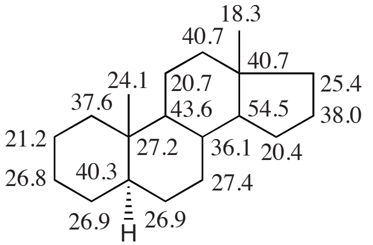
(^{a} in D_2O, ^{b} in DMSO)

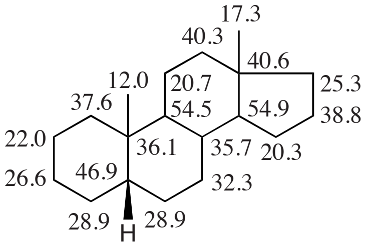


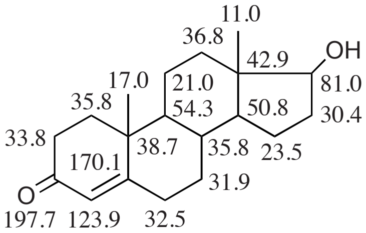
(in D_2O)

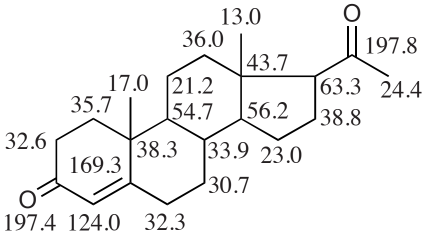
4.13.4 Steroids

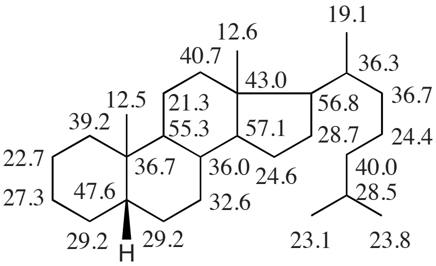
13C Chemical Shifts (δ in ppm)

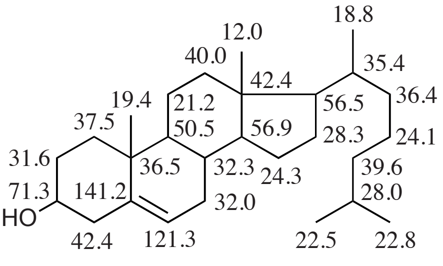










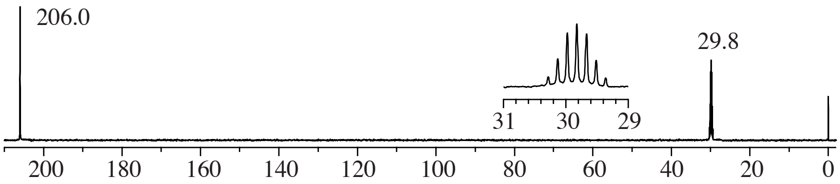


Natural Products

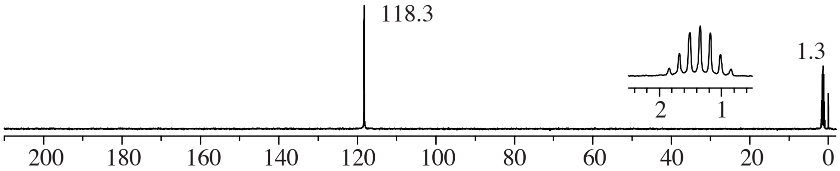
4.14 Spectra of Solvents and Reference Compounds

4.14.1 ^{13}C NMR Spectra of Common Deuterated Solvents (125 MHz, δ in ppm)

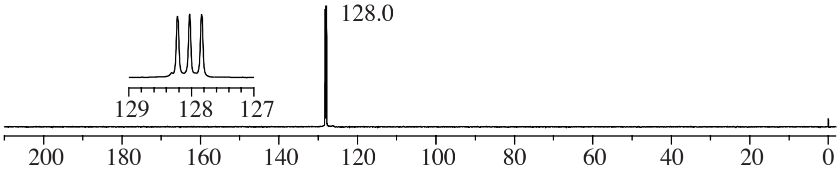
Acetone-d_6



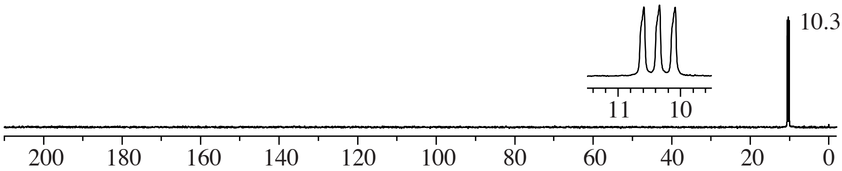
Acetonitrile-d_3



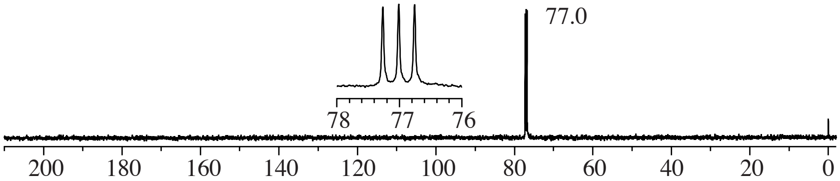
Benzene-d_6



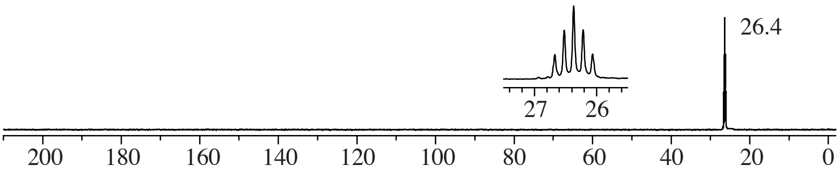
Bromoform-d



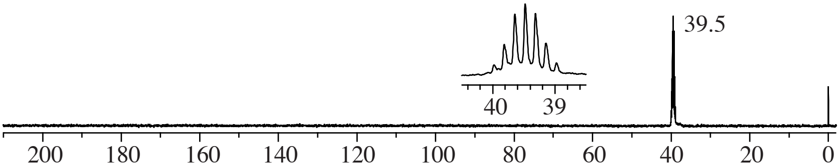
Chloroform-d



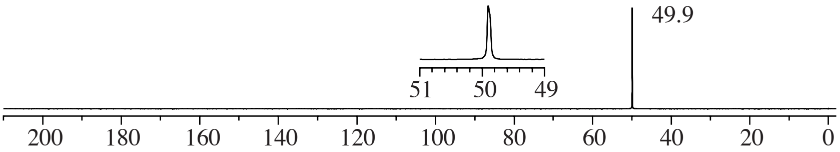
Cyclohexane-d_{12}



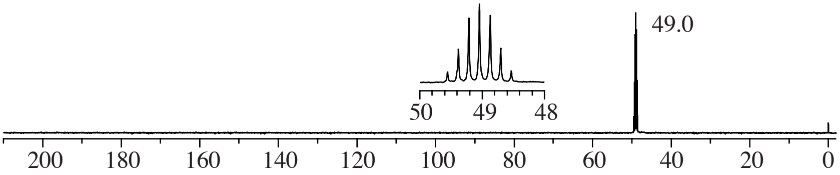
Dimethyl sulfoxide-d_6



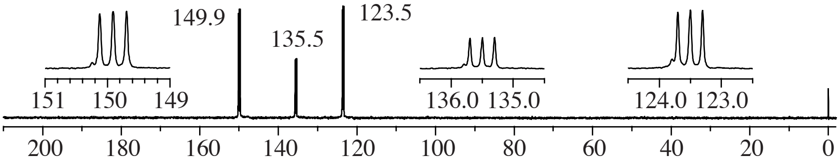
Methanol-d_1



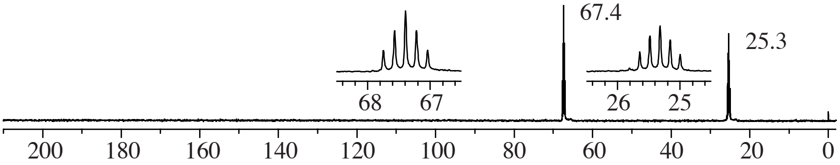
Methanol-d_4



Pyridine-d_5



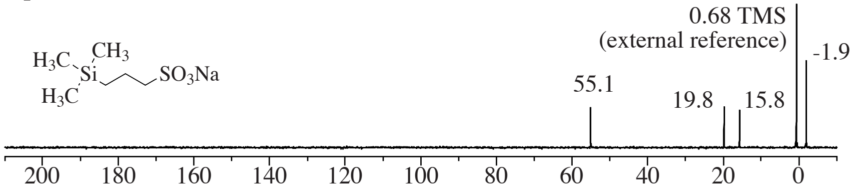
Tetrahydrofuran-d_8



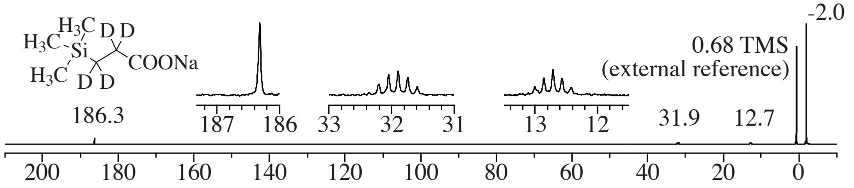
4.14.2 ^{13}C NMR Spectra of Secondary Reference Compounds

Chemical shifts in ^{13}C NMR spectra are usually reported relative to the peak position of tetramethylsilane (TMS), which is added as an internal reference. If TMS is not sufficiently soluble in the sample, the use of a capillary with TMS as external reference is recommended. In this case, owing to the difference in volume susceptibilities, the local magnetic fields in the solvent and reference are different. Therefore, the position of the reference must be corrected. For a D_2O solution in a cylindrical sample with TMS in a capillary, the correction amounts to +0.68 and -0.34 ppm for superconducting and electromagnets, respectively. These values must be subtracted from the ^{13}C chemical shifts relative to the external TMS signal if its position is set to 0.00 ppm. Alternatively, secondary references with $(CH_3)_3SiCH_2$ groups may be used. The following spectra of two secondary reference compounds in D_2O were measured at 125 MHz with a superconducting magnet and TMS as external reference. Chemical shifts are reported in ppm relative to TMS upon correction for the difference in the volume susceptibilities of D_2O. As a result, the peak for the external TMS appears at 0.68 ppm.

3-(Trimethylsilyl)-1-propanesulfonic acid sodium salt (sodium 4,4-dimethyl-4-silapentane-1-sulfonate; DSS)



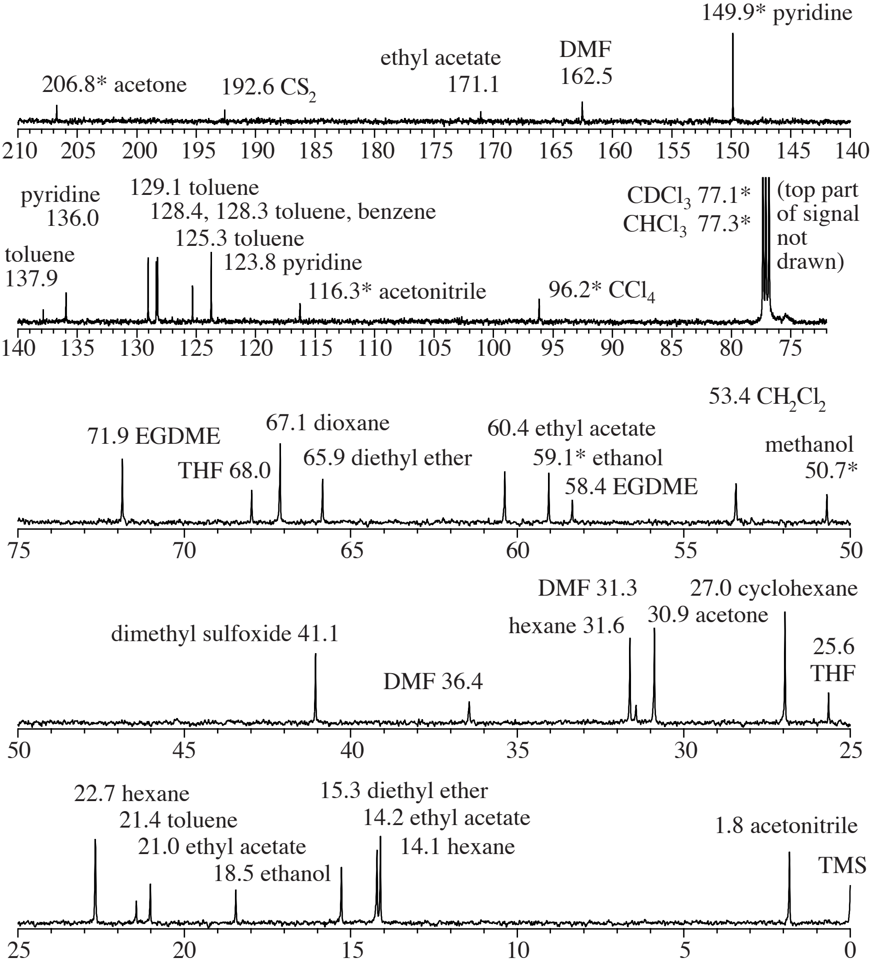
2,2,3,3-D_4-3-(Trimethylsilyl)-propionic acid sodium salt



4.14.3 ^{13}C NMR Spectrum of a Mixture of Common Nondeuterated Solvents

The broad-band-decoupled ^{13}C NMR spectrum (125 MHz, δ in ppm relative to TMS) of a $CDCl_3$ sample with 20 common solvents (0.05–0.4 vol%) shown below serves as a guide to identify possible solvent impurities. Chemical shifts of signals marked with an asterisk (*) may change up to a few ppm if the sample contains solutes having functional groups that can form hydrogen bonds.

DMF: dimethyl formamide; THF: tetrahydrofuran; EGDME: ethylene glycol dimethyl ether.

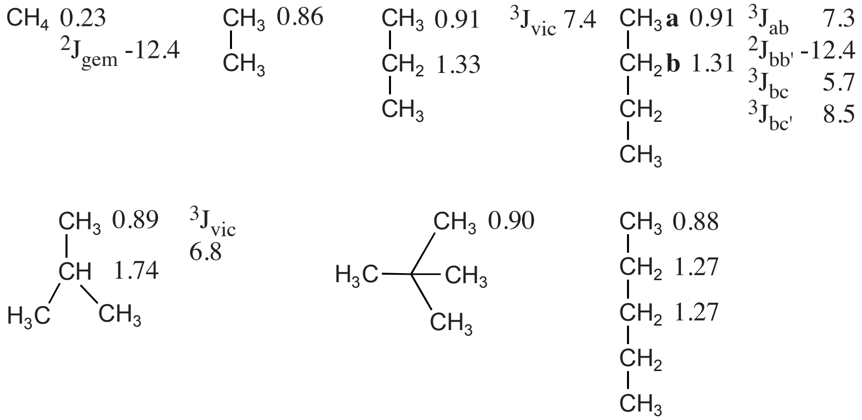


5 ^{1}H NMR Spectroscopy

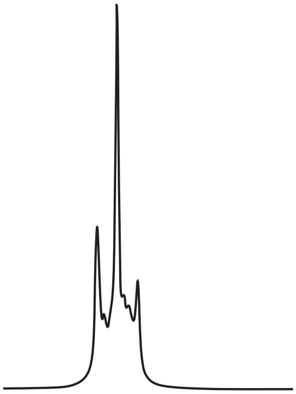
5.1 Alkanes

5.1.1 Chemical Shifts

^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)



In long-chain alkanes, the methyl groups at δ ca. 0.8 ppm typically show distorted triplets due to second-order effects:



^{1}H Chemical Shifts of Monosubstituted Alkanes (δ in ppm)

	Substituent	Methyl	Ethyl		1-Propyl		
		$–CH_3$	$–CH_2$	$–CH_3$	$–CH_2$	$–CH_2$	$–CH_3$
	–H	0.23	0.86	0.86	0.91	1.33	0.91
C	$–CH{=}CH_2$	1.71	2.00	1.00	2.02	1.43	0.91
	– C≡CH	1.80	2.16	1.15	2.10	1.50	0.97
	–phenyl	2.35	2.63	1.21	2.59	1.65	0.95
X	–F	4.27	4.55	1.35	4.30	1.68	0.97
	–Cl	3.06	3.47	1.33	3.47	1.81	1.06
	–Br	2.69	3.37	1.66	3.35	1.89	1.06
	–I	2.16	3.16	1.88	3.16	1.88	1.03
O	–OH	3.48	3.71	1.24	3.59	1.59	0.94
	–O–alkyl	3.24	3.37	1.15	3.27	1.55	0.93
	$–OCH{=}CH_2$	3.16	3.66	1.21			
	–O–phenyl	3.73	3.98	1.38	3.86	1.70	1.05
	$–OCOCH_3$	3.67	4.12	1.26	4.02	1.65	0.95
	–OCO–phenyl	3.88	4.37	1.38	4.25	1.76	1.07
	$–OS(O)_2$–4-tolyl	3.70	4.07	1.30	3.94	1.60	0.95
N	$–NH_2$	2.47	2.66	1.11	2.65	1.46	0.91
	$–NHCH_3$	2.30					
	$–N(CH_3)_2$	2.22	2.32	1.06			
	$–NHCOCH_3$	2.79	3.26	1.14	3.18	1.55	0.96
	$–NO_2$	4.29	4.37	1.58	4.28	2.01	1.03
	–C≡N	1.98	2.35	1.31	2.34	1.70	1.08
	–NC	2.85	3.39	1.28			
S	–SH	2.00	2.44	1.31	2.50	1.63	0.99
	–S–alkyl	2.09	2.49	1.25	2.43	1.59	0.98
	–SS–alkyl	2.30	2.67	1.35	2.63	1.71	1.03
	$–S(O)CH_3$	2.50					
	$–S(O)_2CH_3$	2.84	2.94	2.80			
O=C	–CHO	2.20	2.46	1.13	2.37	1.64	0.97
	$–COCH_3$	2.17	2.44	1.06	2.40	1.60	0.93
	–CO–phenyl	2.55	2.92	1.18	2.86	1.72	1.02
	–COOH	2.10	2.36	1.16	2.31	1.68	1.00
	$–COOCH_3$	2.01	2.32	1.15	2.22	1.65	0.98
	$–CONH_2$	2.02	2.23	1.13	2.19	1.68	0.99
	–COCl	2.66	2.93	1.24	2.87	1.74	1.00

^{1}H Chemical Shifts of Monosubstituted Alkanes (δ in ppm, contd.)

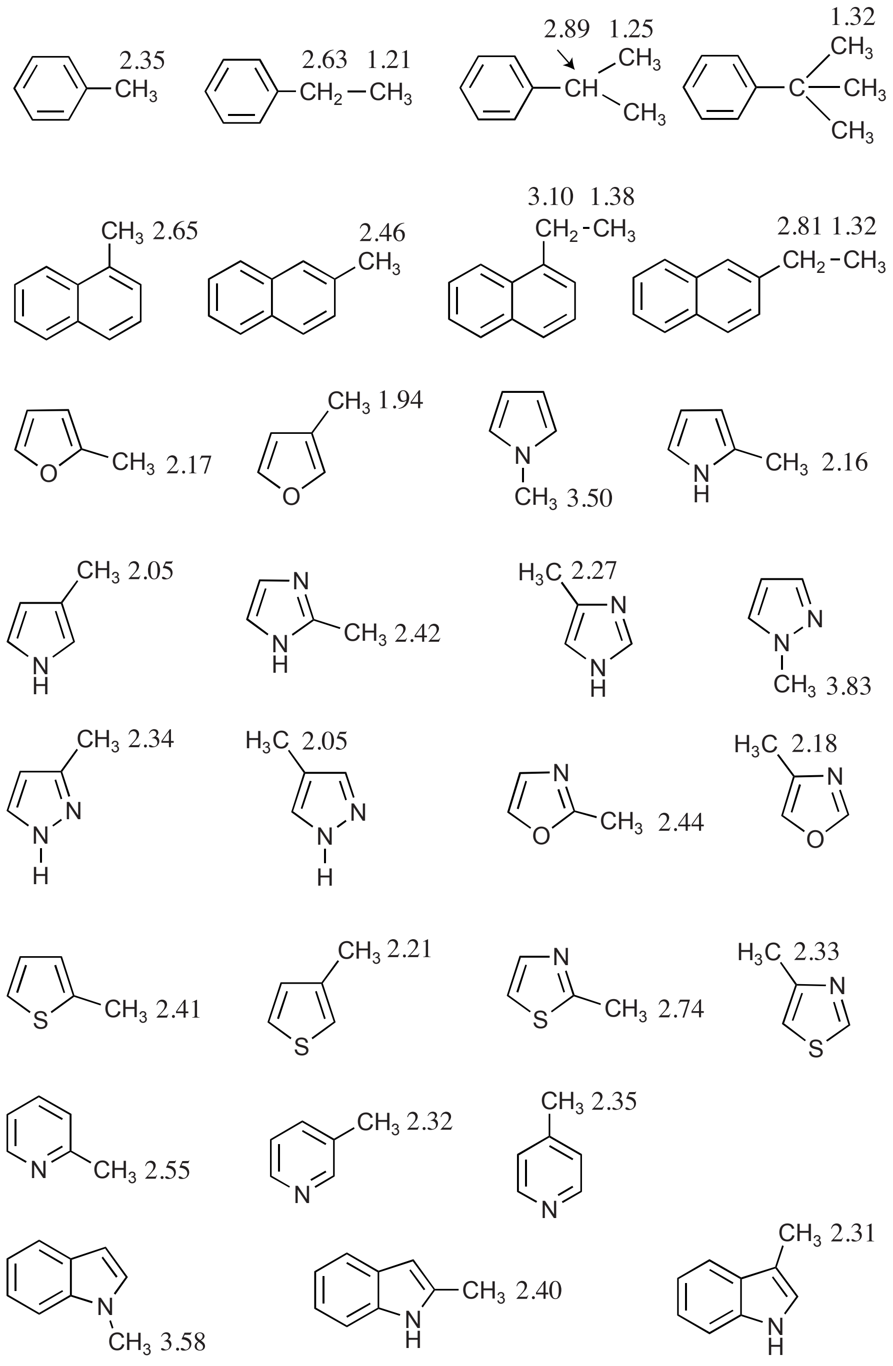
	Substituent	2-Propyl		*n*-Butyl				*tert*-Butyl
		–CH	$–CH_3$	$–CH_2$	$–CH_2$	$–CH_2$	$–CH_3$	$–CH_3$
	–H	1.33	0.91	0.91	1.31	1.31	0.91	0.89
C	$–CH{=}CH_2$			2.06	≈1.5	≈1.2	0.90	1.02
	–C≡C–	2.59	1.15	2.18	1.52	1.41	0.92	1.24
	–phenyl	2.89	1.25	2.61	1.60	1.34	0.93	1.32
X	–F	4.84	1.34	4.34	1.65		0.95	1.34
	–Cl	4.14	1.55	3.42	1.68	1.41	0.92	1.60
	–Br	4.21	1.73	3.42	1.84	1.46	0.93	1.76
	–I	4.24	1.89	3.20	1.80	1.42	0.93	1.95
O	–OH	4.02	1.21	3.64	1.56	1.39	0.94	1.26
	–O–alkyl	3.55	1.08	3.40	1.54	1.38	0.92	1.24
	$–OCH{=}CH_2$	4.06	1.23	3.68	1.61	1.39	0.94	
	–O–phenyl	4.51	1.31	3.94	1.76	1.47	0.97	
	$–OCOCH_3$	4.99	1.23	4.06	1.60	1.39	0.94	1.45
	–OCO–phenyl	5.22	1.37					1.58
	$–OS(O)_2$–4-tolyl	4.70	1.25	4.03	1.62	1.36	0.88	
N	$–NH_2$	3.07	1.03	2.69	1.43	1.35	0.92	1.15
	$–NHCOCH_3$	4.01	1.13	3.21	1.49	1.35	0.92	1.28
	$–NO_2$	4.44	1.53	4.47	2.07	1.50	1.07	1.59
	–C≡N	2.67	1.35	2.34	1.63	1.50	0.96	1.37
	–NC	3.87	1.45					1.44
S	–SH	3.16	1.34	2.52	1.59	1.43	0.92	1.43
	–S–alkyl	2.93	1.25	2.49	1.56	1.42	0.92	1.39
	–SS–alkyl			2.69	1.64	1.42	0.93	1.32
	$–S(O)_2CH_3$	3.13	1.41					1.44
O=C	–CHO	2.39	1.13	2.42	1.59	1.35	0.93	1.08
	$–COCH_3$	2.58	1.11					1.13
	–CO–phenyl	3.58	1.22	2.95	1.72	1.41	0.96	
	–COOH	2.56	1.21	2.35	1.62	1.39	0.93	1.23
	$–COOCH_3$	2.56	1.17	2.31	1.61	1.33	0.92	1.20
	$–CONH_2$	2.44	1.18	2.22	1.60	1.37	0.93	1.22
	–COCl	2.97	1.31	2.88	1.67	1.40	0.93	

Estimation of ^{1}H Chemical Shifts of Substituted Alkanes (δ in ppm)

CH$_3$	CH$_2$	CH
$\delta_{CH_3R^1} = 0.86 + Z_\alpha$	$\delta_{CH_2} = 1.37 + \sum_i Z_{\alpha_i} + \sum_j Z_{\beta_j}$	$\delta_{CH} = 1.50 + \sum_i Z_{\alpha_i} + \sum_j Z_{\beta_j}$
$\delta_{CH_3CR^1R^2R^3} = 0.86 + \sum_i Z_{\beta_i}$		

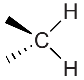
	Substituent (R^1, R^2, R^3)	CH$_3$ Z_α	CH$_3$ Z_β	CH$_2$ Z_α	CH$_2$ Z_β	CH Z_α	CH Z_β
C	–C≤	0.00	0.05	0.00	-0.06	0.17	-0.01
	–C=C<	0.85	0.20	0.63	0.00	0.68	0.03
	–C≡C–	0.94	0.32	0.70	0.13	1.04	
	–phenyl	1.51	0.38	1.22	0.29	1.28	0.38
X	–F	3.41	0.41	2.76	0.16	1.83	0.27
	–Cl	2.20	0.63	2.05	0.24	1.98	0.31
	–Br	1.83	0.83	1.97	0.46	2.44	0.41
	–I	1.30	1.02	1.80	0.53	2.46	0.15
O	–OH	2.53	0.25	2.20	0.15	1.73	0.08
	–O–C≤	2.38	0.25	2.04	0.13	1.85	0.32
	–OC=C<	2.64	0.36	2.63	0.33	2.00	0.30
	–O–phenyl	2.87	0.47	2.61	0.38	2.20	0.50
	–O–CO–	2.81	0.44	2.83	0.24	2.47	0.59
N	–N<	1.61	0.14	1.32	0.22	1.13	0.23
	–N$^+$≤	2.44	0.39	1.91	0.40	1.78	0.56
	–N–CO–	1.88	0.34	1.63	0.22	2.10	0.62
	–NO$_2$	3.43	0.65	3.08	0.58	2.31	
	–C≡N	1.12	0.45	1.08	0.33	1.00	
	–NCS	2.51	0.54	2.20	0.36	1.94	0.60
S	–S–	1.14	0.45	1.23	0.26	1.06	0.31
	–S–CO–	1.41	0.37	1.54	0.63	1.31	0.19
	–S(O)–	1.64	0.36	1.24	0.30	1.25	
	–S(O)$_2$–	1.98	0.42	2.08	0.52	1.50	0.40
	–SCN	1.75	0.66	1.62		1.64	
O=C	–CHO	1.34	0.21	1.07	0.29	0.86	0.22
	–CO–	1.23	0.20	1.12	0.24		
	–COOH	1.22	0.23	0.90	0.23	0.87	0.32
	–COO–	1.15	0.28	0.92	0.35	0.83	0.63
	–CO–N<	1.16	0.28	0.85	0.24	0.94	0.30
	–COCl	1.94	0.22	1.51	0.25		

[1]H Chemical Shifts of Aromatically Substituted Alkanes (δ in ppm)



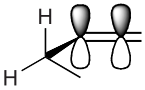
5.1.2 Coupling Constants

Geminal Coupling Constants (^2J in Hz)

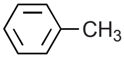
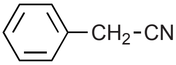


$^2J_{HCH}$ -8 to -18

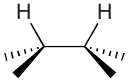
Electronegative substituents cause a decrease in $|J|_{gem}$ while a double or triple bond next to the CH_2 group causes an increase. The latter effect is strongest if one of the C–H bonds is parallel to the π orbitals:



Influence of Substituents on the Geminal Coupling Constant

Compound	J_{gem}	Compound	J_{gem}
CH_4	-12.4	CH_3COCH_3	-14.9
CH_3Cl	-10.8	CH_3COOH	-14.5
CH_2Cl_2	-7.5	CH_3CN	-16.9
CH_3OH	-10.8	$CH_2(CN)_2$	-20.3
	-14.3		-18.5

Vicinal Coupling Constants (^3J in Hz)



conformation not fixed: $^3J_{HCCH} \approx 7$
fixed: $^3J_{HCCH} \approx 0–18$

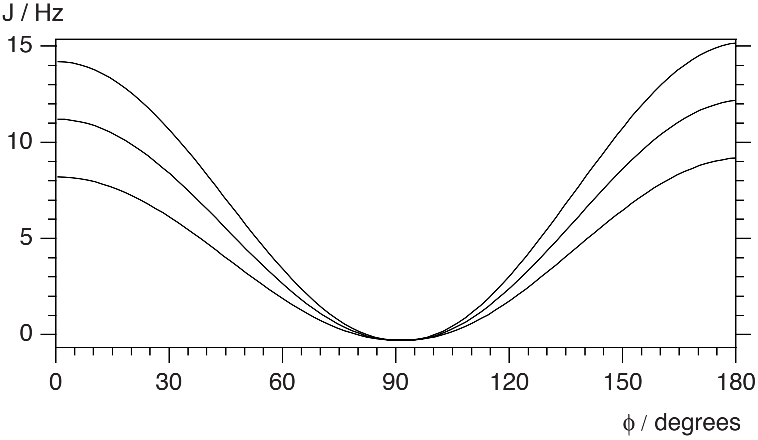
Influence of Substituents on the Vicinal Coupling Constant

Compound	J_{vic}	Compound	J_{vic}	Compound	J_{vic}
CH_3CHF_2	4.5	CH_3CH_2OH	6.9	CH_3CH_2CN	7.6
CH_3CHCl_2	6.1	$(CH_3CH_2)_3O^+ BF_4^-$	7.2	$(CH_3CH_2)_2S$	7.4
CH_3CH_2F	6.9	$(CH_3CH_2)_3N$	7.1	$(CH_3CH_2)_4Si$	8.0
CH_3CH_2Cl	7.2	$(CH_3CH_2)_4N^+ I^-$	7.3	CH_3CH_2Li	8.4

Vicinal coupling constants strongly depend on the dihedral angle, φ (Karplus equation):

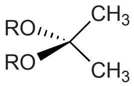
$$^3J = J^0 \cos^2 \phi - 0.3 \qquad 0^o \leq \phi \leq 90^o$$
$$^3J = J^{180} \cos^2 \phi - 0.3 \qquad 90^o \leq \phi \leq 180^o$$

The same relationship between torsional angle and vicinal coupling constant holds for substituted alkanes if appropriate values are used for J^0 and J^{180}. These limiting values depend on the electronegativity and orientation of substituents, the hybridization of carbon atoms, bond lengths, and bond angles.



Long-Range Coupling Constants (|J| in Hz)

Coupling constants through more than three bonds (long-range coupling) in alkanes are generally much smaller than 1 Hz and, thus, not visible in routine 1D NMR spectra. They are, however, much larger than 1 Hz for fixed conformations (e.g., in condensed alicyclic systems, see Chapter 5.4) and in unsaturated compounds (see Chapter 5.2). They are also significant when electronegative substituents are present between the coupling partners, as e.g.:

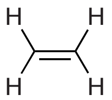


$^4J_{HH}$ 0.7

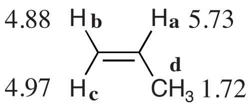
5.2 Alkenes

5.2.1 Substituted Ethylenes

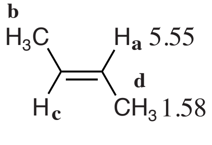
^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)



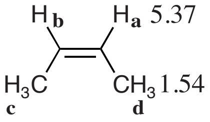
$^2J_{gem}$	2.5
$^3J_{cis}$	11.6
$^3J_{trans}$	19.1



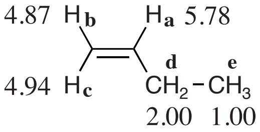
$^3J_{ab}$	10.0
$^3J_{ac}$	16.8
$^3J_{ad}$	6.4
$^2J_{bc}$	2.1
$^4J_{bd}$	-1.3
$^4J_{cd}$	-1.8



$^4J_{ab}$	-1.7
$^3J_{ac}$	15.1
$^3J_{ad}$	6.5
$^5J_{bd}$	1.6



$^3J_{ab}$	10.9
$^4J_{ac}$	-1.8
$^3J_{ad}$	6.8
$^5J_{cd}$	1.2



$^3J_{ab}$	10.3	$^4J_{bd}$	-1.3
$^3J_{ac}$	17.2	$^4J_{cd}$	-1.7
$^3J_{ad}$	6.2		
$^2J_{bc}$	2.0		

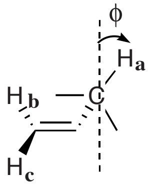
Geminal and Vicinal Couplings of Alkenes (J in Hz)

The values of the coupling constants strongly depend on the electronegativity of the substituents (see Table on pp 166, 167). They decrease with increasing electronegativity and number of electronegative substituents. The same trend holds for the signed values of geminal coupling constants but not for the absolute values because J_{gem} can be positive or negative. Although the total ranges of cis and trans vicinal coupling constants overlap, $J_{trans} > J_{cis}$ always holds for given substituents.

Typical ranges:		
	J_{gem}	-4 to 4
	J_{cis}	4 to 12
	J_{trans}	14 to 19

Coupling Over More than Three Bonds in Alkenes (Long-Range Coupling, J in Hz)

Allylic Coupling

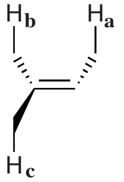


cisoid: $^4J_{ab}$ -3.0 to +2.0
transoid: $^4J_{ac}$ -3.5 to +2.5

In acyclic systems, the coupling constants range from ca. -0.8 to -1.8 Hz and, usually, $|J|_{cisoid}$ is larger than $|J|_{transoid}$. The magnitudes of the coupling constants depend on the conformation. Largest absolute values are observed if the C–H bond of the substituents overlaps with the π electrons of the double bond (ϕ = 0 or 180°):

ϕ	$^4J_{ab}$	$^4J_{ac}$
0°	-3.0	-3.5
90°	+1.8	+2.2
180°	-3.0	-3.5
270°	0.0	0.8

Homoallylic Coupling

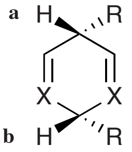


cisoid: $|^5J_{ab}|$ 0–3
transoid: $|^5J_{ac}|$ 0–3

The values of homoallylic coupling constants between methyl groups and of allylic ones are comparable:

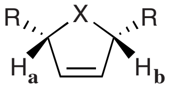
$$^5J_{H_3C\text{–}C=C\text{–}CH_3} \approx {}^4J_{H\text{–}C=C\text{–}CH_3}$$

In acyclic systems, $|J_{cisoid}| < |J_{transoid}|$ usually holds. Large homoallylic coupling constants are occasionally observed in cyclic systems with fixed conformation between the protons:



$^5J_{ab}$ 5–11

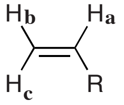
X: CH, N
R: any substituent



$^4J_{ab}$ 0–7

X: O, NH
R: any substituent

^{1}H Chemical Shifts and Coupling Constants of Monosubstituted Ethylenes (δ in ppm, J in Hz)

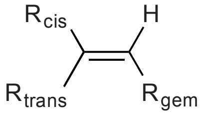


	Substituent R	H_a	H_b	H_c	J_{ab}	J_{ac}	J_{bc}	Other
	–H	5.28	5.28	5.28	11.6	19.1	2.5	
C	$–CH_3$	5.73	4.88	4.97	10.0	16.8	2.1	CH_3 1.72
	$–CH_2CH{=}CH_2$	5.71	4.92	4.95	10.3	16.9	2.2	CH_2 2.72
	$–CH_2$–phenyl	5.89	5.00	5.01	10.0	17.0	1.9	CH_2 3.19
	–cyclopropyl	5.32	4.84	5.04	10.4	17.1	1.8	
	–cyclohexyl	5.79	4.88	4.95	10.5	17.6	1.9	
	$–CF_3$	5.90	5.56	5.85	11.1	17.5	0.2	
	$–CH{=}C{=}CH_2$	6.31	4.99	5.19	10.1	17.2	1.6	
	$–C{\equiv}C–CH_3$	5.62	5.24	5.39	11.1	17.0	2.3	
	–phenyl	6.72	5.20	5.72	11.1	17.9	1.0	
	–2-naphthyl	6.87	5.32	5.86				
	–2-nitrophenyl	7.19	5.45	5.68	10.7	17.4	1.1	
	–3-nitrophenyl	6.74	5.42	5.86	10.9	17.5	0.4	
	–4-nitrophenyl	6.77	5.48	5.90	10.9	17.4	0.8	
	–2-pyridyl	6.84	5.45	6.22	11.3	18.5	1.4	
	–4-pyridyl	6.61	5.42	5.91	10.8	17.6	0.7	
X	–F	6.17	4.03	4.37	4.7	12.8	-3.2	
	–Cl	6.26	5.39	5.48	7.5	14.5	-1.4	
	–Br	6.44	5.97	5.84	7.1	14.9	-1.9	
	–I	6.53	6.23	6.57	7.8	15.9	-1.5	
O	–OH	6.45	3.82	4.18	6.4	14.2	-1.0	
	$–OCH_3$	6.44	3.88	4.03	7.0	14.1	-2.0	CH_3 3.16
	$–OCH{=}CH_2$	6.49	4.21	4.52	6.4	14.0	-1.8	
	–O–phenyl	6.64	4.40	4.74	6.1	13.7	-1.6	
	–OCHO	7.33	4.66	4.96	6.4	13.9	-1.7	CHO 8.07
	$–OCOCH_3$	7.28	4.56	4.88	6.3	14.1	-1.6	CH_3 2.13
	$–OCOCH{=}CH_2$	7.39	4.62	4.96	6.4	14.2	-1.6	
	–OCO–phenyl	7.52	4.67	5.04	6.3	13.8	-1.7	
	$–OP(O)(O–ethyl)_2$	6.58	4.59	4.91	6.0	13.8	-2.1	
N	$–NH_2$	≈6.05	≈3.99	≈4.04				
	$–N^+(CH_3)_3Br^–$	6.50	5.54	5.76	8.2	15.1	-4.3	
	$–NHCOCH_3$	≈7.33	≈4.68	≈4.53				
	$–NO_2$	7.12	5.87	6.55	7.0	14.6	1.4	
	$–C{\equiv}N$	5.69	6.11	6.24	11.8	17.9	0.9	
	–NC	5.90	5.35	5.58	8.6	15.6	-0.5	
	–NCO	6.12	4.77	5.01	7.6	15.2	-0.1	

	Substituent R	H_a	H_b	H_c	J_{ab}	J_{ac}	J_{bc}	Other
S	$-SCH_3$	6.43	5.18	4.95	10.3	16.4	-0.3	CH_3 2.25
	–S–phenyl	6.53	5.32	5.32	9.6	16.7	-0.2	
	$-S(O)CH_3$	6.77	5.92	6.08	9.8	16.7	-0.6	CH_3 2.61
	$-S(O)_2CH_3$	6.76	6.14	6.43	10.0	16.5	-0.5	CH_3 2.96
	$-S(O)_2CH{=}CH_2$	6.67	6.17	6.41	10.0	16.4	-0.6	
	$-S(O)_2OH$	6.73	6.13	6.41	10.2	16.8	-1.2	
	$-S(O)_2OCH_3$	6.57	6.22	6.43	10.1	16.9	-0.6	CH_3 3.85
	$-S(O)_2NH_2$	6.93	5.98	6.17	10.0	16.3	0	NH_2 6.7
	$-S(O)_2NH$–phenyl	6.56	5.86	6.18	10.1	16.7	-0.3	NH 9.07
	$-SF_5$	6.63	5.64	5.96	9.8	16.6	0.4	
	–SCN	6.19	5.70	5.66				
O=C	–CHO	6.37	6.52	6.35	10.0	17.4	1.0	CHO 9.59
	$-COCH_3$	6.30	5.91	6.21	10.7	18.7	1.3	CH_3 2.29
	$-COCH{=}CH_2$	6.67	5.82	6.28	11.0	17.9	1.4	
	–CO–phenyl	7.20	5.81	6.52	9.9	17.7	2.3	
	–COOH	6.15	5.95	6.53	10.5	17.2	1.8	COOH 12.8
	$-COOCH_3$	6.12	5.83	6.41	10.6	17.4	1.5	CH_3 3.77
	$-CONH_2$	6.48	5.71	6.17	7.9	17.3	5.0	NH_2 7.55
	$-CON(CH_3)_2$	6.64	5.55	6.12	9.8	17.0	3.4	
	–COF	6.14	6.25	6.60	10.7	17.3	0.8	
	–COCl	6.35	6.16	6.63	10.6	17.4	0.2	
P	$-P(CH_3)_2$	6.23	5.51	5.39	11.8	18.3	2.0	CH_3 0.95
	$-P(CH{=}CH_2)_2$	6.16	5.64	5.59	11.8	18.4	2.0	
	$-P(phenyl)_2$	7.38	6.31	7.07	12.5	18.2	0	
	$-PCl_2$	7.48	6.68	6.64	11.7	18.6	0.4	
	$-P(O)(phenyl)_2$	6.72	6.21	6.25	12.9	18.9	1.8	
	$-PSCl_2$	6.42	5.90	6.13	11.0	17.5	0.3	
	$-P(S)(CH_3)_2$	6.60	6.14	6.26	11.8	17.9	1.8	
	$-P(S)(phenyl)_2$	6.82	6.17	6.34	11.7	17.9	1.6	
M	–Li	7.29	6.65	5.91	19.3	23.9	7.1	
	–MgBr	6.66	6.15	5.51	17.7	23.3	7.6	
	$-Si(CH_3)_3$	6.11	5.88	5.63	14.6	20.2	3.8	CH_3 0.06
	$-Sn(CH{=}CH_2)_3$	6.39	6.21	5.75	13.4	20.7	3.1	
	$-Pb(CH{=}CH_2)_3$	6.70	6.19	5.46	12.2	19.8	2.1	
	–HgBr	6.45	5.92	5.52	11.9	18.7	3.1	

C=C

Estimation of ^{1}H Chemical Shifts of Substituted Ethylenes (δ in ppm)



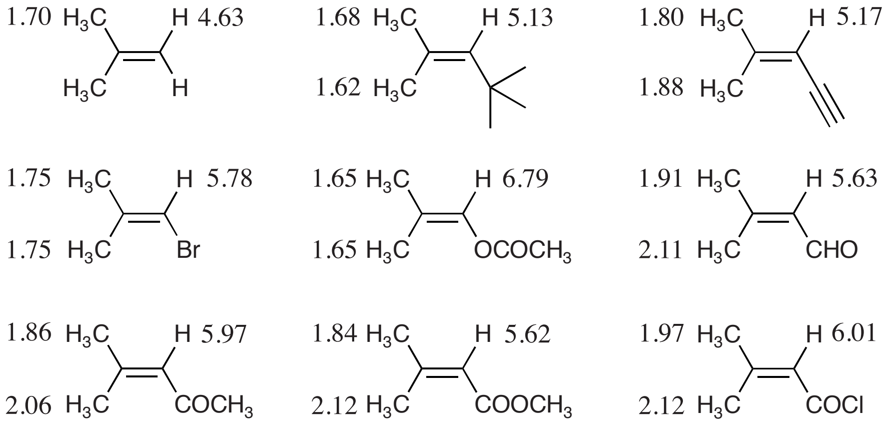
$$\delta_{C=CH} = 5.25 + Z_{gem} + Z_{cis} + Z_{trans}$$

	Substituent R	Z_{gem}	Z_{cis}	Z_{trans}
	–H	0.00	0.00	0.00
C	–alkyl	0.45	-0.22	-0.28
	–alkyl ring[1]	0.69	-0.25	-0.28
	–CH_2–aromatic	1.05	-0.29	-0.32
	–CH_2X, X: F, Cl, Br	0.70	0.11	-0.04
	–CHF_2	0.66	0.32	0.21
	–CF_3	0.66	0.61	0.32
	–CH_2O–	0.64	-0.01	-0.02
	–CH_2N<	0.58	-0.10	-0.08
	–CH_2CN	0.69	-0.08	-0.06
	–CH_2S–	0.71	-0.13	-0.22
	–CH_2CO–	0.69	-0.08	-0.06
	–C=C<	1.00	-0.09	-0.23
	–C=C< conjugated[2]	1.24	0.02	-0.05
	–C≡C–	0.47	0.38	0.12
	–aromatic	1.38	0.36	-0.07
	–aromatic, fixed[3]	1.60	–	-0.05
	–aromatic, *o*-substituted	1.65	0.19	0.09
X	–F	1.54	-0.40	-1.02
	–Cl	1.08	0.18	0.13
	–Br	1.07	0.45	0.55
	–I	1.14	0.81	0.88
O	–OC≤ (*sp*^{3})	1.22	-1.07	-1.21
	–OC= (*sp*^{2})	1.21	-0.60	-1.00
	–OCO–	2.11	-0.35	-0.64
	–OP(O)$(OCH_2CH_3)_2$	1.33	-0.34	-0.66
N	–NR_2; R: H, C≤ (*sp*^{3})	0.80	-1.26	-1.21
	–NR–; R: C= (*sp*^{2})	1.17	-0.53	-0.99
	–NCO–R	2.08	-0.57	-0.72
	–N=N–phenyl	2.39	1.11	0.67
	–NO_2	1.87	1.30	0.62
	–C≡N	0.27	0.75	0.55

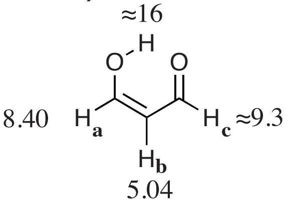
	Substituent R	Z_{gem}	Z_{cis}	Z_{trans}
S	–S–	1.11	-0.29	-0.13
	–S(O)–	1.27	0.67	0.41
	$–S(O)_2–$	1.55	1.16	0.93
	–SCO–	1.41	0.06	0.02
	–SCN	0.94	0.45	0.41
	–SF	1.68	0.61	0.49
O=C	–CHO	1.02	0.95	1.17
	–CO–	1.10	1.12	0.87
	–CO– conjugated[2]	1.06	0.91	0.74
	–COOH	0.97	1.41	0.71
	–COOH conjugated[2]	0.80	0.98	0.32
	–COOR	0.80	1.18	0.55
	–COOR conjugated[2]	0.78	1.01	0.46
	–CON<	1.37	0.98	0.46
	–COCl	1.11	1.46	1.01
	$–P(O)(OCH_2CH_3)_2$	0.66	0.88	0.67

1) The increment "alkyl ring" is to be used if the substituent and the double bond are part of a cyclic structure.
2) The increment "conjugated" is to be used if either the double bond or the substituent is conjugated to other substituents.
3) The increment "aromatic, fixed" is to be used if the double bond conjugated to an aromatic ring is part of a fused ring (such as in 1,2-dihydronaphthalene).

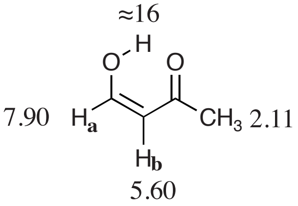
Influence of cis- *and* trans-*Substituents on the* ^{1}*H Chemical Shift of Methyl Groups at the Double Bond in Isobutenes (δ in ppm)*



¹H Chemical Shifts and Coupling Constants of Enols (δ in ppm, J in Hz)



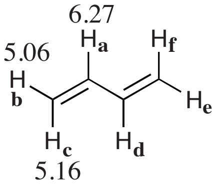
$^3J_{ab}$ 9.7
$^3J_{bc}$ ≈8



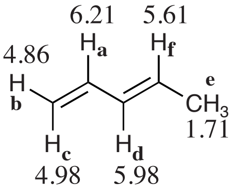
$^3J_{ab}$ 5.1

5.2.2 Conjugated Dienes

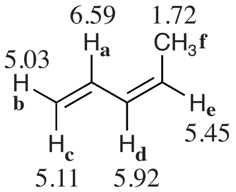
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)



$^3J_{ab}$	10.2	$^2J_{bc}$	1.8
$^3J_{ac}$	17.1	$^5J_{be}$	1.3
$^3J_{ad}$	10.4	$^5J_{bf}$	0.6
$^4J_{ae}$	-0.9	$^5J_{cf}$	0.7
$^4J_{af}$	-0.8		



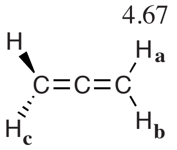
$^3J_{ab}$	10.2	$^2J_{bc}$	1.9	$^6J_{ce}$	-0.7
$^3J_{ac}$	16.9	$^4J_{bd}$	-0.8	$^5J_{cf}$	0.7
$^3J_{ad}$	10.3	$^6J_{be}$	-0.7	$^4J_{de}$	-1.6
$^5J_{ae}$	0.4	$^5J_{bf}$	0.7	$^3J_{df}$	15.1
$^4J_{af}$	-0.8	$^4J_{cd}$	-0.8	$^3J_{ef}$	6.6



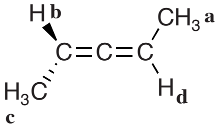
$^3J_{ab}$	10.2	$^2J_{bc}$	2.1	$^5J_{ce}$	0.7
$^3J_{ac}$	16.9	$^4J_{bd}$	-0.8	$^6J_{cf}$	-0.6
$^3J_{ad}$	10.9	$^5J_{be}$	-0.7	$^3J_{de}$	10.8
$^4J_{ae}$	-1.1	$^6J_{bf}$	0.7	$^4J_{df}$	-1.8
$^5J_{af}$	0.2	$^4J_{cd}$	-0.8	$^3J_{ef}$	7.0

5.2.3 Allenes

^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

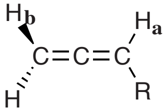


$^2J_{ab}$ -9
$^4J_{ac}$ -6



$^5J_{ab}$ 3.2
$^6J_{ac}$ 0
$^3J_{ad}$ 6.8
$^4J_{bd}$ -6.4

^{1}H Chemical Shifts and Coupling Constants of Monosubstituted Allenes (δ in ppm, J in Hz)



Substituent R	H_a	H_b	J_{ab}	Other
–H	4.67	4.67	-9.0	
$–CH_3$	4.94	4.50	-6.7	CH_3 (**c**) 1.59, $^3J_{ac}$ 7.2, $^5J_{bc}$ 3.4
$–CH_2CH_3$	5.03	4.55	-6.8	CH_2 (**c**) 1.95, $^3J_{ac}$ 6.2, $^5J_{bc}$ 3.5
$–CH_2Cl$	5.4.3	4.92	-6.6	CH_2 (**c**) 4.11, $^3J_{ac}$ 7.7, $^5J_{bc}$ 2.2
$–CH{=}CH_2$	5.96	4.92	-6.6	CH (**c**) 6.31, $^3J_{ac}$ 10.4, $^5J_{bc}$ 1.1[a]
–Cl	5.76	5.17	-6.1	
–Br	5.85	4.83	-6.1	
–I	5.63	4.48	-6.3	
–phenyl	5.91	4.92		
$–OCH_3$	6.77	5.48	-5.9	
$–COCH_3$	5.77	5.25	-6.4	
–C≡N	4.97	5.04		
$–Si(CH_3)_3$	4.92	4.31		
$–SiCl_3$	5.35	4.92	-5.9	
$–SnCl_3$	4.98	4.11	-7.2	

[a] $=CH_2$, H_{cis} (**d**) 5.19, $^4J_{ad}$ -0.8, $^6J_{bd}$ -1.5, $^3J_{cd}$ 17.2; CH_2, H_{trans} (**e**) 4.99, $^4J_{ae}$ -0.9, $^6J_{be}$ -1.8, $^3J_{ce}$ 10.1, $^2J_{de}$ 1.6

C≡C

5.3 Alkynes

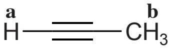
^{1}H Chemical Shifts of Substituted Alkynes (δ in ppm)

$$R-C\equiv C-H_a$$

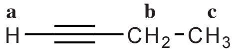
	Substituent R	H_a
	–H	1.91
C	$-CH_3$	1.91
	$-CH_2CH_3$	1.97
	$-C(CH_3)_3$	2.07
	$-CF_3$	2.95
	$-CH{=}CH_2$	3.07
	$-C\equiv CH$	2.16
	–phenyl	3.07
	–1-naphthyl	3.43
X	–F	1.74
	–Cl	2.05
	–Br	2.32
	–I	2.34
O	$-OCH_2CH_3$	1.48
	$-OCH{=}CH_2$	2.04
	–O–phenyl	2.07
N	$-N(CH_2CH_3)_2$	2.30
	$-N(\text{phenyl})_2$	2.86
	$-C\equiv N$	2.63

	Substituent R	H_a
S	$-SCH_2CH_3$	2.79
	$-SCH{=}CH_2$	3.26
	–S–phenyl	3.28
	$-S(O)_2$–*n*-butyl	3.95
O=C	$-COCH_3$	3.65
	–CO–phenyl	3.48
	–COOH	3.17
	$-COOCH_2CH_3$	2.90
	$-CONH_2$	3.05
Si	$-Si(CH_3)_3$	2.34
	$-Si(\text{phenyl})_3$	2.47
P	$-P(CH_2CH_3)_2$	2.85
	$-P(\text{phenyl})_2$	3.22
	$-P(O)(CH_2CH_3)_2$	3.33
	$-P(O)(\text{phenyl})_2$	3.48

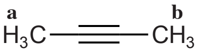
^{1}H,^{1}H Coupling Constants of Substituted Alkynes (J in Hz)



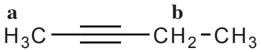
$^4J_{ab}$ 2.9



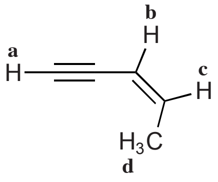
$^4J_{ab}$ 2.6
$^5J_{ac}$ 0
$^3J_{bc}$ 7.4



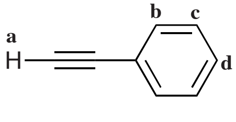
$|^5J_{ab}|$ 2.7



$|^5J_{ab}|$ 2.5



$^4J_{ab}$ 2.0
$^5J_{ac}$ 1.0
$^6J_{ad}$ 0.6
$^3J_{bc}$ 10.5
$^4J_{bd}$ 1.6
$^3J_{cd}$ 6.5



$^5J_{ab}$ 0.3
$^6J_{ac}$ -0.1
$^7J_{ad}$ 0.2

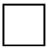
5.4 Alicyclics

¹H Chemical Shifts and Coupling Constants of Saturated Alicyclic Hydrocarbons (δ in ppm, J in Hz)

0.20 $^2J_{gem}$ -4.3; $^3J_{cis}$ 9.0; $^3J_{trans}$ 5.6

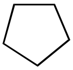
In derivatives: $^2J_{gem}$ -3 to -9; $^3J_{cis}$ 6 to 12; $^3J_{trans}$ 2 to 9

Throughout: $J_{cis} > J_{trans}$



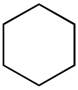
1.94

In derivatives: $^2J_{gem}$ -10 to -17; $^3J_{cis}$ 4 to 12; $^3J_{trans}$ 2 to 10; $^4J_{cis}$ ≈ 0; $^4J_{trans}$ ≈ -1



1.51

In derivatives: $^2J_{gem}$ -8 to -18; $^3J_{cis}$ 5 to 10; $^3J_{trans}$ 5 to 10

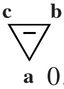


1.44

At -100 °C: H_{ax} 1.12; H_{eq} 1.60

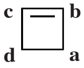
In derivatives: $^2J_{gem}$ -11 to -14; $^3J_{ax,ax}$ 8 to 13; $^3J_{eq,ax}$ 2 to 6; $^3J_{eq,eq}$ 2 to 5

Generally: $J_{eq,ax} \approx J_{eq,eq} + 1$



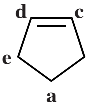
b 7.01; a 0.92

In derivatives: $^3J_{ab}$ 1.5 to 2.0; $^3J_{bc}$ 0.5 to 1.5



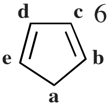
b 5.95; a 2.57

$^2J_{gem}$	-13.7	$^3J_{ad,cis}$	1.8
$^3J_{ab}$	1.0	$^3J_{ad,trans}$	4.6
$^4J_{ac}$	-0.3	$^3J_{bc}$	2.8



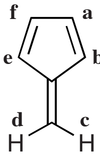
c 5.66; b 2.27; a 1.79

$^2J_{gem,a}$	-12.8	$^4J_{bd}$	-2.3
$^3J_{ab,cis}$	9.3	$^5J_{be,cis}$	2.1
$^3J_{ab,trans}$	5.7	$^5J_{be,trans}$	3.0
$^2J_{gem,b}$	-16.1	$^3J_{cd}$	5.8
$^3J_{bc}$	2.3		



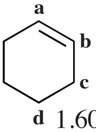
c 6.43; b 6.28; a 2.80

$^3J_{ab}$	1.3	$^4J_{bd}$	1.1
$^4J_{ac}$	-1.5	$^5J_{be}$	2.0
$^3J_{bc}$	5.0	$^3J_{cd}$	1.9



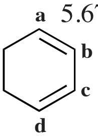
a 6.53; b 6.23; c 5.89

$^3J_{ab}$	5.1	$^4J_{bc}$	-0.2
$^5J_{ac}$	0.5	$^4J_{bd}$	-0.4
$^5J_{ad}$	1.4	$^4J_{be}$	2.0
$^4J_{ae}$	1.3	$^2J_{cd}$	0.1
$^3J_{af}$	2.0		



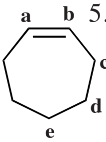
b 5.69; c 1.99; d 1.60

$^3J_{ab}$ ≈10; $^3J_{bc}$ 1.5



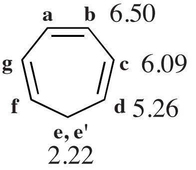
a 5.67; b 5.79; 2.08

$^3J_{ab}$ 9.7; $^4J_{ac}$ 1.0; $^5J_{ad}$ 1.1; $^3J_{bc}$ 5.1

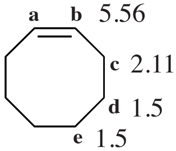


b 5.79; c 2.12; d 1.50; e 1.75

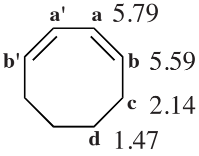
$^3J_{ab}$ ≈10; $^3J_{bc}$ 3.7



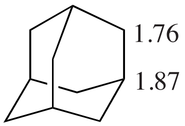
$^{3}J_{ab}$ 11.2 $^{5}J_{cg}$ -0.6
$^{4}J_{ac}$ 0.8 $^{3}J_{de}$ 6.7
$^{3}J_{bc}$ 5.5 $^{2}J_{ee'}$ -13.0
$^{3}J_{cd}$ 8.9
$^{5}J_{cf} = {}^{4}J_{df} = {}^{5}J_{dg} = 0$

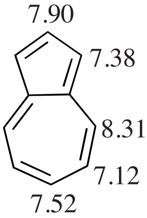


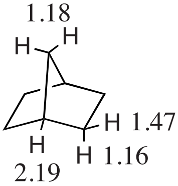
$^{3}J_{ab}$ ≈10
$^{3}J_{bc}$ 3.7

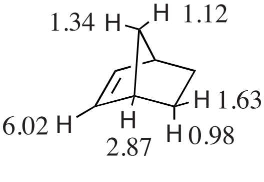


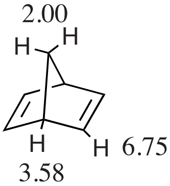
$^{3}J_{ab}$ 11.3
$^{3}J_{aa'}$ 4.1
$^{4}J_{ab'}$ -0.6
$^{5}J_{bb'}$ 0.5

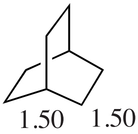


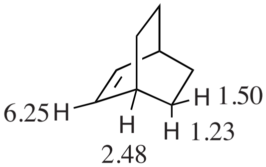


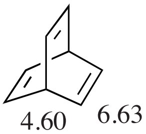




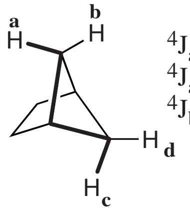




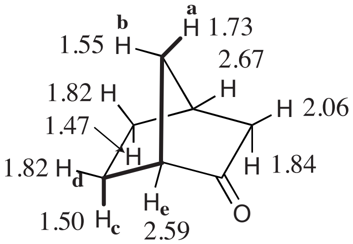




In condensed alicyclics, couplings over four bonds are often observed. Such long-range couplings are particularly large if the arrangement of the bonds between the two protons is W-shaped (cf. J_{ac} vs. J_{ad} and J_{bd} below left and J_{ac} vs. J_{bc} below right). Owing to the rigid arrangement, vicinal coupling constants (^{3}J) may assume unusually small values when the torsional angles are close to 90° (J_{ce} below right).

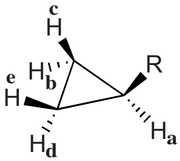


$^{4}J_{ac}$ ≈7
$^{4}J_{ad}$ ≈0
$^{4}J_{bd}$ ≈0



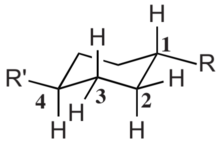
$^{4}J_{ac}$ 2.3
$^{4}J_{bc}$ -0.1
$^{3}J_{ce}$ 0.1
$^{3}J_{de}$ 4.7

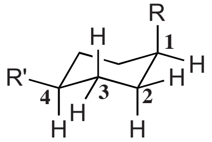
^{1}H Chemical Shifts and Coupling Constants of Monosubstituted Cyclopropanes (δ in ppm, J in Hz)



	Substituent R	H_a	$H_{b;d}$	$H_{c;e}$	$^3J_{ab}$	$^3J_{ac}$	$^2J_{bc}$	$^3J_{bd}$	$^3J_{be}$	$^3J_{ce}$
	–H	0.20	0.20	0.20	9.0	5.6	-4.3	9.0	5.6	9.0
C	$–CH_3$	1.00	0.35	0.15						
	$–CH_2OH$	1.14	0.40	0.30						
	$–CH=CH_2$	1.35	0.64	0.34	8.2	4.9	-4.5	9.3	6.2	9.0
	–phenyl	1.83	0.89	0.65	9.5	6.3	-4.5	9.5	5.2	8.9
X	–F	4.32	0.69	0.27	5.9	2.4	-6.7	10.8	7.7	12.0
	–Cl	2.55	0.87	0.74	7.0	3.6	-6.0	10.3	7.1	10.6
	–Br	2.83	0.96	0.81	7.1	3.8	-6.1	10.2	7.0	10.5
	–I	2.31	1.04	0.76	7.5	4.4	-5.9	9.9	6.6	10.0
O	–OH	3.35	0.40	0.48	6.2	2.9	-5.4	10.3	6.8	10.9
N	$–NH_2$	2.23	0.32	0.20	6.6	3.6	-4.3	9.7	6.2	9.9
	$–NH_3^+$	1.06	0.52	0.34						
	$–NO_2$	4.21	1.13	1.60	7.0	3.4	-5.5	10.1	8.3	11.3
	–C≡N	1.29	0.96	1.04	8.4	5.1	-4.7	9.2	7.1	9.5
O=C	–CHO	1.79	0.99	1.03	8.0	4.6	-4.5	8.8	7.0	9.6
	$–COCH_3$	1.83	0.77	0.93	7.9	4.6	-3.5	9.2	7.0	9.5
	–CO–cyclopropyl	1.70	0.56	1.02	7.9	4.6	-3.5	9.1	7.0	9.5
	–CO–phenyl	2.65	1.01	1.23						
	–COOH	1.59	0.91	1.05	8.0	4.6	-4.0	9.3	7.1	9.7
	$–COOCH_3$	1.61	0.86	0.98	8.0	4.6	-3.4	8.8	6.9	9.6
	$–CONH_2$	1.39	0.70	0.95						
	–COF	1.66	1.11	1.20	8.0	4.6	-4.5	10.1	7.5	9.3
	–COCl	2.11	1.18	1.28	7.9	4.4	-4.5	9.2	7.6	10.0
M	–Li	-2.53	0.43	-0.12	10.3	9.1	-1.6	7.7	3.2	6.5
	–MgBr	-2.04	0.25	-0.13	11.0	8.5	-1.7	7.8	3.5	6.6
	$–B(cyclopropyl)_2$	-0.25	0.61	0.66	8.9	5.8	-3.3	8.2	5.9	8.4
	$–Si(cyclopropyl)_3$	-0.67	0.49	0.36	9.7	6.9	-3.4	8.4	5.1	8.1
	$–P^+(phenyl)_3$	3.28	1.82	0.63						
	–Hg–cyclopropyl	0.00	0.75	0.47	9.6	6.9	-3.7	8.5	4.8	7.9

^{1}H Chemical Shifts and Coupling Constants of Equatorially and Axially Substituted Cyclohexanes (δ in ppm, J in Hz)





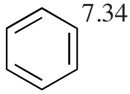
	Eq. substituent R	$H_{1,ax}$	$H_{2,ax}$	$H_{2,eq}$	$H_{3,ax}$	$H_{3,eq}$	$H_{4,ax}$	$H_{4,eq}$
	–D*	1.12	1.12	1.60	1.12	1.60	1.12	1.60
C	–C*	1.27	0.81	1.57	1.15	1.60	1.06	1.58
	–C≡C*	2.25	1.36	1.98	1.20	1.73	1.17	1.67
	–phenyl*	2.46						
X	–F*	4.49	1.42	2.15	1.28	1.86	1.12	1.65
	–Cl*	3.88	1.58	2.22	1.33	1.84	1.18	1.68
	–Br*	4.09	1.75	2.33	1.35	1.80	1.22	1.72
	–I*	4.18	1.97	2.45	1.36	1.67	1.30	1.80
O	–OH**	3.52	1.22	2.01	1.05	1.78	0.97	
	$–OCOCH_3$*	4.74	1.72	1.85	1.35	1.41	1.25	1.55
N	$–NH_2$**	2.55	1.03	1.89	1.03	1.76	0.96	
	$–NHCOCH_3$**	3.67	1.07	2.01	1.11	1.78	1.01	
	$–NO_2$*	4.38	2.23	1.85	1.38	1.85	1.28	1.67
	–C≡N**	2.31	1.53	2.16	0.98	1.86	1.03	
S	–SH*	2.79	1.34	2.01	1.31	1.75	1.22	1.61
	$–COOCH_3$*	2.30	1.44	1.90	1.27	1.75	1.24	1.64

	Ax. substituent R	$H_{1,eq}$	$H_{2,ax}$	$H_{2,eq}$	$H_{3,ax}$	$H_{3,eq}$	$H_{4,ax}$	$H_{4,eq}$
	–D*	1.60	1.12	1.60	1.12	1.60	1.12	1.60
C	–C*	1.93	1.37	1.40	1.39	1.34	1.06	1.58
	–C≡C*	2.87	1.48	1.78				
	–phenyl*	3.16		2.42				
X	–F*	4.94	1.43	2.03	1.63	1.75	1.28	1.58
	–Cl*	4.59	1.76	2.00	1.77	1.75	1.26	1.75
	–Br*	4.80	1.81	2.08	1.79	1.60	1.24	1.78
	–I*	4.96	1.53	2.06	1.72	1.62	1.26	1.73
O	–OH**	4.03	1.49	1.83	1.35	1.54	0.99	
	$–OCOCH_3$*	5.31	1.49	2.51				
N	$–NH_2$**	3.15	1.54	1.65	1.27	1.53	0.96	
	$–NHCOCH_3$**	4.11	1.51	1.85	1.03	1.66	1.04	
	$–NO_2$**	4.43	1.6	2.6				
	–C≡N**	2.96	1.54	2.00	1.50	1.70	1.20	
S	–SH**	3.43	1.5	1.9				

* R': –H; ** R': –*tert*-butyl

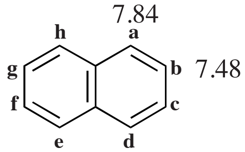
5.5 Aromatic Hydrocarbons

^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)



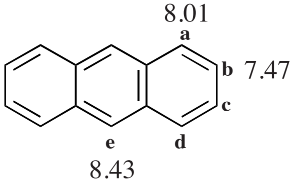
In derivatives:

$^3J_{ortho}$	6.5–8.5
$^4J_{meta}$	1.0–3.0
$^5J_{para}$	0.0–1.0



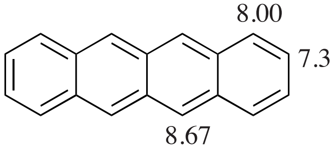
In derivatives:

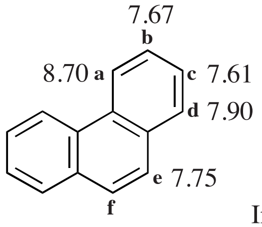
$^3J_{ab}$	8–9	$^6J_{af}$	≈-0.1
$^4J_{ac}$	1–2	$^5J_{ag}$	≈0.2
$^5J_{ad}$	≈1	$^4J_{ah}$	≈-0.5
$^3J_{bc}$	5–7	$^7J_{bf}$	≈0.3
$^5J_{ae}$	≈0.9	$^6J_{bg}$	≈0.1



In derivatives:

$^3J_{ab}$	8.5–9.5
$^4J_{ac}$	0.8–1.5
$^5J_{ad}$	0.6–0.9
$^5J_{ae}$	≈0.8
$^3J_{bc}$	6.5–8.0
$^4J_{de}$	≈0.4

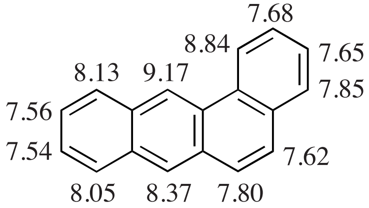


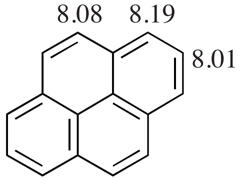


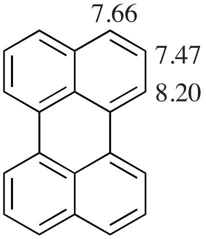
$^3J_{ab}$	8.4
$^4J_{ac}$	1.2
$^5J_{ad}$	0.7
$^3J_{bc}$	7.2
$^4J_{bd}$	1.3
$^3J_{cd}$	8.1

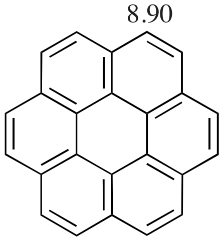
In derivatives:

$^3J_{ef}$ ≈9

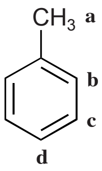




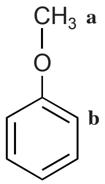




Weak long-range couplings between aromatic protons and aliphatic substituents are usually not resolved but lead to a characteristic broadening of the corresponding lines.

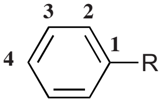


$^4J_{ab}$	-0.7
$^5J_{ac}$	0.3
$^6J_{ad}$	-0.6



$^5J_{ab}$ ≈0.8

Effect of Substituents on ^{1}H Chemical Shifts of Monosubstituted Benzenes (in ppm)

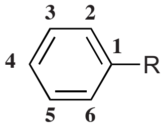


$$\delta_{H_i} = 7.34 + Z_i$$

	Substituent R	Z_2	Z_3	Z_4
C	$-CH_3$	-0.17	-0.09	-0.17
	$-CH_2CH_3$	-0.14	-0.05	-0.18
	$-CH(CH_3)_2$	-0.13	-0.08	-0.18
	$-C(CH_3)_3$	0.05	-0.04	-0.18
	$-CF_3$	0.19	-0.07	0.00
	$-CCl_3$	0.55	-0.07	-0.09
	$-CH_2OH$	-0.07	-0.07	-0.07
	$-CH{=}CH_2$	0.08	-0.02	-0.09
	–CH=CH–phenyl (*trans*)	0.16	0.00	-0.15
	–C≡CH	0.16	-0.01	-0.01
	–C≡C–phenyl	0.20	-0.04	-0.07
	–phenyl	0.22	0.06	-0.04
	–2-pyridyl	0.73	0.09	0.02
X	–F	-0.31	-0.03	-0.21
	–Cl	-0.01	-0.06	-0.12
	–Br	0.15	-0.12	-0.06
	–I	0.36	-0.24	-0.02
O	–OH	-0.51	-0.10	-0.41
	$-OCH_3$	-0.44	-0.05	-0.40
	$-OCH_2CH{=}CH_2$	-0.45	-0.13	-0.43
	–O–phenyl	-0.33	-0.02	-0.25
	$-OCOCH_3$	-0.26	0.03	-0.12
	–OCO–phenyl	-0.12	0.10	-0.06
	$-OS(O)_2CH_3$	-0.05	0.07	-0.01
N	$-NH_2$	-0.67	-0.20	-0.59
	$-NHCH_3$	-0.73	-0.16	-0.64
	$-N(CH_3)_2$	-0.60	-0.10	-0.62
	$-N(\text{phenyl})_2$	-0.26	-0.10	-0.34
	$-N^+(CH_3)_3\ I^-$	0.72	0.40	0.34
	–NHCHO (*trans* to O)	-0.25	0.03	-0.13
	–NHCHO (*cis* to O)	-0.20	0.21	-0.01
	$-N(CH_3)CHO$	-0.16	0.07	-0.05
	$-NHCOCH_3$	0.15	-0.02	-0.23
	$-NHCSNH_2$	0.14	0.07	-0.14

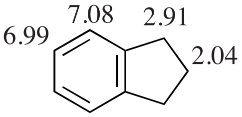
	Substituent R	Z_2	Z_3	Z_4
	$-NHNH_2$	-0.60	-0.08	-0.55
	–N=N–phenyl	0.67	0.20	0.20
	–NO	0.55	0.29	0.35
	$-NO_2$	0.93	0.26	0.39
	–C≡N	0.32	0.14	0.28
	–NCS	-0.11	0.04	-0.02
S	–SH	-0.08	-0.16	-0.22
	$-SCH_3$	-0.08	-0.10	-0.24
	–S–phenyl	-0.06	-0.20	-0.26
	–S–S–phenyl	0.13	-0.05	-0.10
	$-S(O)-CH=CH_2$	0.28	0.15	0.15
	–S(O)–phenyl	0.29	0.09	0.13
	$-S(O)_2CH_3$	0.70	0.37	0.41
	$-S(O)_2OCH_3$	0.60	0.26	0.28
	$-S(O)_2Cl$	0.68	0.27	0.37
	$-S(O)_2NH_2$	0.51	0.28	0.24
C=O	–CHO	0.54	0.19	0.29
	$-COCH_3$	0.62	0.12	0.22
	$-COCH_2CH_3$	0.61	0.11	0.21
	–CO–phenyl	0.56	0.12	0.23
	–CO–(2-pyridyl)	0.86	0.11	0.20
	–COOH	0.79	0.14	0.28
	$-COOCH_3$	0.70	0.09	0.21
	$-COOCH(CH_3)_2$	0.73	0.11	0.20
	–COO–phenyl	0.87	0.18	0.30
	$-CONH_2$	0.48	0.11	0.19
	–COF	0.71	0.21	0.38
	–COCl	0.77	0.15	0.35
	–COBr	0.70	0.15	0.32
	–CH=N–phenyl	0.64	0.24	0.24
M	–Li	0.77	0.26	-0.29
	–MgBr	0.40	-0.19	-0.26
	–Mg–phenyl	-0.49	0.18	0.25
	$-Si(CH_3)_3$	0.19	0.00	0.00
	$-Si(phenyl)_2Cl$	0.32	0.07	0.12
	$-SiCl_3$	0.52	0.20	0.20
	$-P(phenyl)_2$	0.0	0.0	0.0
	$-P(O)(OCH_3)_2$	0.46	0.14	0.22
	$-Pb^+(phenyl)\ Cl^-$	0.30	0.49	0.61
	–Zn–phenyl	-0.36	0.02	0.05
	–Hg–phenyl	0.06	0.10	-0.10

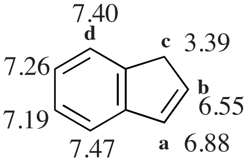
^{1}H-^{1}H Coupling Constants in Selected Monosubstituted Benzenes (|J| in Hz)



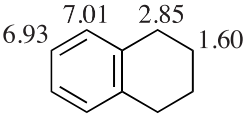
	Substituent R	J_{23}	J_{24}	J_{25}	J_{26}	J_{34}	J_{35}
C	$-CH_3$	7.7	1.3	0.6	2.0	7.5	1.5
	$-CH{=}CH_2$	7.8	1.1	0.6	1.9	7.4	1.5
	$-C{\equiv}CH$	7.8	1.3	0.6	1.7	7.6	1.3
	–phenyl	7.8	1.2	0.6	2.0	7.5	1.4
X	–F	8.4	1.1	0.4	2.7	7.5	1.8
	–Cl	8.1	1.1	0.5	2.3	7.5	1.7
	–Br	8.0	1.1	0.5	2.2	7.4	1.8
	–I	7.9	1.1	0.5	1.9	7.5	1.8
O	–OH	8.2	1.1	0.5	2.7	7.4	1.7
	$-OCH_3$	8.3	1.0	0.4	2.7	7.4	1.8
	–O–phenyl	8.3	1.1	0.5	2.6	7.4	1.7
	$-OCOCH_3$	8.2	1.1	0.5	2.5	7.5	1.7
N	$-NH_2$	8.0	1.1	0.5	2.5	7.4	1.6
	$-NHCOCH_3$	8.2	1.2	0.5	2.4	7.4	1.5
	$-NO_2$	8.4	1.2	0.5	2.4	7.5	1.5
	$-C{\equiv}N$	7.8	1.3	0.7	1.8	7.7	1.3
S	–SH	7.9	1.2	0.6	2.1	7.5	1.5
	$-S(O)_2OCH_3$	8.0	1.2	0.6	2.0	7.6	1.4
O=C	–CHO	7.7	1.3	0.6	1.8	7.5	1.3
	$-COCH_3$	8.0	1.3	0.6	1.8	7.5	1.3
	–COOH	7.9	1.3	0.6	1.9	7.4	1.4
	$-COOCH_3$	7.9	1.4	0.6	1.8	7.5	1.3
	$-CONH_2$	7.9	1.2	0.6	2.0	7.5	1.3
	–COCl	8.0	1.2	0.6	2.0	7.5	1.4
M	–Li	6.7	1.5	0.8	0.7	7.4	1.3
	–MgBr	6.9	1.5	0.7	0.7	7.4	1.4
	$-P(phenyl)_2$	7.6	1.2	0.6	1.7	7.4	1.4
	$-PO(OCH_3)_2$	7.7	1.4	0.6	1.6	7.6	1.4
	–Zn–phenyl	6.6	2.1	0.7	0.8	7.4	1.5
	–Hg–phenyl	7.5	1.4	0.6	1.1	7.5	1.5

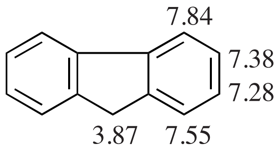
^{1}H Chemical Shifts and Coupling Constants of Condensed Aromatic-Alicyclic Hydrocarbons (δ in ppm, J in Hz)

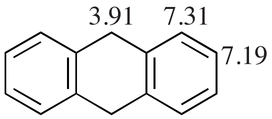


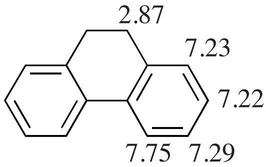


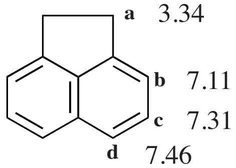
$^3J_{ab}$	5.5
$^4J_{ac}$	2.0
$^5J_{ad}$	0.7
$^3J_{bc}$	1.9



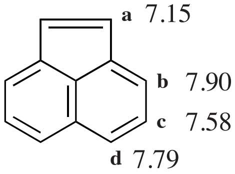




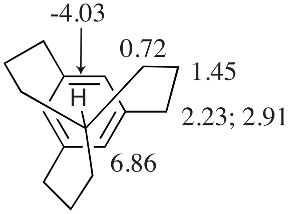




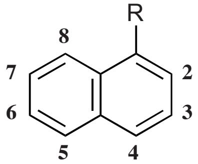
$^4J_{ab}$	1.5
$^5J_{ac}$	0
$^6J_{ad}$	0.5
$^3J_{bc}$	6.7
$^4J_{bd}$	1.2
$^3J_{cd}$	8.1



$^4J_{ab}$	0
$^5J_{ac}$	0
$^6J_{ad}$	0
$^3J_{bc}$	7
$^4J_{bd}$	0.6
$^3J_{cd}$	8



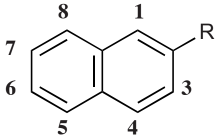
Effect of Substituents in Position 1 on the ^{1}H Chemical Shifts of Monosubstituted Naphthalenes (in ppm)



for R: H $\delta_{H_1}, \delta_{H_4}, \delta_{H_5}, \delta_{H_8} = 7.84$
$\delta_{H_2}, \delta_{H_3}, \delta_{H_6}, \delta_{H_7} = 7.48$

	Substituent R	H-2	H-3	H-4	H-5	H-6	H-7	H-8
C	$-CH_3$	-0.20	-0.14	-0.17	-0.03	-0.03	0.00	0.12
	$-CH_2CH_3$	-0.15	-0.08	-0.15	0.00	-0.02	0.01	0.21
	$-CH_2C\equiv CH$	0.09	-0.23	-0.23	-0.17	-0.13	-0.03	0.52
	$-CH_2Cl$	-0.10	-0.18	-0.11	-0.07	-0.05	0.02	0.22
	$-CF_3$	0.51	-0.01	0.01	0.06	0.07	0.13	0.35
	$-CH_2OH$	-0.07	-0.10	-0.09	-0.01	-0.02	0.01	0.18
	$-CH_2NH_2$	-0.14	-0.13	-0.14	-0.05	-0.07	-0.03	0.10
	$-C\equiv CH$	0.22	-0.14	-0.08	-0.08	-0.04	0.05	0.51
	–phenyl	-0.11	-0.04	-0.06	0.05	-0.13	-0.07	-0.02
X	–F	-0.35	-0.10	-0.23	0.00	0.05	0.03	0.27
	–Cl	0.06	-0.14	-0.12	-0.02	0.02	0.09	0.42
	–Br	0.29	-0.25	-0.04	-0.01	0.04	0.11	0.39
	–I	0.20	-0.46	0.13	-0.22	-0.09	-0.02	0.18
O	–OH	-0.75	-0.22	-0.42	-0.05	-0.03	-0.01	0.32
	$-OCH_3$	-0.84	-0.25	-0.55	-0.18	-0.12	-0.13	0.33
	–O–phenyl	-0.53	-0.10	-0.22	0.03	0.04	0.00	0.37
	$-OCOCH_3$	-0.31	-0.05	-0.27	-0.14	-0.23	-0.09	-0.01
N	$-NH_2$	-0.76	-0.22	-0.55	-0.07	-0.05	-0.08	-0.09
	$-N(CH_3)_2$	-0.46	-0.13	-0.36	-0.06	-0.03	-0.06	0.38
	$-NHCOCH_3$	0.24	0.01	-0.12	0.08	0.04	0.08	0.27
	$-NO_2$	0.71	0.02	0.24	0.08	0.04	0.16	0.69
	$-C\equiv N$	0.42	0.04	0.23	0.07	0.14	0.21	0.38
	–NCO	-0.25	-0.13	-0.18	-0.03	0.05	0.01	0.21
	–NCS	-0.13	-0.13	-0.14	-0.04	0.03	0.08	0.21
C=O	–CHO	0.46	0.07	0.20	0.05	0.06	0.17	1.39
	$-COCH_3$	0.40	-0.02	0.09	0.01	0.02	0.09	0.90
	–COOH	0.72	0.14	0.34	0.20	0.13	0.20	1.09
	$-COOCH_3$	0.65	-0.08	0.08	-0.05	-0.02	0.09	1.09
	–COCl	1.03	0.06	0.22	0.02	0.07	0.17	0.88

Effect of Substituents in Position 2 on the ^{1}H Chemical Shifts of Monosubstituted Naphthalenes (in ppm)



for R: H $\delta_{H_1}, \delta_{H_4}, \delta_{H_5}, \delta_{H_8} = 7.84$

$\delta_{H_2}, \delta_{H_3}, \delta_{H_6}, \delta_{H_7} = 7.48$

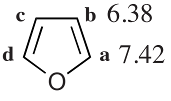
	Substituent R	H-1	H-3	H-4	H-5	H-6	H-7	H-8
C	$-CH_3$	-0.24	-0.18	-0.11	-0.06	-0.09	-0.05	-0.10
	$-CH_2CH_3$	-0.22	-0.14	-0.08	-0.05	-0.08	-0.06	-0.08
	$-CH(CH_3)_2$	-0.24	-0.15	-0.12	-0.10	-0.12	-0.10	-0.10
	$-CF_3$	0.28	0.14	0.06	-0.10	0.09	0.06	-0.10
	$-CH_2OH$	-0.13	-0.08	-0.07	-0.05	-0.04	-0.03	-0.08
	$-CH=CH_2$	-0.11	0.14	-0.06	-0.06	-0.06	-0.04	-0.06
	–C≡CH	0.19	0.04	-0.05	-0.03	0.02	0.02	-0.03
	–phenyl	0.20	0.25	0.06	0.01	0.02	-0.02	0.05
X	–Cl	-0.04	-0.08	-0.10	-0.05[a]	-0.03[b]	-0.01[b]	-0.12[a]
	–Br	0.14	0.05	-0.16	-0.12	-0.02	0.00	-0.06
O	–OH	-0.72	-0.39	-0.10	-0.09	-0.16	-0.06	-0.18
	$-OCH_3$	-0.76	-0.33	-0.14	-0.10	-0.14	-0.06	-0.14
	–O–phenyl	-0.53	-0.22	-0.01	-0.02	-0.08	-0.04	-0.15
	$-OCOCH_3$	-0.30	-0.27	-0.04	-0.04	-0.04	-0.02	-0.08
N	$-NH_2$	-0.93	-0.62	-0.23	-0.19	-0.27	-0.15	-0.27
	$-N(CH_3)_2$	-1.07	-0.49	-0.30	-0.29	-0.39	-0.24	-0.33
	$-NHCOCH_3$	0.33	-0.02	-0.10	-0.11	-0.09	-0.06	-0.09
	$-NO_2$	0.90	0.70	0.05	0.05	0.19	0.15	0.14
	–C≡N	0.40	0.13	0.08	0.06	0.19	0.13	0.06
S	–SH	-0.14	-0.19	-0.17	-0.11	-0.09	-0.06	-0.19
O=C	–CHO	0.44	0.45	0.05	0.03	0.14	0.08	0.12
	$-COCH_3$	0.58	0.51	0.01	0.01	0.08	0.03	0.01
	–CO–phenyl	0.42	0.46	0.09	0.06	0.13	0.06	0.06
	–COOH	0.83	0.57	0.20	0.19	0.20	0.16	0.31
	$-COOCH_3$	0.66	0.50	-0.08	-0.07	-0.01	-0.05	0.00
	–COCl	0.85	0.58	0.22	0.32	0.17	0.21	0.20

[a] interchangeable; [b] interchangeable

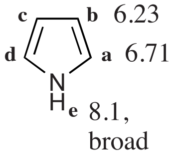
5.6 Heteroaromatic Compounds

5.6.1 Non-Condensed Heteroaromatic Rings

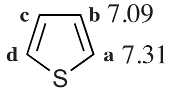
¹H Chemical Shifts and Coupling Constants (δ in ppm, |J| in Hz)



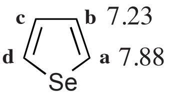
$^3J_{ab}$ 1.8
$^4J_{ac}$ 0.9
$^4J_{ad}$ 1.5
$^3J_{bc}$ 3.4



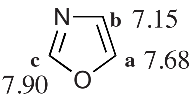
$^3J_{ab}$ 2.6
$^4J_{ac}$ 1.3
$^4J_{ad}$ 2.1
$^3J_{ae}$ 2.6
$^3J_{bc}$ 3.5
$^4J_{be}$ 2.6



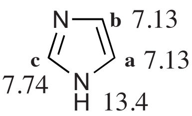
$^3J_{ab}$ 4.8
$^4J_{ac}$ 1.0
$^4J_{ad}$ 2.8
$^3J_{bc}$ 3.5



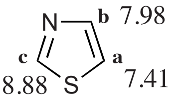
$^3J_{ab}$ 5.4
$^4J_{ac}$ 1.5
$^4J_{ad}$ 2.3
$^3J_{bc}$ 3.7



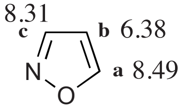
$^3J_{ab}$ 0.8
$^4J_{ac}$ 0.5
$^4J_{bc}$ 0.0



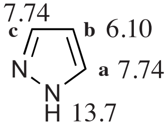
$^3J_{ab}$ 1–2
$^4J_{ac}$ 1.0
$^4J_{bc}$ 1.0



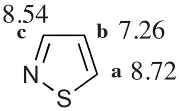
$^3J_{ab}$ 3.2
$^4J_{ac}$ 1.9
$^4J_{bc}$ 0.0



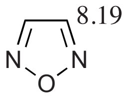
$^3J_{ab}$ 1.7
$^4J_{ac}$ 0.3
$^3J_{bc}$ 1.8

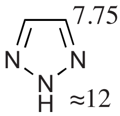


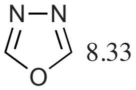
$^3J_{ab}$ 2.1
$^4J_{ac}$ 0.0
$^3J_{bc}$ 2.1

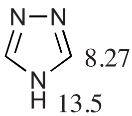


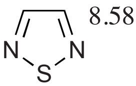
$^3J_{ab}$ 4.7
$^4J_{ac}$ 0.4
$^3J_{bc}$ 1.7

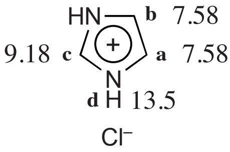




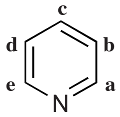






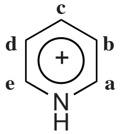


$^4J_{ac}$ 1.4
$^3J_{ad} + {}^4J_{bd} = 4.4$
$^3J_{cd}$ 2.4



Solvent:	$CDCl_3$	DMSO
a	8.59	8.58
b	7.25	7.38
c	7.62	7.75

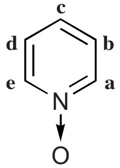
$^3J_{ab}$ 6.0, $^4J_{ac}$ 1.9, $^5J_{ad}$ 0.9, $^4J_{ae}$ 0.4, $^3J_{bc}$ 7.6, $^4J_{bd}$ 1.6



Solvent:	$CDCl_3$*	DMSO**
a	9.00	8.98
b	7.97	8.14
c	8.43	8.67

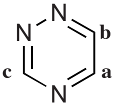
* p-toluylsulfonate
** HSO_3^-

$^3J_{ab}$ 6.0, $^4J_{ac}$ 1.6, $^5J_{ad}$ 0.8, $^4J_{ae}$ 1.0, $^3J_{bc}$ 7.9, $^4J_{bd}$ 1.4

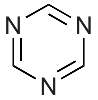


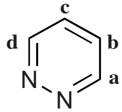
(in acetone)

$^3J_{ab}$ 6.5, $^4J_{ac}$ 1.1, $^5J_{ad}$ 0.6, $^4J_{ae}$ 1.9, $^3J_{bc}$ 7.7, $^4J_{bd}$ 2.1

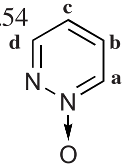


$^3J_{ab}$ 2.7, $^4J_{ac}$ 0.0, $^5J_{bc}$ 2.2

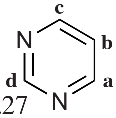




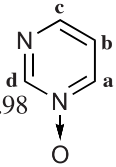
$^3J_{ab}$ 4.9, $^4J_{ac}$ 2.0, $^5J_{ad}$ 3.5, $^3J_{bc}$ 8.4



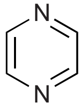
$^3J_{ab}$ 5.3, $^4J_{ac}$ 1.0, $^5J_{ad}$ 1.0, $^3J_{bc}$ 8.0, $^4J_{bd}$ 2.5, $^3J_{cd}$ 6.5

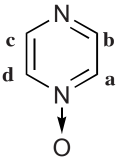


$^3J_{ab}$ 5.0, $^4J_{ac}$ 2.5, $^4J_{ad}$ 0, $^5J_{bd}$ 1.5



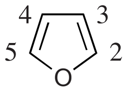
$^3J_{ab}$ 6.8, $^4J_{ac}$ 1.6, $^4J_{ad}$ 2.0, $^3J_{bc}$ 4.9, $^5J_{bd}$ 1.0, $^4J_{cd}$ 0





$^3J_{ab}$ 4.1, $^5J_{ac}$ 0.8, $^4J_{ad}$ 0.6, $^4J_{bc}$ 0.4

Effect of Substituents on the ^{1}H Chemical Shifts of Monosubstituted Furans (in ppm)

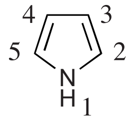


$$\delta_{H\text{-}2} = 7.42 + Z_{i2}$$
$$\delta_{H\text{-}3} = 6.38 + Z_{i3}$$
$$\delta_{H\text{-}4} = 6.38 + Z_{i4}$$
$$\delta_{H\text{-}5} = 7.42 + Z_{i5}$$

	Substituent	**H_3**	**H_4**	**H_5**	**H_2**	**H_4**	**H_5**
		in position 2 or 5			in position 3 or 4		
		Z_{23} Z_{54}	Z_{24} Z_{53}	Z_{25} Z_{52}	Z_{32} Z_{45}	Z_{34} Z_{43}	Z_{35} Z_{42}
C	$-CH_3$	-0.45	-0.15	-0.17	-0.25	-0.17	-0.12
	$-CH_2CH_3$	-0.42	-0.12	-0.14			
	$-CH_2OH$	-0.12	-0.07	-0.05	-0.07	0.00	-0.06
	$-CH_2SH$	-0.22	-0.09	-0.09			
	$-CH_2SCH_3$	-0.21	-0.09	-0.08			
	$-CH{=}CHCOCH_3$ (*trans*)	0.29	0.11	0.08			
X	–Br	-0.23	-0.17	-0.17			
	–I	0.04	-0.21	-0.05	-0.17	-0.04	-0.26
O	$-OCH_3$	-1.26	-0.14	-0.57	-0.50	-0.36	-0.41
N	$-NO_2$	1.13	0.47	0.47			
	–C≡N	0.48	-0.02	-0.04	0.41	0.14	-0.06
S	$-SCH_3$	0.05	0.01	0.13	-0.22	-0.13	-0.19
	–SCN	0.32	-0.02	0.06	0.15	0.11	-0.01
O=C	–CHO	0.92	0.25	0.31	0.92	0.47	0.19
	$-COCH_3$	0.81	0.16	0.18	0.42	0.28	-0.16
	–COCO–2-furyl	1.26	0.27	0.37			
	–COOH	0.97	0.19	0.24	0.70	0.40	0.03
	$-COOCH_3$	0.81	0.14	0.18	0.60	0.37	0.01
	–COCl	1.14	0.32	0.46			
M	$-P(-x\text{-furyl})_2$	0.25[a]	-0.12[a]	0.03[a]	-0.16[b]	-0.10[b]	-0.09[b]
	$-P(O)(-x\text{-furyl})_2$	0.76[a]	0.15[a]	0.30[a]	0.14[b]	0.19[b]	0.31[b]
	$-P(S)(-x\text{-furyl})_2$	0.77[a]	0.12[a]	0.27[a]	0.10[b]	0.18[b]	0.30[b]
	$-P^+(CH_3)(2\text{-furyl})_2\ I^-$	1.53	0.49	0.77			
	–HgCl				-0.09	0.02	0.25
	–Hg–x-furyl	0.18[a]	0.24[a]	0.47[a]	-0.10[b]	0.10[b]	-0.10[b]

[a] x = 2, [b] x = 3

Effect of Substituents on the ^{1}H Chemical Shifts of Monosubstituted Pyrroles (in ppm)



$\delta_{H\text{-}1} \approx 8$, broad, solvent-dependent

$\delta_{H\text{-}2} = 6.71 + Z_{i2}$

$\delta_{H\text{-}3} = 6.23 + Z_{i3}$

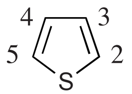
$\delta_{H\text{-}4} = 6.23 + Z_{i4}$

$\delta_{H\text{-}5} = 6.71 + Z_{i5}$

Substituent in position 1	**H_3** Z_{12} Z_{15}	**H_4** Z_{13} Z_{14}
$–CH_3$	-0.13	-0.11
$–CH_2CH_3$	-0.16	-0.12
$–CH_2CH_2CN$	-0.05	-0.07
$–CH_2$–phenyl	-0.12	-0.04
–phenyl	0.36	0.11
$–N(CH_3)_2$	0.11	-0.19
$–COCH_3$	0.56	0.12
–CO–phenyl	0.57	0.18
$–Si(CH(CH_3)_2)_3$	0.08	0.08

	Substituent	**H_3** in position 2 or 5 Z_{23} Z_{54}	**H_4** in position 2 or 5 Z_{24} Z_{53}	**H_5** in position 2 or 5 Z_{25} Z_{52}	**H_2** in position 3 or 4 Z_{32} Z_{45}	**H_4** in position 3 or 4 Z_{34} Z_{43}	**H_5** in position 3 or 4 Z_{35} Z_{42}
C	$–CH_3$	5.72	5.89	6.36	-0.33	-0.16	-0.26
N	$–NO_2$	7.11	6.29	7.05	1.06	0.24	0.43
	–C≡N	6.88	6.28	7.13	0.83	0.23	0.51
S	$–SCH_3$	6.23	6.10	6.72	0.18	0.05	0.10
	–SCN	6.53	6.15	6.90	0.48	0.10	0.28
O=C	–CHO	7.01	6.34	7.18	0.78	0.11	0.47
	$–COCH_3$	6.93	6.26	7.06	0.70	0.03	0.35
	$–COOCH_3$	6.84	6.18	6.91	0.79	0.13	0.29

Effect of Substituents on the ^{1}H Chemical Shifts of Monosubstituted Thiophenes (in ppm)

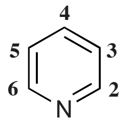


$$\delta_{H\text{-}2} = 7.31 + Z_{i2}$$
$$\delta_{H\text{-}3} = 7.09 + Z_{i3}$$
$$\delta_{H\text{-}4} = 7.09 + Z_{i4}$$
$$\delta_{H\text{-}5} = 7.31 + Z_{i5}$$

	Substituent	H_3 in position 2 or 5 Z_{23} Z_{54}	H_4 in position 2 or 5 Z_{24} Z_{53}	H_5 in position 2 or 5 Z_{25} Z_{52}	H_2 in position 3 or 4 Z_{32} Z_{45}	H_4 in position 3 or 4 Z_{34} Z_{43}	H_5 in position 3 or 4 Z_{35} Z_{42}
C	$-CH_3$	-0.34	-0.20	-0.24	-0.45	-0.22	-0.15
	–C≡C	0.02	-0.29	-0.23			
	–phenyl				0.11	0.28	0.05
	–2-thienyl	0.08	-0.09	-0.11			
	–2-pyridyl	0.48	0.01	0.06			
X	–F	-0.78	-0.54	-0.86	-0.80	-0.40	-0.31
	–Cl	-0.30	-0.35	-0.39	-0.25	-0.17	-0.09
	–Br	-0.05	-0.23	-0.10	-0.23	-0.21	-0.21
	–I	0.11	-0.34	-0.01	-0.05	-0.13	-0.30
O	–OH*	-0.85	0.44	-3.21			
	$-OCH_3$	-0.93	-0.41	-0.82	-1.10	-0.36	-0.17
N	$-NH_2$	-1.08	-0.58	-0.96	-1.36	-0.66	-0.36
	$-NO_2$	0.69	-0.16	0.19	0.84	0.47	-0.08
	–C≡N	0.34	-0.13	0.17	0.52	0.07	0.04
S	–SH	-0.13	-0.33	-0.18	-0.33	-0.33	-0.21
	$-SCH_3$	-0.16	-0.31	-0.16	-0.44	-0.23	-0.14
	$-S(O)_2CH_3$	0.90	0.07	0.68	0.85	0.35	0.35
	$-S(O)_2Cl$	0.60	-0.07	0.34			
	–SCN	0.17	-0.18	0.17	0.14	-0.08	-0.06
O=C	–CHO	0.69	0.13	0.47	0.81	0.44	0.06
	$-COCH_3$	0.60	0.03	0.32	0.74	0.45	0.01
	–CO–phenyl	0.55	0.06	0.40			
	–COOH	0.67	-0.05	0.29	0.93	0.48	0.03
	$-COOCH_3$	0.70	0.00	0.22	0.67	0.34	-0.16
	$-CONHNH_2$	0.63	0.04	0.41	-7.31	-7.09	-7.31
	–COCl	0.75	-0.07	0.33	0.94	0.37	-0.08

* Keto form

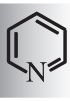
Effect of Substituents on the ^{1}H Chemical Shifts of Monosubstituted Pyridines (in ppm)



$\delta_{H\text{-}2} = 8.59 + Z_{i2}$
$\delta_{H\text{-}3} = 7.25 + Z_{i3}$
$\delta_{H\text{-}4} = 7.62 + Z_{i4}$
$\delta_{H\text{-}5} = 7.25 + Z_{i5}$
$\delta_{H\text{-}6} = 8.59 + Z_{i6}$

	Substituent in position 2 or 6	H_3 Z_{23} Z_{65}	H_4 Z_{24} Z_{63}	H_5 Z_{25} Z_{63}	H_6 Z_{26} Z_{62}
C	$-CH_3$	-0.11	-0.08	-0.15	-0.11
	$-CH_2CH_3$	-0.09	0.01	-0.15	0.03
	$-CH_2$–phenyl	0.03	-0.06	0.04	-0.04
	$-CH_2OH$	0.14	0.03	-0.08	-0.14
	$-CH{=}CH_2$	-0.07	-0.14	-0.23	-0.12
	–phenyl	0.42	0.02	-0.09	0.07
	–2-pyridyl	1.27	0.04	-0.11	0.00
X	–F	-0.30	0.16	-0.05	-0.36
	–Cl	0.09	0.02	0.00	-0.10
	–Br	0.26	-0.06	0.03	-0.23
	–I	0.49	-0.29	0.04	-0.23
O	–OH*	-0.63	-0.13	-0.93	-1.17
	$-OCH_3$	-0.51	-0.10	-0.41	-0.43
N	$-NH_2$	-0.76	-0.24	-0.63	-0.54
	$-NHCH_2CH_3$	-0.87	-0.22	-0.69	-0.52
	$-N(CH_3)_2$	-0.77	-0.23	-0.73	-0.44
	$-NHNH_2$	-0.55	-0.17	-0.58	-0.48
	$-NHCOCH_3$	1.00	0.09	-0.19	-0.32
	–NHN=CH–2-pyridyl	0.21	-0.01	-0.42	-0.36
	$-NO_2$	0.93	0.44	0.45	0.00
	–C≡N	0.52	0.26	0.35	0.15
S	–SH	0.34	-0.20	-0.42	-0.91
O=C	–CHO	0.73	0.26	0.31	0.21
	$-COCH_3$	0.80	0.22	0.24	0.10
	–CO–phenyl	0.81	0.27	0.25	0.13
	–COOH	0.87	0.41	0.44	0.17
	$-COOCH_3$	0.91	0.24	0.27	0.17
	$-CONH_2$	0.98	0.24	0.22	-0.01
	–CH=N–NH–2-pyridyl	0.76	0.05	-0.06	-0.03
	$-Si(CH_3)_3$	0.15	-0.22	-0.24	0.09

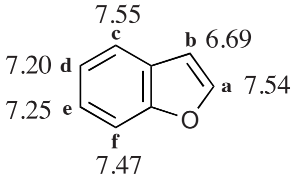
* Keto form (2-pyridone)



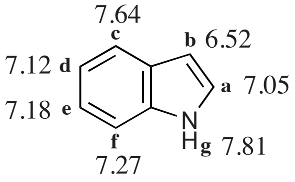
	Substituent	H_2	H_4	H_5	H_6	H_2	H_3
		in position 3 or 5				in position 4	
		Z_{32} Z_{56}	Z_{34} Z_{54}	Z_{35} Z_{53}	Z_{36} Z_{52}	Z_{42} Z_{46}	Z_{43} Z_{45}
C	$-CH_3$	-0.15	-0.16	-0.07	-0.17	-0.13	-0.13
	$-CH_2CH_3$	-0.13	-0.14	-0.06	-0.17	-0.12	-0.14
	$-CH_2$–phenyl	-0.08	-0.18	-0.04	-0.14	0.00	-0.15
	–phenyl	0.25	0.20	0.08	-0.03		
	$-CH=CH_2$					-0.12	-0.08
X	–F	-0.05	-0.21	0.04	-0.13	-0.07	-0.03
	–Cl	0.09	0.00	0.05	-0.05	0.00	0.05
	–Br	0.09	0.18	-0.04	-0.07	0.09	0.35
O	–OH	-0.31	-0.29	0.06	-0.50		
	$-OCH_3$	-0.27	-0.37	-0.04	-0.40	-0.16	-0.42
	$-OCOCH_3$	-0.15	-0.15	0.08	-0.13		
N	$-NH_2$	-0.51	-0.65	-0.20	-0.60	-0.15	-0.74
	$-N(CH_3)_2$					-0.38	-0.77
	$-NHCOCH_3$	0.37	0.50	0.06	-0.16	-0.19	0.16
	$-C\equiv N$	0.32	0.38	0.25	0.26	0.24	0.32
S	–S–phenyl					0.05	-0.16
	$-S(O)_2OH$	0.70	1.14	0.81	0.70		
O=C	–CHO	0.52	0.58	0.30	0.28	0.31	0.49
	$-COCH_3$	0.58	0.61	0.20	0.20	0.21	0.50
	–CO–phenyl					0.23	0.35
	–COOH	0.54	0.57	0.20	0.24	0.20	0.45
	$-COOCH_3$	0.64	0.67	0.16	0.19	0.19	0.61
	–COO–phenyl					0.24	0.75
	$-CONH_2$	0.49	0.50	0.15	0.15		
	$-CSNH_2$	0.68	0.67	0.24	0.26	0.35	0.68
	–CH=NOH	0.39	0.43	0.19	0.15	0.06	0.32
	$-Si(CH_3)$	0.08	0.00	-0.21	-0.11	-0.08	0.01

5.6.2 Condensed Heteroaromatic Rings

¹H Chemical Shifts and Coupling Constants (δ in ppm, |J| in Hz)

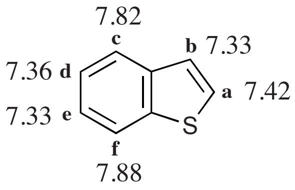


$^3J_{ab}$ 2.5
$^5J_{bf}$ 0.9
$^3J_{cd}$ 7.9
$^4J_{ce}$ 1.2
$^5J_{cf}$ 0.8
$^3J_{de}$ 7.3
$^4J_{df}$ 0.9
$^3J_{ef}$ 8.4
(all other coupling constants negligible)

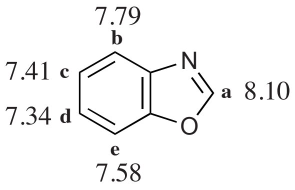


(chemical shifts in $CDCl_3$, coupling constants in acetone)

$^3J_{ab}$ 3.1
$^3J_{ag}$ 2.5
$^5J_{bf}$ 0.7
$^4J_{bg}$ 2.0
$^3J_{cd}$ 7.8
$^4J_{ce}$ 1.2
$^5J_{cf}$ 0.9
$^5J_{cg}$ 0.8
$^3J_{de}$ 7.1
$^4J_{df}$ 1.3
$^3J_{ef}$ 8.1
(all other coupling constants negligible)

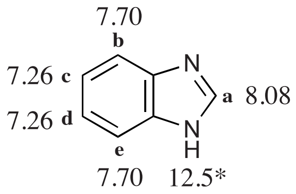


$^3J_{ab}$ 5.5
$^5J_{bf}$ 0.8
$^3J_{cd}$ 8.0
$^4J_{ce}$ 1.1
$^5J_{cf}$ 0.9
$^3J_{de}$ 7.2
$^4J_{df}$ 1.0
$^3J_{ef}$ 8.0
(all other coupling constants negligible)



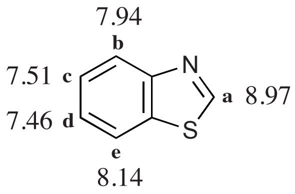
(chemical shifts in $CDCl_3$, coupling constants in acetone)

$^5J_{ab}$ 0.2
$^6J_{ac}$ -0.1
$^6J_{ad}$ 0.4
$^5J_{ae}$ 0.0
$^3J_{bc}$ 8.2
$^4J_{bd}$ 1.0
$^5J_{be}$ 0.7
$^3J_{cd}$ 7.4
$^4J_{ce}$ 1.2
$^3J_{de}$ 8.3



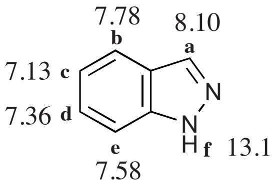
* in DMSO

$^3J_{bc}$ 8.2
$^4J_{bd}$ 1.4
$^5J_{be}$ 0.7
$^3J_{cd}$ 7.1
(all other coupling constants negligible)



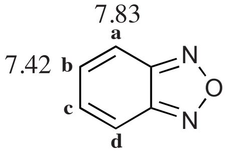
(chemical shifts in $CDCl_3$, coupling constants in acetone)

$^5J_{ab}$ 0.1
$^6J_{ac}$ -0.2
$^6J_{ad}$ 0.4
$^5J_{ae}$ 0.1
$^3J_{bc}$ 8.2
$^4J_{bd}$ 1.1
$^5J_{be}$ 0.6
$^3J_{cd}$ 7.2
$^4J_{ce}$ 1.1
$^3J_{de}$ 8.2

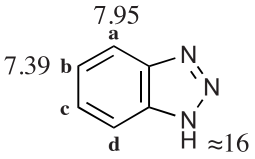


(in DMSO)

$^5J_{ae}$ 0.8
$^3J_{bc}$ 7.8
$^4J_{bd}$ 1.2
$^5J_{be}$ 1.0
$^3J_{cd}$ 7.0
$^4J_{ce}$ 1.2
$^3J_{de}$ 7.9
(all other coupling constants negligible)

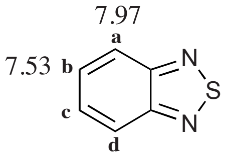


$^3J_{ab}$ 9.1
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.9
$^3J_{bc}$ 6.4

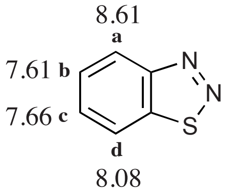


$^3J_{ab}$ 8.3
$^4J_{ac}$ 1.0
$^5J_{ad}$ 0.9
$^3J_{bc}$ 7.0

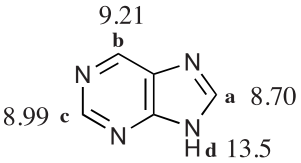
(chemical shifts in $CDCl_3$, coupling constants in acetone)



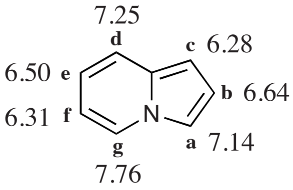
$^3J_{ab}$ 8.9
$^4J_{ac}$ 1.2
$^5J_{ad}$ 0.9
$^3J_{bc}$ 6.7



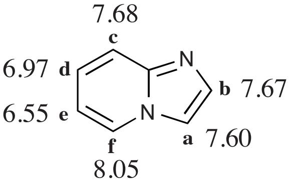
$^3J_{ab}$ 8.4
$^4J_{ac}$ 1.0
$^5J_{ad}$ 0.8
$^3J_{bc}$ 7.0
$^4J_{bd}$ 1.0
$^3J_{cd}$ 7.9



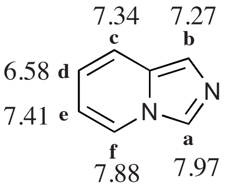
(in DMSO)



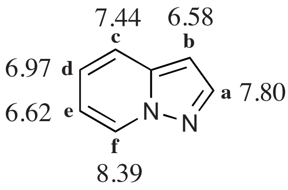
$^3J_{ab}$ 2.7
$^4J_{ac}$ 1.2
$^5J_{ad}$ 0.5
$^3J_{bc}$ 3.9
$^6J_{bf}$ 0.5
$^5J_{cg}$ 1.0
$^3J_{de}$ 9.0
$^4J_{df}$ 1.0
$^5J_{dg}$ 1.2
$^3J_{ef}$ 6.4
$^4J_{eg}$ 1.0
$^3J_{fg}$ 6.8
(all other coupling constants negligible)



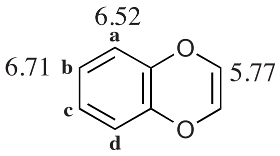
$^3J_{ab}$ 1.2
$^4J_{af}$ 0.7
$^3J_{cd}$ 9.0
$^4J_{ce}$ 1.3
$^5J_{cf}$ 1.3
$^3J_{de}$ 6.6
$^4J_{df}$ 1.3
$^3J_{ef}$ 6.8
(all other coupling constants negligible)



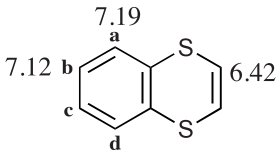
$^5J_{ac}$ 1.0
$^5J_{bf}$ 0.5
$^3J_{cd}$ 9.2
$^4J_{ce}$ 1.1
$^5J_{cf}$ 1.1
$^3J_{de}$ 6.4
$^4J_{df}$ 0.9
$^3J_{ef}$ 7.1
(all other coupling constants negligible)



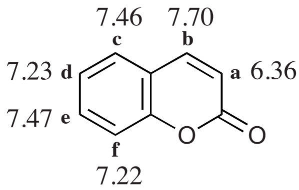
$^3J_{ab}$ 2.2
$^6J_{ad}$ 0.5
$^5J_{bf}$ 0.9
$^3J_{cd}$ 8.9
$^4J_{ce}$ 1.2
$^5J_{cf}$ 1.0
$^3J_{de}$ 6.8
$^4J_{df}$ 1.0
$^3J_{ef}$ 6.9
(all other coupling constants negligible)



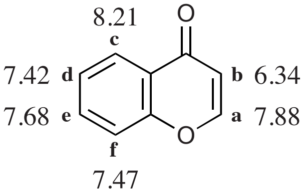
$^3J_{ab}$ 7.9
$^4J_{ac}$ 1.5
$^5J_{ad}$ 0.4
$^3J_{bc}$ 7.9



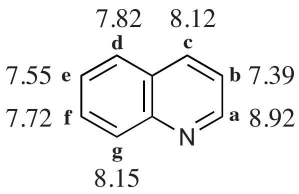
$^3J_{ab}$ 7.8
$^4J_{ac}$ 1.3
$^5J_{ad}$ 1.1
$^3J_{bc}$ 7.1



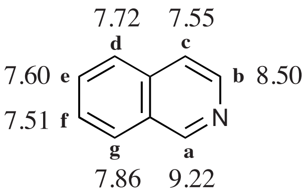
$^3J_{ab}$ 9.6
$^3J_{cd}$ 7.7
$^4J_{ce}$ 1.6
$^3J_{de}$ 7.4
$^4J_{df}$ 1.1
$^3J_{ef}$ 8.4
(all other coupling constants negligible)



$^3J_{ab}$ 6.0
$^3J_{cd}$ 8.0
$^4J_{ce}$ 1.8
$^5J_{cf}$ 0.5
$^3J_{de}$ 7.0
$^4J_{df}$ 1.1
$^3J_{ef}$ 8.4

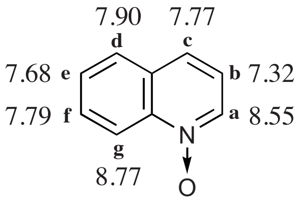


$^3J_{ab}$ 4.2
$^4J_{ac}$ 1.8
$^3J_{bc}$ 8.2
$^5J_{cg}$ 0.8
$^3J_{de}$ 8.2
$^4J_{df}$ 1.4
$^5J_{dg}$ 0.7
$^3J_{ef}$ 6.9
$^4J_{eg}$ 1.2
$^3J_{fg}$ 8.5
(all other coupling constants negligible)



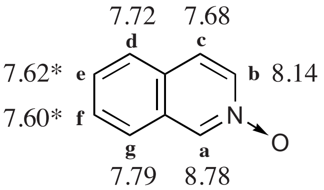
$^5J_{ac}$ 1.0
$^5J_{ad}$ 0.9
$^3J_{bc}$ 5.8
$^5J_{cg}$ 0.9
$^3J_{de}$ 8.3
$^4J_{df}$ 1.2
$^5J_{dg}$ 0.8
$^3J_{ef}$ 6.9
$^4J_{eg}$ 1.2
$^3J_{fg}$ 8.3

(all other coupling constants negligible)



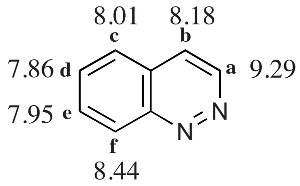
$^3J_{ab}$ 6.1
$^4J_{ac}$ 1.0
$^3J_{bc}$ 8.5
$^5J_{cg}$ 0.9
$^3J_{de}$ 8.2
$^4J_{df}$ 1.2
$^5J_{dg}$ 0.3
$^3J_{ef}$ 7.2
$^4J_{eg}$ 1.4
$^3J_{fg}$ 8.6

(all other coupling constants negligible)

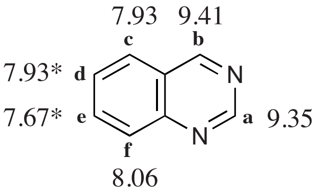


$^4J_{ab}$ 1.7
$^3J_{bc}$ 7.0

* assignment uncertain

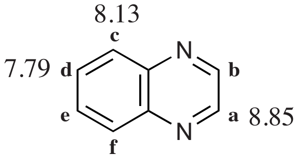


$^3J_{ab}$ 5.9
$^5J_{bf}$ 0.8
$^3J_{cd}$ 7.8
$^4J_{ce}$ 1.5
$^5J_{cf}$ 0.8
$^3J_{de}$ 6.9
$^4J_{df}$ 1.3
$^3J_{ef}$ 8.6

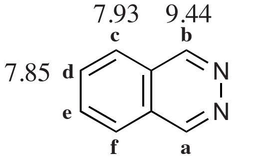


* assignment uncertain

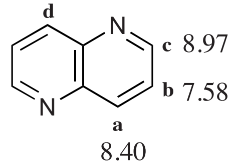
$^4J_{ab}$ 0.0
$^5J_{bf}$ 0.5
$^3J_{cd}$ 8.0
$^4J_{ce}$ 1.3
$^5J_{cf}$ 0.9
$^3J_{de}$ 7.0
$^4J_{df}$ 1.3
$^3J_{ef}$ 8.6



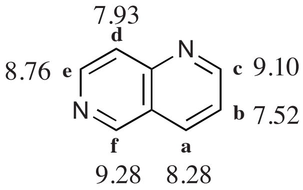
$^3J_{cd}$ 8.4
$^4J_{ce}$ 1.4
$^5J_{cf}$ 0.7
$^3J_{de}$ 6.9



$^5J_{ac}$ 0.4
$^3J_{cd}$ 8.2
$^4J_{ce}$ 1.2
$^5J_{cf}$ 0.6
$^3J_{de}$ 6.8

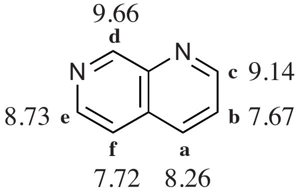


$^3J_{ab}$ 8.0
$^4J_{ac}$ 1.8
$^5J_{ad}$ 0.6
$^3J_{bc}$ 4.1



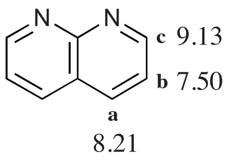
$^3J_{ab}$ 8.2
$^4J_{ac}$ 1.9
$^5J_{ad}$ 0.9
$^3J_{bc}$ 4.1
$^3J_{de}$ 6.0
$^5J_{df}$ 0.9

(all other coupling constants negligible)

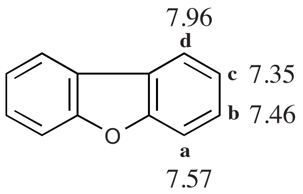


$^3J_{ab}$ 8.4
$^4J_{ac}$ 1.6
$^5J_{ad}$ 0.9
$^3J_{bc}$ 4.2
$^5J_{df}$ 0.9
$^3J_{ef}$ 5.6

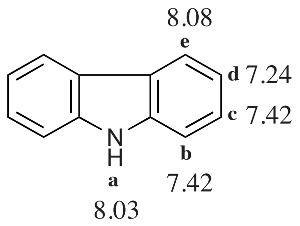
(all other coupling constants negligible)



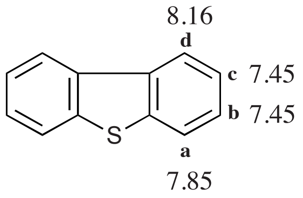
$^3J_{ab}$ 8.2
$^4J_{ac}$ 2.0
$^3J_{bc}$ 4.3



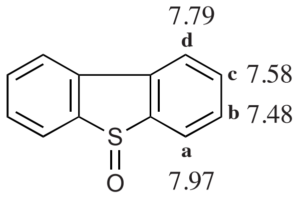
$^3J_{ab}$ 8.5
$^4J_{ac}$ 0.9
$^5J_{ad}$ 0.6
$^3J_{bc}$ 7.3
$^4J_{bd}$ 1.3
$^3J_{cd}$ 7.6



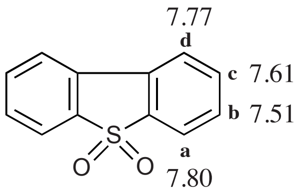
$^5J_{ae}$ 0.7
$^3J_{bc}$ 8.1
$^4J_{bd}$ 1.0
$^5J_{be}$ 0.7
$^3J_{cd}$ 7.2
$^4J_{ce}$ 1.2
$^3J_{de}$ 7.9



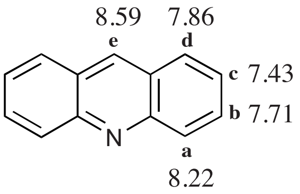
$^3J_{ab}$ 8.0
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.7
$^3J_{bc}$ 7.2
$^4J_{bd}$ 1.2
$^3J_{cd}$ 8.1



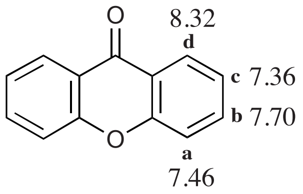
$^3J_{ab}$ 7.7
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.7
$^3J_{bc}$ 7.5
$^4J_{bd}$ 1.1
$^3J_{cd}$ 7.6



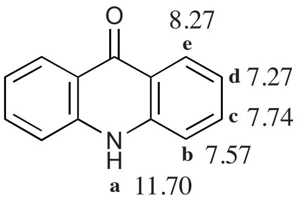
$^3J_{ab}$ 7.7
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.6
$^3J_{bc}$ 7.5
$^4J_{bd}$ 1.0
$^3J_{cd}$ 7.8



$^3J_{ab}$ 8.9
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.8
$^5J_{ae}$ 0.9
$^3J_{bc}$ 6.7
$^4J_{bd}$ 1.5
$^3J_{cd}$ 8.4
$^4J_{de}$ -0.5



$^3J_{ab}$ 8.4
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.5
$^3J_{bc}$ 7.1
$^4J_{bd}$ 1.8
$^3J_{cd}$ 8.0



$^5J_{ae}$ 0.4
$^3J_{bc}$ 8.6
$^4J_{bd}$ 1.0
$^5J_{be}$ 0.4
$^3J_{cd}$ 7.0
$^4J_{ce}$ 1.4
$^3J_{de}$ 8.2

5.7 Halogen Compounds

5.7.1 Fluoro Compounds

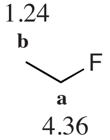
^{19}F (natural abundance 100%) has a spin quantum number I of 1/2. The signals of ^{1}H atoms are split by coupling to ^{19}F up to a distance of about four bonds.

^{1}H Chemical Shifts and Coupling Constants (δ in ppm, |J| in Hz)

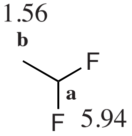
4.27 CH_3F $^{2}J_{HF}$ 46.4

5.45 CH_2F_2 $^{2}J_{HF}$ 50.2

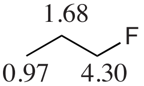
6.25 CHF_3 $^{2}J_{HF}$ 79.2

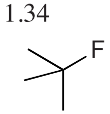


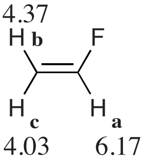
$^{2}J_{aF}$ 46.4
$^{3}J_{bF}$ 25.2
$^{3}J_{ab}$ 6.9



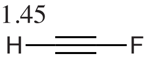
$^{2}J_{aF}$ 57.3
$^{3}J_{bF}$ 20.9
$^{3}J_{ab}$ 4.5



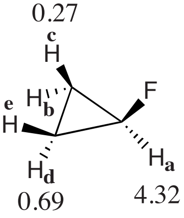




$^{2}J_{aF}$ 87.4 $^{3}J_{ab}$ 12.8
$^{3}J_{bF}$ 20.1 $^{3}J_{ac}$ 4.7
$^{3}J_{cF}$ 52.4 $^{2}J_{bc}$ -3.2

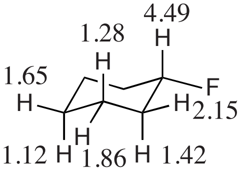


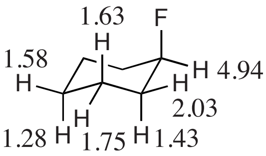
$^{3}J_{HF}$ 15.3

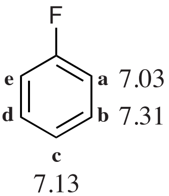


(in benzene/$CFCl_3$)

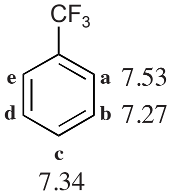
$^{2}J_{aF}$ 64.9
$^{3}J_{bF}$ 9.9
$^{3}J_{cF}$ 21.0
$^{3}J_{ab}$ 5.9
$^{3}J_{ac}$ 2.4
$^{2}J_{bc}$ -6.7
$^{3}J_{bd}$ 10.8
$^{3}J_{be}$ 7.7
$^{3}J_{ce}$ 12.0



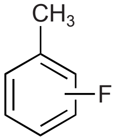




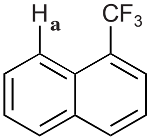
$^{3}J_{aF}$ 8.9 $^{3}J_{ab}$ 8.4
$^{4}J_{bF}$ 5.7 $^{4}J_{ac}$ 1.1
$^{5}J_{cF}$ 0.2 $^{5}J_{ad}$ 0.4
$^{4}J_{ae}$ 2.7
$^{3}J_{bc}$ 7.5
$^{4}J_{bd}$ 1.8



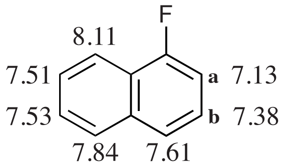
$^{4}J_{aF}$ -0.8 $^{3}J_{ab}$ 7.9
$^{5}J_{bF}$ 0.8 $^{4}J_{ac}$ 1.2
$^{6}J_{cF}$ -0.7 $^{5}J_{ad}$ 0.6
$^{4}J_{ae}$ 2.0
$^{3}J_{bc}$ 7.6
$^{4}J_{bd}$ 1.3



Couplings with CH_3:
^4J_o 2.5
^5J_m 0.0
^6J_p 1.5



$^5J_{aF}$ 2.2



$^3J_{aF}$ 10.7
$^4J_{bF}$ 5.4

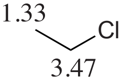
5.7.2 Chloro Compounds

^1H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

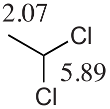
3.06 CH_3Cl

5.33 CH_2Cl_2

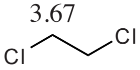
7.26 $CHCl_3$



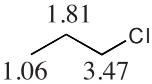
^3J 7.2

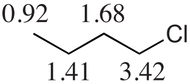


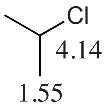
^3J 6.1



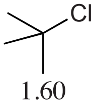
^3J 6.8

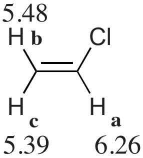




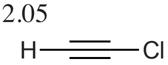


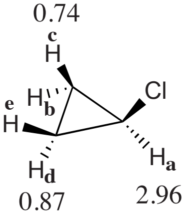
^3J 6.4



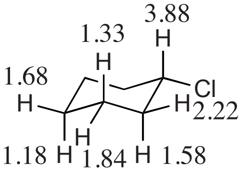


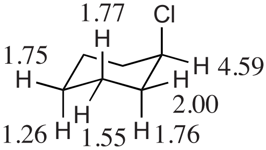
$^3J_{ab}$ 14.5
$^3J_{ac}$ 7.5
$^2J_{bc}$ -1.4

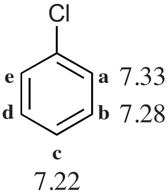




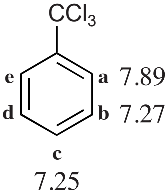
$^3J_{ab}$ 7.0
$^3J_{ac}$ 3.6
$^2J_{bc}$ -6.0
$^3J_{bd}$ 10.3
$^3J_{be}$ 7.1
$^3J_{ce}$ 10.6





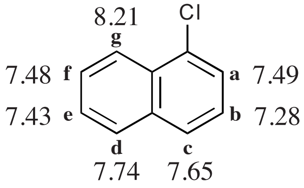


$^3J_{ab}$ 8.1
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.5
$^4J_{ae}$ 2.3
$^3J_{bc}$ 7.5
$^4J_{bd}$ 1.7

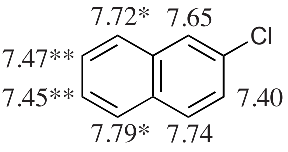


$^3J_{ab}$ 8.1
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.5
$^4J_{ae}$ 2.4
$^3J_{bc}$ 7.5
$^4J_{bd}$ 1.4

Hal



$^{3}J_{ab}$ 7.7
$^{4}J_{ac}$ 1.1
$^{3}J_{bc}$ 7.8
$^{5}J_{cg}$ 0.7
$^{3}J_{de}$ 8.1
$^{4}J_{df}$ 1.0
$^{5}J_{dg}$ 0.2
$^{3}J_{ef}$ 6.8
$^{4}J_{eg}$ 0.6
$^{3}J_{fg}$ 7.8



*; **: assignments interchangeable

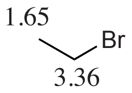
5.7.3 Bromo Compounds

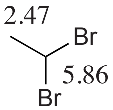
^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

2.68 CH_3Br

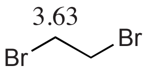
4.94 CH_2Br_2

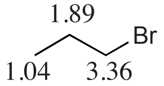
6.82 $CHBr_3$

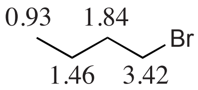


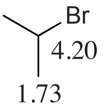


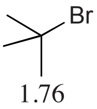
^{3}J 6.4

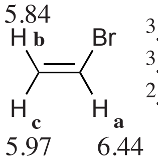




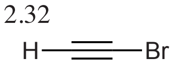


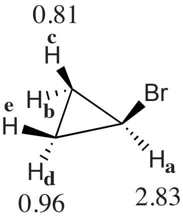




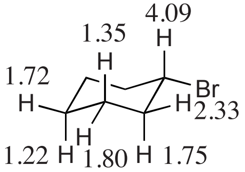


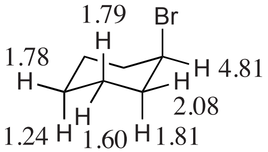
$^{3}J_{ab}$ 14.9
$^{3}J_{ac}$ 7.1
$^{2}J_{bc}$ -1.9

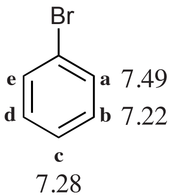




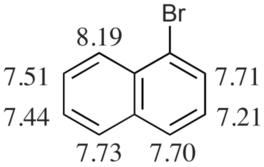
$^{3}J_{ab}$ 7.1
$^{3}J_{ac}$ 3.8
$^{2}J_{bc}$ -6.1
$^{3}J_{bd}$ 10.2
$^{3}J_{be}$ 7.0
$^{3}J_{ce}$ 10.5

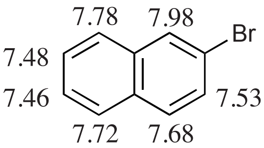






$^{3}J_{ab}$ 8.0
$^{4}J_{ac}$ 1.1
$^{5}J_{ad}$ 0.5
$^{4}J_{ae}$ 2.2
$^{3}J_{bc}$ 7.4
$^{4}J_{bd}$ 1.8





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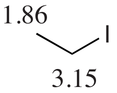
5.7.4 Iodo Compounds

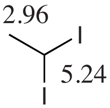
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

2.16
CH_3I

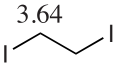
3.90
CH_2I_2

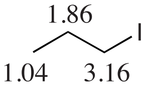
4.91
CHI_3

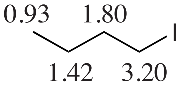


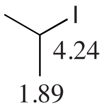


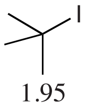
^3J 7.0

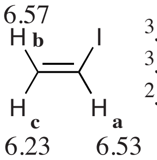




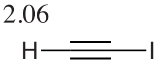


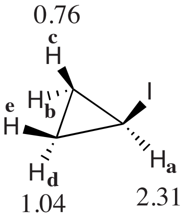




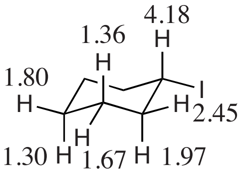


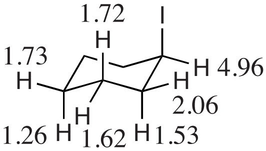
$^3J_{ab}$ 15.9
$^3J_{ac}$ 7.8
$^2J_{bc}$ -1.5

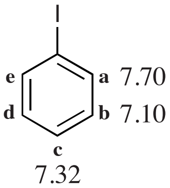




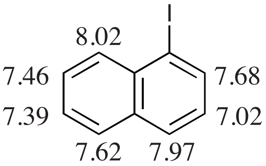
$^3J_{ab}$ 7.5
$^3J_{ac}$ 4.4
$^2J_{bc}$ -5.9
$^3J_{bd}$ 9.9
$^3J_{be}$ 6.6
$^3J_{ce}$ 10.0







$^3J_{ab}$ 8.0
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.5
$^4J_{ae}$ 2.2
$^3J_{bc}$ 7.4
$^4J_{bd}$ 1.8



Hal

5.8 Alcohols, Ethers, and Related Compounds

5.8.1 Alcohols

^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

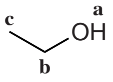
Aliphatic and alicyclic alcohols:	δ_{OH} = 0.5–3.0 (in DMSO: 4–6) ppm
Phenols:	δ_{OH} = 4.0–8.0 (in DMSO: 8–12) ppm

Hydrogen bonds strongly deshield hydroxyl protons. The position of the signal may depend heavily on the experimental conditions including the concentration of the sample. If a compound contains several kinds of hydroxyl protons (–OH, –COOH, H_2O), in general only one signal at an average position is seen because of rapid exchange. In dimethyl sulfoxide (DMSO) as solvent, this exchange in most cases is so slow that isolated signals are obtained. In this case, the chemical shifts of hydroxyl protons are characteristic. However, if the sample contains strong acids or amine bases, the exchange rate increases and, also in DMSO, a single signal at an average position is observed. Frequently, intermediate exchange rates lead to very broad signals extending over several ppm and, therefore, sometimes not discernible in routine spectra.

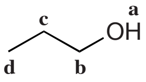
As a consequence of fast intermolecular exchange of the hydroxyl protons, their coupling with the protons on the adjacent carbon atoms is usually not observed. However, in very pure (acid-free) solutions or in DMSO, the exchange is sufficiently slow so that the H–O–C–H couplings become visible. Their dependence on the conformation is analogous to that shown by the H–C–C–H couplings (Chapter 5.1.2). In case of fast rotation: $^3J_{HOCH}$ ≈ 5 Hz. In cyclohexanols, the vicinal coupling constants for axial hydroxyl protons (3.0–4.2 Hz) are lower than those of equatorial ones (4.2–5.7 Hz).

b **a**
CH_3OH

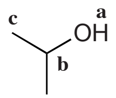
	$CDCl_3$	DMSO	D_2O
a	1.13	4.05	
b	3.49	3.17	3.34
		$^3J_{ab}$ 5.2	



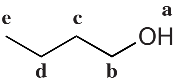
	$CDCl_3$	DMSO	D_2O
a	1.51	4.31	
b	3.71	3.44	3.65
c	1.24	1.06	1.17
		$^3J_{ab}$ 4.8	
		$^3J_{bc}$ 6.9	



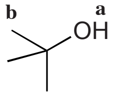
	$CDCl_3$	DMSO	D_2O
a	1.51	4.31	
b	3.59	3.34	3.61
c	1.59	1.42	1.57
d	0.94	0.84	0.89



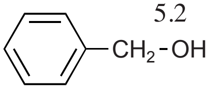
	CDCl$_3$	DMSO	D$_2$O
a	1.36	4.30	
b	4.04	3.78	4.02
c	1.22	1.04	1.17
		$^3J_{ab}$ 6.2	



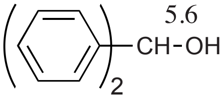
	CDCl$_3$	DMSO	D$_2$O
a	1.50	4.30	
b	3.64	3.38	3.61
c	1.56	1.40	1.51
d	1.39	1.30	1.35
e	0.94	0.87	0.91



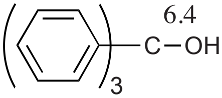
	CDCl$_3$	DMSO	D$_2$O
a	1.37	4.19	
b	1.28	1.11	1.24



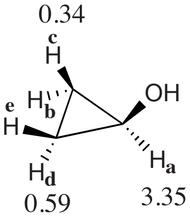
(in DMSO)



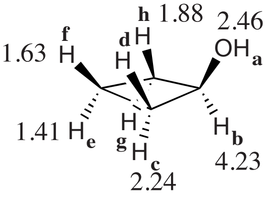
(in DMSO)



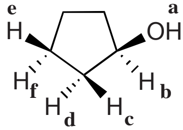
(in DMSO)



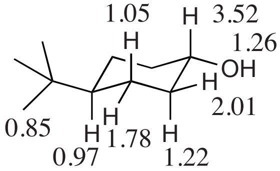
$^3J_{ab}$	6.2
$^3J_{ac}$	2.9
$^2J_{bc}$	-5.4
$^3J_{bd}$	10.3
$^3J_{be}$	6.8
$^3J_{ce}$	10.9



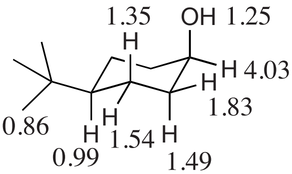
$^3J_{bc}$	7.0	$^3J_{cf}$	2.3
$^3J_{bd}$	8.1	$^4J_{cg}$	5.2
$^4J_{be}$	0	$^3J_{de}$	10.4
$^4J_{bf}$	-1.1	$^3J_{df}$	9.7
$^2J_{cd}$	-11.0	$^4J_{dg}$	-0.9
$^3J_{ce}$	7.9	$^4J_{dh}$	0
		$^3J_{ef}$	-11.0



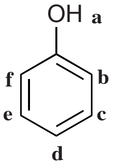
	CDCl$_3$	DMSO
a	1.28	4.33
b	4.32	4.09
c	1.56	1.44
d	1.76	1.61
e	1.76	1.61
f	1.56	1.44



Derivatives in DMSO:
δ_{OH} 4.0–4.5
$J_{CH,OH}$ 4.2–5.7

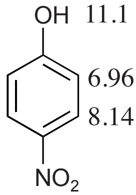


Derivatives in DMSO:
δ_{OH} 3.8–4.2
$J_{CH,OH}$ 3.0–4.2

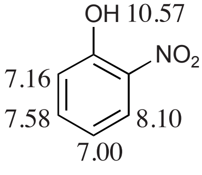


	CDCl$_3$	DMSO
a	4.69	9.29
b	6.83	6.75
c	7.24	7.15
d	6.93	6.76

$^3J_{bc}$	8.0
$^4J_{bd}$	1.1
$^5J_{be}$	0.5
$^4J_{bf}$	2.2
$^3J_{cd}$	7.4
$^4J_{ce}$	1.8



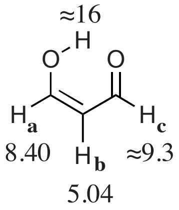
in DMSO



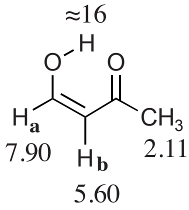
in CDCl$_3$

O

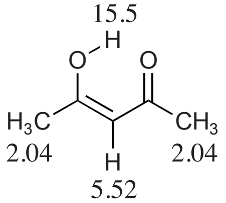
¹H Chemical Shifts and Coupling Constants of Enols (δ in ppm, J in Hz)



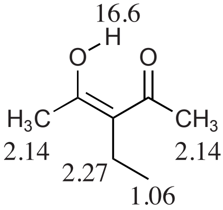
$^3J_{ab}$ 9.7
$^3J_{bc}$ ≈8

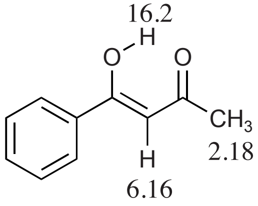


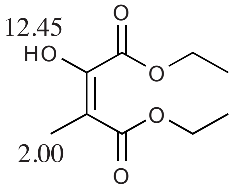
$^3J_{ab}$ 5.1



(in $CDCl_3$, partly enolized; for the keto form, see Chapter 5.11.2)



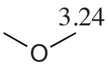




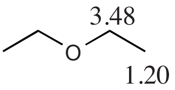
(in $CDCl_3$, partly enolized)

5.8.2 Ethers

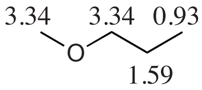
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

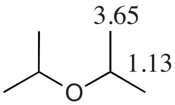


$^2J_{gem}$ -10.6

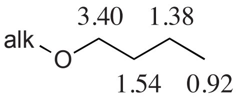


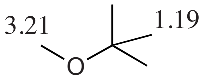
$^3J_{vic}$ 7.0

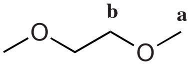




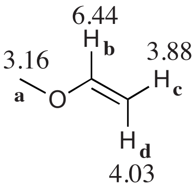
$^3J_{vic}$ 6.1





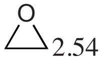


	$CDCl_3$	DMSO	D_2O
a	3.40	3.24	3.37
b	3.55	3.43	3.60



$^4J_{ab}$ 0.3
$^3J_{bc}$ 7.0
$^3J_{bd}$ 14.1
$^2J_{cd}$ -2.0

^{1}H Chemical Shifts and Coupling Constants of Cyclic Ethers (δ in ppm, J in Hz)

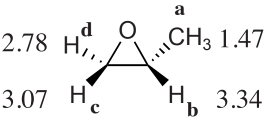


In derivatives:

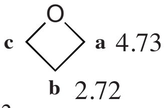
$^2J_{gem}$	5–6
$^3J_{cis}$	4.5
$^3J_{trans}$	3.1

Throughout:

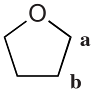
$J_{cis} > J_{trans}$



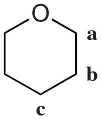
$^3J_{ab}$	5.2
$^3J_{bc}$	4.2
$^3J_{bd}$	3.3
$^2J_{cd}$	4.2



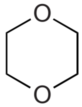
$^2J_{a,gem}$	-5.8
$^2J_{b,gem}$	-11.0
$^3J_{cis}$	8.7
$^3J_{trans}$	6.6
$\lvert ^4J_{ac} \rvert$	<0.3



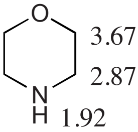
	$CDCl_3$	DMSO
a	3.74	3.60
b	1.85	1.76

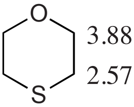


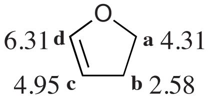
	$CDCl_3$	DMSO
a	3.65	3.53
b	1.57	1.47
c	1.64	1.58



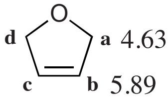
$CDCl_3$	DMSO	D_2O
3.71	3.57	3.75



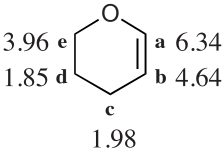




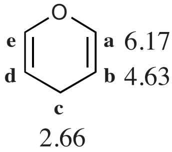
$^3J_{ab,cis}$	8.3
$^3J_{ab,trans}$	10.7
$^3J_{bc}$	2.5
$^4J_{bd}$	2.6
$^3J_{cd}$	2.6



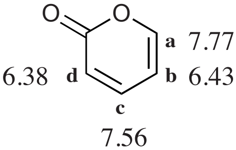
$^3J_{ab}$	1.6
$^4J_{ac}$	-2.5
$^4J_{ad,cis}$	7.1
$^4J_{ad,trans}$	4.6
$^3J_{bc}$	6.3



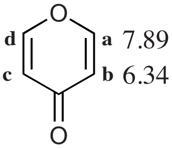
$^3J_{ab}$	6.2
$^4J_{ac}$	2.0
$^3J_{bc}$	3.8
$^4J_{bd}$	0.6



$^3J_{ab}$	7.0
$^4J_{ac}$	1.7
$^4J_{ae}$	1.5
$^3J_{bc}$	3.4



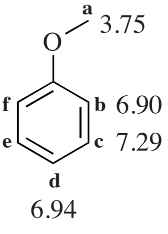
$^3J_{ab}$	5.0
$^4J_{ac}$	2.4
$^5J_{ad}$	1.2
$^3J_{bc}$	6.3
$^4J_{bd}$	1.5
$^3J_{cd}$	9.4



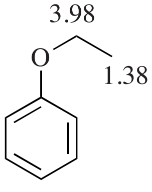
$^3J_{ab}$	6.0
$^5J_{ac}$	0.3
$^4J_{ad}$	2.7
$^4J_{bc}$	1.1

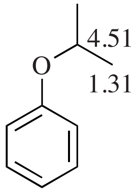
O

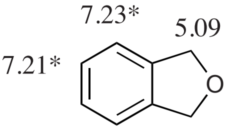
^{1}H Chemical Shifts and Coupling Constants of Aromatic Ethers (δ in ppm, J in Hz)



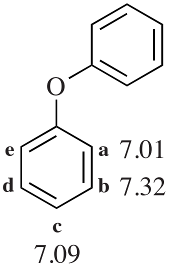
$^{5}J_{ab}$ ≈0.8
$^{3}J_{bc}$ 8.3
$^{4}J_{bd}$ 1.0
$^{5}J_{be}$ 0.4
$^{4}J_{bf}$ 2.7
$^{3}J_{cd}$ 7.4
$^{4}J_{ce}$ 1.8





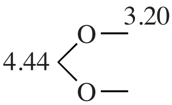


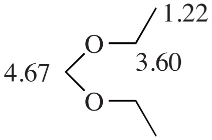
* assignment uncertain

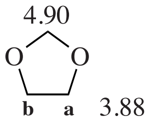


$^{3}J_{ab}$ 8.3
$^{4}J_{ac}$ 1.1
$^{5}J_{ad}$ 0.5
$^{4}J_{ae}$ 2.6
$^{3}J_{bc}$ 7.4
$^{4}J_{bd}$ 1.7

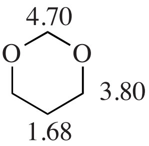
^{1}H Chemical Shifts and Coupling Constants of Acetals, Ketals, and Ortho Esters (δ in ppm, J in Hz)

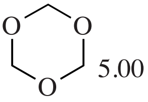


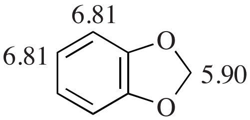




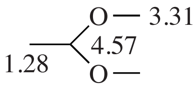
$^{2}J_{a,gem}$ -7.5
$^{3}J_{ab,cis}$ 7.3
$^{3}J_{ab,trans}$ 6.0

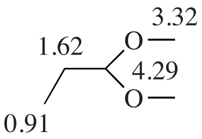


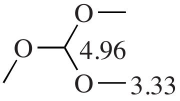


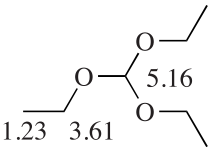


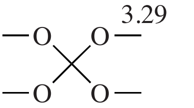
$^{2}J_{gem}$ 1.5

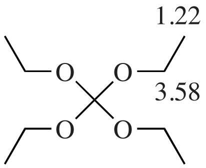












5.9 Nitrogen Compounds

5.9.1 Amines

Amine and Ammonium Protons (δ in ppm, |J| in Hz)

Chemical shifts of amine protons lie around 0.5–6 ppm depending on solvent, concentration, and hydrogen bonding. Those of ammonium protons are found between ca. 7 and 12 ppm. Neighboring H bond acceptors lead to deshielding in all cases.

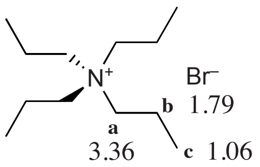
			in $CDCl_3$	in DMSO
Amines:	δ_{NH_2}, δ_{NH}	aliphatic	<1–2	2–4
		aromatic	3–4	4–7
Ammonium:	$\delta_{NH_3^+}$, $\delta_{NH_2^+}$, δ_{NH^+}	aliphatic	7–11	7–11
		aromatic	8–12	8–12

Coupling of amine protons with vicinal H atoms is usually not seen in aliphatic amines because of their rapid intermolecular exchange. However, for =C–NH–CH moieties (enamines, aromatic amines, amides, etc.), the exchange rate is slower and splitting (or line broadening at intermediate rates) is often observed. The H–C–N–H coupling depends on the conformation in a similar way as the H–C–C–H coupling (see Chapter 5.1.2). For N–CH_3 and N–CH_2 groups: $^3J_{HCNH} \approx 5–6$ Hz.

In acidic media (e.g., in trifluoroacetic acid as solvent), the exchange of the ammonium protons is slowed down to such an extent that the vicinal coupling H–N^+–C–H generally becomes observable. In other media, signals are usually broad owing to intermediate exchange rates.

The signals of amine and especially of ammonium protons are often broadened additionally because the ^{14}N–^1H coupling is only partly eliminated by the quadrupole relaxation of ^{14}N (spin quantum number, I = 1; natural abundance, 99.6%; $^1J_{NH} \approx 60$ Hz). This line broadening has no effect on the vicinal H–C–N–H coupling so that sharp multiplets can be observed for neighboring H atoms even when the NH proton exhibits a broad signal. In ammonium compounds of high symmetry, the quadrupole relaxation is slow and the coupling with ^{14}N leads to triplets of equal intensity for all three lines.

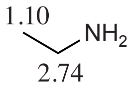
NH_4^+ $^1J_{NH}$ 52.8

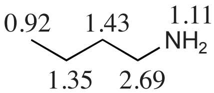


$^2J_{aN}$ 0.5
$^3J_{bN}$ 1.6
$^4J_{cN}$ 0.0

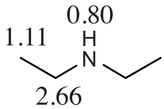
^{1}H Chemical Shifts and Coupling Constants of Amines and Ammonium Salts (δ in ppm, J in Hz)

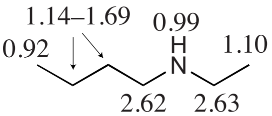
2.47
CH_3NH_2



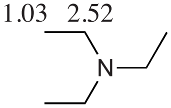


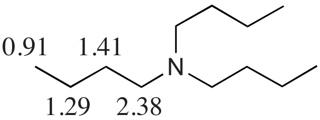
≈2.3
$(CH_3)_2NH$





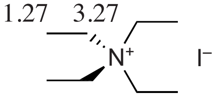
2.22
$(CH_3)_3N$
$^2J_{gem}$ -11.7



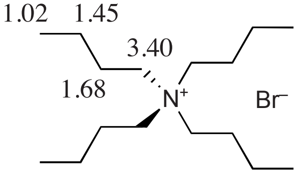


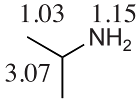
3.21 (in D_2O)
$(CH_3)_4N^+ I^-$

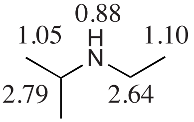
3.46 (in $CDCl_3$)
$(CH_3)_4N^+ AcO^-$

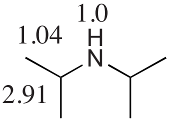


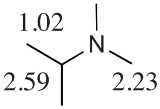
(in D_2O)

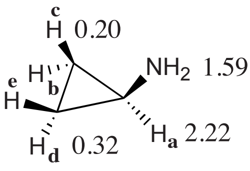






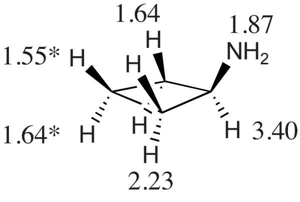




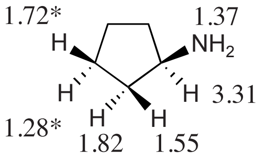


(neat)

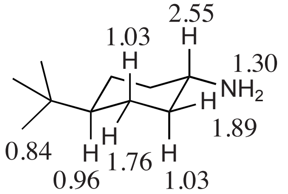
$^3J_{ab}$ 6.6
$^3J_{ac}$ 3.6
$^2J_{bc}$ -4.3
$^3J_{bd}$ 9.7
$^3J_{be}$ 6.2
$^3J_{ce}$ 9.9

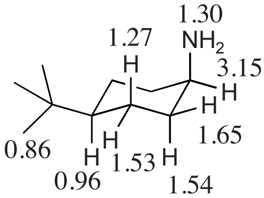


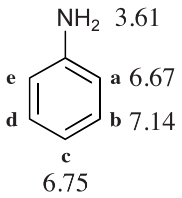
* assignment uncertain



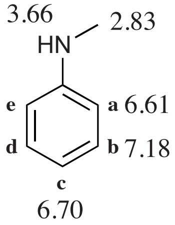
* assignment uncertain



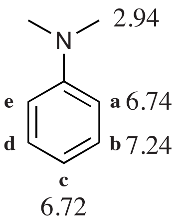




$^3J_{ab}$ 8.0
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.5
$^4J_{ae}$ 2.5
$^3J_{bc}$ 7.4
$^4J_{bd}$ 1.6

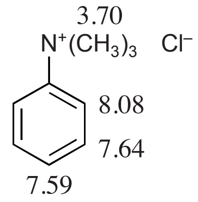


$^3J_{ab}$ 8.2
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.4
$^4J_{ae}$ 2.5
$^3J_{bc}$ 7.3
$^4J_{bd}$ 1.7

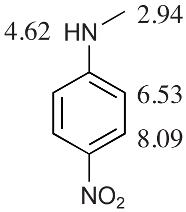


$^3J_{ab}$ 8.4
$^4J_{ac}$ 1.0
$^5J_{ad}$ 0.4
$^4J_{ae}$ 2.8
$^3J_{bc}$ 7.3
$^4J_{bd}$ 1.8

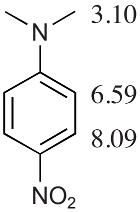
δ_{NH_2} in DMSO: 4.94 δ_{NH} in DMSO: 5.52



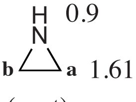
(in DMSO)



δ_{NH} in DMSO: 7.32

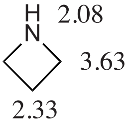


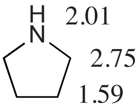
¹H Chemical Shifts and Coupling Constants of Cyclic Amines (δ in ppm, J in Hz)

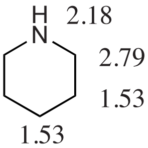


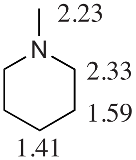
(neat)

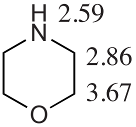
$^2J_{gem}$ ≈1
$^3J_{ab,cis}$ 6.3
$^3J_{ab,trans}$ 3.8

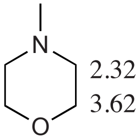


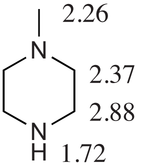








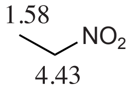




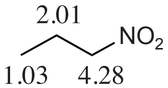
5.9.2 Nitro and Nitroso Compounds

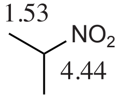
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

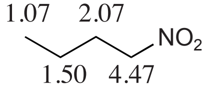
4.34
CH_3NO_2

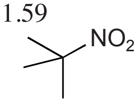


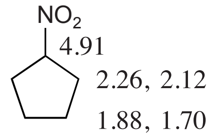
J_{vic} 7.4

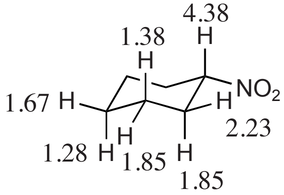


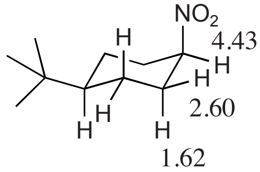


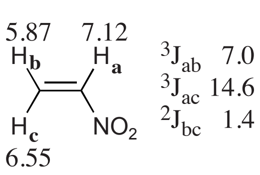


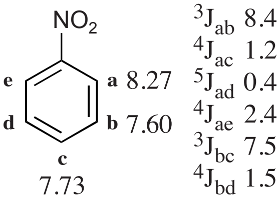


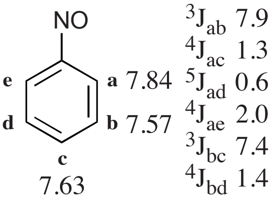






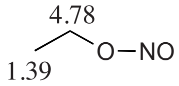


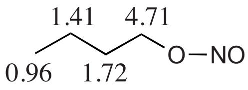


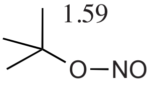


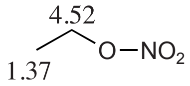
5.9.3 Nitrites and Nitrates

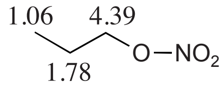
¹H Chemical Shifts (δ in ppm)

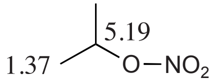












N

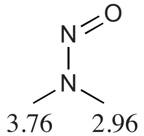
5.9.4 Nitrosamines, Azo and Azoxy Compounds

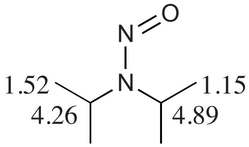
¹H Chemical Shifts (δ in ppm)

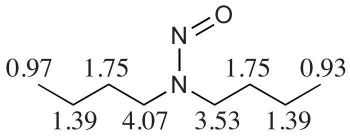
Owing to hindered rotation around the N–NO bond, corresponding protons in *cis* and *trans* positions have different chemical shifts in the neighborhood of the N=O group.

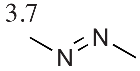
In general: $\delta_{cis} < \delta_{trans}$ for α-CH_3, α-CH_2, and β-CH_3

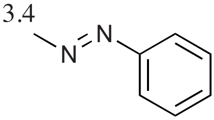
$\delta_{cis} > \delta_{trans}$ for α-CH

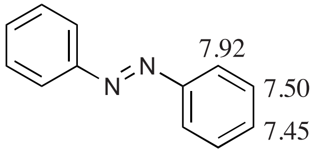


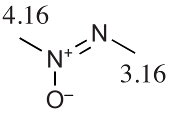


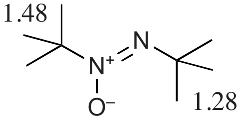


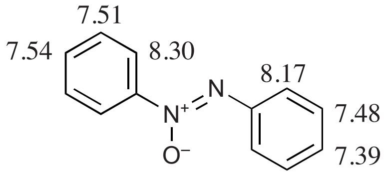






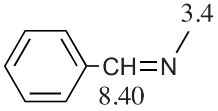


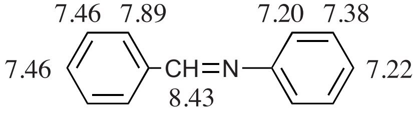




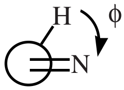
5.9.5 Imines, Oximes, Hydrazones, and Azines

¹H Chemical Shifts (δ in ppm)

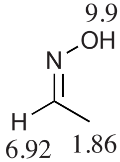


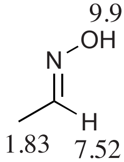


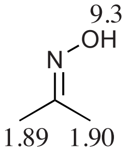
In aldoximes and ketoximes, the chemical shift difference between *syn* and *anti* protons at the α-CH group, $\Delta\delta = \delta_{syn} - \delta_{anti}$, depends on the dihedral angle, $\phi_{H\text{–}C\text{–}C=N}$:

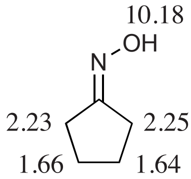


ϕ	$\Delta\delta = \delta_{syn} - \delta_{anti}$
0°	1
60°	0
115°	-0.3

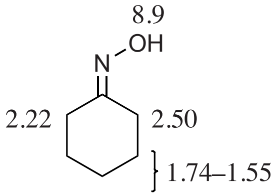


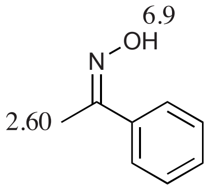


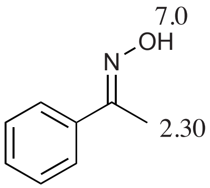


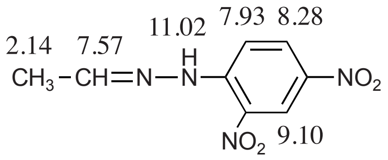


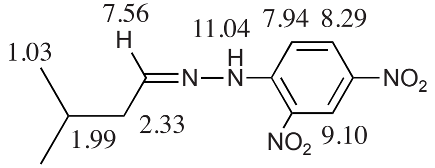
(in DMSO)

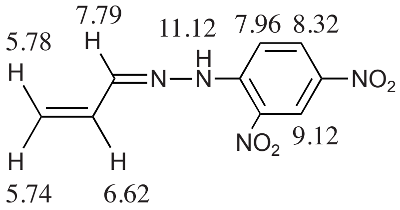


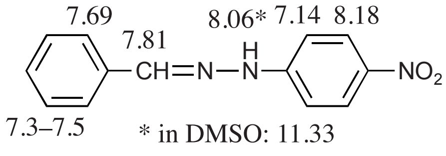


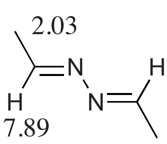


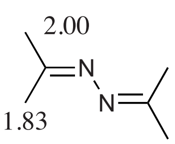


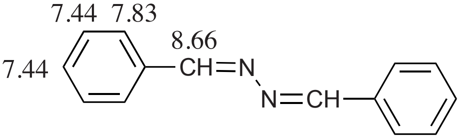








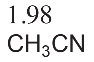


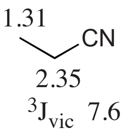


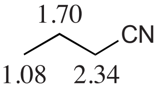
5.9.6 Nitriles and Isonitriles

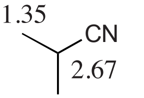
¹H Chemical Shifts and Coupling Constants of Nitriles (δ in ppm, J in Hz)

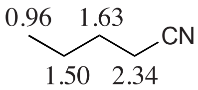
N

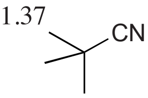


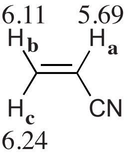




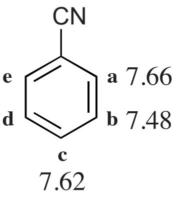




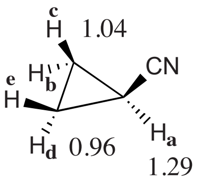




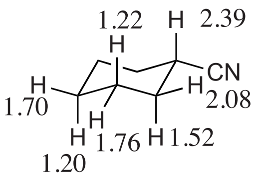
$^3J_{ab}$ 11.8
$^3J_{ac}$ 17.9
$^2J_{bc}$ 0.9

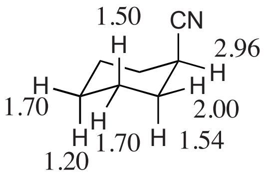


$^3J_{ab}$ 7.8
$^4J_{ac}$ 1.3
$^5J_{ad}$ 0.7
$^4J_{ae}$ 1.8
$^3J_{bc}$ 7.7
$^4J_{bd}$ 1.3



$^3J_{ab}$ 8.4
$^3J_{ac}$ 5.1
$^2J_{bc}$ -4.7
$^3J_{bd}$ 9.2
$^3J_{be}$ 7.1
$^3J_{ce}$ 9.5



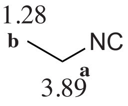


^1H Chemical Shifts and Coupling Constants of Isonitriles (δ in ppm, |J| in Hz)

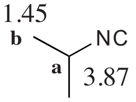
Because of the symmetrical electron distribution around the N atom, the quadrupole relaxation of the nitrogen nucleus is so slow that the ^{14}N-^1H coupling becomes observable and leads to triplets with relative intensities of 1:1:1 (spin quantum number of ^{14}N: I = 1; natural abundance, 99.6%):

$-\overset{b}{CH_2}-\overset{a}{CH_2}-{}^{14}NC$ $|^2J_{aN}|$ 1.8–2.8; $|^3J_{bN}|$ 1.5–3.5

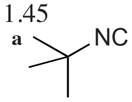
2.85 (a) CH_3NC $|^2J_{aN}|$ 2.3



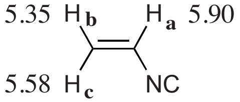
$|^2J_{aN}|$ 2.0
$|^3J_{bN}|$ 2.4
$|^3J_{ab}|$ 7.3



$|^2J_{aN}|$ 1.8
$|^3J_{bN}|$ 2.6
$|^3J_{ab}|$ 7.0



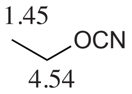
$|^3J_{aN}|$ 2.1



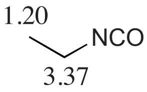
$|^2J_{aN}|$ 2.3 $^3J_{ab}$ 8.6
$|^3J_{bN}|$ 6.1 $^3J_{ac}$ 15.6
$|^3J_{cN}|$ 3.1 $^2J_{bc}$ -0.5

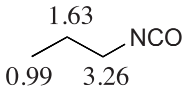
5.9.7 Cyanates, Isocyanates, Thiocyanates, and Isothiocyanates

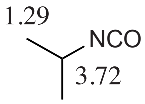
^1H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

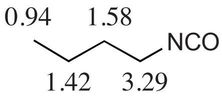


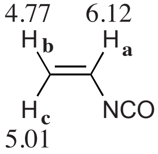
3.02
CH_3NCO





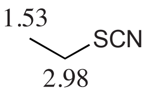


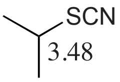




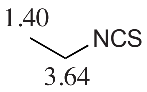
$^3J_{ab}$ 7.6
$^3J_{ac}$ 15.2
$^2J_{bc}$ -0.1

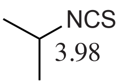
2.60
CH_3SCN

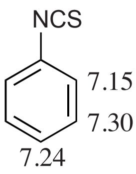




3.37
CH_3NCS







N

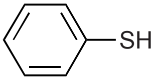
5.10 Sulfur Compounds

5.10.1 Thiols

¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

Typical ranges of SH chemical shifts:

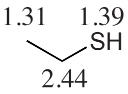
alk–SH 1–2

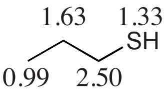


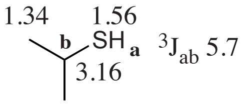
2–4

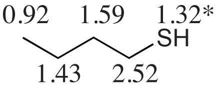
The exchange with other SH, OH, NH, or COOH protons is generally so slow that the chemical shift is characteristic and the vicinal coupling with SH protons becomes visible (5–9 Hz in aliphatic systems with fast rotation).

2.00 1.26
CH_3SH J_{vic} 7.4
(in benzene)

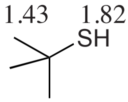


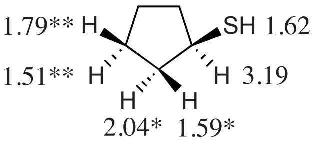




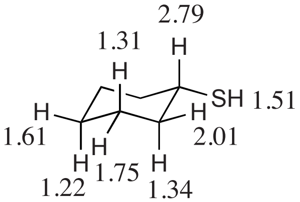


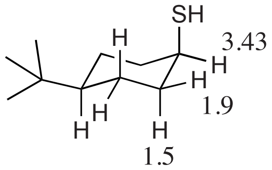
* in DMSO: 2.17

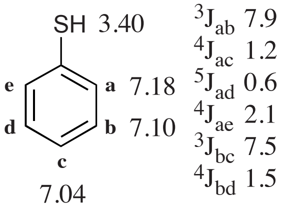




*; **: assignments uncertain



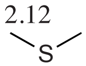


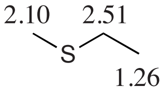


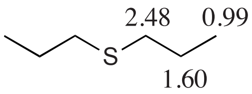
S

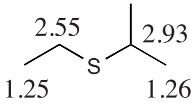
5.10.2 Sulfides

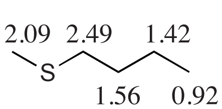
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

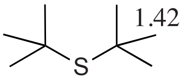


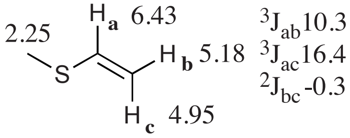


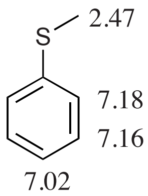


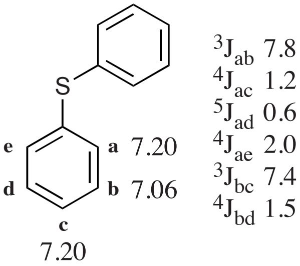




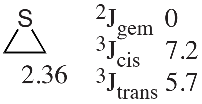


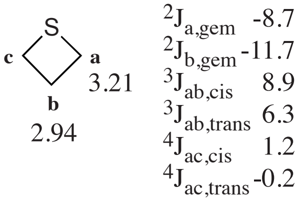


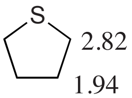


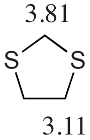


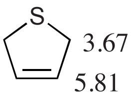
¹H Chemical Shifts and Coupling Constants of Cyclic Sulfides (δ in ppm, J in Hz)

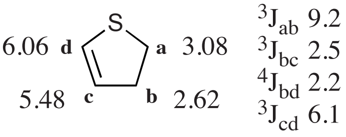


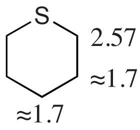


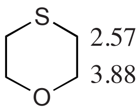


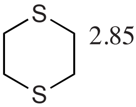


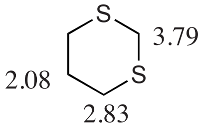


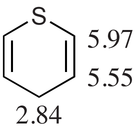


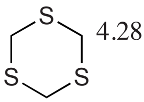










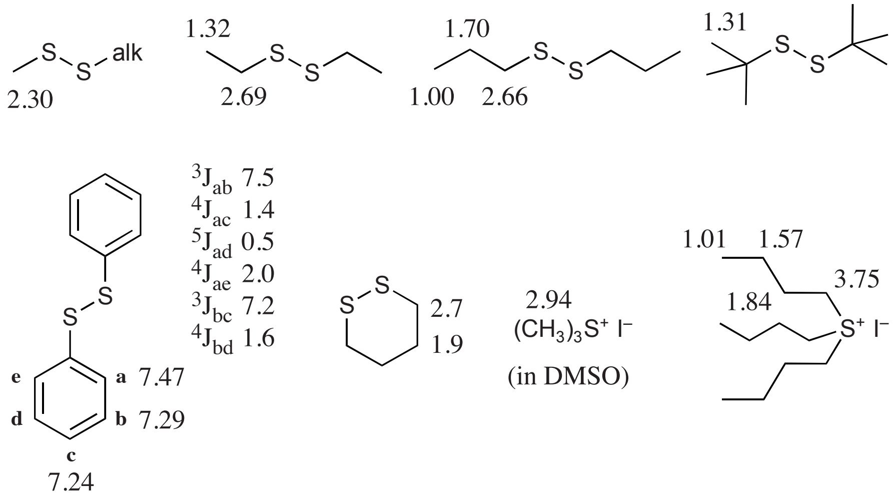


(in DMSO)

S

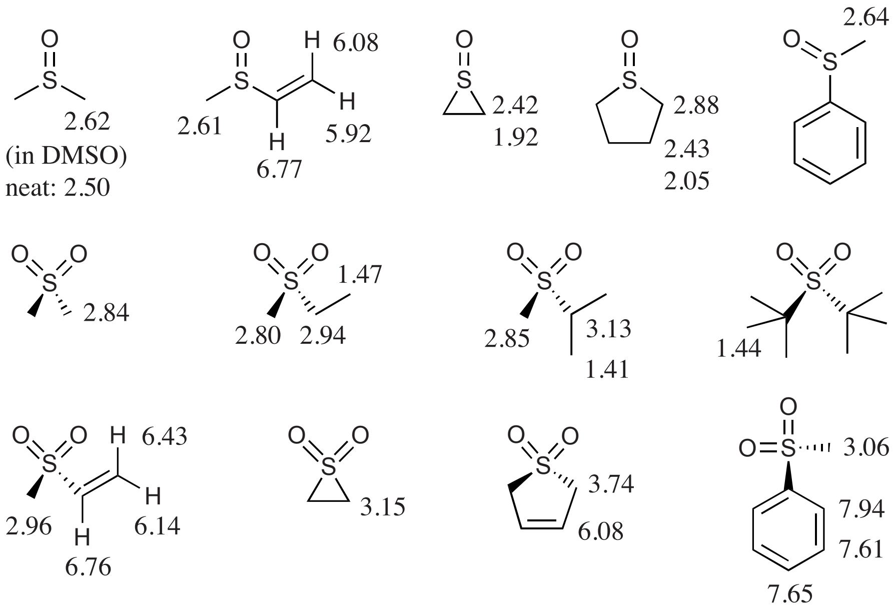
5.10.3 Disulfides and Sulfonium Salts

^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)



5.10.4 Sulfoxides and Sulfones

^{1}H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)



5.10.5 Sulfonic, Sulfurous, and Sulfuric Acids and Derivatives

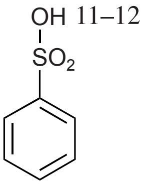
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

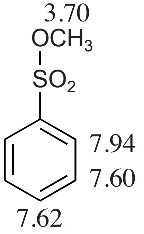
1.42 3.17 11.1
$CH_3-CH_2-SO_2-OH$

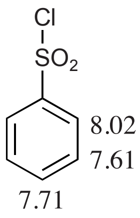
Apparently lower δ_{OH} values in DMSO due to fast exchange with H_2O

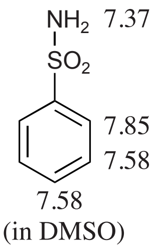
3.02 3.91
$CH_3-SO_2-O-CH_3$

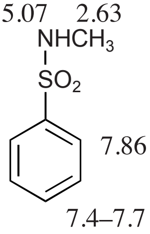
2.78 2.86
$CH_3-SO_2-N(CH_3)_2$

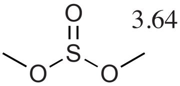


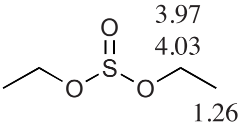




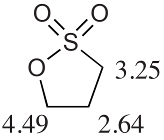


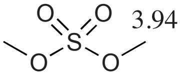


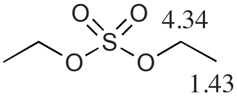


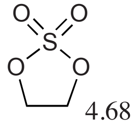


$^2J_{gem}$ -10.3
$^3J_{vic}$ 6.7
$^3J'_{vic}$ 6.9





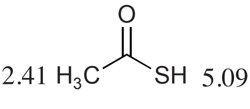


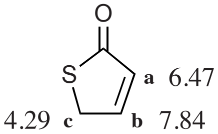


S

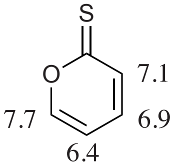
5.10.6 Thiocarboxylate Derivatives

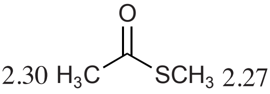
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

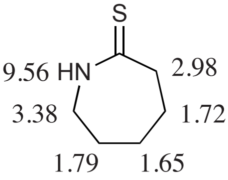




$^3J_{ab}$ 5.9
$^4J_{ac}$ 2.0
$^3J_{bc}$ 2.8





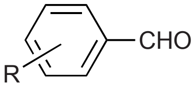


5.11 Carbonyl Compounds

5.11.1 Aldehydes

¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

alk—CHO	9–10	$^3J_{vic}$ 0–3
alken—CHO	9–10	$^3J_{vic}$ ≈8

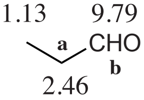


ortho-substituted: 10–10.5

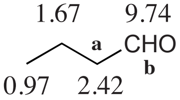
meta-, *para*-substituted: 9.5–10.2

9.60
$CH_2{=}O$
(in TMS)
$|^2J_{gem}|$ 42.4

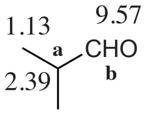
2.20 9.79
$H_3C{-}CHO$
a **b**
$^3J_{ab}$ 3.0



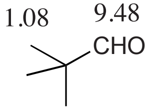
$^3J_{ab}$ 1.4

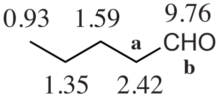


$^3J_{ab}$ 2.0

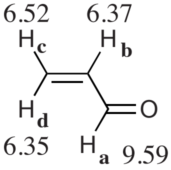


$^3J_{ab}$ 1.1

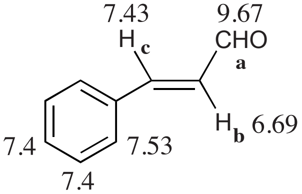




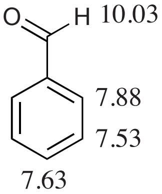
$^3J_{ab}$ 1.9

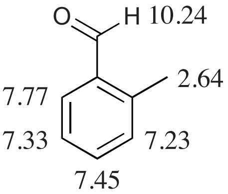


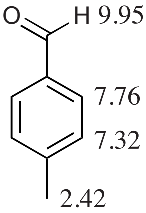
$^3J_{ab}$ 4.7	$^3J_{bc}$ 10.0
$^4J_{ac}$ <1	$^4J_{bd}$ 17.4
$^4J_{ad}$ <1	$^2J_{cd}$ 1.0

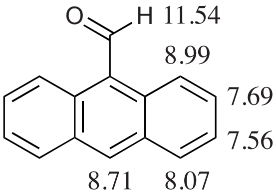


$^3J_{ab}$ 7.8
$^3J_{bc}$ 16.0





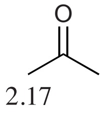


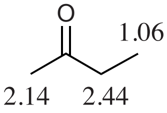


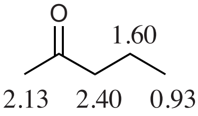
C = X

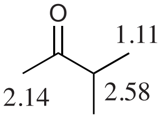
5.11.2 Ketones

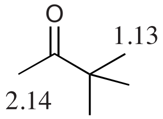
¹H Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

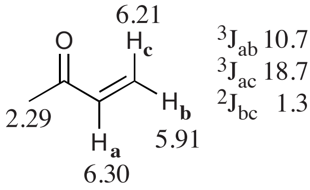


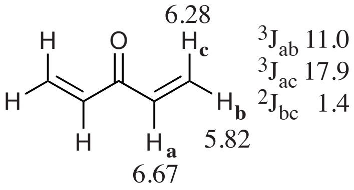


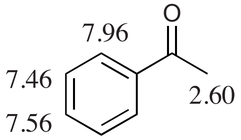


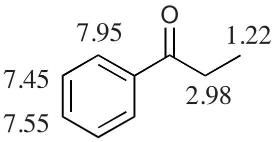


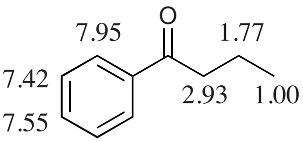


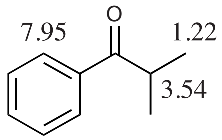


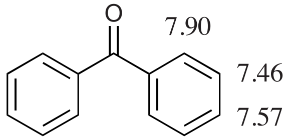


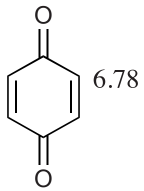


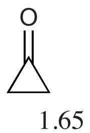


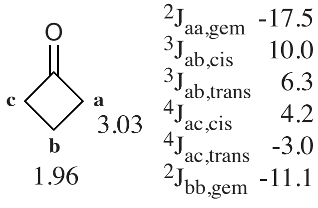


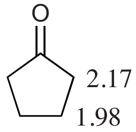


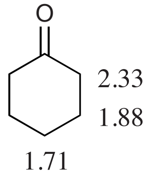


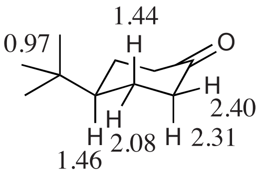


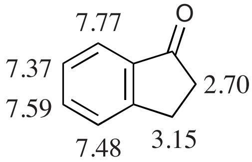


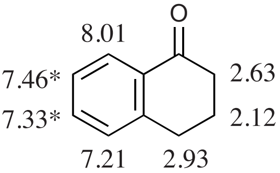






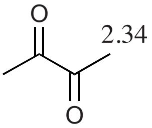


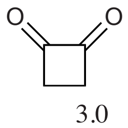


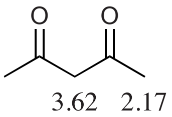


* assignment uncertain

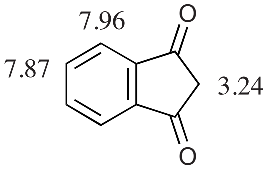
^{1}H Chemical Shifts of Diketones (δ in ppm)





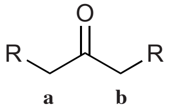


For the enol form, see Chapter 5.8.1



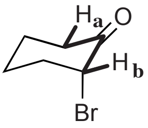
Long-Range Coupling in Ketones (|J| in Hz)

For fixed conformations, the coupling over the C=O group is often detectable for W-arrangement of the coupling path.



|$^4J_{ab}$| 0–0.5

No W-arrangement



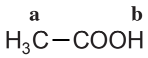
|$^4J_{ab}$| 1.0

W-arrangement

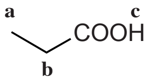
5.11.3 Carboxylic Acids and Carboxylates

^{1}H Chemical Shifts (δ in ppm)

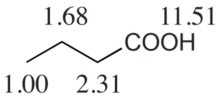
	$CDCl_3$	DMSO	D_2O
a	8.05	8.13	8.26
b	10.85	12.50	

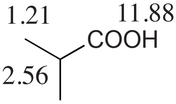


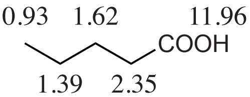
	$CDCl_3$	DMSO
a	2.10	1.91
b	11.51	11.91

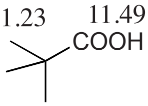


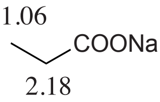
	$CDCl_3$	DMSO
a	1.16	1.00
b	2.39	2.21
c	10.35	11.90



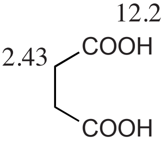




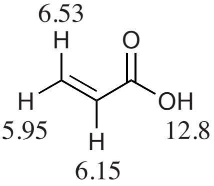


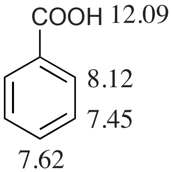


(in D_2O)



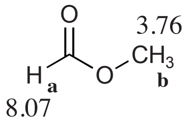
(in DMSO)



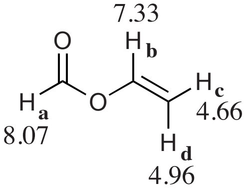


5.11.4 Esters and Lactones

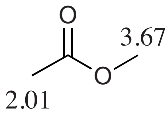
^{1}H Chemical Shifts and Coupling Constants of Carboxylic Acid Esters (δ in ppm, J in Hz)

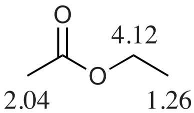


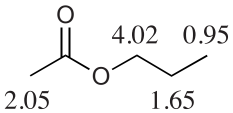
$|^4J_{ab}|$ 0.8

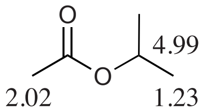


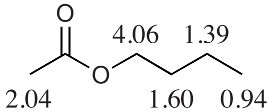
$^4J_{ab}$ -0.7 $^3J_{bc}$ 6.4
$^5J_{ac}$ 1.6 $^3J_{bd}$ 13.9
$^5J_{ad}$ 0.8 $^2J_{cd}$ -1.7

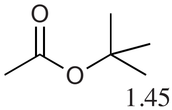


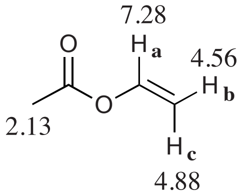




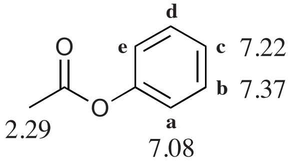




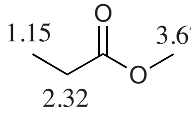


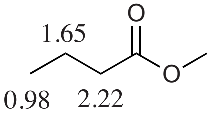


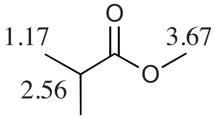
$^3J_{ab}$ 6.3
$^3J_{ac}$ 14.1
$^2J_{bc}$ -1.6

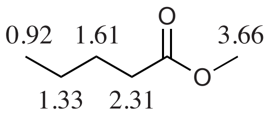


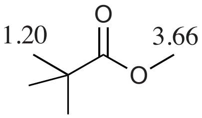
$^3J_{ab}$ 8.2
$^4J_{ac}$ 1.1
$^5J_{ad}$ 0.5
$^4J_{ae}$ 2.5
$^3J_{bc}$ 7.5
$^4J_{bd}$ 1.7

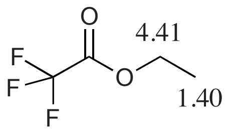


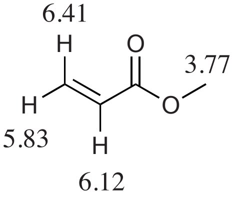


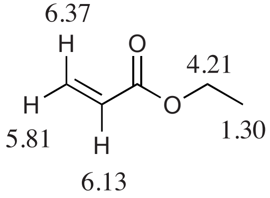


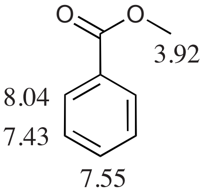


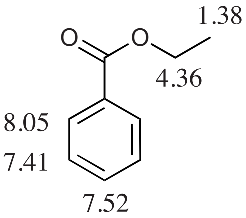


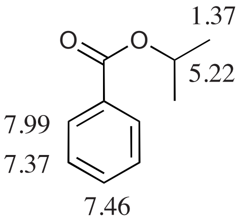


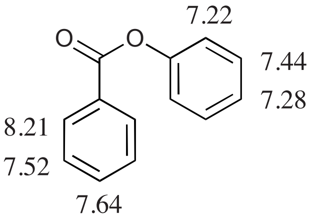


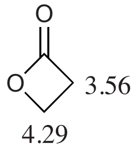


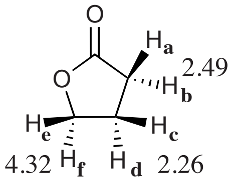




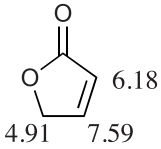


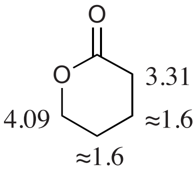


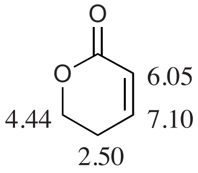


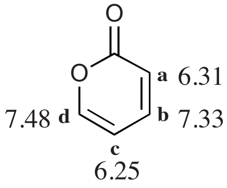


$^2J_{ab}$	-17.5	$^2J_{cd}$	-12.7
$^3J_{ac}$	9.5	$^3J_{ce}$	7.9
$^3J_{ad}$	6.9	$^3J_{cf}$	6.3
$^4J_{ae}$	0.3	$^2J_{ef}$	-8.8
$^4J_{af}$	-0.5		







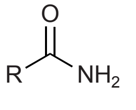


$^3J_{ab}$	9.4
$^4J_{ac}$	1.5
$^5J_{ad}$	1.3
$^3J_{bc}$	6.3
$^4J_{bd}$	2.4
$^3J_{cd}$	5.0

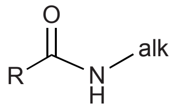
C = X

5.11.5 Amides and Lactams

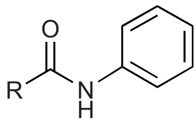
Amide Protons (δ in ppm, J in Hz)



5–7
R: alk or ar



6–8.5
R: alk or ar



7.5–10
R: alk or ar

Higher values in DMSO or with H bond acceptors in the neighborhood.

The signals of the NH protons are often broad because the ^{14}N–^{1}H coupling is only partly eliminated by the quadrupole relaxation of ^{14}N (spin quantum number, I = 1; $^{1}J_{NH} \approx 60$). In primary amides, the hindered rotation around the CO–N bond is another reason for line broadening. At slow rotation, the chemical shifts of the two primary amide protons differ by about 0.4–1 ppm. Therefore, at intermediate rotation rates, line widths of up to 1 ppm may be observed.

Due to the slow intermolecular exchange of amide protons, their coupling to neighboring hydrogen atoms is usually detectable. The splitting of the C–H signal is clearly observed even in those cases where the signal of the NH proton is broad and featureless. The H–N–C–H coupling depends on the conformation in a similar way as the H–C–C–H coupling (see Chapter 5.1.2). For N–CH_3 and N–CH_2 groups: $^{3}J_{HNCH} \approx 7$ Hz.

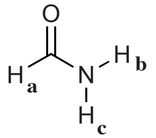
Tertiary Alkylamides

The rotation around the CO–N bond is usually so slow that, for identical substituents, two separate signals are observed for *cis* and *trans* positions. With different N-substituents, two separate pairs of signals are observed for the two conformers. In general, the following relationships hold:

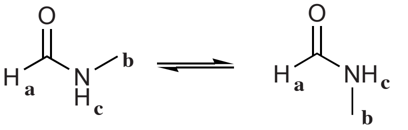
for N**CH₃**, NCH_2**CH₃**, and NCH(**CH₃**)$_2$	$\delta_{cis\ to\ O} \leq \delta_{trans\ to\ O}$
for N**CH**$(CH_3)_2$ and NC(**CH₃**)$_3$	$\delta_{trans\ to\ O} \leq \delta_{cis\ to\ O}$
for N**CH₂**	$\delta_{cis\ to\ O} \approx \delta_{trans\ to\ O}$

Formamides (δ in ppm, J in Hz)

In the more stable conformer of monosubstituted formamides, the substituent occupies the *cis* position relative to the carbonyl oxygen. In the more stable conformer of asymmetrically disubstituted formamides, the larger substituent occupies the *trans* position relative to the carbonyl oxygen.



	$CDCl_3$	DMSO
a	8.23	7.98
b	5.80	7.14
c	5.48	7.41

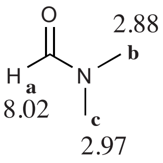


≈ 90% in $CDCl_3$

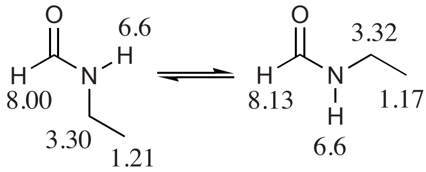
	$CDCl_3$	DMSO
a	8.19	8.01
b	2.86	2.59
c	5.55	7.90

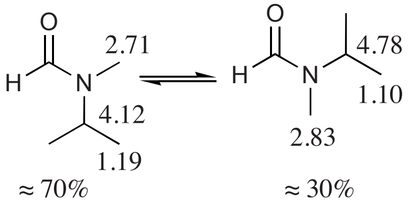
≈ 10% in $CDCl_3$

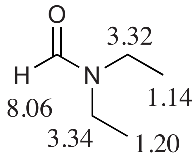
	$CDCl_3$	DMSO
a	8.06	7.81
b	2.94	2.72
c	5.86	7.90

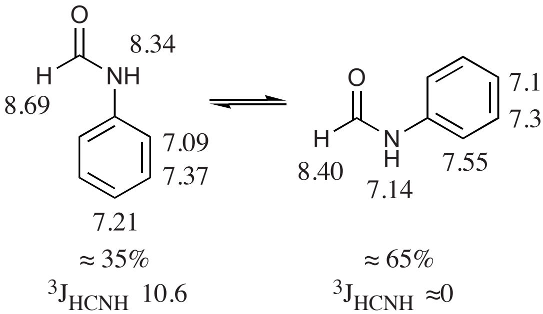


$^{4}J_{ab} \approx 0.3$
$^{4}J_{ac} \approx 0.7$



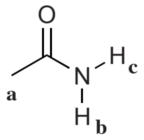




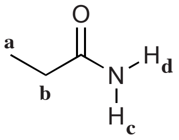


Amides of Aliphatic Carboxylic Acids (δ in ppm, J in Hz)

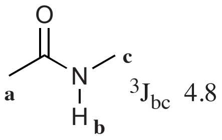
In *monosubstituted* acetamides, the substituent of the only observable conformer is *cis* to the carbonyl oxygen. In *disubstituted* acetamides, the more stable conformer has the larger substituent *cis* to the carbonyl oxygen.



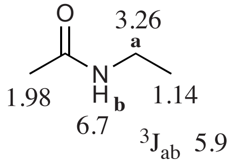
	$CDCl_3$	DMSO
a	2.03	1.76
b	5.42	7.30
c	5.42	6.70

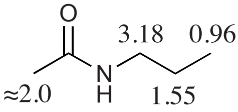


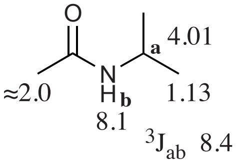
	$CDCl_3$	DMSO
a	1.17	0.97
b	2.26	2.04
c	5.38	7.16
d	6.14	6.62

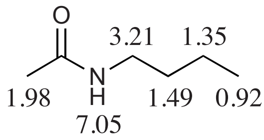


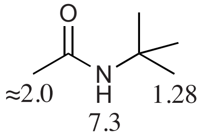
	$CDCl_3$	DMSO
a	1.98	1.78
b	5.53	7.70
c	2.80	2.50

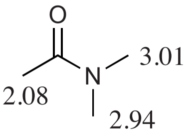


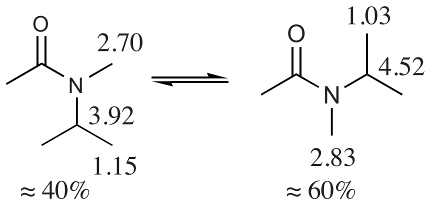


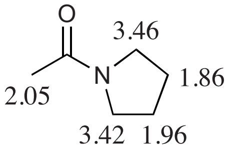


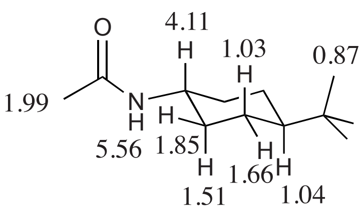


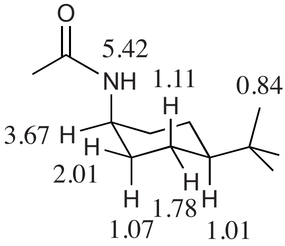


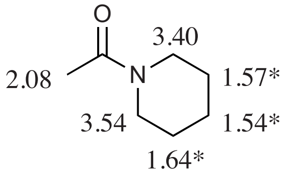




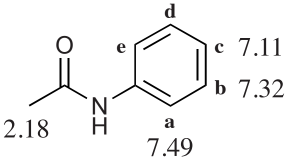






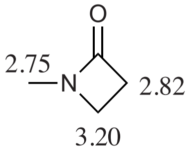


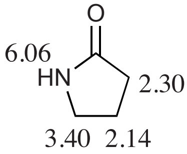
* assignment uncertain



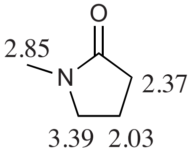
$^3J_{ab}$ 8.2
$^4J_{ac}$ 1.2
$^5J_{ad}$ 0.5
$^4J_{ae}$ 2.4
$^3J_{bc}$ 7.4
$^4J_{bd}$ 1.5

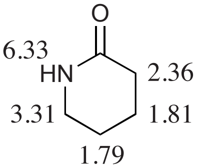
Lactams (δ in ppm, J in Hz)



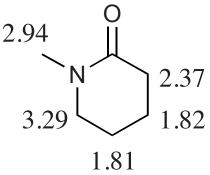


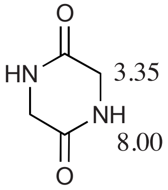
$^3J_{HNCH_2}$ 0.8



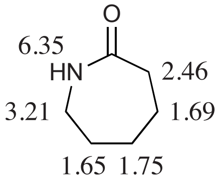


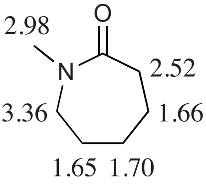
C = X





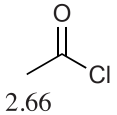
$^3J_{HNCH_2}$ 2.2

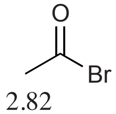


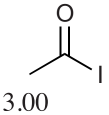


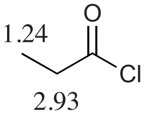
5.11.6 Miscellaneous Carbonyl Derivatives

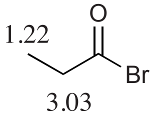
Carboxylic Acid Halides (δ in ppm, J in Hz)

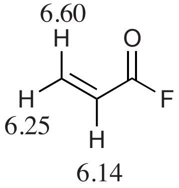


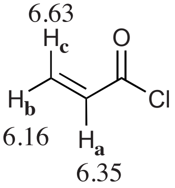




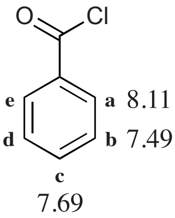






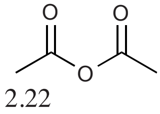


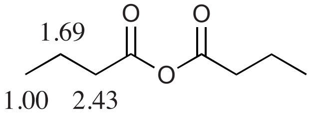
$^3J_{ab}$ 10.6
$^3J_{ac}$ 17.4
$^2J_{bc}$ 0.2

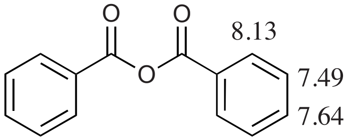


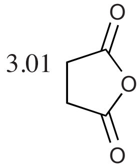
$^3J_{ab}$ 8.0
$^4J_{ac}$ 1.2
$^5J_{ad}$ 0.6
$^4J_{ae}$ 2.0
$^3J_{bc}$ 7.5
$^4J_{bd}$ 1.4

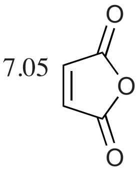
Carboxylic Acid Anhydrides (δ in ppm)

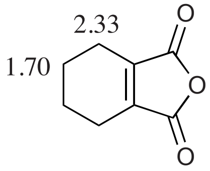




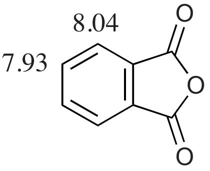




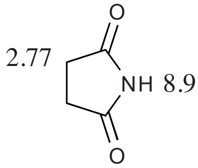


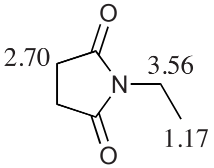


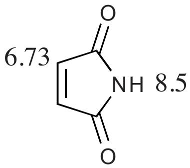
(in DMSO)

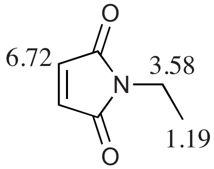


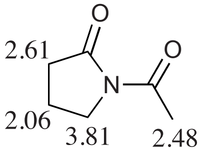
Carboxylic Acid Imides (δ in ppm, J in Hz)

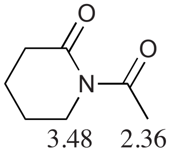


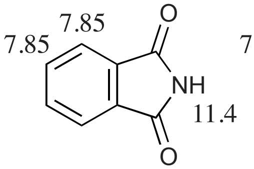


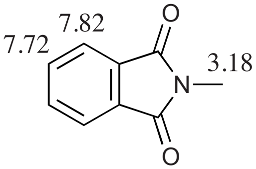




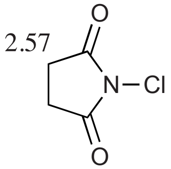




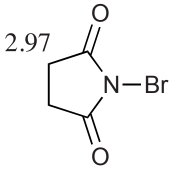


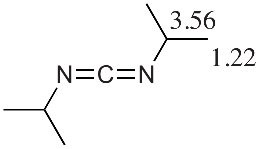


(in DMSO)

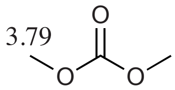


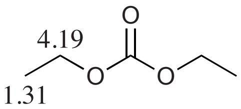
(in DMSO)

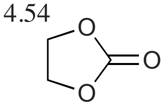


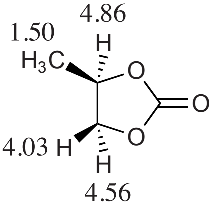


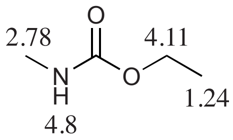
Carbonic Acid Derivatives (δ in ppm, J in Hz)

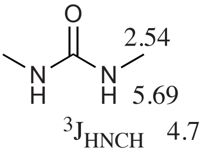




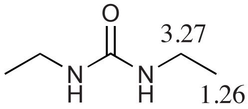


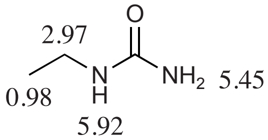




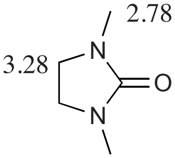


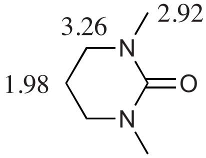
(in DMSO)

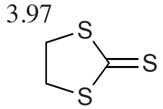


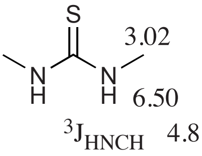


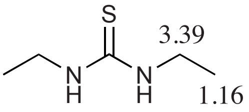
(in DMSO)











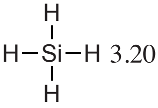
(in D_2O)

5.12 Miscellaneous Compounds

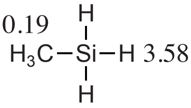
5.12.1 Compounds with Group IV Elements

Silicon Compounds (δ in ppm, J in Hz)

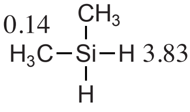
Coupling with silicon: The isotope ^{29}Si (natural abundance, 4.7 %) has a spin quantum number I of 1/2. Doublets with the corresponding intensity ("Si satellites") are usually observed. Typical coupling constants: $^{1}J_{HSi}$ -150 to -380 Hz
$^{2}J_{HCSi}$ 5 to 10 Hz



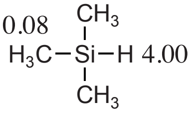
$^{1}J_{HSi}$ -202.5



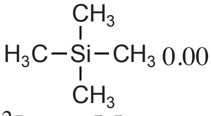
$^{1}J_{HSi}$ -202.5
$^{3}J_{HSiCH}$ 4.7



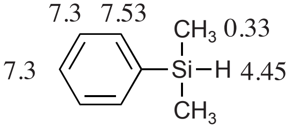
$^{1}J_{HSi}$ -202.5
$^{2}J_{HCSi}$ 7.7
$^{3}J_{HSiCH}$ 4.3



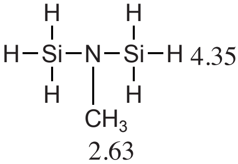
$^{1}J_{HSi}$ -190
$^{2}J_{HCSi}$ 7.5
$^{3}J_{HSiCH}$ 3.8



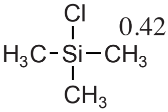
$^{2}J_{HCSi}$ 5.5

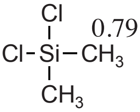


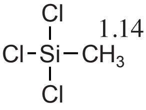
$^{1}J_{HSi}$ -189
$^{2}J_{HCSi}$ 7.6
$^{3}J_{HSiCH}$ 3.6

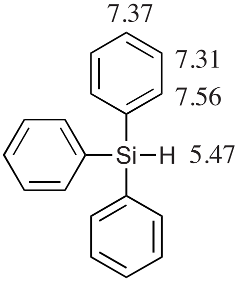


$^{1}J_{HSi}$ -209.4

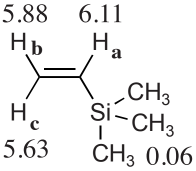




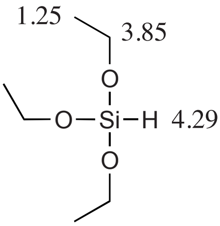


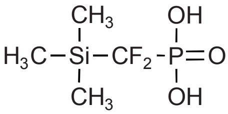


$^{1}J_{HSi}$ -199.1



$^{3}J_{ab}$ 14.6
$^{3}J_{ac}$ 20.2
$^{2}J_{bc}$ 3.8



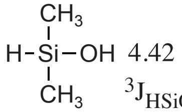


pH-dependent shift of the methyl H atoms relative to DSS. Between pH 4.3 and 8.2 [1]:

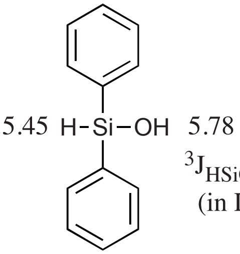
$$pH = 6.246 - \log\left[\frac{0.224 - \delta}{\delta - 0.193}\right]$$

P Si

The silanol hydrogen is exchangeable with D_2O. Slow intermolecular exchange is observed in DMSO as solvent so that the vicinal coupling in H–Si–O–H is detectable ($^3J_{HSiOH} \approx$ 2–7 Hz).

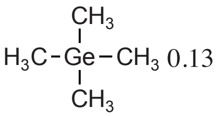


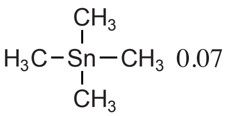
$^3J_{HSiOH}$ 1.8 (in DMSO)



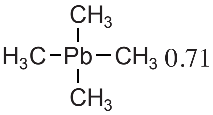
$^3J_{HSiOH}$ 2.0 (in DMSO)

Germanium, Tin, and Lead Compounds (δ in ppm, J in Hz)





$^2J_{HC^{117}Sn}$ 51.9 (7.6 %)
$^2J_{HC^{119}Sn}$ 54.3 (8.6 %)



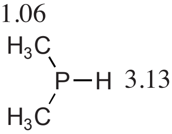
5.12.2 Phosphorus Compounds

^{31}P (natural abundance, 100%) has a spin quantum number I of 1/2. Couplings to protons through up to 5 bonds are usually observed.

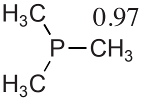
Phosphines and Phosphonium Compounds (δ in ppm, J in Hz)

PH_3 1.79
$^1J_{HP}$ 184.9

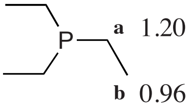
0.98 2.63
$H_3C{-}PH_2$
$^1J_{HP}$ 186.4
$^2J_{HCP}$ 4.1
$^3J_{HCPH}$ 8.2



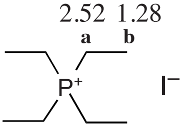
$^1J_{HP}$ 191.6
$^2J_{HCP}$ 3.6
$^3J_{HCPH}$ 7.7



$^2J_{HCP}$ 2.1

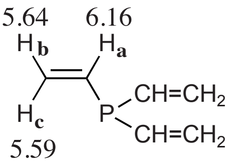


$^2J_{aP}$ 13.7
$^3J_{bP}$ 0.5
$^3J_{ab}$ 7.6

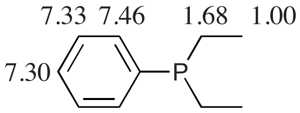


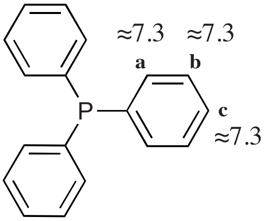
$^2J_{aP}$ 12.8
$^3J_{bP}$ 18.0
$^3J_{ab}$ 7.6

P Si

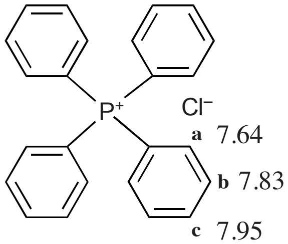


$^{2}J_{aP}$ 11.7
$^{3}J_{bP}$ 30.2
$^{3}J_{cP}$ 13.6
$^{3}J_{ab}$ 11.8
$^{3}J_{ac}$ 18.4
$^{2}J_{bc}$ 2.0



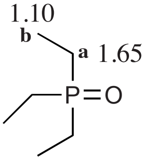


$^{3}J_{aP}$ 7.5
$^{4}J_{bP}$ 1.4
$^{5}J_{cP}$ 0.7

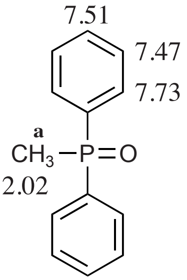


$^{3}J_{aP}$ 13.1
$^{4}J_{bP}$ 3.7
$^{5}J_{cP}$ 2.1

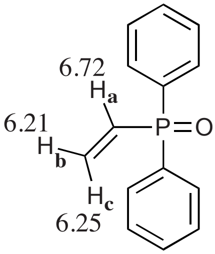
Phosphine Oxides and Sulfides (δ in ppm, J in Hz)



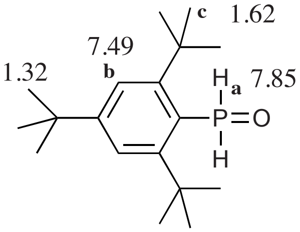
$^{3}J_{aP}$ 11.9
$^{4}J_{bP}$ 16.3
$^{5}J_{ab}$ 7.8



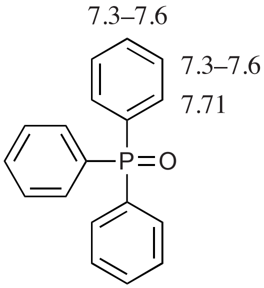
$^{3}J_{aP}$ 13.2

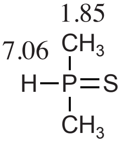


$^{2}J_{aP}$ 25.9 $^{3}J_{ab}$ 12.9
$^{3}J_{bP}$ 41.8 $^{3}J_{ac}$ 18.9
$^{3}J_{cP}$ 22.3 $^{2}J_{bc}$ 1.8

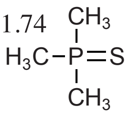


$^{1}J_{aP}$ 481.7
$^{4}J_{bP}$ 0
$^{5}J_{cP}$ 0

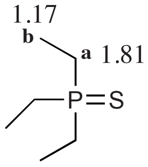




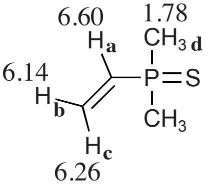
$^{1}J_{HP}$ 14.3
$^{2}J_{HCP}$ 4.4



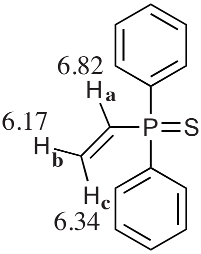
$^{2}J_{HCP}$ 13.0



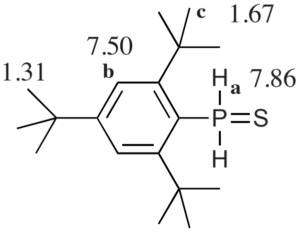
$^{2}J_{aP}$ 11.3
$^{3}J_{bP}$ 18.1
$^{3}J_{ab}$ 7.5



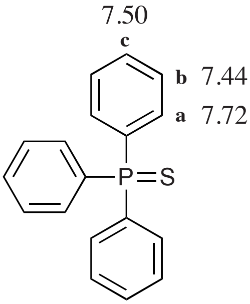
$^{2}J_{aP}$ 25.9
$^{3}J_{bP}$ 45.3
$^{3}J_{cP}$ 25.4
$^{2}J_{dP}$ 13.5
$^{3}J_{ab}$ 11.8
$^{3}J_{ac}$ 17.9
$^{2}J_{bc}$ 1.8



$^{2}J_{aP}$ 24.9 $^{3}J_{ab}$ 11.7
$^{3}J_{bP}$ 47.0 $^{3}J_{ac}$ 17.9
$^{3}J_{cP}$ 25.5 $^{2}J_{bc}$ 1.6

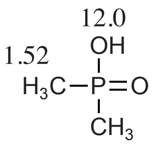


$^{1}J_{aP}$ 466.2
$^{4}J_{bP}$ 4.8
$^{5}J_{cP}$ 1.1

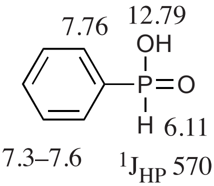


$^{3}J_{aP}$ 13.3
$^{4}J_{bP}$ 3.1
$^{5}J_{cP}$ 2.1

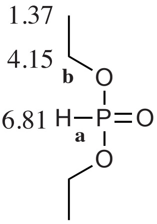
Phosphinic and Phosphonic Acid Derivatives (δ in ppm, J in Hz)



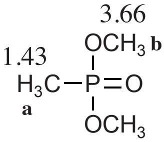
$^{2}J_{HP}$ 14.4



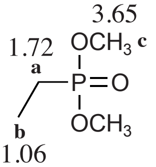
$^{1}J_{HP}$ 570.4



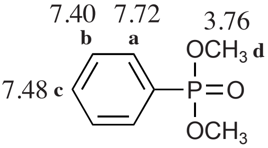
$^{1}J_{aP}$ 691.7
$^{3}J_{bP}$ 7.1



$^{2}J_{aP}$ 17.3
$^{3}J_{bP}$ 11.0

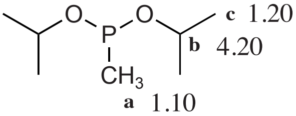


$^{2}J_{aP}$ -18.0
$^{3}J_{bP}$ 19.5
$^{3}J_{cP}$ 10.0
$^{3}J_{ab}$ 7.5

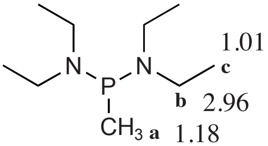


$^{3}J_{aP}$ 13.3
$^{4}J_{bP}$ 4.1
$^{5}J_{cP}$ 1.2
$^{3}J_{dP}$ 11.1

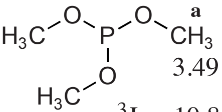
Phosphonous and Phosphorous Acid Derivatives (δ in ppm, J in Hz)



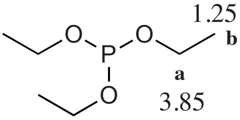
$^2J_{aP}$ 9.7
$^3J_{bP}$ 9.5
$^4J_{cP}$ 6.0



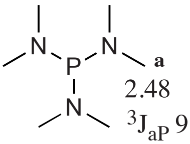
$^2J_{aP}$ 8.7
$^3J_{bP}$ 8.0
$^4J_{cP}$ ≈1.0
$^3J_{bc}$ 7.5



$^3J_{aP}$ 10.8

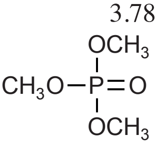


$^3J_{aP}$ 8.0
$^4J_{bP}$ 0.6
$^3J_{ab}$ 7.1

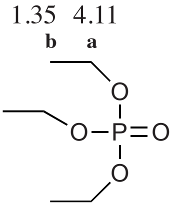


$^3J_{aP}$ 9.1

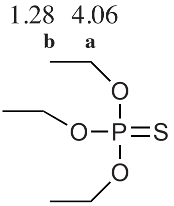
Phosphoric Acid Derivatives (δ in ppm, J in Hz)



$^3J_{HP}$ 11.0

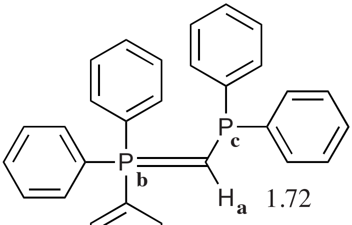


$^3J_{aP}$ 8.0
$^4J_{bP}$ 0.9
$^3J_{ab}$ 7.1

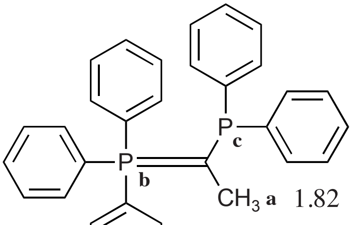


$^3J_{aP}$ 10.0
$^4J_{bP}$ 0.7
$^3J_{ab}$ 7

Phosphorus Ylids (δ in ppm, J in Hz)



$^2J_{ab}$ 12.7
$^2J_{ac}$ -1.2

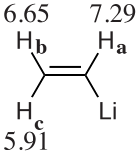


$^3J_{ab}$ 15.9
$^3J_{ac}$ 3.6

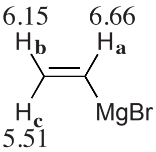
5.12.3 Miscellaneous Compounds

Organometallic Compounds (δ in ppm, J in Hz)

$Li-CH_3$ -1.32 (in benzene)
-1.74 (in ether)

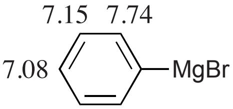


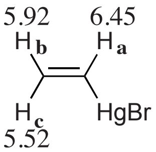
$^3J_{ab}$ 19.3
$^3J_{ac}$ 23.9
$^2J_{bc}$ 7.1



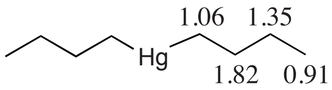
$^3J_{ab}$ 17.7
$^3J_{ac}$ 23.3
$^2J_{bc}$ 7.6

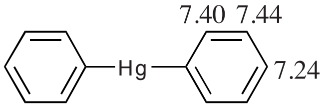
(in THF)





$^3J_{ab}$ 11.9
$^3J_{ac}$ 18.7
$^2J_{bc}$ 3.1

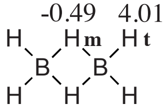




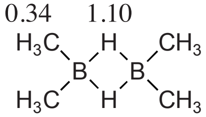
Boron Compounds (δ in ppm, J in Hz)

1.50 2.26
$BH_3 \cdot CH_3SCH_3$

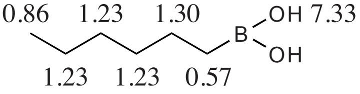
$^1J_{H^{11}B}$ 104
(^{11}B: 80.4%, I = 3/2)



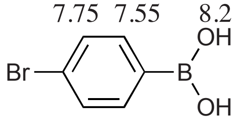
$^2J_{mt}$ 7.5
$^1J_{m^{11}B}$ 46.2
$^1J_{t^{11}B}$ 133.0



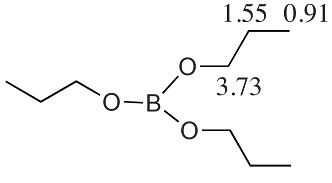
$^3J_{HH}$ 3.0

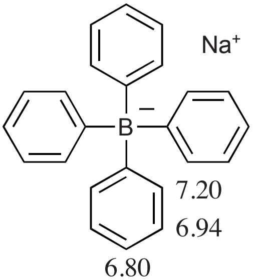


(in DMSO)



(in DMSO)





(in DMSO)

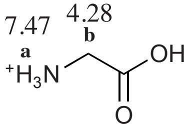
5.12.4 References

[1] M.D. Reily, L.C. Robosky, M.L. Manning, A. Butler, J.D. Baker, R.T. Winters, DFTMP, an NMR reagent for assessing the near-neutral pH of biological samples, *J. Am. Chem. Soc.* **2006**, *128*, 12360.

5.13 Natural Products

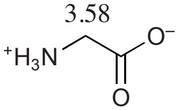
5.13.1 Amino Acids

Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

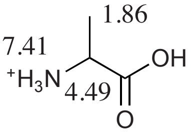


$^{3}J_{ab}$ 5.7

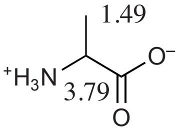
(in TFA)



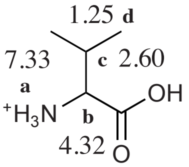
(in D_2O)



(in TFA)

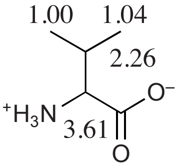


(in D_2O)

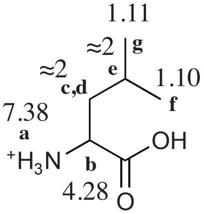


$^{3}J_{ab}$ 5.7
$^{3}J_{bc}$ 4.2
$^{3}J_{cd}$ 6.9

(in TFA)

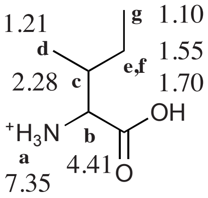


(in D_2O)



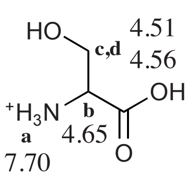
$^{3}J_{ab}$ 5.5
$^{3}J_{bc}$ ≈6.7
$^{3}J_{bd}$ ≈6.7
$^{3}J_{eg}$ 6.1
$^{3}J_{ef}$ 5.7

(in TFA)



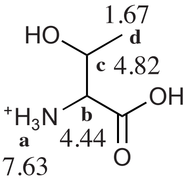
$^{3}J_{ab}$	5.5	$^{3}J_{cf}$	8.4
$^{3}J_{bc}$	3.6	$^{2}J_{ef}$	-13.6
$^{3}J_{cd}$	7.0	$^{3}J_{eg}$	7.0
$^{3}J_{ce}$	6.1	$^{3}J_{fg}$	7.0

(in TFA)



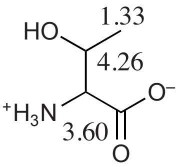
$^{3}J_{ab}$ ≈6
$^{3}J_{bc}$ 4.0*
$^{3}J_{bd}$ 4.0*
$^{2}J_{cd}$ -13.5

(in TFA)
* average value



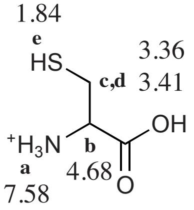
$^{3}J_{ab}$ 5.5
$^{3}J_{bc}$ 4.5
$^{3}J_{cd}$ 6.5

(in TFA)



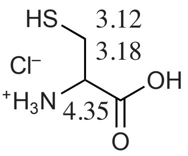
(in D_2O)

Natural Products

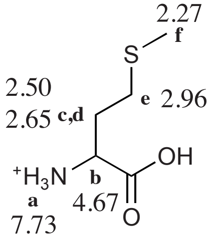


$^3J_{ab}$ 5.3 $^2J_{cd}$ -15.5
$^3J_{bc}$ 5.0* $^3J_{ce}$ 9.1*
$^3J_{bd}$ 5.0* $^3J_{de}$ 9.1*

(in TFA)
* average value

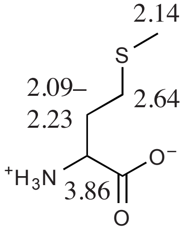


(in D_2O)

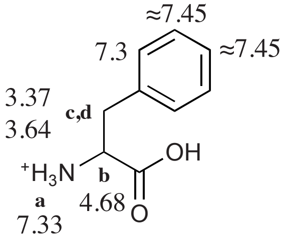


$^3J_{ab}$ 5.5 $^2J_{cd}$ -15.7
$^3J_{bc}$ 7.7 $^3J_{ce}$ 6.5*
$^3J_{bd}$ 4.4 $^3J_{de}$ 6.5*

(in TFA)
* average value

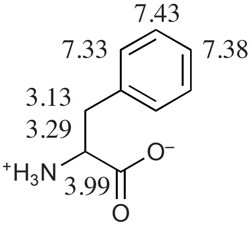


(in D_2O)

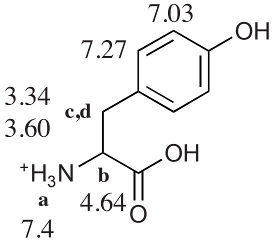


$^3J_{bc}$ 8.5
$^3J_{bd}$ 4.5
$^2J_{cd}$ -14.5

(in TFA)

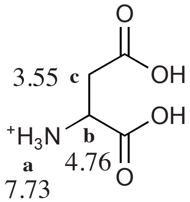


(in D_2O)

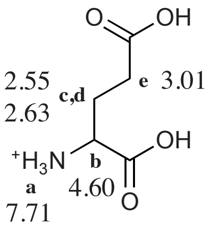


$^3J_{bc}$ 8.5
$^3J_{bd}$ 4.5
$^2J_{cd}$ -15.0

(in TFA)

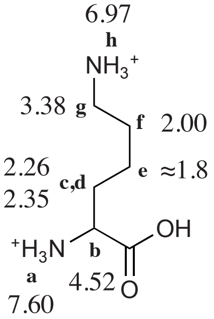


$^3J_{ab}$ 5.1
$^3J_{bc}$ 4.7
(in TFA)



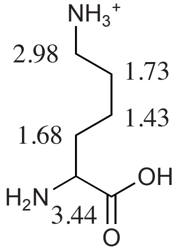
$^3J_{ab}$ 5.5 $^2J_{cd}$ -15.5
$^3J_{bc}$ 5.6* $^3J_{ce}$ 6.2*
$^3J_{bd}$ 5.6* $^3J_{de}$ 6.2*

(in TFA)
* average value

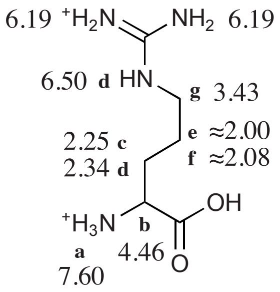


$^3J_{ab}$ 5.8 $^2J_{cd}$ -15.0
$^3J_{bc}$ 5.6* $^3J_{ce}$ 6.0*
$^3J_{bd}$ 5.6* $^3J_{de}$ 6.0*

(in TFA)
* average value

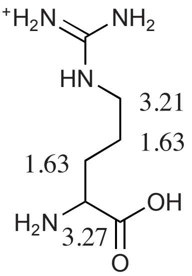


(in D_2O)

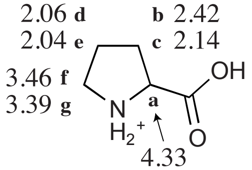


$^3J_{ab}$	5.5
$^3J_{bc}$	5.3*
$^3J_{bd}$	5.3*
$^2J_{cd}$	≈ -15.0
$^3J_{eg}$	6.5*
$^3J_{fg}$	6.5*
$^3J_{gh}$	5.3

(in TFA)
* average value

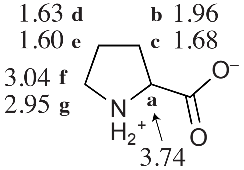


(in D_2O)



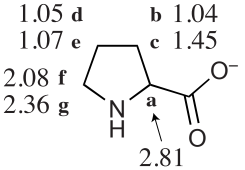
$^3J_{ab}$	8.5	$^3J_{cd}$	5.5
$^3J_{ac}$	6.5	$^3J_{ce}$	7.5
$^2J_{bc}$	-13.5	$^2J_{de}$	-13.0
$^3J_{bd}$	7.5	$^3J_{df}$	5.5
$^3J_{be}$	5.5	$^3J_{dg}$	7.5
$^4J_{bf}$	-0.4	$^3J_{ef}$	7.5
$^4J_{bg}$	0.0	$^3J_{eg}$	5.5
		$^2J_{fg}$	-11.0

(in D_2O, pH 2.0)



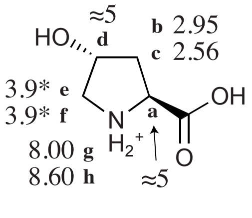
$^3J_{ab}$	8.4	$^3J_{cd}$	5.6
$^3J_{ac}$	6.2	$^3J_{ce}$	7.8
$^2J_{bc}$	-13.5	$^2J_{de}$	-13.0
$^3J_{bd}$	7.6	$^3J_{df}$	5.7
$^3J_{be}$	5.4	$^3J_{dg}$	7.9
$^4J_{bf}$	-0.4	$^3J_{ef}$	7.9
$^4J_{bg}$	0.0	$^3J_{eg}$	5.7
		$^2J_{fg}$	-11.0

(in D_2O, pH 7.0)



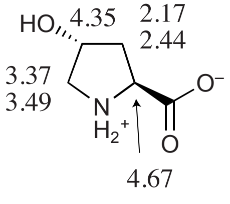
$^3J_{ab}$	8.6	$^3J_{cd}$	6.7
$^3J_{ac}$	6.6	$^3J_{ce}$	8.5
$^2J_{bc}$	-12.0	$^2J_{de}$	-11.0
$^3J_{bd}$	8.1	$^3J_{df}$	5.5
$^3J_{be}$	5.9	$^3J_{dg}$	8.1
$^4J_{bf}$	-0.6	$^3J_{ef}$	7.7
$^4J_{bg}$	0.0	$^3J_{eg}$	5.7
		$^2J_{fg}$	-10.5

(in D_2O, pH 13.0)

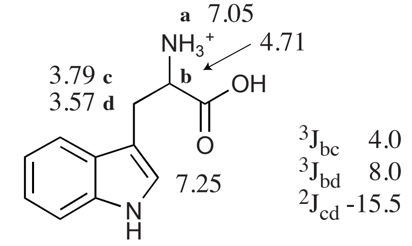


$^3J_{ab}$	8.2
$^3J_{ac}$	10.4
$^2J_{bc}$	-15.0
$^3J_{bd}$	<2
$^3J_{cd}$	4.2

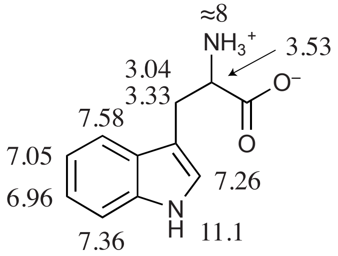
(in TFA)
* average value



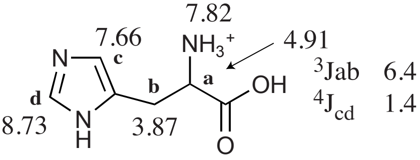
(in D_2O)



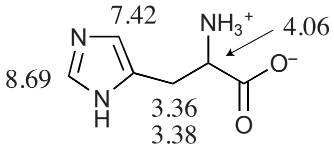
(in TFA)



(in DMSO)



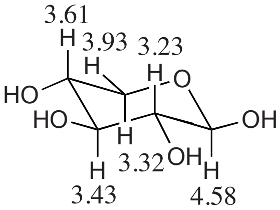
(in TFA)



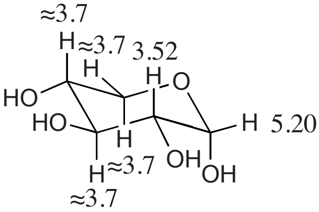
(in D_2O)

5.13.2 Carbohydrates

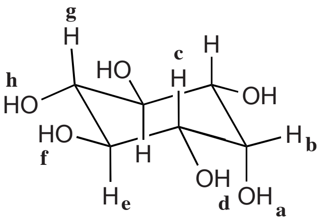
Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)



(in D_2O)

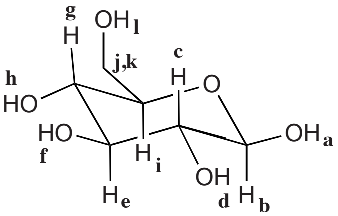


(in D_2O)



	D_2O	DMSO
a		4.55
b	4.07	3.72
c	3.54	3.14
d		4.51
e	3.64	3.37
f		4.46
g	3.27	2.93
h		4.31

Natural Products

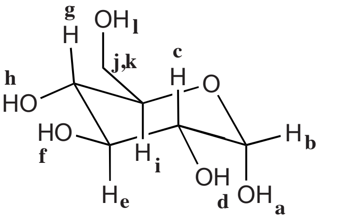


(in DMSO ca. 80%)

	D_2O*	DMSO
a		6.58
b	4.51	4.27
c	3.13	2.89
d		4.84
e	3.37	3.10
f		4.84
g	3.30	3.10
h		4.84
i	3.35	3.04
j	3.60	3.42
k	3.75	3.66
l		4.48

D_2O		DMSO	
$^3J_{bc}$	7.8	$^3J_{ab}$	6.5
$^3J_{ce}$	9.5	$^3J_{cd}$	4.5–6
$^3J_{eg}$	9.5	$^3J_{ef}$	4.5–6
$^3J_{gi}$	9.5	$^3J_{gh}$	4.5–6
$^3J_{ij}$	2.8	$^3J_{jl}$	5.5
$^3J_{ik}$	5.7	$^3J_{kl}$	6.0
$^2J_{jk}$	-12.8		

(* relative to internal acetone at δ = 2.12)

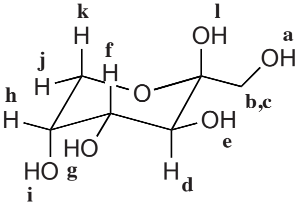


(in DMSO ca. 20%)

	D_2O*	DMSO
a		6.20
b	5.09	4.91
c	3.41	3.10
d		4.84
e	3.61	3.42
f		4.64
g	3.29	3.04
h		4.77
i	3.72	3.57
j	3.72	3.57
k	3.63	3.42
l		4.37

D_2O		DMSO	
$^3J_{bc}$	3.6	$^3J_{ab}$	4.5
$^3J_{ce}$	9.5	$^3J_{cd}$	6.8
$^3J_{eg}$	9.5	$^3J_{ef}$	4.8
$^3J_{gi}$	9.5	$^3J_{gh}$	5.5
$^3J_{ij}$	2.8	$^3J_{jl}$	5.7
$^3J_{ik}$	5.7	$^3J_{kl}$	6.2
$^2J_{jk}$	-12.8		

(* relative to internal acetone at δ = 2.12)

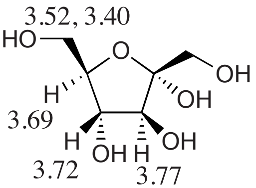


	D_2O[a]	DMSO[b]
a		4.48
b	3.68	3.39
c	3.53	3.25
d	3.76	3.55
e		4.23
f	3.86	3.58
g		4.38
h	3.96	3.62
i		4.32
j	4.00	3.77
k	3.68	3.41
l		5.14

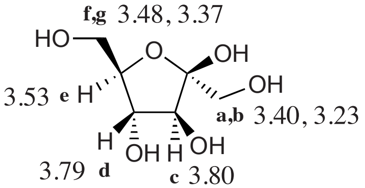
DMSO	
$^3J_{ab}$	7.4[c]
$^3J_{ac}$	5.4[c]
$^2J_{bc}$	-11.3[d]
$^3J_{de}$	6.8[c]
$^3J_{df}$	10.1[d]
$^3J_{fg}$	5.8[c]
$^3J_{fh}$	4.0[d]
$^3J_{hi}$	3.8[c]
$^3J_{hj}$	1.9[d]
$^3J_{hk}$	1.6[d]
$^2J_{jk}$	-12.1[d]

[a] 25% β-D
[b] 75% β-D

[c] at 25 °C
[d] at 70 °C



(in DMSO, 20% α-D)

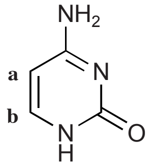


(in DMSO, 55% β-D)

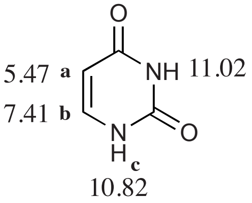
$^2J_{ab}$	-11.0
$^3J_{cd}$	7.1
$^3J_{de}$	5.9
$^3J_{ef}$	2.3
$^3J_{eg}$	3.6
$^2J_{fg}$	-11.3

5.13.3 Nucleotides and Nucleosides

Chemical Shifts and Coupling Constants (δ in ppm, J in Hz)

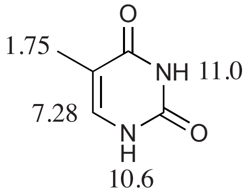


	D_2O	DMSO
a	5.97	5.57
b	7.50	7.31
J_{ab}	7	

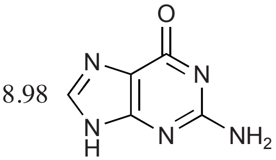


(in DMSO)

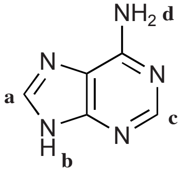
$^3J_{ab}$ 7.5
$^3J_{bc}$ 5.7



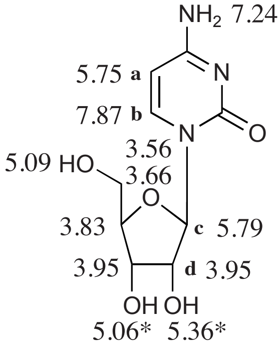
(in DMSO)



(in TFA)

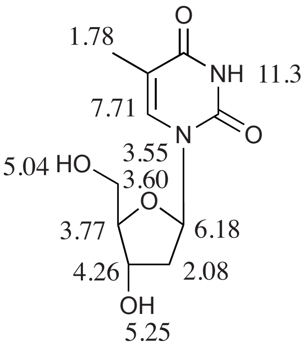


	D_2O	DMSO	$CDCl_3$	TFA
a	8.62	8.12	8.11	8.88
b			12.8	
c	8.57	8.10	8.14	9.31
d			7.09	

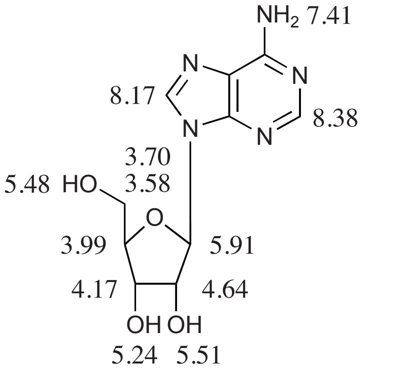


(in DMSO)
* interchangeable

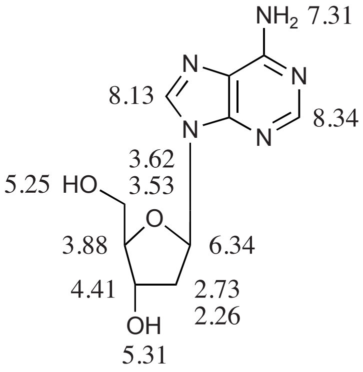
$^2J_{ab}$ 7.4
$^3J_{cd}$ 3.8



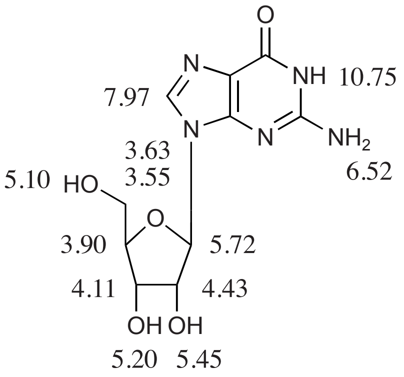
(in DMSO)



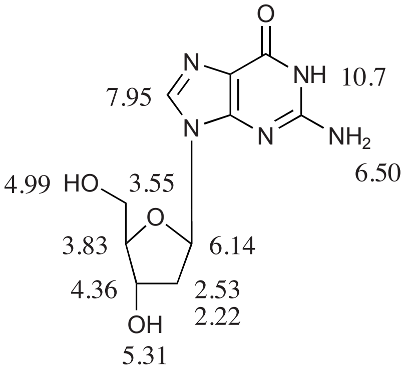
(in DMSO)



(in DMSO)



(in DMSO)



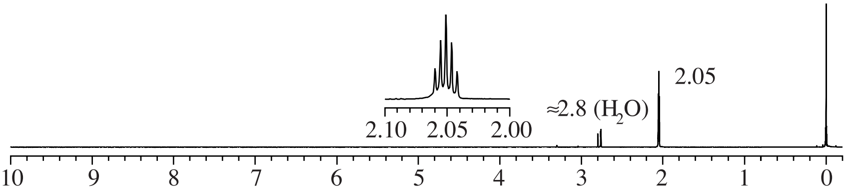
(in DMSO)

5.14 Spectra of Solvents and Reference Compounds

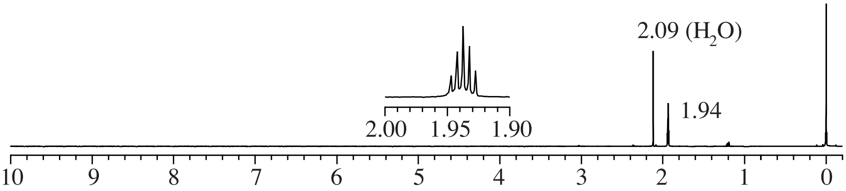
5.14.1 ^{1}H NMR Spectra of Common Deuterated Solvents

500 MHz; ≈1 000 data points per 1 ppm; δ in ppm relative to TMS

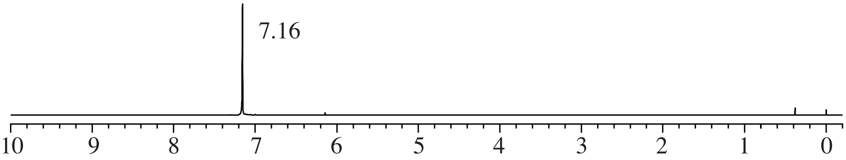
Acetone-d_6



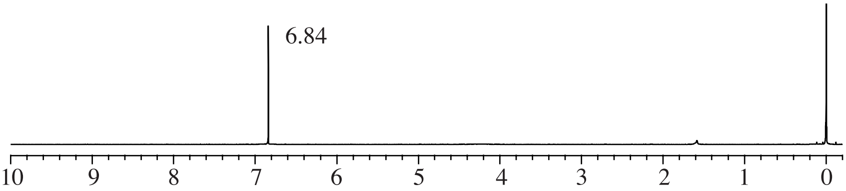
Acetonitrile-d_3



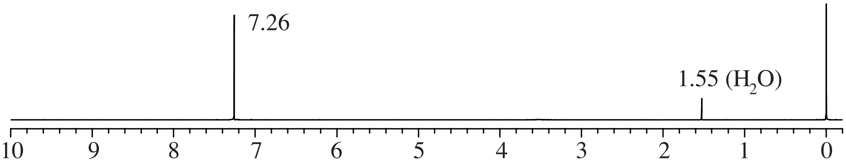
Benzene-d_6



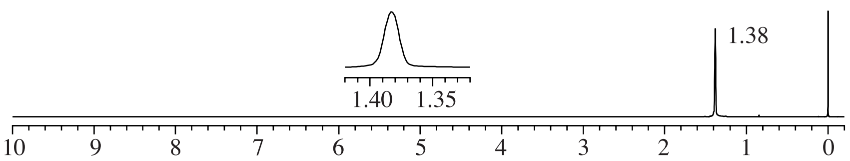
Bromoform-d



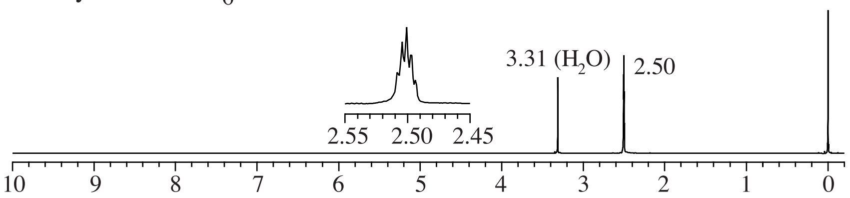
Chloroform-d



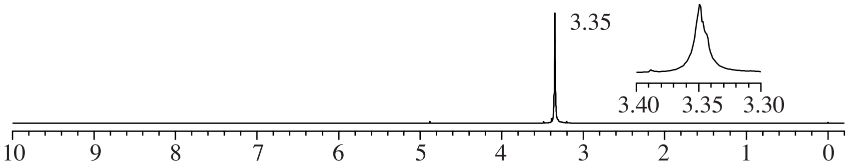
Cyclohexane-d_{12}



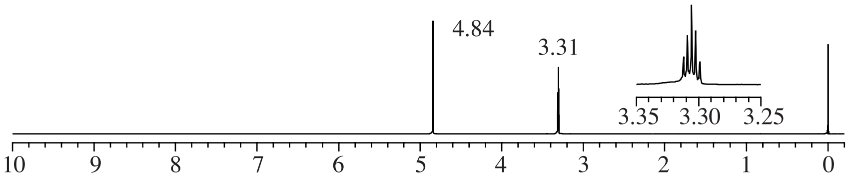
Dimethyl sulfoxide-d_6



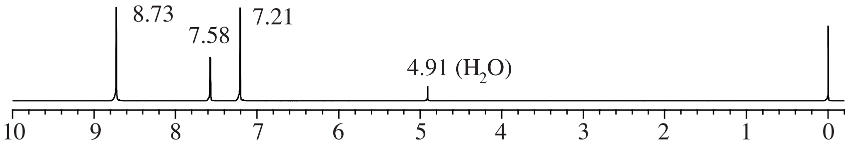
Methanol-d_1



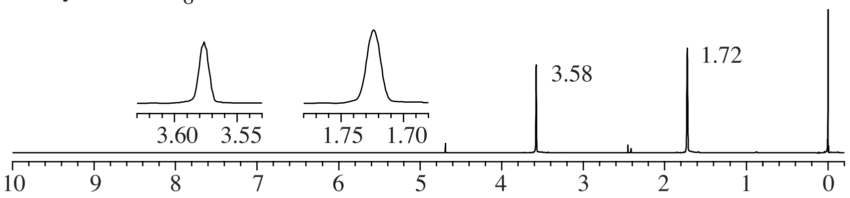
Methanol-d_4



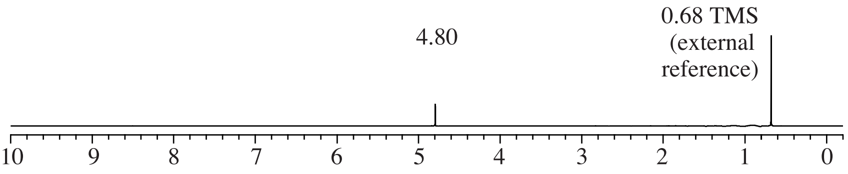
Pyridine-d_5



Tetrahydrofuran-d_8



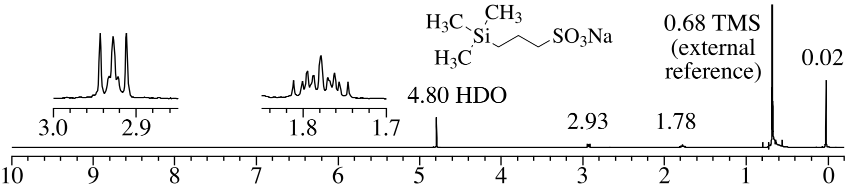
Water-d_2



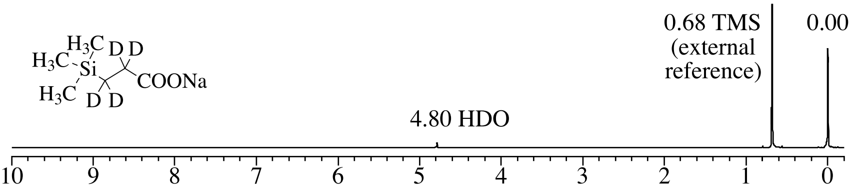
5.14.2 ^{1}H NMR Spectra of Secondary Reference Compounds

Chemical shifts in ^{1}H NMR spectra are usually reported relative to the peak position of tetramethylsilane (TMS) added to the sample as an internal reference. If TMS is not sufficiently soluble, a capillary with TMS may be used as external reference. In this case, owing to the different volume susceptibilities, the local magnetic fields in the sample and reference differ, and the peak position of the reference must be corrected. For a D_2O solution in a cylindrical sample and neat TMS in a capillary, the correction amounts to +0.68 and -0.34 ppm for superconducting and electromagnets, respectively. These values must be subtracted from the chemical shifts relative to the external TMS signal if its position is set to 0.00 ppm. Alternatively, secondary references with $(CH_3)_3SiCH_2$ groups may be used. The following spectra of two such secondary reference compounds in D_2O were measured at 500 MHz with TMS as external reference. Chemical shifts are reported in ppm relative to TMS upon correction for the difference in the volume susceptibilities of D_2O and TMS. As a result, the peak for the external TMS appears at 0.68 ppm.

3-(Trimethylsilyl)-1-propanesulfonic acid sodium salt (sodium 4,4-dimethyl-4-silapentane-1-sulfonate; DSS)

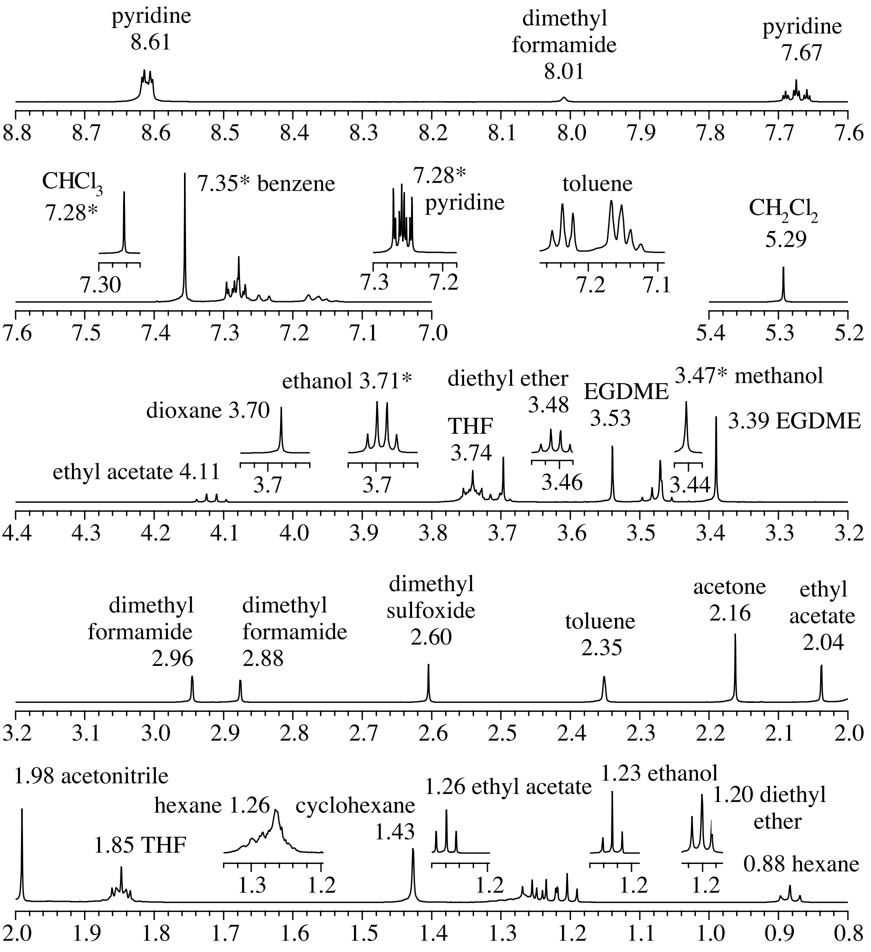


2,2,3,3-D$_4$-3-(Trimethylsilyl)propionic acid sodium salt



5.14.3 ^{1}H NMR Spectrum of a Mixture of Common Nondeuterated Solvents

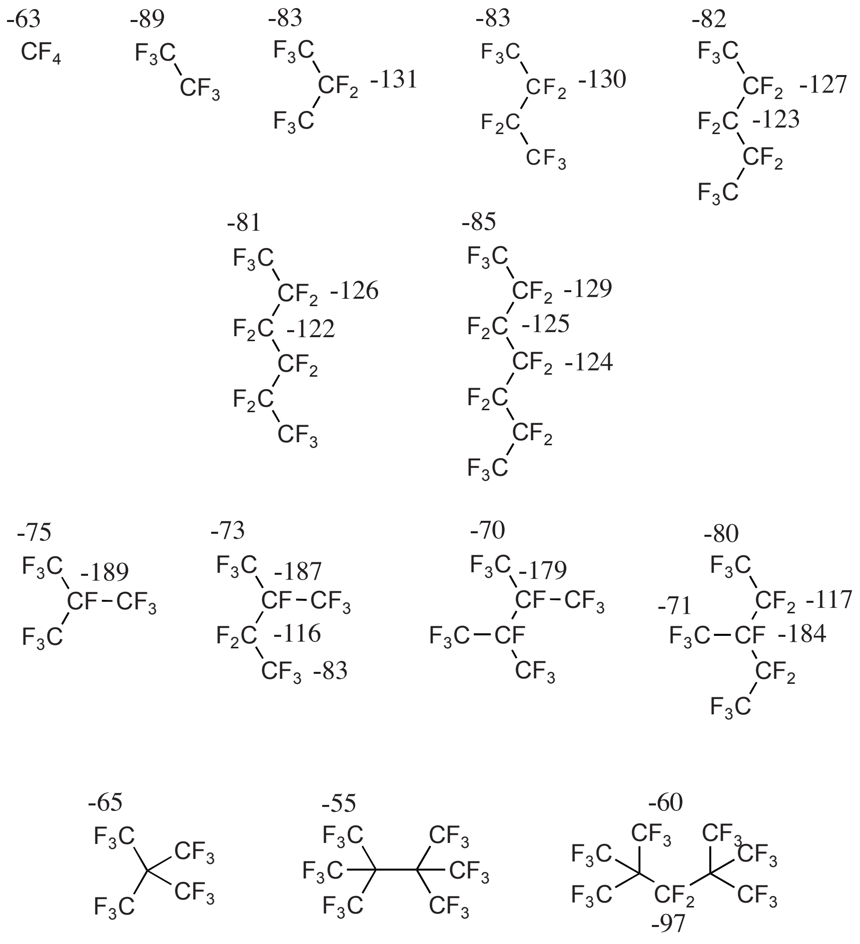
The following ^{1}H NMR spectrum (500 MHz, δ in ppm relative to TMS) of $CDCl_3$ containing 18 common solvents (0.05–0.4 vol%) is shown as a guide for the identification of possible impurities. Where the signals of several solvents overlap, insets show signals for the individual compounds from separate spectra. Peaks in these insets are labeled with the corresponding chemical shifts from their main spectrum but their values may differ by up to 0.03 ppm. Signals that are particularly prone to vary in their position are marked with *. THF: tetrahydrofuran; EGDME: ethylene glycol dimethyl ether.



6 Heteronuclear NMR Spectroscopy

6.1 ^{19}F NMR Spectroscopy

6.1.1 ^{19}F Chemical Shifts of Perfluoroalkanes (δ in ppm relative to $CFCl_3$)



^{19}F Chemical Shifts of CF_3 Groups (δ in ppm)

	Substituent	δ
	–H	-78
C	$–CH_3$	-62
	$–CH_2CH_3$	-70
	–*n*-C_7H_{15}	-67
	$–CH_2OH$	-78
	$–CH_2NH_2$	-72
	$–CH_2COOH$	-64
	$–CH_2CH_2$–1-pyridinium	-75
	$–C(CF_3)_3$	-65
	$–CF_3$	-89
	$–CF_2CF_3$	-83
	–perfluorocyclohexyl	-70
	$–CCl_3$	-82
	$–CH{=}CH_2$	-67
	$–C{\equiv}CH$	-56
	–phenyl	-64
	$–C_6F_5$	-55
	–4-nitrophenyl	-64
	–4-aminophenyl	-62
	$–C_6(CF_3)_5$	-53
	–1-naphthyl	-75
	–2-naphthyl	-73
	–2-pyridyl	-68
	–3-pyridyl	-62
	–4-pyridyl	-65

	Substituent	δ
X	–F	-63
	–Cl	-29
	–Br	-18
	–I	-5
O	–OH	-55
	–O–cyclohexyl	-58
	$–O–CF_3$	-58
	–O–phenyl	-58
	$–O–CO–CO–O–CF_3$	-31
N	$–NH_2$	-49
	$–C{\equiv}N$	-53
	–NC	-51
S	–SH	-32
	$–S–CF_3$	-39
	$–SS–CF_3$	-47
	$–SO_3H$	-79
	$–S(O)_2$–phenyl	-79
C=O	$–COCF_3$	-85
	–CO–phenyl	-58
	–COOH	-77
	$–COO^-$	-74
	$–COOCH_2CH_3$	-74
	–COF	-76
P	$–P(O)(OCH_2CH_3)_2$	-73
	$–P(CF_3)_2$	-51
	$–P^+(phenyl)_3$	-58

^{19}F Chemical Shifts of CHF_2 Groups (δ in ppm)

Substituent	δ	Substituent	δ
$-H$	-144	$-CCl_3$	-122
$-CH_3$	-110	–phenyl	-111
$-CH_2CH_3$	-120	$-O-CH_3$	-88
$-CH_2CH_2CH_3$	-117	$-O-CF_3$	-86
$-CH_2$–phenyl	-115	$-C{\equiv}N$	-120
$-CF_3$	-141	–S–phenyl	-121
$-CF_2CF_3$	-138	–COOH	-127
–cyclohexyl	-126	$-P(CF_3)_2$	-126

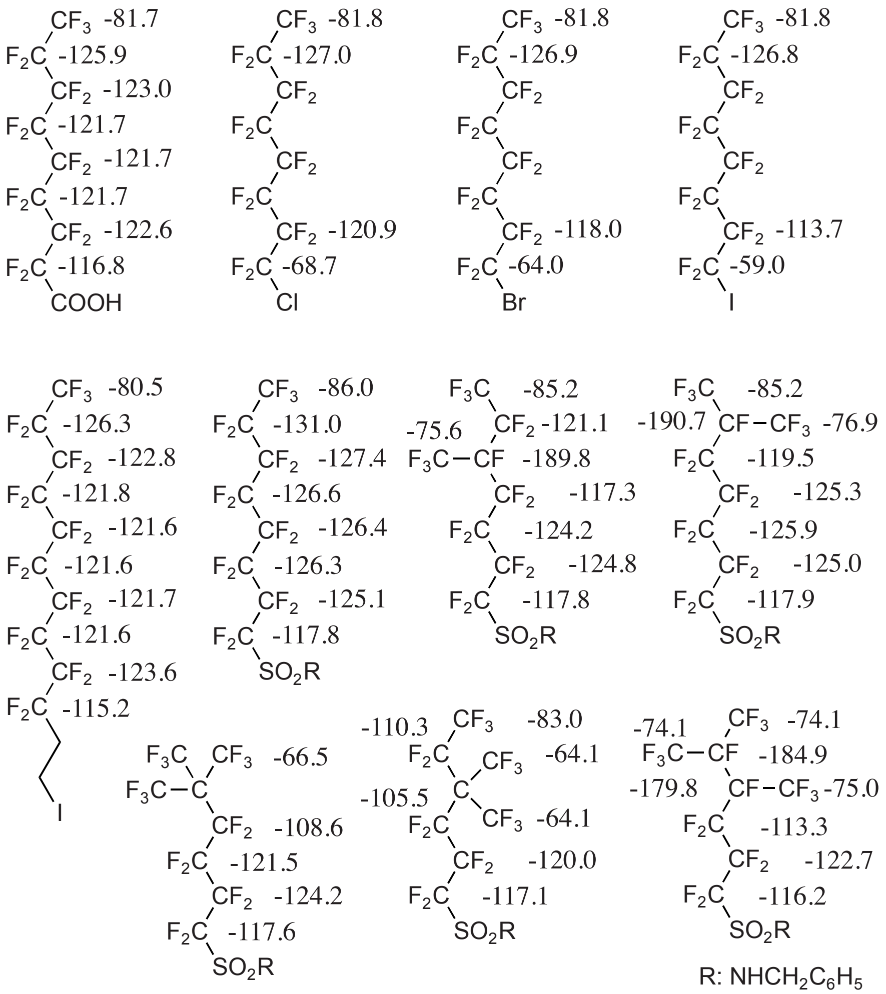
^{19}F Chemical Shifts of CH_2F Groups (δ in ppm)

Substituent	δ	Substituent	δ
$-H$	-268	$-CCl_3$	-198
$-CH_3$	-212	$-CH{=}CH_2$	-216
$-CH_2CH_3$	-212	$-C{\equiv}CH$	-218
$-CH_2CH_2CH_3$	-219	–phenyl	-206
$-CH_2CH_2CH_2CH_3$	-219	$-C{\equiv}N$	-251
$-CH_2OH$	-226	–CO–phenyl	-226
$-CH_2$–phenyl	-216	–COOH	-229
$-CF_3$	-241	$-COO^-$	-218
$-CF_2CF_3$	-243		

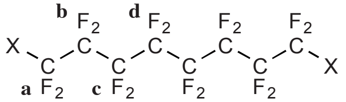
^{19}F Chemical Shifts of CF_2R_2, $CHFR_2$, and CFR_3 Groups (δ in ppm)

Substituent	CF_2R_2	$CHFR_2$	CFR_3
$-CH_3$	-85	-165	-131
$-CH_2CH_3$	-92	-183	-156
$-CF_3$	-132	-77	-189
–phenyl	-89	-167	-127
–Cl	-7	-81	0

^{19}F Chemical Shifts of Monosubstituted Perfluoroalkanes (δ in ppm) [1]

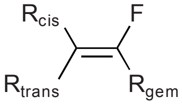


Halogen Bonding (δ in ppm) [2]



X	in cyclohexane				in pyridine			
	a	**b**	**c**	**d**	**a**	**b**	**c**	**d**
–F	-81.1	-121.7	-122.5	-126.1	-80.9	-122.1	-122.8	-126.2
–Br	-65.1	-118.1	-122.6	-123.2	-67.7	-117.9	-122.2	-122.8
–I	-60.0	-114.6	-122.5	-123.3	-71.6	-115.2	-122.0	-122.8

6.1.2 Estimation of ^{19}F Chemical Shifts of Substituted Fluoroethylenes (δ in ppm relative to $CFCl_3$) [3]

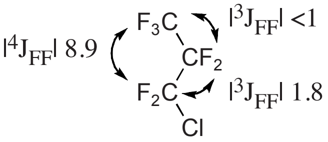


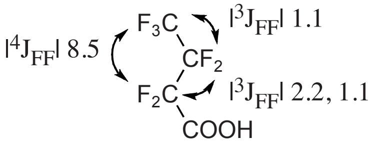
$$\delta_{C=CF} = -133.9 + Z_{cis} + Z_{trans} + Z_{gem} + S_{cis/trans} + S_{cis/gem} + S_{trans/gem}$$

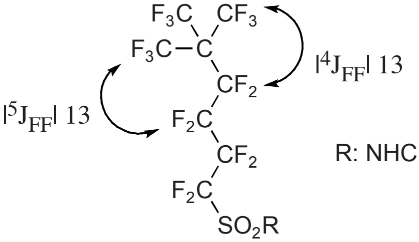
Substituent R	Z_{cis}	Z_{trans}	Z_{gem}
–H	-7.4	-31.3	49.9
$–CH_3$	-6.0	-43.0	9.5
$–CF_3$	-25.3	-40.7	54.3
$–CH{=}CH_2$	–	–	47.7
$–CF{=}CF_2$	-23.8	-38.9	44.7
–phenyl	-15.7	-35.1	38.7
–F	0.0	0.0	0.0
–Cl	-16.5	-29.4	–
–Br	-17.7	-40.0	–
–I	-21.3	-46.3	17.4
$–OC_2H_5$	-77.5	–	84.2
–COF	-46.5	-56.8	54.1
$–SCH_3$	-25.1	-43.7	16.6

Substituent	Substituent	$S_{cis/trans}$	$S_{cis/gem}$	$S_{trans/gem}$
–H	–H	-26.6	–	2.8
–H	$–CF_3$	-21.3	–	–
–H	$–CH_3$	–	11.4	–
–H	$–OCH_2CH_3$	-47.0	–	–
–H	–phenyl	-4.8	–	5.2
$–CF_3$	–H	-7.5	-10.6	12.5
$–CF_3$	$–CF_3$	-5.9	-5.3	-4.7
$–CF_3$	$–CH_3$	17.0	–	–
$–CF_3$	–phenyl	-15.6	–	-23.4
$–CH_3$	–H	–	-12.2	–
$–CH_3$	$–CF_3$	–	-13.8	-8.9
$–CH_3$	–phenyl	–	-19.5	-19.5
$–OCH_2CH_3$	–H	-5.1	–	–
–phenyl	–H	–	–	20.1
–phenyl	$–CF_3$	-23.2	–	–

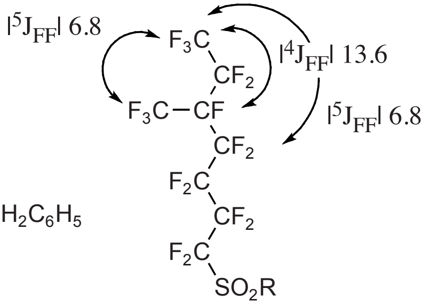
6.1.3 Coupling Constants in Fluorinated Alkanes and Alkenes (J_{FF} in Hz)

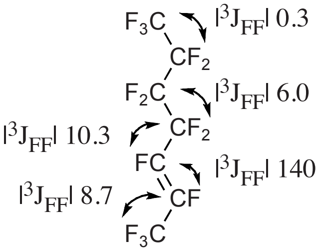


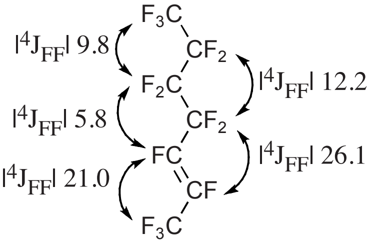


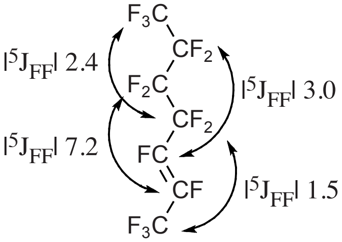


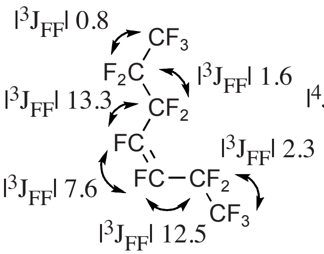
R: $NHCH_2C_6H_5$

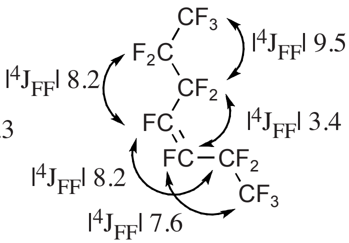


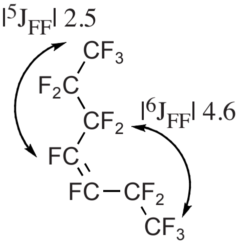


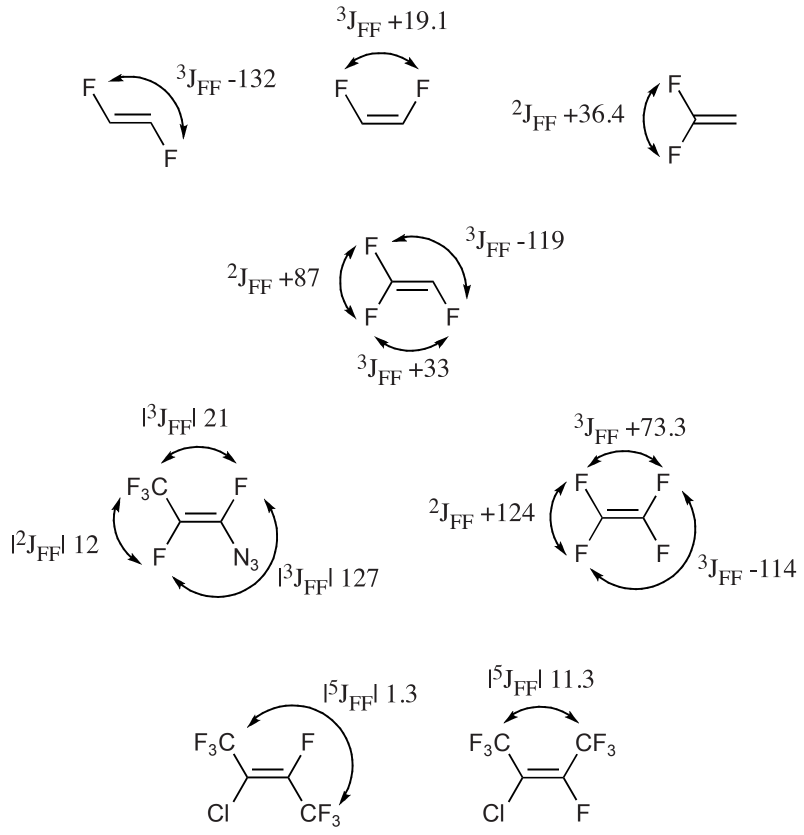




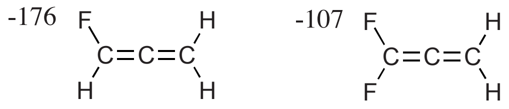


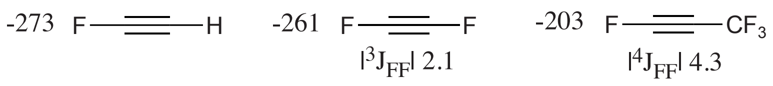




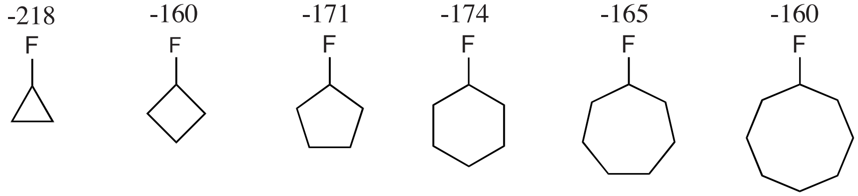


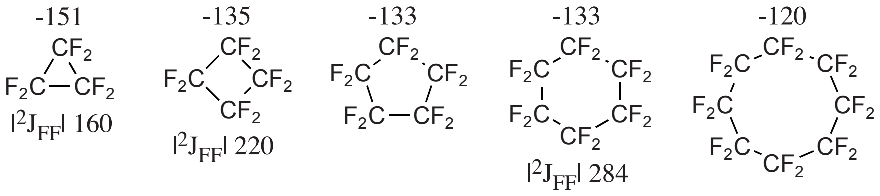
6.1.4 ^{19}F Chemical Shifts of Allenes and Alkynes (δ in ppm relative to $CFCl_3$, $|J_{FF}|$ in Hz)

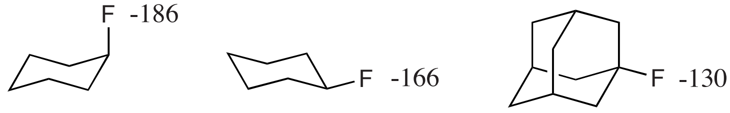


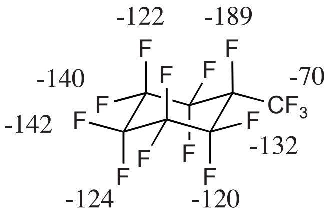


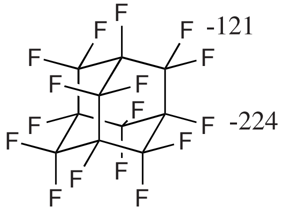
6.1.5 ^{19}F Chemical Shifts and Coupling Constants of Fluorinated Alicyclics (δ in ppm relative to $CFCl_3$, $|J_{FF}|$ in Hz)

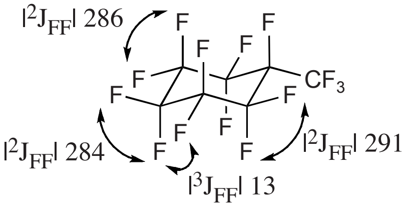


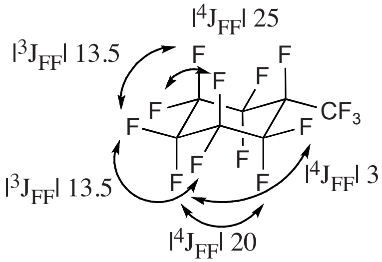


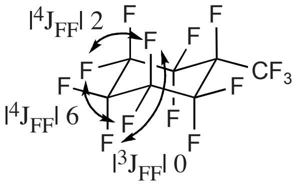






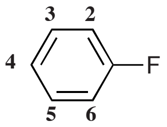






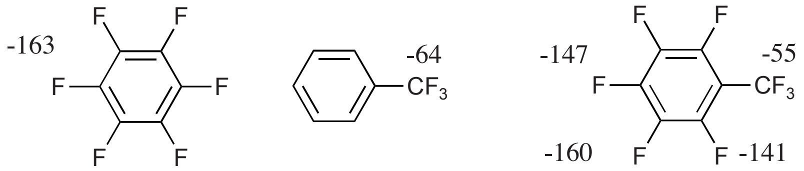
6.1.6 ^{19}F Chemical Shifts and Coupling Constants of Aromatics and Heteroaromatics (δ in ppm relative to $CFCl_3$)

Estimation of ^{19}F Chemical Shifts of Substituted Fluorobenzenes [4]



$$\delta_F = -113.9 + \sum Z_i$$

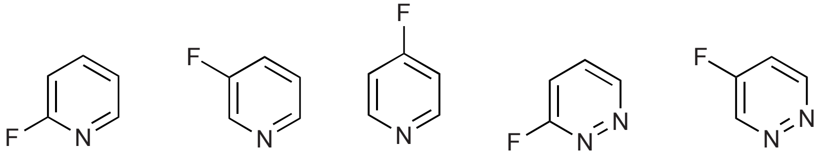
	Substituent	$Z_{2,6}$	$Z_{3,5}$	Z_4
C	$-CH_3$	-3.9	-0.4	-3.6
	$-CF_3$	0.4	3.1	5.8
	$-CH=CH_2$	-4.4	0.7	-0.6
	$-C\equiv CH$	–	–	3.3
X	–F	-23.2	2.0	-6.6
	–Cl	-0.3	3.5	-0.7
	–Br	7.6	3.5	0.1
	–I	19.9	3.6	1.4
O	–OH	-23.5	0.0	-13.3
	$-OCH_3$	-18.9	-0.8	-9.0
	$-OCOCH_3$	–	–	-3.7
N	$-NH_2$	-22.9	-1.3	-17.4
	$-NHCOOCH_3$	–	0.1	-7.1
	$-NHCONH_2$	–	0.9	-8.1
	$-N_3$	-11.4	2.8	-0.3
	$-NO_2$	-5.6	3.8	9.6
	$-C\equiv N$	6.9	4.1	10.1
	–NCO	-9.2	2.3	-2.2
S	–SH	10.0	0.9	-3.5
	$-SCH_3$	6.5	1.2	-4.5
	$-S(O)_2F$	7.5	5.8	13.8
	$-S(O)_2-CF_3$	9.5	5.5	-14.3
	$-S(O)_2OCH_2CH_3$	–	3.7	9.1
O=C	–CHO	-7.4	2.1	10.3
	$-COCH_3$	2.5	1.8	7.6
	–COOH	2.3	1.1	6.5
	$-COOCH_3$	3.3	3.8	7.1
	$-CONH_2$	0.5	-0.8	3.4
	–COF	-14.8	3.0	6.2
	–COCl	3.4	3.5	12.9
	$-B(OH)_2$	6.8	0.8	2.1
	$-Si(CH_3)_3$	13.8	0.3	1.6



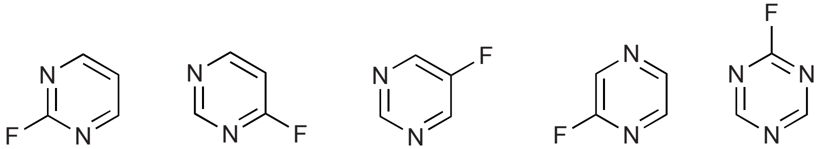
Estimation of ^{19}F Chemical Shifts of Substituted Pyridines, Pyrimidines, Pyrazines, and Triazines (δ in ppm) [5]

$$\delta_F = Y + \sum Z_i$$

To estimate the ^{19}F chemical shifts of substituted 6-ring heteroaromatics, the same increments, Z_i, can be used as for substituted fluorobenzenes (see preceding page). However, different base values, Y (as given below), apply depending on the number and position of nitrogens and the position of the fluorine substituent in question:

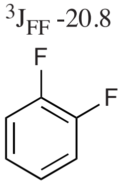


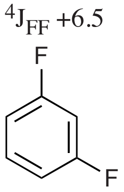
Y: -47.8 -3.0 -18.3 -45.7 -3.5

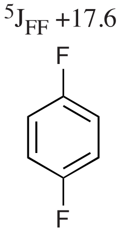


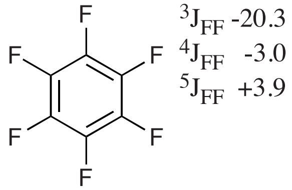
Y: -67.7 -52.1 -18.7 -41.8 -74.6

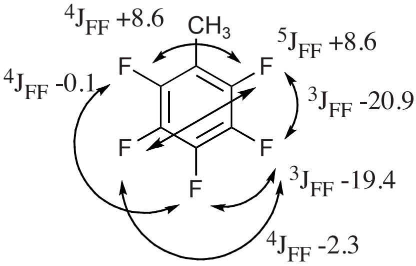
Coupling Constants in Aromatics and Heteroaromatics (J_{FF} in Hz)

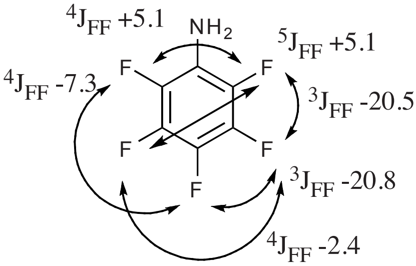


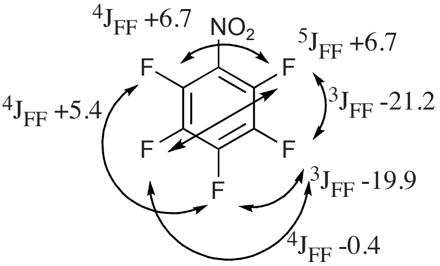


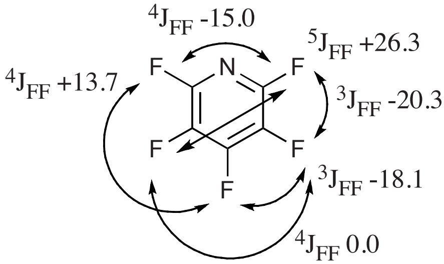


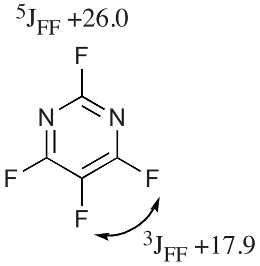




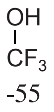


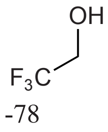


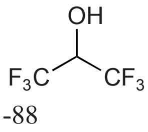


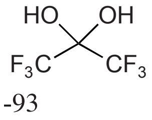


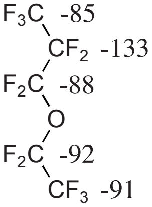
6.1.7 ^{19}F Chemical Shifts of Alcohols and Ethers (δ in ppm relative to $CFCl_3$)

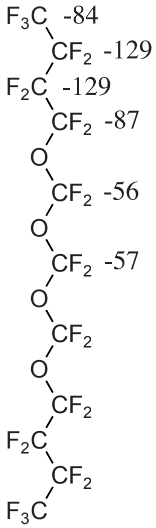


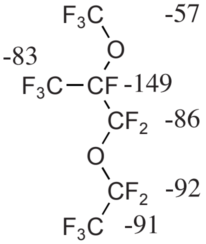


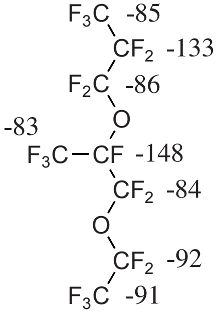


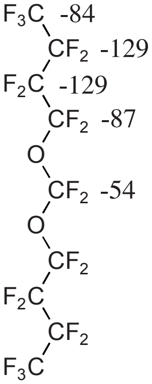


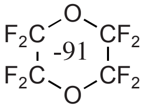


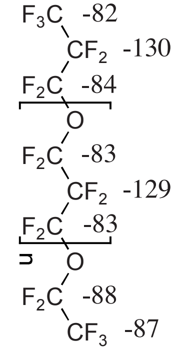


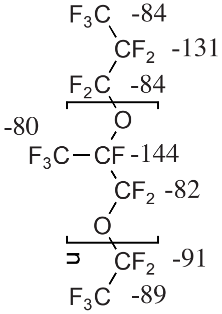


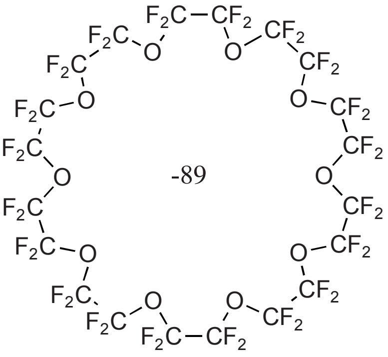


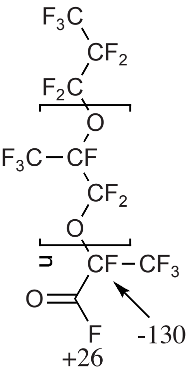


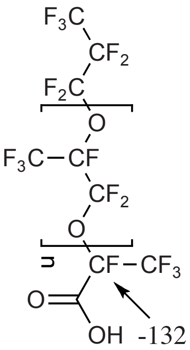




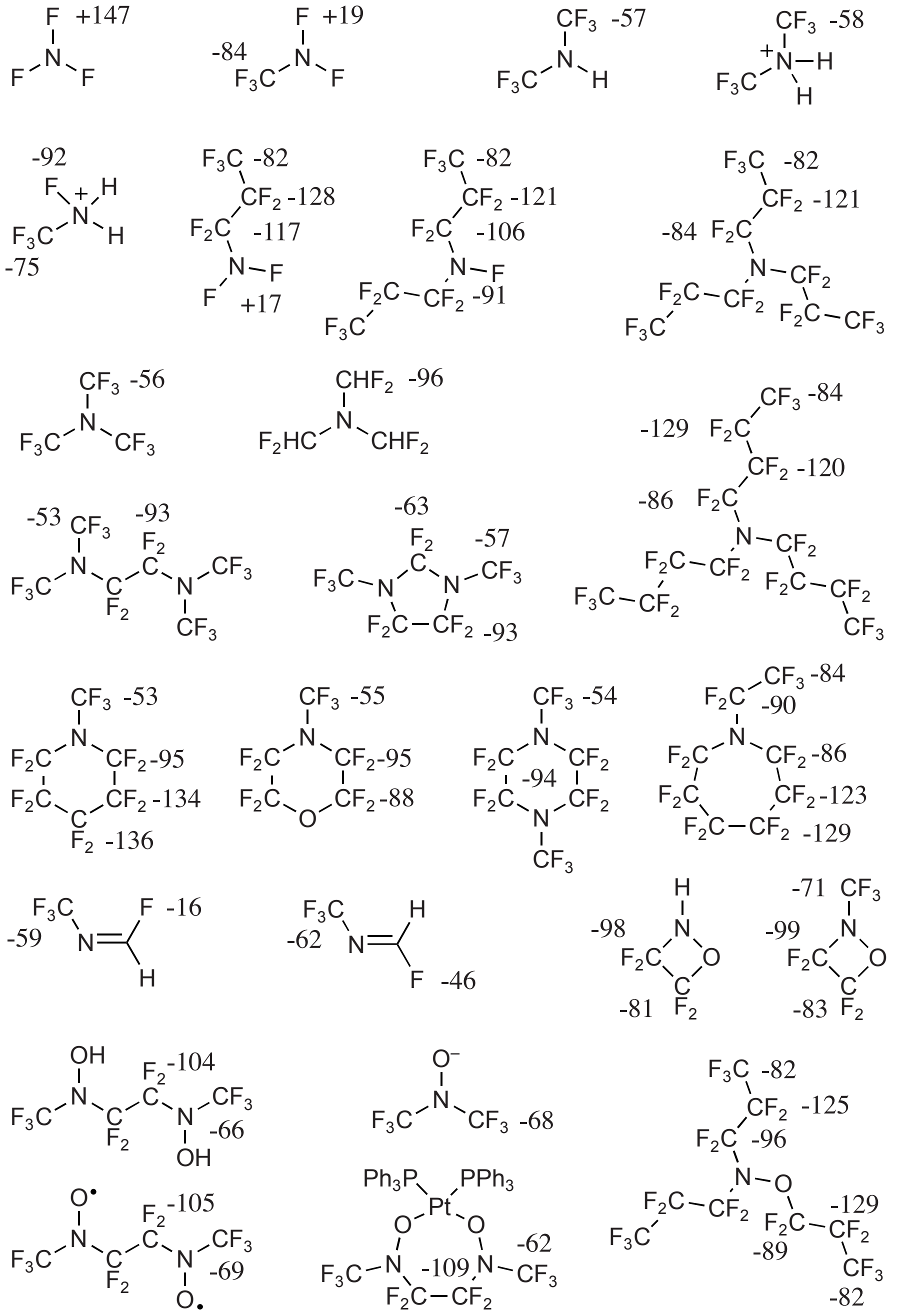




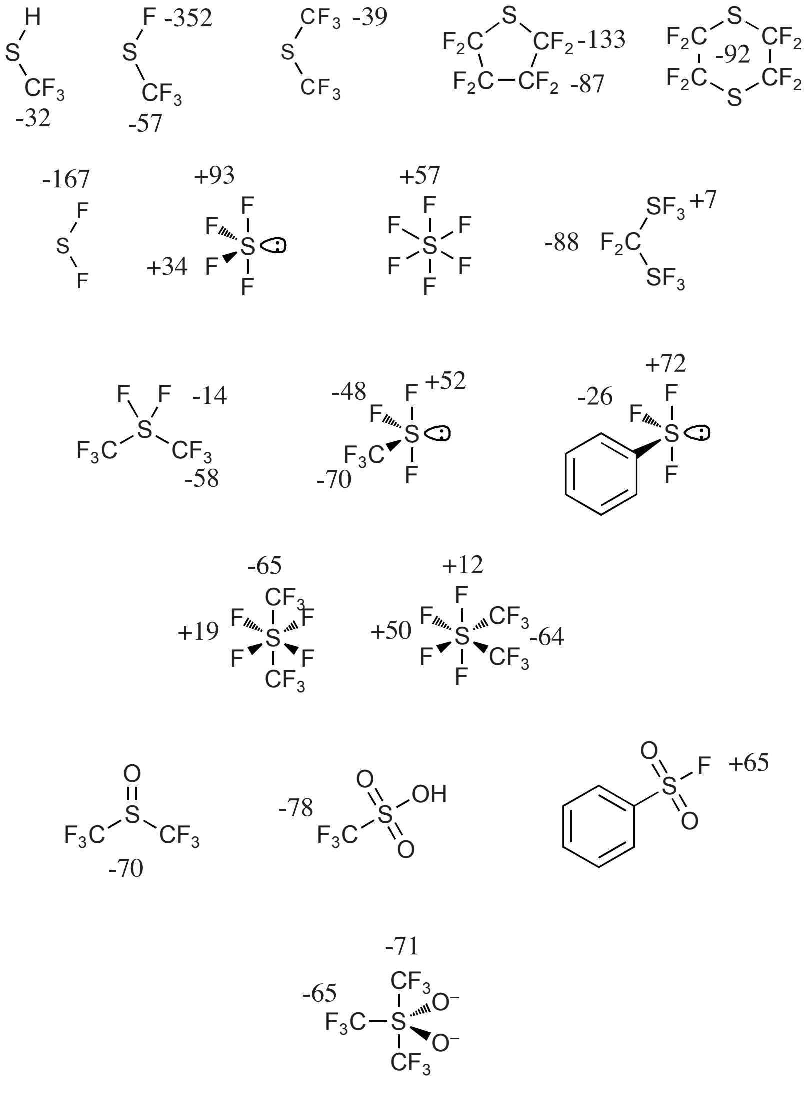




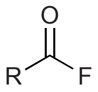
6.1.8 ^{19}F Chemical Shifts of Fluorinated Amine, Imine, and Hydroxylamine Derivatives (δ in ppm relative to $CFCl_3$)



6.1.9 ^{19}F Chemical Shifts of Sulfur Compounds (δ in ppm relative to $CFCl_3$)

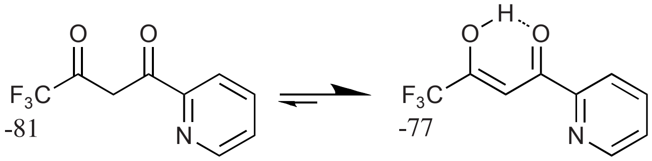


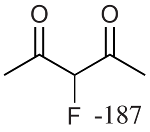
6.1.10 ^{19}F Chemical Shifts of Carbonyl and Thiocarbonyl Compounds (δ in ppm relative to $CFCl_3$)

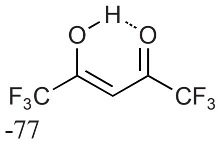


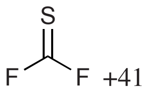
Substituent R	δ
–H	+41
$–CH_3$	+49
$–C(CH_3)_3$	+22
$–CH_2F$	+26
$–CF_3$	+15
$–CF(CF_3)_3$	+31
$–CH{=}CH_2$	+24

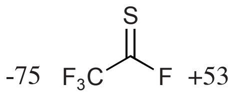
Substituent R	δ
–phenyl	+17
–F	-23
$–NH–CH_2CH_2CH_3$	-16
–O–cyclohexyl	-8
–O–phenyl	-17
–S–phenyl	+47



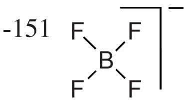


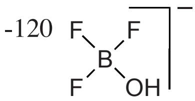


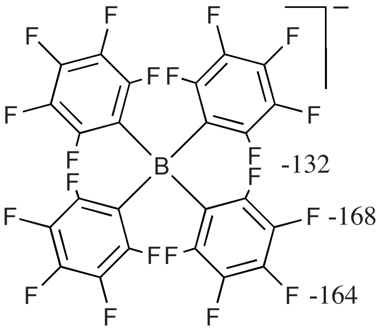


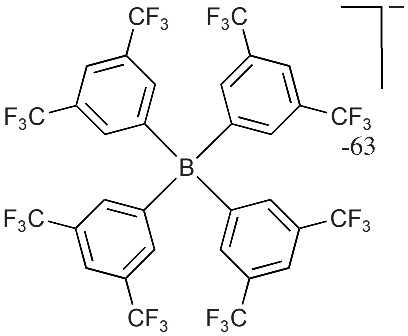


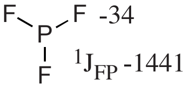
6.1.11 ^{19}F Chemical Shifts of Fluorinated Boron, Phosphorus, and Silicon Compounds (δ in ppm relative to $CFCl_3$, J_{FP} in Hz)

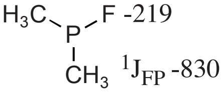


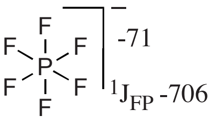


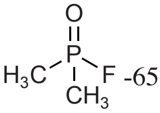


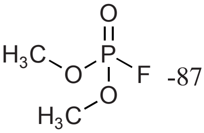


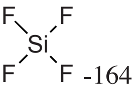


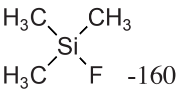




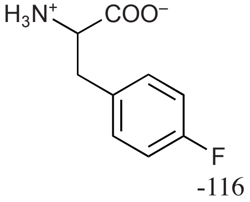


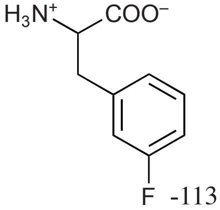


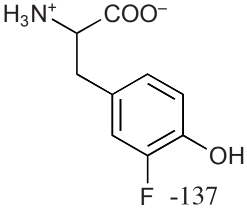


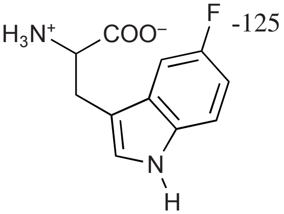


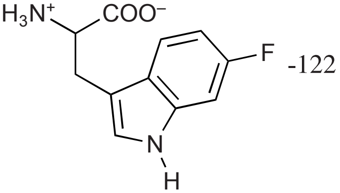
6.1.12 ^{19}F Chemical Shifts of Natural Product Analogues (δ in ppm relative to $CFCl_3$, J_{FF} in Hz)

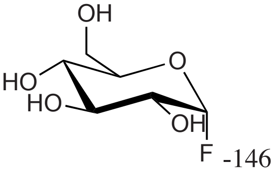


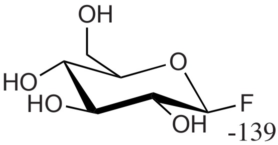


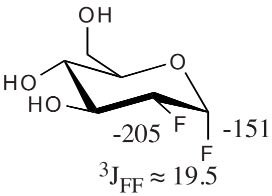


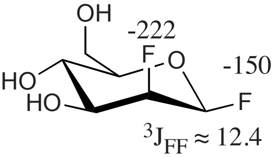


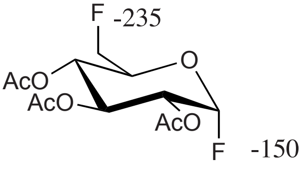


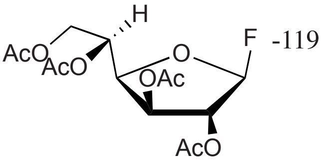


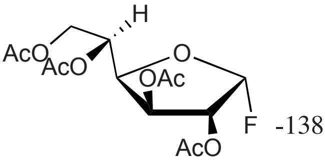


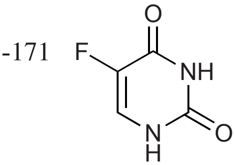


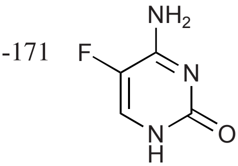












6.1.13 References

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6.2 ^{31}P NMR Spectroscopy

6.2.1 ^{31}P Chemical Shifts of Tricoordinated Phosphorus, $PR^1R^2R^3$ (δ in ppm relative to H_3PO_4)

	Substituent R^1	R^2	R^3	δ
H_2	$-H$	$-H$	$-H$	-235
	$-CH_3$	$-H$	$-H$	-164
	$-CH_2CH_3$	$-H$	$-H$	-127
	–phenyl	$-H$	$-H$	-124
H	$-CH_3$	$-CH_3$	$-H$	-99
	$-CH_2CH_3$	$-CH_2CH_3$	$-H$	-55
	–phenyl	–phenyl	$-H$	-41
	$-OCH_3$	$-OCH_3$	$-H$	171
C	$-CH_3$	$-CH_3$	$-CH_3$	-63
	$-CH_2CH_3$	$-CH_2CH_3$	$-CH_2CH_3$	-20
	$-CH_2CH_2CH_3$	$-CH_2CH_2CH_3$	$-CH_2CH_2CH_3$	-33
	$-CH(CH_3)_2$	$-CH(CH_3)_2$	$-CH(CH_3)_2$	20
	$-C(CH_3)_3$	$-C(CH_3)_3$	$-C(CH_3)_3$	62
	–phenyl	$-CH_3$	$-CH_3$	-48
	–phenyl	–phenyl	$-CH_3$	-28
	–phenyl	–phenyl	–phenyl	-6
X	$-CH_3$	$-CH_3$	$-F$	185
	$-CH_3$	$-CH_3$	$-Cl$	92
	$-CH_3$	$-CH_3$	$-Br$	88
	$-CH_3$	$-F$	$-F$	244
	$-CH_3$	$-Cl$	$-Cl$	192
	$-CH_3$	$-Br$	$-Br$	184
	$-CH_3$	$-I$	$-I$	131
	$-F$	$-F$	$-F$	97
	$-Cl$	$-Cl$	$-Cl$	220
	$-Br$	$-Br$	$-Br$	227
	$-I$	$-I$	$-I$	178
O	$-OCH_3$	$-CH_3$	$-CH_3$	91
	$-OCH_3$	$-OCH_3$	$-CH_3$	183
	$-OCH_3$	$-OCH_3$	$-OCH_3$	140
	$-OCH_2CH_3$	$-OCH_2CH_3$	$-OCH_2CH_3$	138
N	$-N(CH_3)_2$	$-CH_3$	$-CH_3$	39
	$-N(CH_3)_2$	–phenyl	–phenyl	65
	$-N(CH_3)_2$	$-N(CH_3)_2$	$-CH_3$	86
	$-N(CH_3)_2$	$-N(CH_3)_2$	$-N(CH_3)_2$	123
S	$-SCH_3$	$-SCH_3$	$-SCH_3$	125

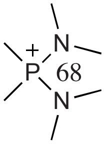
6.2.2 ^{31}P Chemical Shifts of Tetracoordinated Phosphonium Compounds (δ in ppm relative to H_3PO_4)

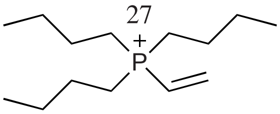
^{31}P Chemical Shifts of Symmetrically Substituted Phosphonium Compounds, PR_4^+

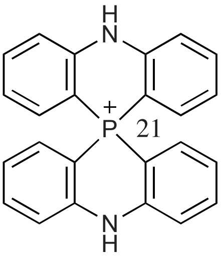
Substituent R	δ	Substituent R	δ
$-CH_3$	25	–*n*-butyl	34
$-CH_2CH_3$	41	–phenyl	23
–*n*-propyl	31	$-OCH_3$	5

^{31}P Chemical Shifts of Triphenylphosphonium Compounds, $P(phenyl)_3R^+$

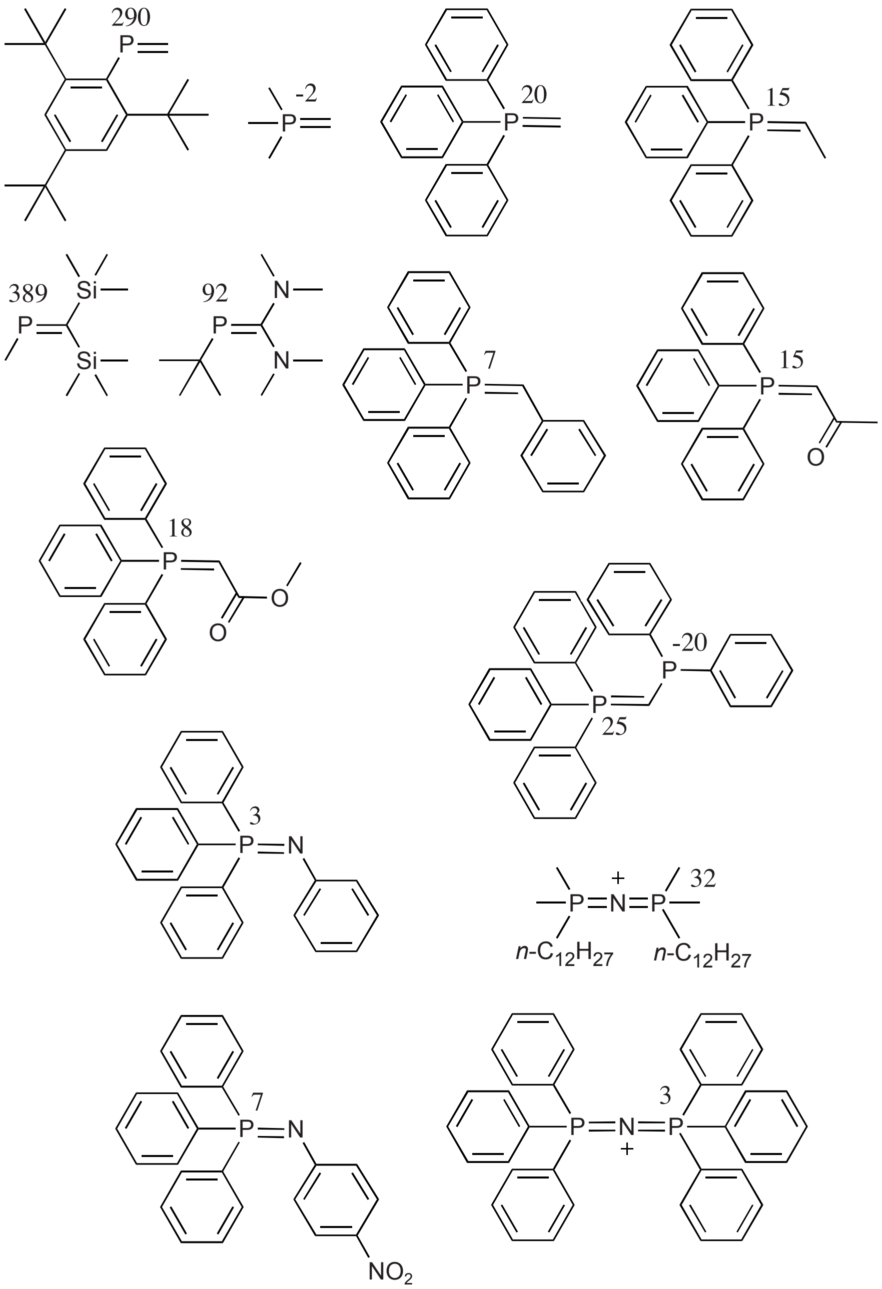
Substituent R	δ	Substituent R	δ
$-CH_3$	23	$-CH{=}CH_2$	19
$-CH_2CH_3$	26	$-CH{=}C{=}CH_2$	19
$-CH_2Cl$	24	–C≡C–phenyl	5
$-CH_2OH$	18	$-NH_2$	36
$-CH_2COCH_3$	26	$-N(CH_3)_2$	48
$-CH_2COOCH_2CH_3$	21	$-OCH_2CH_3$	62





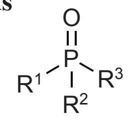


6.2.3 ^{31}P Chemical Shifts of Compounds with a P=C or P=N Bond (δ in ppm relative to H_3PO_4)



6.2.4 ^{31}P Chemical Shifts of Tetracoordinated P(=O) and P(=S) Compounds (δ in ppm relative to H_3PO_4)

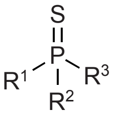
^{31}P Chemical Shifts of Tetracoordinated P(=O) Compounds



Substituent R^1		R^2	R^3	δ
C	$-CH_3$	$-CH_3$	–H	63
	$-CH_3$	$-CH_3$	$-CH_3$	41
	$-CH_2CH_3$	$-CH_2CH_3$	$-CH_2CH_3$	48
	–phenyl	–phenyl	–phenyl	27
X	$-CH_3$	$-CH_3$	–F	66
	$-CH_3$	$-CH_3$	–Cl	65
	$-CH_3$	$-CH_3$	–Br	51
	$-CH_2CH_3$	$-CH_2CH_3$	–Cl	77
	–phenyl	–phenyl	–Cl	43
	$-CH_3$	–F	–F	27
	$-CH_3$	–Cl	–Cl	44
	$-CH_3$	–Br	–Br	9
	$-CH_2CH_3$	–Cl	–Cl	55
	–F	–F	–F	-36
	–Cl	–Cl	–Cl	2
	–Br	–Br	–Br	-103
N	$-N(CH_3)_2$	$-N(CH_3)_2$	$-N(CH_3)_2$	24
O	–H	–H	$-OCH_3$	19
	$-CH_3$	–H	–OH	35
	$-CH_3$	$-CH_3$	–OH	31
	$-CH_3$	$-CH_3$	$-OCH_3$	52
	–phenyl	–phenyl	–OH	29
	–phenyl	–phenyl	$-OCH_3$	32
	$-CH_3$	–Cl	$-OCH_2CH_3$	40
	–Cl	–Cl	$-OCH_3$	6
	–F	–F	$-OCH_2CH_3$	-21
2 O	–H	$-OCH_3$	$-OCH_3$	11
	$-CH_3$	–OH	–OH	31
	$-CH_3$	$-OCH_3$	$-OCH_3$	32
	$-CCl_3$	$-OCH_2CH_3$	$-OCH_2CH_3$	7
	–phenyl	–OH	–OH	18
	–phenyl	$-OCH_3$	$-OCH_3$	21
	–Cl	$-OCH_2CH_3$	$-OCH_2CH_3$	3

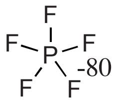
	Substituent R^1	R^2	R^3	δ
3 O	–OH	–OH	–OH	0
	$–OCH_3$	$–OCH_3$	$–OCH_3$	0
	$–OCH_2CH_3$	$–OCH_2CH_3$	$–OCH_2CH_3$	-1
	$–OCH(CH_3)_2$	$–OCH(CH_3)_2$	$–OCH(CH_3)_2$	-13
	–O–phenyl	–OH	–OH	-4
	–O–phenyl	–O–phenyl	–OH	-11
	–O–phenyl	–O–phenyl	–O–phenyl	-18
S	–S–*n*-butyl	–S–*n*-butyl	–OH	37
	–S–*n*-butyl	–S–*n*-butyl	– S–*n*-butyl	62

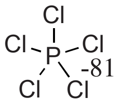
^{31}P Chemical Shifts of Tetracoordinated P(=S) Compounds

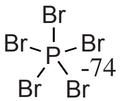


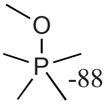
	Substituent R^1	R^2	R^3	δ
C	$–CH_3$	$–CH_3$	$–CH_3$	59
	$–CH_2CH_3$	$–CH_2CH_3$	$–CH_2CH_3$	53
	–phenyl	–phenyl	–phenyl	43
X	$–CH_3$	$–CH_3$	–Cl	87
	–phenyl	–phenyl	–Cl	80
	$–CH_3$	$–CH_3$	–Br	63
	$–CH_2CH_3$	–F	–F	111
	$–CH_3$	–Cl	–Cl	81
	$–CH_2CH_3$	–Cl	–Cl	95
	–F	–F	–F	32
	–Cl	–Cl	–Cl	29
	–Br	–Br	–Br	-112
	–Br	–I	–I	-315
N	$–N(CH_2CH_3)_2$	$–N(CH_2CH_3)_2$	$–N(CH_2CH_3)_2$	78
O	$–CH_3$	$–OCH_3$	$–OCH_3$	100
	$–OCH_2CH_3$	$–OCH_2CH_3$	$–OCH_2CH_3$	68
S	$–CH_3$	–S–*n*-propyl	–S–*n*-propyl	78
	–S–*n*-butyl	–S–*n*-butyl	$–OCH_3$	111
	–S–*n*-propyl	–S–*n*-propyl	–S–*n*-propyl	93

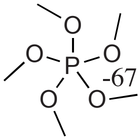
6.2.5 ^{31}P Chemical Shifts of Penta- and Hexacoordinated Phosphorus Compounds (δ in ppm relative to H_3PO_4)

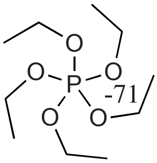


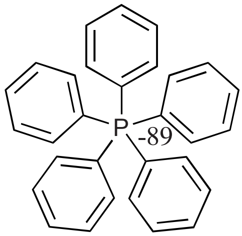


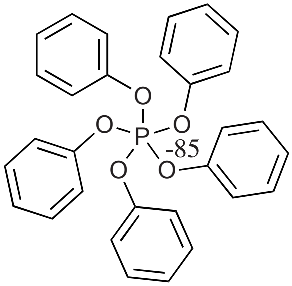


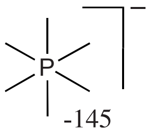


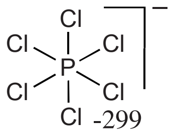


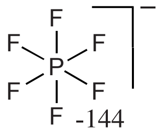


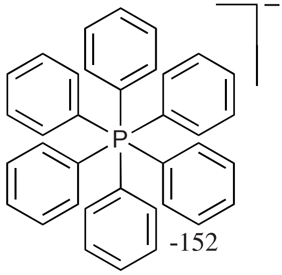


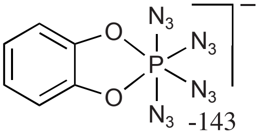




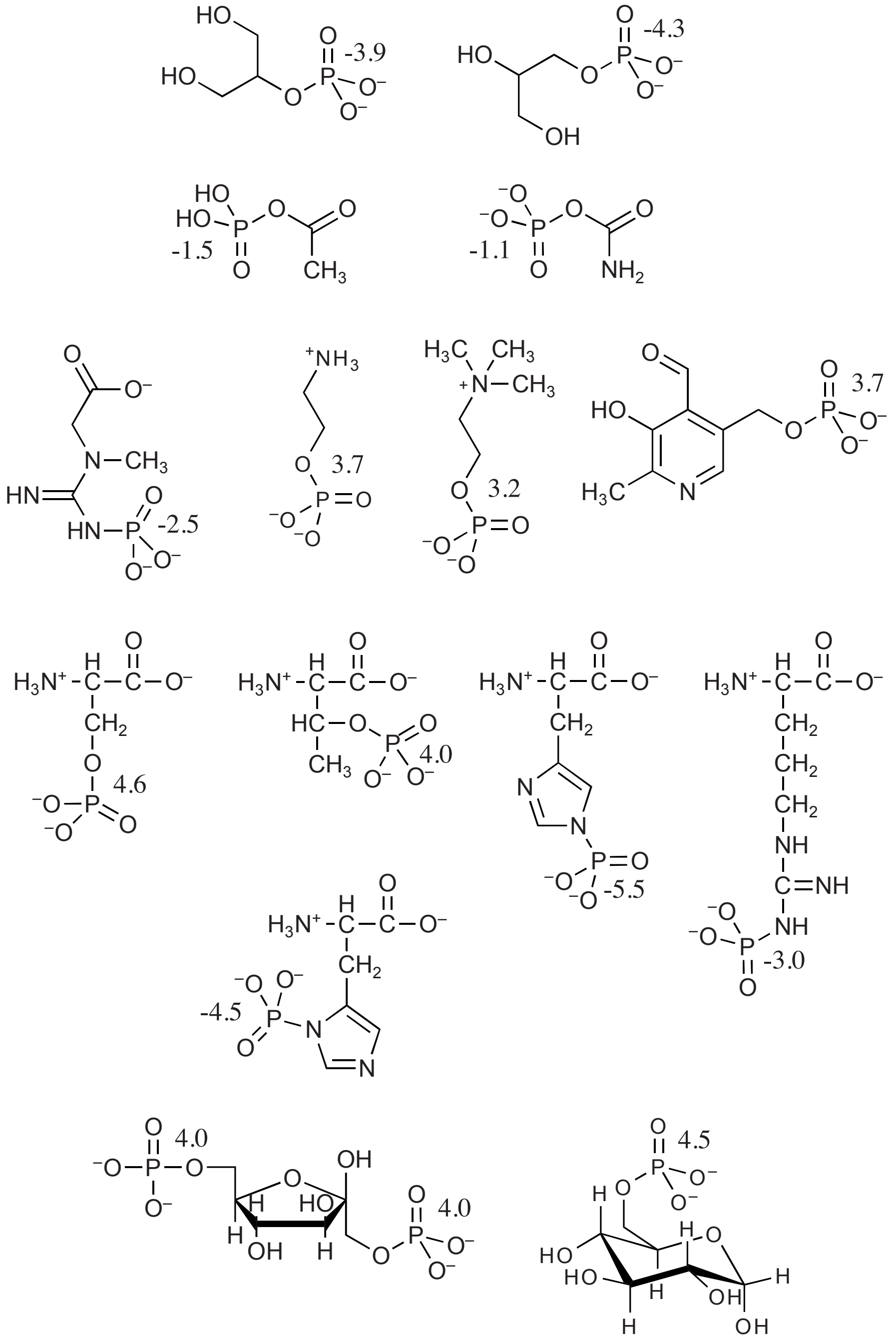


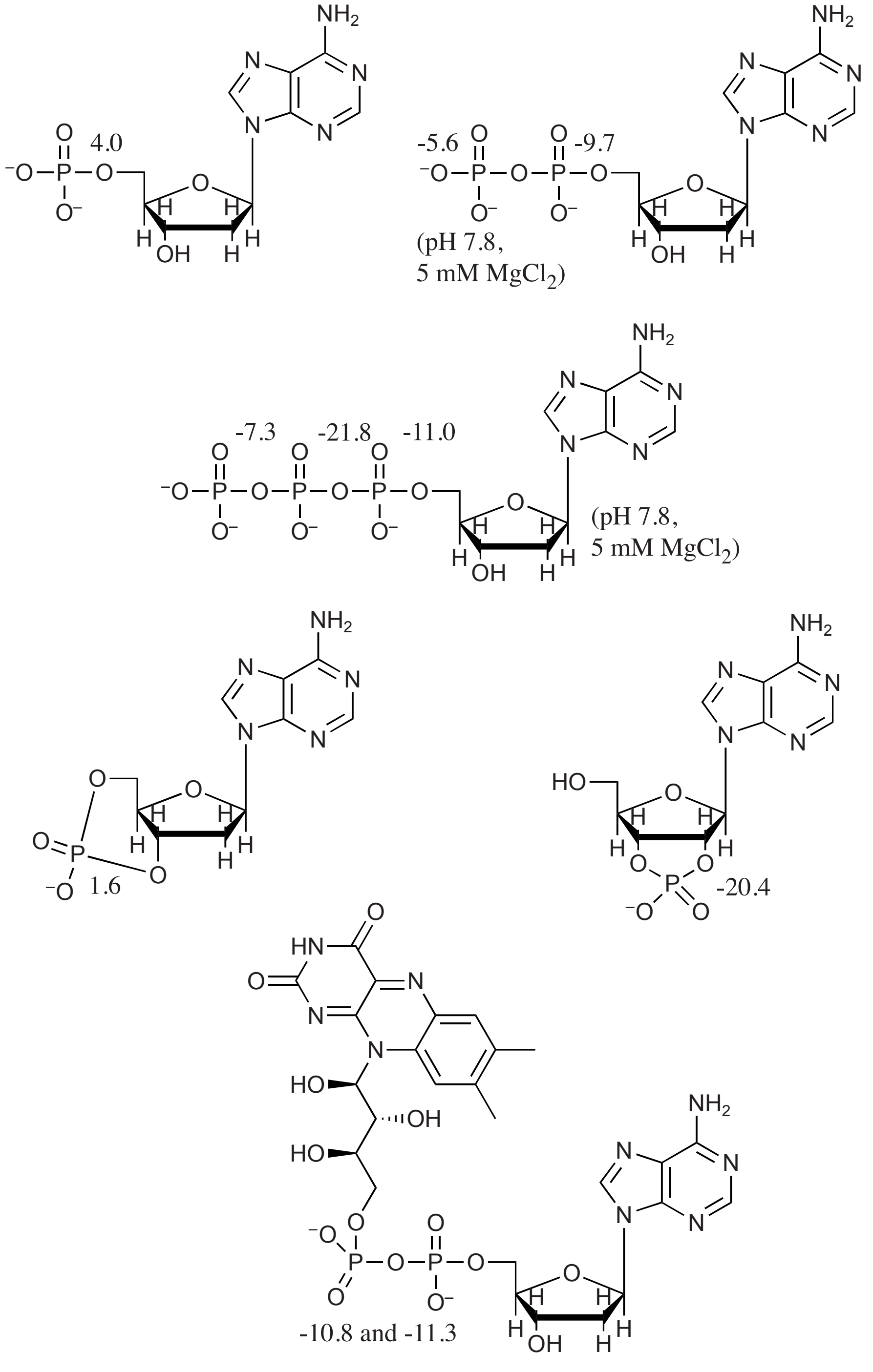






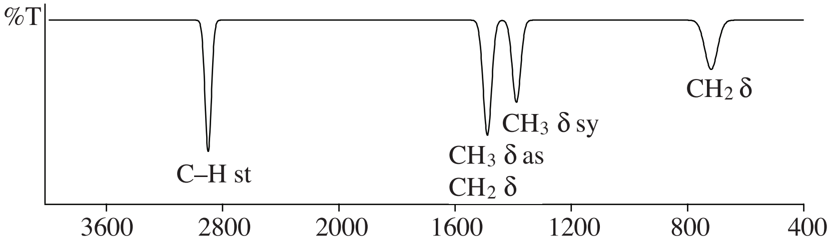
6.2.6 ^{31}P Chemical Shifts of Natural Phosphorus Compounds (δ in ppm relative to H_3PO_4)



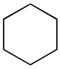

4.0
-5.6
-9.7
(pH 7.8,
5 mM MgCl2)
-7.3
-21.8
-11.0
(pH 7.8,
5 mM MgCl2)
1.6
-20.4
-10.8 and -11.3

7 IR Spectroscopy

7.1 Alkanes



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

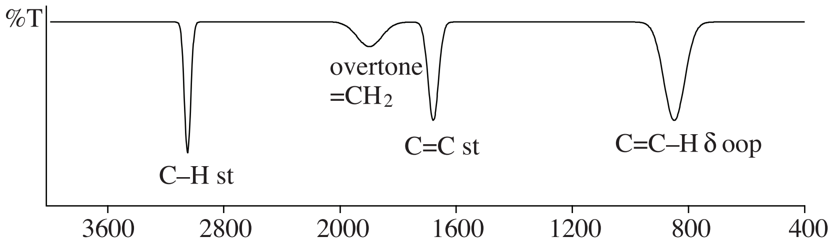
Assignment	Range	Comments
C–H st	3000–2840	Intensity variable, often multiplet
	Beyond normal range:	
	2850–2815	CH_3–O, methyl ethers
	2880–2830	CH_2–O, ethers
	2880–2835, 2780–2750	O–CH_2–O, methylenedioxy
	≈2820	O–CH–O, acetals: weak
	3050–3000	
	2900–2800, 2780–2750	CH=O, aldehydes: Fermi resonance
	2820–2780	CH_3–N, CH_2–N; amines
	3100–3050, 3035–2995	
	2930–2915, 2900–2850	 cyclohexanes: weak, comb at ≈2700
	3080–2900	CH–hal st

Assignment	Range	Comments
CH_3 δ as	1470–1430	Medium, coincides with CH_2 δ
	Beyond normal range:	
	1440–1400	CH_3–C=O, methyl ketones, acetals, CH_3–C=C
CH_2 δ	1475–1450	Medium, coincides with CH_3 δ as
	Beyond normal range:	
	≈1440	CH_2–C=C CH_2–C≡C
	≈1425	CH_2–C=O, CH_2– C≡N, CH_2–X (X: hal, NO_2, S, P)
CH_3 δ sy	1395–1365	Medium. Doublet in compounds with geminal methyl groups:
	≈1385, ≈1370	$CH(CH_3)_2$, of equal intensity (γ: 1175–1140, d)
	≈1385, ≈1365	$C(CH_3)_2$, 1385 weaker than 1365 (γ: 1220–1190, often d)
	≈1390, ≈1365	$C(CH_3)_3$, of equal intensity, sometimes triplet (γ: 1250–1200, d) $N(CH_3)_2$, no doublet Solid-state spectra: sometimes doublet also in the absence of geminal methyl groups
	Beyond normal range:	
	1325–1310	SO_2–CH_3
	1330–1290	S–CH_3, sulfides
	1310–1280	P–CH_3
	1275–1260	Si–CH_3, strong, sharp
CH_3 γ	1250–800	Intensity variable, of no practical significance. Strong band in compounds with geminal methyl groups:
	1175–1140	$CH(CH_3)_2$, doublet
	1220–1190	$C(CH_3)_2$, generally doublet
	1250–1200	$C(CH_3)_3$, doublet, often not resolved
	Beyond normal range:	
	≈765	$SiCH_3$
	≈855, ≈800	$Si(CH_3)_2$
	≈840, ≈765	$Si(CH_3)_3$

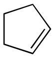
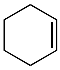
Assignment	Range	Comments
CH_2 γ	770–720	Medium, sometimes doublet
		$C–(CH_2)_n–C$: for $n > 4$ at ≈720; for $n < 4$ at higher wavenumbers; in cyclohexanes at ≈890, weaker
	Beyond normal range:	
	1060–800	Cycloalkanes, numerous bands, unreliable
C–D st	2200–2080	In general, substitution of L by isotope L': $\tilde{\nu}_{X-L'} = \tilde{\nu}_{X-L}\sqrt{\frac{1/m_x + 1/m_{L'}}{1/m_x + 1/m_L}}$

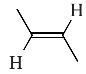
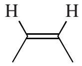
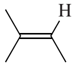
7.2 Alkenes

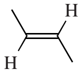
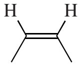
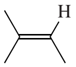
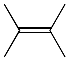
7.2.1 Monoenes



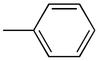
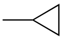
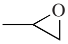
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
=CH_2 st	3095–3075	Medium, often multiple bands
=CH st	3040–3010	Medium, often multiple bands CH st in aromatic hydrocarbons and three-membered rings fall into the same range
	In cyclic compounds:	
	≈3075	
	≈3060	
	≈3045	
	≈3020	
=CH δ ip	1420–1290	Of no practical significance
=CH δ oop	1005–675	A number of bands
	In the same range: ar CH δ oop, C–O–C γ, and C–N–C γ in saturated heterocyclics, OH δ oop in carboxylic acids, NH γ, NO st, SO st, CH_2 γ, CF st, CCl st	

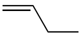
Assignment	Range	Comments		
Subranges:	C=C	C=C–C=O	C=C–OR	C=C–O–C=O
CH=CH_2	1005–985 920–900 (overtone at 1850–1800)	≈980 ≈960 ≈810	≈960 ≈815	≈950 ≈870
C=CH_2	900–880 (overtone at 1850–1780)	≈940 ≈810	≈795	
	990–960	≈975	≈960	≈950
	725–675	≈820		
	840–800	≈820		

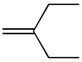
Assignment	Range	Comments
C=C st	1690–1635	Of variable intensity, weak for highly symmetric compounds, strong for N–C=C and O–C=C
	Subranges:	
	1650–1635	CH=CH_2
	1660–1640	C=CH_2
	1690–1665	 Weak
	1665–1635	
	1690–1660	 Weak, often absent
	1690–1650	 Weak, often absent
	Beyond normal range:	
	down to ≈1590	C=C–X with X: O, N, S; of higher intensity; in vinyl ethers often doublet due to rotational isomers

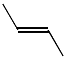
At lower frequency if conjugated with:

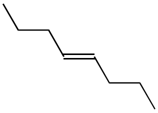
C=C	≈1650 ≈1600		≈1630
C≡C	≈1600		≈1640
C≡N	≈1620		≈1640
C=O	≈1630		

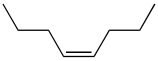
Examples ($\tilde{\nu}$ in cm^{-1})

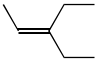

1645
994
912

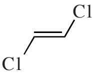

1647
889
669


1682
972
963

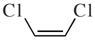

1670
968


1650
709

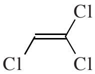

1667
825



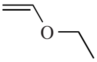
1575
826
761



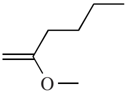
1595
848
714



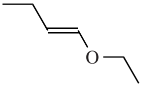
1587
929
835
780



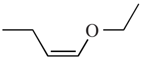
neat:	CCl_4:
1610	1634
	1608
987	964
810	943



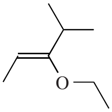
1655
1592
958
793



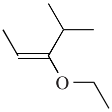
1670
1652
937
925



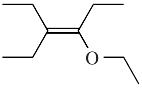
1663



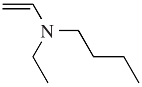
1660



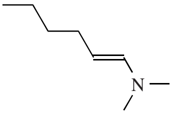
1673



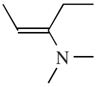
1663



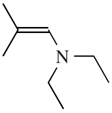
1628



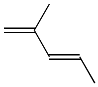
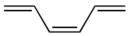
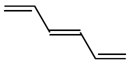
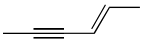
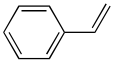
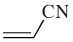
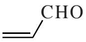
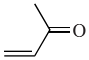
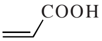
1650



1640

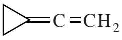


1662

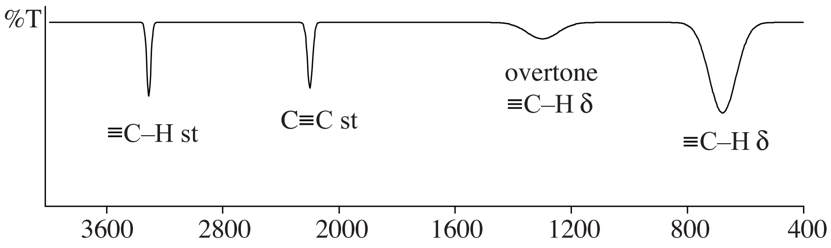
	1652 1612		1830 1621 987 818		1800 1621 941 899
	1607 (2270)		1636	 	1645 1612
 	1618 (1704)	 	1618 (1684)	 	1635 1615 (1730) (1706)
 	1637 (1735)				

7.2.2 Allenes

Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
(C=C)=C–H st	3050–2950	
C=C=C st as	1950–1930	Strong, doublet in X–C=C=CH_2 if X other than alkyl Ring strain increases frequency:  ≈2020
C=C=C st sy	1075–1060	Weak, absent with highly symmetric substitution. In Raman, strong
(C=C)=CH_2 δ oop	≈850	Strong; overtone at ≈1700 (weak)

7.3 Alkynes

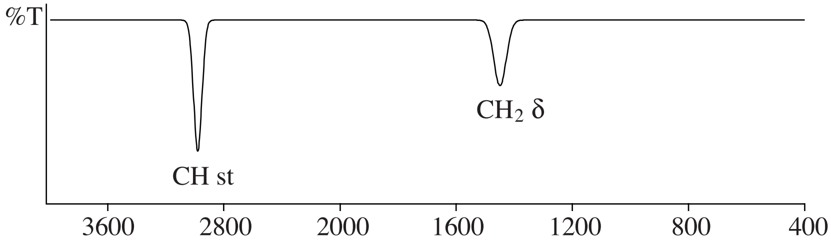


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

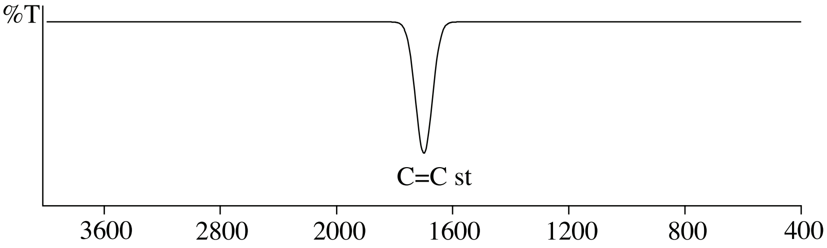
Assignment	Range	Comments
≡C–H st	3340–3250	Strong, sharp; in the same region also OH st, NH st
C≡C st	2260–2100	Weak, sharp. In Raman, strong
	Beyond normal range:	
		R–C≡C–H; at the lower end of the cited range
		R–C≡C–R; usually 2 bands (Fermi resonance), often missing if symmetrical, strong in Raman
	Subranges:	
	≈2120	C–C≡C–H
	≈2220	C–C≡C–C
	≈2240	C–C≡C–CN
	≈2240	C–C≡C–COOH
	≈2240, ≈2140	C–C≡C–$COOCH_3$
	In the same range: C≡Z st, X=Y=Z st, Si–H st	
≡C–H δ	700–600	Strong, broad; overtone at 1370–1220 (broad, weak)

7.4 Alicyclics

Cyclic Alkanes



Cyclic Alkenes



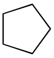
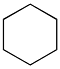
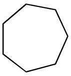
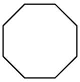
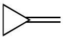
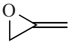
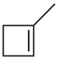
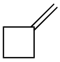
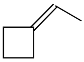
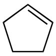
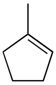
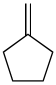
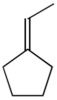
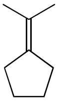
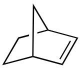
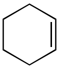
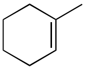
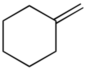
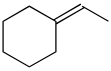
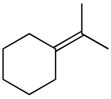
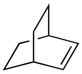
The other vibrations are similar to those in noncyclic alkenes and cyclic alkanes.

Typical Ranges ($\tilde{\nu}$ in cm^{-1})

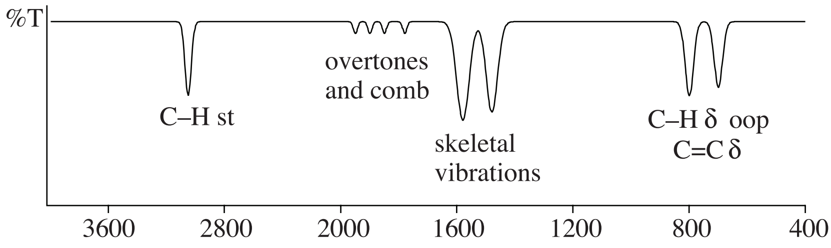
Assignment	Range	Comments
C–H st	3090–2860	Strong
H–C–H δ	1470–1430	Weak
C=C st	1780–1610	Varies with ring size and substitution

Twisting and wagging CH_2 as well as C–C st do not significantly differ from the corresponding vibrations in noncyclic compounds and are of limited diagnostic value.

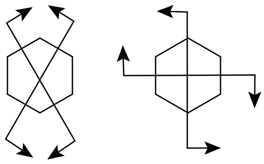
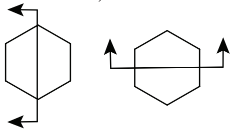
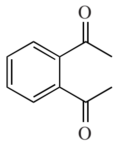
***Examples* (ṽ *in cm*$^{-1}$)**

	3090 3019 2933 1434		2974 2896 1450		2951 2871 1455
	2920 2860 1447		2933 2865 1462		2941 1471 1451
	≈1640		≈1780		≈1650
	≈1570		≈1640		≈1680
	≈1690		≈1610		≈1660
	≈1660		≈1670		≈1690
	≈1570		≈1650		≈1675
	≈1650		≈1665		≈1670
	≈1615				

7.5 Aromatic Hydrocarbons

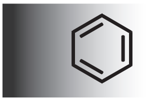


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

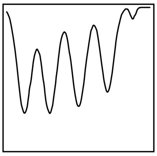
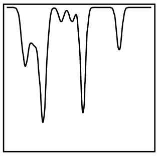
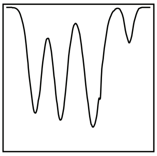
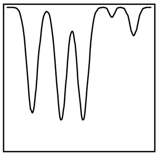
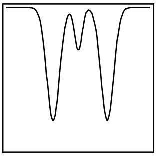
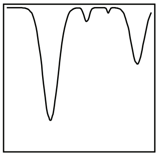
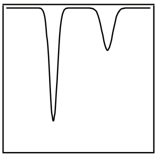
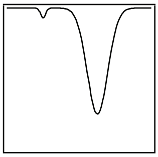
Assignment	Range	Comments
ar C–H st	3080–3030	Often numerous bands; in the same range also CH st of alkenes and small rings
ar C–C	1625–1575	 Medium, often doublet; generally weak in benzene derivatives having a center of symmetry in the ring
		In the same range: C=C st, C=N st, C=O st, N=O st, C–C in heterocyclics, NH δ
	1525–1450	Medium, often doublet:  Weak in: 
		In the same range: C=O st, N=O st, C–C in heterocyclics, B–N st, CH_3 δ, CH_2 δ, NH δ
comb	2000–1650	Very weak; useful for determining substitution patterns in 6-membered aromatic rings
		In the same range: C=O st, B–H···B st, N^+–H st, H_2O δ
ar C–H δ ip	1250–950	Numerous bands of variable intensity; of no practical significance. May be very strong in Raman and, thereby, indicative of substitution type

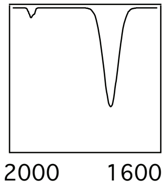
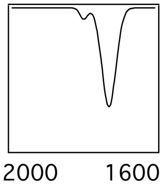
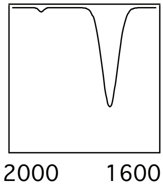
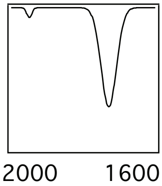
Assignment	Range	Comments
ar C–H δ oop	900–650	One or more strong bands; useful for determining substitution patterns in 6-membered aromatic rings. In Raman, generally weak
	In the same range:	=C–H δ oop, C–O–C γ and C–N–C γ in saturated heterocyclics, OH δ oop in carboxylic acids, NH δ, N–O st, S–O st, CH_2 γ, C–F δ, C–Cl st

Determination of Substitution Patterns in 6-Membered Aromatic Rings: Position and Shape of Bands Related to the Number of Adjacent H Atoms ($\tilde{\nu}$ in cm^{-1})

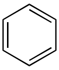
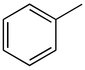
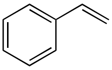
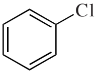
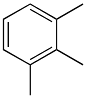
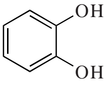


Not to be used for ring systems with strongly conjugated substituents such as C=O, NO_2, C≡N.

Comb, overtones	Substitution type; CH δ oop, ar C–C γ	Comb, overtones	Substitution type; CH δ oop, ar C–C γ
 2000 1600	mono- ≈900 770–730 710–690	 2000 1600	*o*-di- 770–735
 2000 1600	*m*-di- 900–860 865–810 810–750 725–680	 2000 1600	*vic*-tri- 800–770 780–760 720–685
 2000 1600	1,2,4-tri- 900–860 860–800 730–690	 2000 1600	*p*-di- 860–780
 2000 1600	1,2,3,4-tetra- 860–780	 2000 1600	1,3,5-tri- 900–840 850–800 730–675

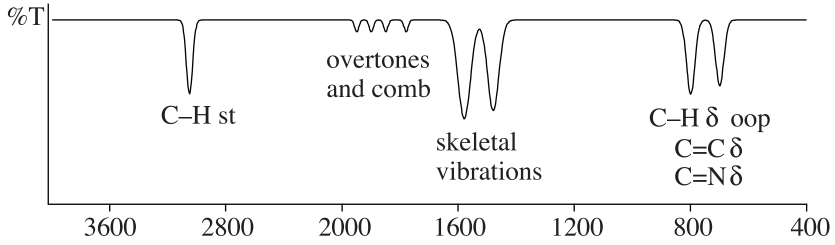
Comb, overtones	Substitution type; CH δ oop, ar C–C γ	Comb, overtones	Substitution type; CH δ oop, ar C–C γ
	1,2,3,5-tetra- 900–840		1,2,4,5-tetra- 900–840
	penta- 900–840		hexa- –

Examples ($\tilde{\nu}$ in cm^{-1})

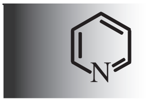
	3080 3040 1968 1818		3021 1945 1862 1808 1739		3086
	3080		3040 1915 1845 1775		1927 1887 1764

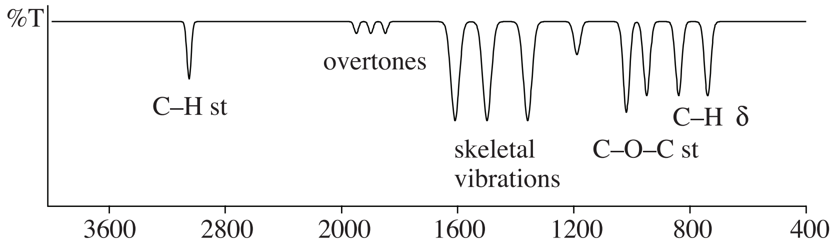
7.6 Heteroaromatic Compounds

Pyridines

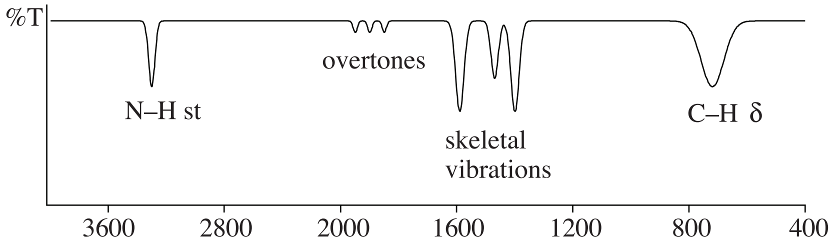


Furans





Pyrroles

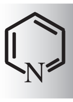


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

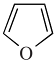
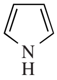
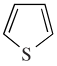
Assignment	Range	Comments
N–H st	3450–3200	Medium, narrow; shifted by formation of hydrogen bonds
Overtones	2100–1800	Weak, characteristic
Ring skeleton	1610–1360	Strong, sharp bands
C–H δ	1000–700	Strong, broad; difficult to identify
C–H st	3100–3000	Medium, sharp
CO–C st	1190–990	Medium or strong; of variable intensity

Pyridines:

The frequencies of pyridines are very similar to those observed in benzenes. The nitrogen atom behaves like a substituted carbon atom in benzenes.

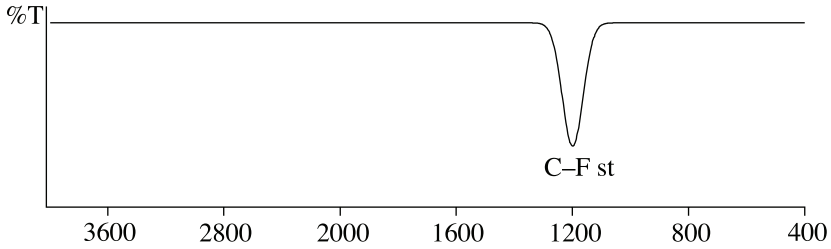


5-Ring Heteroaromatics

			
NH st free		3500–3400	
NH st H-bonded		3400–2800	
CH st	≈3100	≈3100	≈3100
Ring skeleton: intensity variable, generally multiplets	1610–1560 1510–1475	1590–1560 1540–1500	1535–1515 1455–1410
CH δ oop: generally strong	990–725	770–710	935–700

7.7 Halogen Compounds

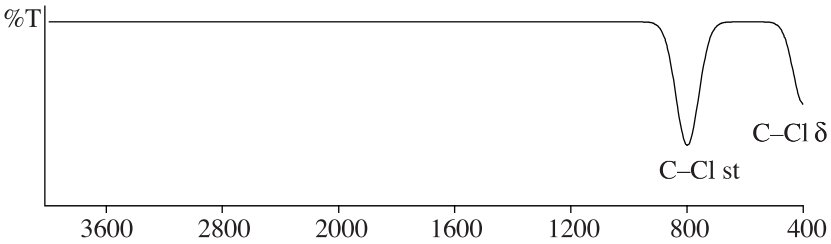
7.7.1 Fluoro Compounds



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C–F st	1400–1000	Strong, often more than one band (rotational isomers), often not resolved. In Raman, weak to medium
	Subranges:	
	1100–1000	al CF_2 (FC–H st: 3080–2990)
	1150–1000	al CF_2
	1350–1100	al CF_3
	1350–1150	C=CF
	≈1745	C=CF_2 st
	1250–1100	ar CF
	In the same range: strong bands for C–O st, NO_2 st sym, C=S st, S=O st	
CF_2	780–680	Medium or weak, assignment uncertain
CF_3	780–680	(C–F δ?)
S–F st	815–755	Strong
P–F st	1110–760	
Si–F st	980–820	
B–F st	1500–800	

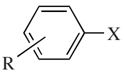
7.7.2 Chloro Compounds



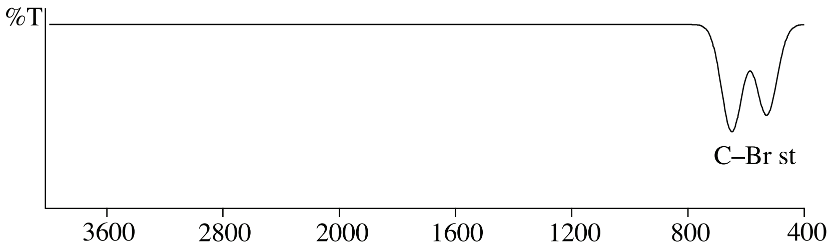
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C–Cl st	830– <600	Strong, often broad (rotational isomers), absent in chloroaromatics
C–Cl δ	400–280	Of medium strength and width
Other	1100–1020	Strong, narrow or of medium width; chloroaromatics
P–Cl st	<600	
Si–Cl st	<625	
B–Cl st	1100–650	

In disubstituted halobenzenes, characteristic skeletal vibrations:

	X	*ortho*	*meta*	*para*
	Cl	1055–1035	1080–1075	1095–1090
	Br	1045–1030	1075–1065	1075–1070
	I			1060–1055

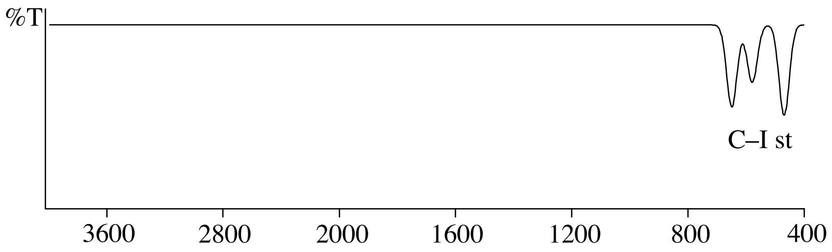
7.7.3 Bromo Compounds



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C–Br st	700–500	Strong, of medium width; absent in bromoaromatics
C–Br δ	350–250	Of medium strength and width
Other	1080–1000	Strong, narrow or of medium width; bromoaromatics

7.7.4 Iodo Compounds



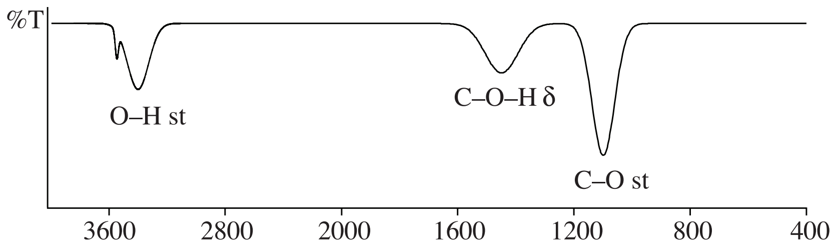
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C–I st	650–450	Strong, two or more bands
C–I δ	300–50	Of medium strength and width

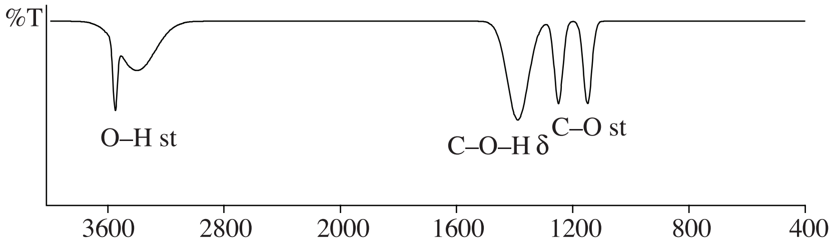
7.8 Alcohols, Ethers, and Related Compounds

7.8.1 Alcohols and Phenols

Alcohols



Phenols

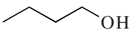


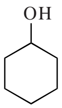
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

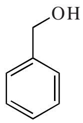
Assignment	Range	Comments
O–H st	3650–3200	Of variable intensity. In Raman, generally weak
	Subranges:	
	3650–3590	Free OH; sharp
	3550–3450	H-bonded OH; broad
	3500–3200	Polymer OH; broad, often numerous bands
	Beyond normal range:	
	3200–2500	Enols, chelates; often very broad
	In the same range: NH st, ≡CH st (≈3300, sharp), H_2O	
O–H δ ip	1450–1200	Medium, of no practical significance

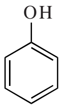
Assignment	Range	Comments
C–O st	1260–970	Strong, often doublet
	Subranges:	
	1075–1000	CH_2–OH
	1125–1000	CH–OH
	1210–1100	C–OH
	1275–1150	ar C–OH
	In the same range: C–F st, C–N st, N–O st, P–O st, C=S st, S=O st, P=O st, Si–O st, Si–H δ	
O–H δ oop	<700	Medium, of no practical significance

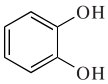
Examples ($\tilde{\nu}$ in cm^{-1})

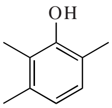

3250
1430
1075
1050


3335
1350

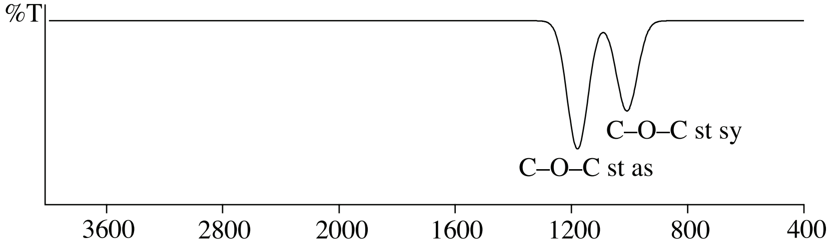

3290
1430
1020


3215
1368
1220


3450
1370
1260
1195


3460
1315
1237
1210

7.8.2 Ethers, Acetals, and Ketals

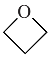
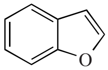
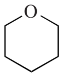
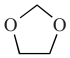
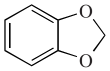
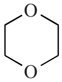
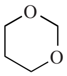


In acetals and ketals, the C–O stretching vibrations are split into 3, sometimes even 4 to 5 bands.

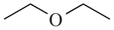
Acetals have an additional band due to a special C–H δ vibration.

The C–H st vibration frequency is especially low for OCH_3 st (2850–2815) and OCH_2 st (2880–2835).

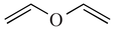
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C–O–C st as	1310–1000	Strong, sometimes split
	Subranges for noncyclic ethers:	
	1150–1085	CH_2–O–CH_2
	1170–1115	CH–O–CH, often split
	1225–1180	C=C–O–al C
	1275–1200	ar C–O–al C
	Subranges for cyclic ethers:	
	1280 sy 870 as	
	≈1030 sy ≈980 as	
	≈1070 sy ≈915 as	
	≈1235	
	≈1100 as ≈815 sy	
	≈950	 ketals, acetals: 4 to 5 bands
	≈925	
	1024, 1086 as ≈880 sy	
	≈800	 in acetals: C–H st, ≈2820, weak
C–O–C st sy	1055–870	Strong, sometimes multiple bands
	Subranges for noncyclic ethers:	
	1125–1080	C=C–O–al C, medium
	1075–1020	ar C–O–al C, medium
	In the same range: strong bands for C–O st, C–F st, C–N st, N–O st, P–O st, C=S st, S=O st, P=O st, Si–O st, Si–H δ	

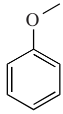
Examples ($\tilde{\nu}$ in cm^{-1})



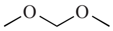
1136
935
917



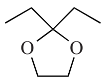
1225
1218
1211
1003



1250
1040

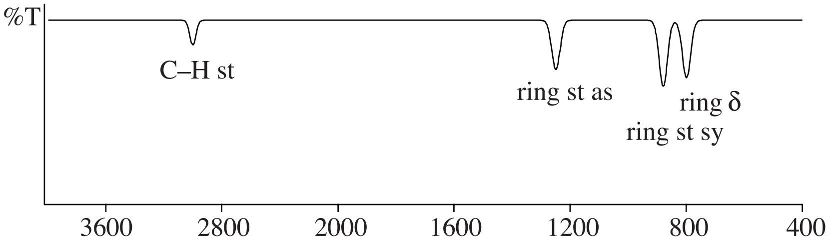


1188
1138
1111
1046



1172
1132
1077
1057
1038

7.8.3 Epoxides

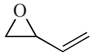


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

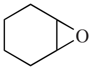
Assignment	Range	Comments
C–H st	3050–2990	Frequency higher than normally found in alkanes
ring st as	1280–1230	Variable intensity
ring st sy	950–815	Variable intensity
ring δ	880–750	Variable intensity

Examples ($\tilde{\nu}$ in cm^{-1})

1280
870

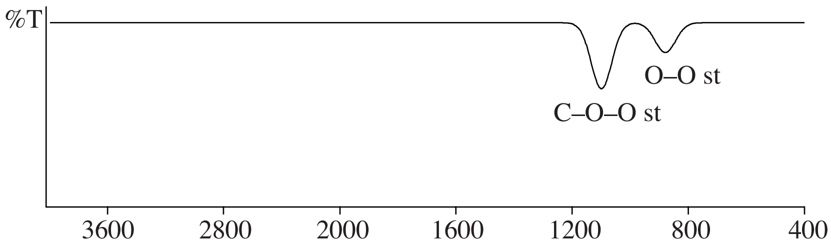


1230 sy
885 as
845 δ



1260 sy
890 as
780 δ

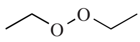
7.8.4 Peroxides and Hydroperoxides

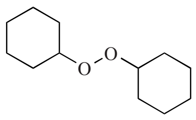


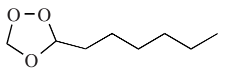
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
O–O–H st	3450–3200	Of variable intensity
	Subranges:	
	≈3450	Free OOH; H-bonded: ≈30 cm^{-1} higher than in corresponding alcohols
	In the same range: OH st, NH st, ≡CH st, H_2O	
C–O–O st	1200–1000	Strong, ≈20 cm^{-1} lower than in corresponding alcohols
	In the same range: strong bands for C–O st, C–F st, C–N st, N–O st, P–O st, C=S st, S=O st, P=O st, Si–O st, Si–H δ	
O–O st	1000–800	Medium or weak, often doublet, assignment uncertain
Also:	1760–1745	C=O st in peracids
	1820–1770	C=O st in diacylperoxides (two bands)

Examples ($\tilde{\nu}$ in cm^{-1})


1017
880

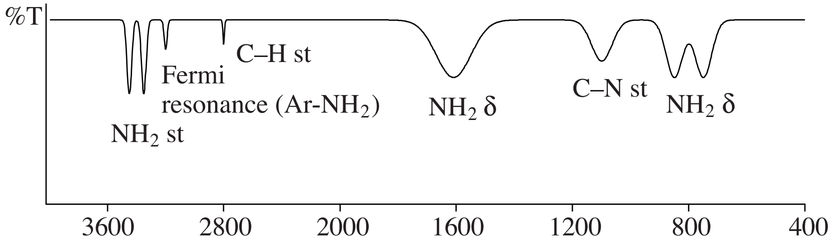

1070
1060
943


1100
852

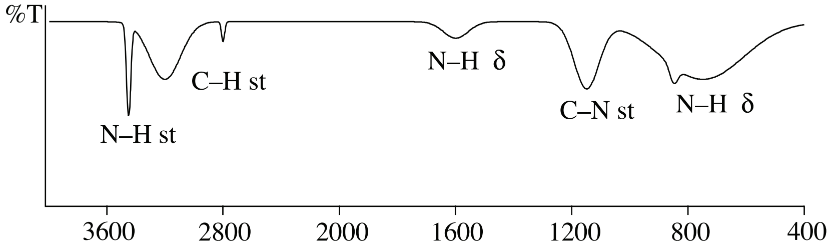
7.9 Nitrogen Compounds

7.9.1 Amines and Related Compounds

Primary Amines

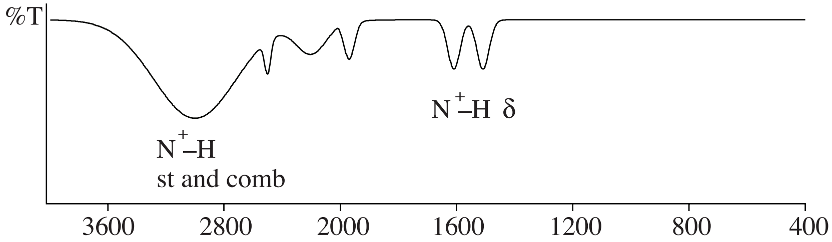


Secondary Amines



N

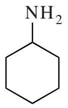
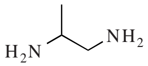
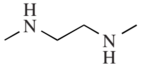
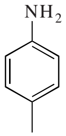
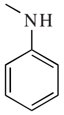
Ammonium



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

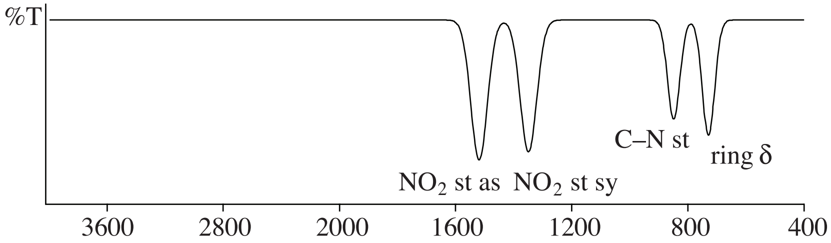
Assignment	Range	Comments
NH_2 st	3500–3300	Of variable intensity, generally 2 sharp bands, $\Delta\tilde{\nu}$ = 65–75
		At lower wavenumbers (<3200) and broader if H-bonded. Free and H-bonded forms often simultaneously observed
		In primary aromatic amines, additional combination band at ≈3200
	In the same range: OH st, ≡CH st	
NH st	3450–3300	Of variable intensity, only one band
		At lower wavenumbers (<3200) and broader if H-bonded. Free and H-bonded forms often simultaneously observed
	In the same range: OH st, ≡CH st, H_2O	
NH_3^+ st	3000–2000	Medium, broad, highly structured
	3000–2700	Major maximum, comb: ≈2000
NH_2^+ st	3000–2000	Medium, broad, highly structured
	3000–2700	Major maximum
NH^+ st	3000–2000	Medium, broad, highly structured
	2700–2250	Major maximum
	In the same range: OH st, NH st, CH st, SH st, PH st, SiH st, BH st, X=Y=Z st, X≡Y st	
NH_2 δ	1650–1590	Medium or weak
NH δ	1650–1550	Weak
NH_3^+ δ	1600–1460	Medium, often more than one band; weak in aliphatic amines
NH_2^+ δ	1600–1460	Medium, often more than one band; weak in aliphatic amines
NH^+ δ	1600–1460	Medium, often more than one band; weak in aliphatic amines
C–N st	1400–1000	Medium, of no practical significance
NH_2 δ	850–700	Medium or weak; 2 bands in primary amines
NH δ	850–700	Medium or weak
P–N–C st	1110–930	
	770–680	

Examples ($\tilde{\nu}$ in cm^{-1})

CH_3-NH_2	3470 3360 1622		3357 3278 3200 sh 1605		3356 3274 3175 1650
	3279		3487 3405		3416 3386 1322 1266

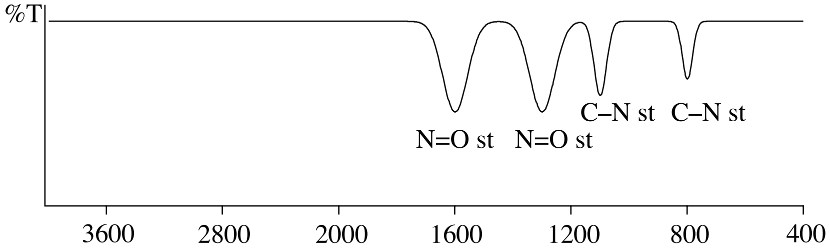
7.9.2 Nitro and Nitroso Compounds

Nitro Compounds



N

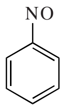
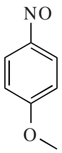
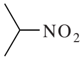
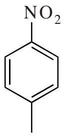
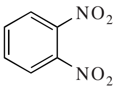
Nitroso Compounds



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

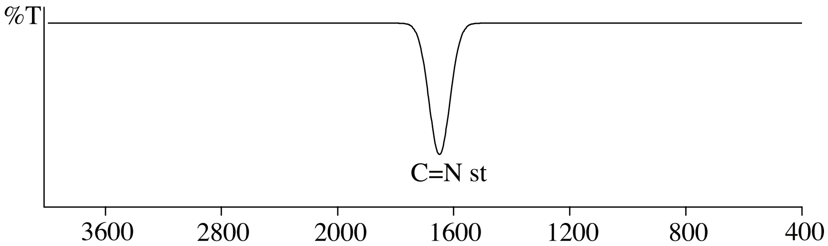
Assignment	Range	Comments
NO_2 st as	1660–1490	Very strong, of medium width. In Raman, of weak to medium intensity
	Subranges:	
	1660–1625	O–NO_2, nitrates; missing in Raman
	1570–1540	C–NO_2, aliphatic nitro compounds
	1560–1490	C–NO_2, aromatic nitro compounds
	1630–1530	N–NO_2, nitramines
NO_2 st sy	1390–1260	Strong, of medium width
	Subranges:	
	1285–1270	O–NO_2, nitrates
	1390–1340	C–NO_2, aliphatic nitro compounds
	1360–1310	C–NO_2, aromatic nitro compounds; often 2 bands
	1315–1260	N–NO_2, nitramines
	In nitrates also:	
	≈870	N–O st, strong
	≈760	NO_2 γ
	≈700	NO_2 δ
Ring δ	760–705	Strong; modified deformation of aromatic ring
N=O st	1680–1450	Very strong, in monomers
	1420–1250	Very strong, in dimers
	Subranges:	
	1680–1650	O–NO (nitrites) *trans*; 1625–1610: *cis*
	1585–1540	C–NO, aliphatic *C*-nitroso compounds
	1510–1490	C–NO, aromatic *C*-nitroso compounds
	≈1450	N–NO, *N*-nitroso compounds
	In nitrites also:	
	3300–3200, ≈2500, 2300–2250	comb
	≈800	N–O st *trans*; *cis*: very weak
	≈600	O–NO δ *trans*; *cis*: ≈650
C–N st	≈850	C–NO, aliphatic *C*-nitroso compounds; coupled with other vibrations
	≈1100	C–NO, aromatic *C*-nitroso compounds
N–N st	≈1040	*N*-Nitroso compounds

Examples (ṽ in cm⁻¹)

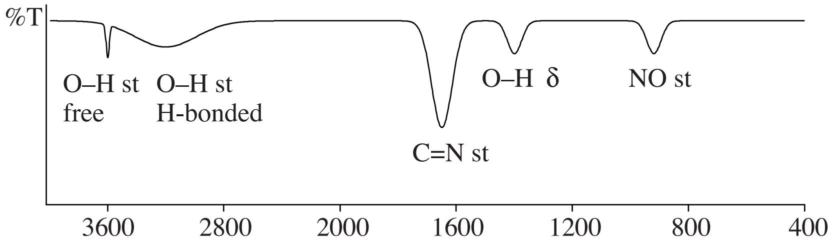
Compound	ṽ	Compound	ṽ	Compound	ṽ
CH_3-NO	1564 842		1506 1110 810		1497 1112 858
	1524 1359 851		1527 1351 853 720		1506 1351 1261 873 748

7.9.3 Imines and Oximes

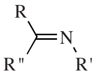
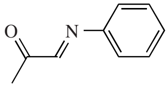
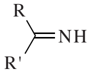
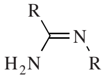
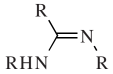
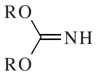
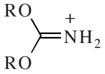
Imines



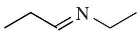
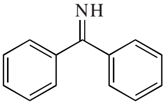
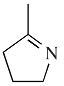
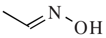
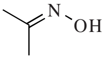
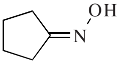
Oximes



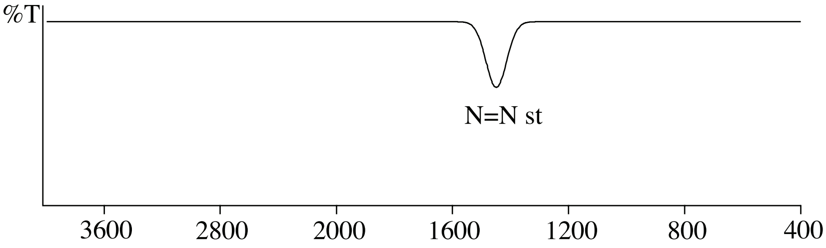
Typical Ranges (ṽ in cm⁻¹)

Assignment	Range	Comments
C=N st	1690–1520	Generally strong. In Raman, generally strong
	Subranges:	
	≈1670	R–CH=N–R' R, R': al
	≈1645	R–CH=N–R' R or R': conjugated
	≈1630	R–CH=N–R' R, R': conjugated
	≈1655	 R, R', R'': al
	≈1645	R: conjugated
	≈1635	R, R': conjugated
	≈1555	 Additional band: ≈1655 C=O st
	≈1645	 R, R': al
	≈1625	R, R': conjugated
	1685–1580	 Additional band at 1540–1515 in: 
	1670–1600	CH=N–N=CH
	1690–1645	 Additional bands: NH st: ≈3300, C–O st: ≈1325, ≈1100
	1680–1635	 Additional bands: NH_2^+ st: ≈3000 NH_2^+ δ: 1590–1540
	2050–2000	C=C=N; ketimines, very strong, sometimes doublet
	1580–1520	Quinone oximes: C=O st 1680–1620
	1685–1650	Aliphatic oximes
	1650–1615	Aromatic oximes
	1690–1645	O–C=N
	1640–1605	S–C=N
	1640–1580	S–S–C=N
OH st	3600–2700	Strong
	Subranges:	
	≈3600	Free
	3300–3100	H-bonded, broad
	≥ ≈2700	Quinone oximes, more than one band
OH δ	1475–1315	Of no practical significance
N–O st	1050–400	Of no practical significance

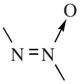
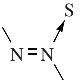
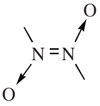
Examples (ṽ in cm⁻¹)

	1667		1603		1637
	1675		1672 (solid) 1662 (gas)		1684

7.9.4 Azo, Azoxy, and Azothio Compounds

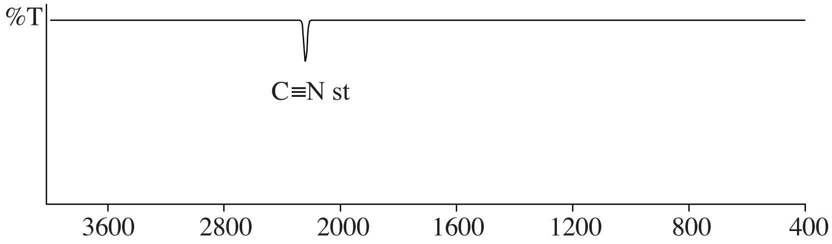


Typical Ranges (ṽ in cm⁻¹)

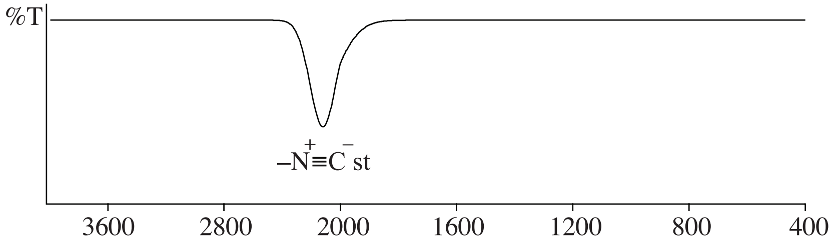
Assignment	Range	Comments
N=N st	1580–1400	Very weak, missing in compounds of high symmetry. In Raman, generally strong
	1480–1450 1335–1315	 st as (mainly N=N st) st sy (mainly N–O st)
	≈1450 ≈1060	 st as (mainly N=N st) st sy (mainly N–S st)
	1410–1175	 Dimers of *C*-nitroso compounds
	Subranges:	
	1290–1175	Aliphatic *trans*
	1425–1385, 1345–1320	Aliphatic *cis*
	1300–1250	Aromatic *trans*
	≈1410, ≈1395	Aromatic *cis*

7.9.5 Nitriles and Isonitriles

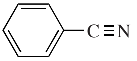
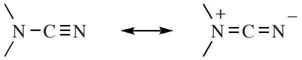
Nitriles



Isonitriles

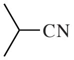
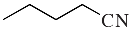
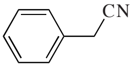
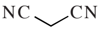
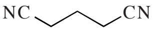
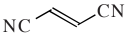
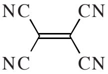
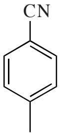
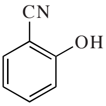


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C≡N st	2260–2240	Medium to strong, sharp; for O–CH_2–C≡N, N–CH_2–C≡N: of low intensity or absent. In Raman, of medium to high intensity
	Beyond normal range:	
	2240–2215	C=C–C≡N
	2240–2215	
	2240–2230	XC–C≡N, X: Cl, Br, I
	≈2275	–CF_2–C≡N
	2225–2175	
	2210–2185	>N–C=C–C≡N
	2200–2070	$C≡N^-$
$-N^+≡C^-$	2150–2110	Strong

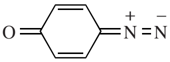
N

Examples ($\tilde{\nu}$ in cm^{-1})

	2222		2235		2252
	2273		2235		2252
	2257 2222		2245		2220
NaCN, KCN	2080–2070	AgCN	2178	NH_2–CN	2268

7.9.6 Diazo Compounds

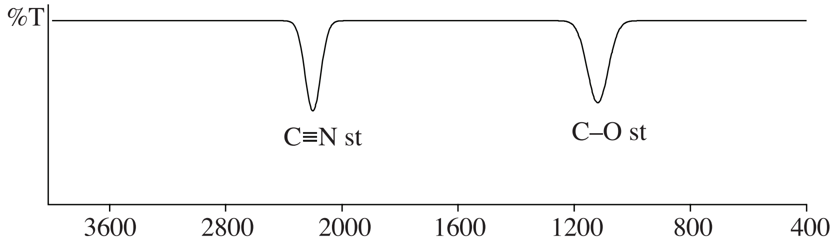
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments	
–N⁺≡N st	2310–2130	Medium, frequency depends on anion	
	In the same range: C≡C st, X=Y=Z st as, NH^+ st, PH st, POH st, SiH st, BH st		
$C=N^+=N^-$	2050–2010	Very strong	
	Subranges:		
	2050–2035	R–CH=N^+=N^-, R: al or ar	
	2035–2010	R_2–C=N^+=N^-, R: al or ar	
	Beyond normal range:		
	2100–2050	R–CO–C=N^+=N^-	C=O st ≈1645 (R: al) C=O st ≈1615 (R: ar) $C=N^+=N^-$ st sy: ≈1350, strong
	2180–2010		C=O st 1655–1560

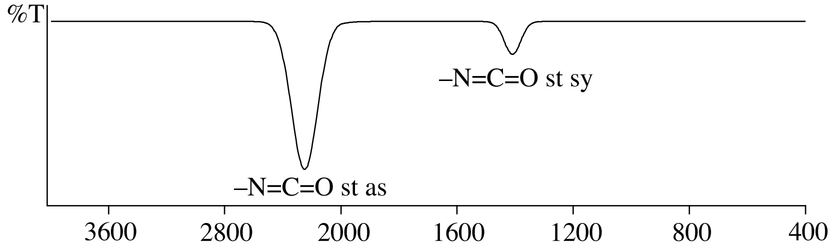
N

7.9.7 Cyanates and Isocyanates

Cyanates



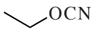
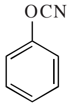
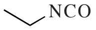
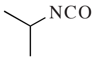
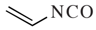
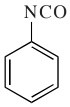
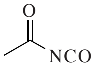
Isocyanates



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

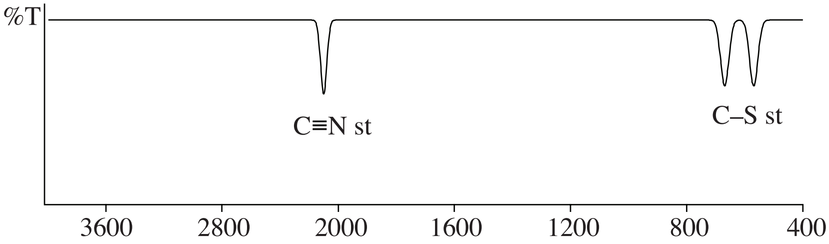
Assignment	Range	Comments
OC≡N st	2260–2130	Medium to strong
	2220–2130	$(OC{\equiv}N)^-$ st as
	1335–1290	$(OC{\equiv}N)^-$ st sy
C–O st	1200–1080	Strong
N=C=O st as	2280–2230	Strong, sharp. In Raman, weak or absent
	≈2300	$-CF_2-NCO$
N=C=O st sy	1450–1380	Weak
	Beyond normal range:	
	2220–2130	$(N{=}C{=}O)^-$

Examples ($\tilde{\nu}$ in cm^{-1})

CH_3–OCN	2248		2248 2282		2235 2261 2282
CH_3–NCO	2265		2280		2270
	2256 (1629 C=C)		2267		2246

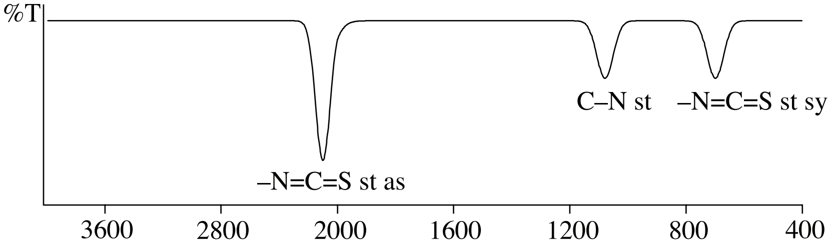
7.9.8 Thiocyanates and Isothiocyanates

Thiocyanates



N

Isothiocyanates

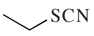


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
SC≡N st	2170–2130	Medium, sharp
	2090–2020	$(SC{\equiv}N)^-$
C–S st	750-550	Often doublet
N=C=S st as	2200–2050	Very strong, generally doublet, Fermi resonance
N=C=S st sy	950–650	
	≈950	aliphatic –N=C=S
	700–650	aromatic –N=C=S
	Beyond normal range:	
	2090–2020	$(N{=}C{=}S)^-$
C–N st	1090–1075	

Examples ($\tilde{\nu}$ in cm^{-1})

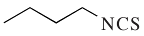
CH_3–SCN 2157



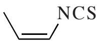
2158

2170

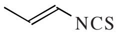
CH_3–NCS neat: 2206, 2114; in CCl_4: 2221, 2106, 2077



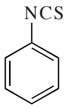
2173
2097
2068



2105



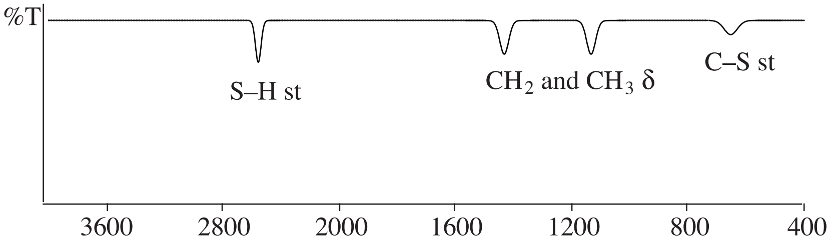
2062



neat: 2090
in CCl_4: 2065
in $CHCl_3$: 2112

7.10 Sulfur Compounds

7.10.1 Thiols and Sulfides

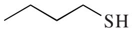


Typical Ranges (ṽ in cm⁻¹)

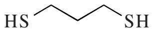
Assignment	Range	Comments
S–H st	2600–2540	Often weak, narrow. In Raman, strong
S–H δ	915–800	Weak, of no practical significance
C–S st	710–570	Weak, broad, of no practical significance. In Raman, strong
S–S st	≈500	Weak, of no practical significance
Also:	≈2880	(S–)CH_3 st as
	≈2860	(S–)CH_2 st as
	≈1430	(S–)CH_3 δ as
	1330–1290	(S–)CH_3 δ sy
	≈1425	(S–)CH_2 δ
	815–755	S–F st, strong
	≈630	S–N st in S–N=O
	725–550	S–C in S– C≡N, often doublet

S

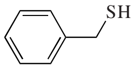
Examples (ṽ in cm⁻¹)



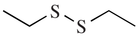
2950



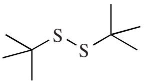
2525



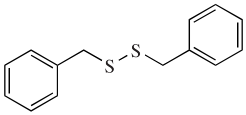
2585



698
668
641



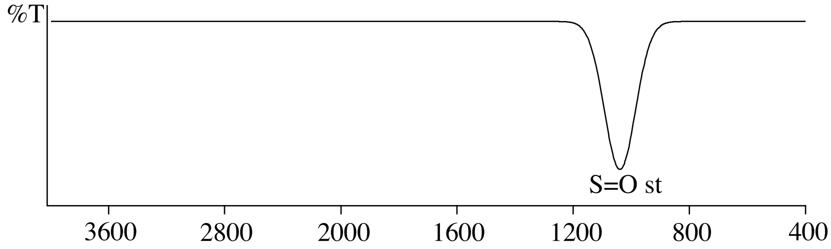
566



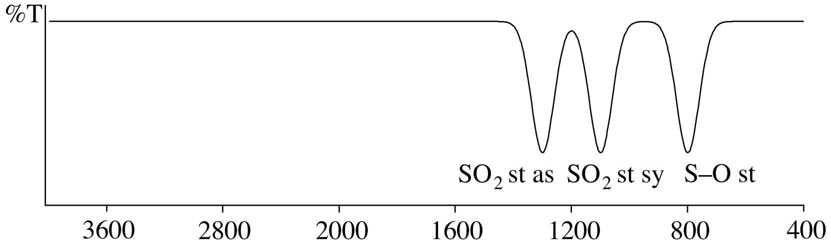
662

7.10.2 Sulfoxides and Sulfones

Sulfoxides

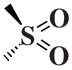


Sulfones

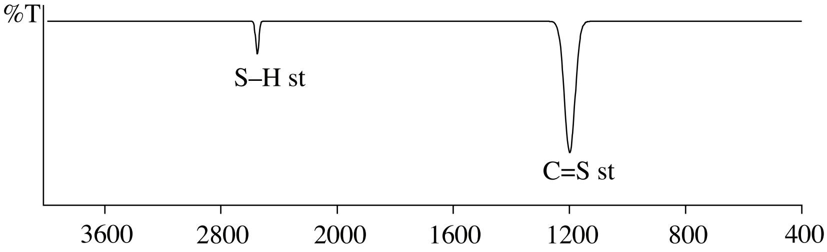


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments	
S=O st	1225–980	Strong, sometimes multiple bands. In Raman, weak to medium	
	Subranges:		
	1060–1015	R–SO–R	
	≈1100	R–SO–OH	S–O st 870–810 OH st free ≈3700, H-bonded ≈2900, ≈2500
	≈1135	R–SO–OR	S–O st 740–720, 710–690
	1225–1195	RO–SO–OR	
	≈1135	R–SO–Cl	
	≈1030, ≈980	R–SO_2^-	
	≈1100, ≈1050	R=SO	N=SO: ≈1250, ≈1135

Assignment	Range	Comments	
 **st as**	1420–1300	Very strong; in Raman, often missing	
st sy	1200–1000	Very strong; in Raman, strong	
	Subranges:		
	1370–1290, 1170–1110	$R–SO_2–R$	
	1375–1350, 1185–1165	$R–SO_2–OR$	
	≈1340, ≈1150	$R–SO_2–SR$	
	1415–1390, 1200–1185	$RO–SO_2–OR$	
	1365–1315, 1180–1150	$R–SO_2–N$	N–H st: 3330–3250; N–H δ: ≈1570; S–N st: 910–900
	1410–1375, 1205–1170	$R–SO_2$–hal	
	1355–1340, 1165–1150	$R–SO_2–OH$	O–H st, H-bonded: ≈2900, ≈2400 hydrated: 2800–1650, broad
	1250–1140, 1070–1030	$R–SO_3^-$	
	1315–1220, 1140–1050	$RO–SO_3^-$	
S–O st	870–690	Of variable intensity, weak in sulfites	

7.10.3 Thiocarbonyl Derivatives

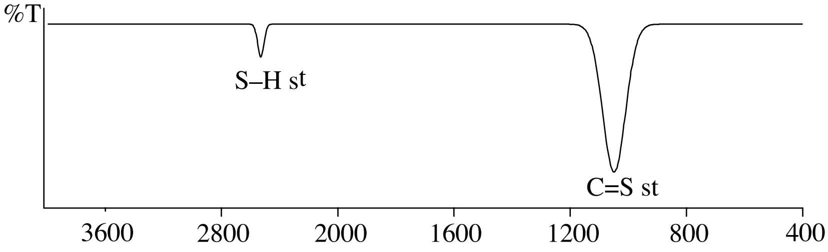


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

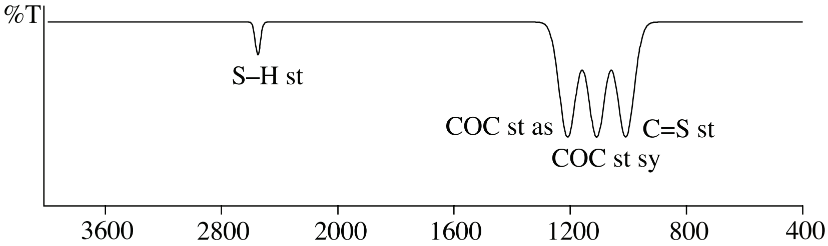
Assignment	Range	Comments	
C=S st	1275–1030	Strong, narrow. In Raman, strong	
	Subranges:		
	1075–1030	Thioketones	
	1210–1080	Thioesters	
	≈1215	Dithiocarboxylic acids	SH st: ≈2550 SH δ: ≈860
	1125–1075	Thiocarboxylic acid fluoride	perfluorinated: 1130–1105
	1100–1065	Thiocarboxylic acid chloride	perchlorinated: 1100–1075
	1140–1090	Thioamides and thiolactams	C–N st: 1535–1520 NH δ: 1380–1300
Also:	750–580	P=S st	

7.10.4 Thiocarbonic Acid Derivatives

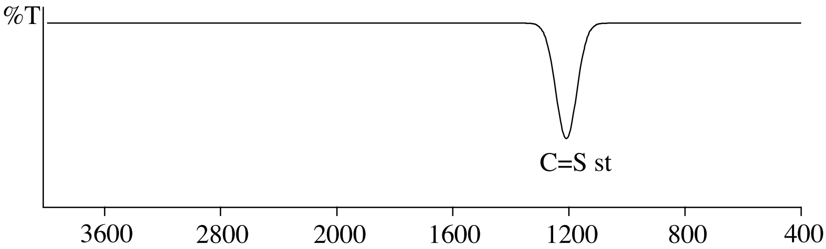
Trithiocarbonates



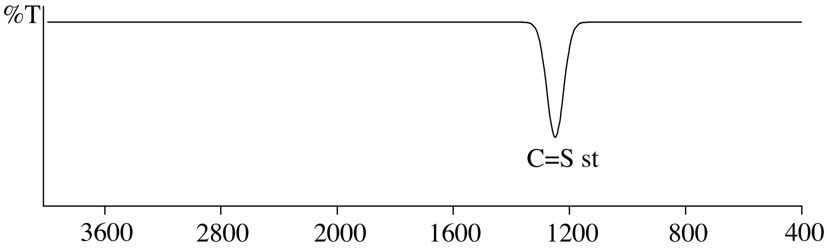
Xanthates



Thiocarbonates



Thioureas

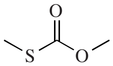
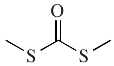
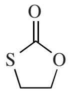
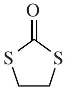
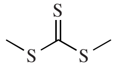
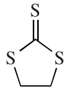
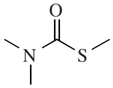
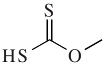
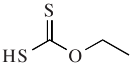
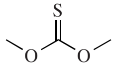
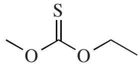
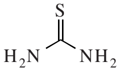
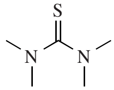
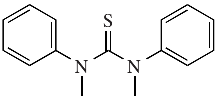


S

Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments	
S–H st	2560–2510	Weak, narrow	trithiocarbonates
	2600–2500	Weak, narrow	xanthates
C=S st	1100–1020	Very strong	trithiocarbonates
	1070–1000	Strong	xanthates
	1250–1180	Strong	thiocarbonates
	1400–1100	Strong	thioureas
COC st as	1260–1140	Strong	xanthates
COC st sy	1150–1090	Strong to medium	xanthates

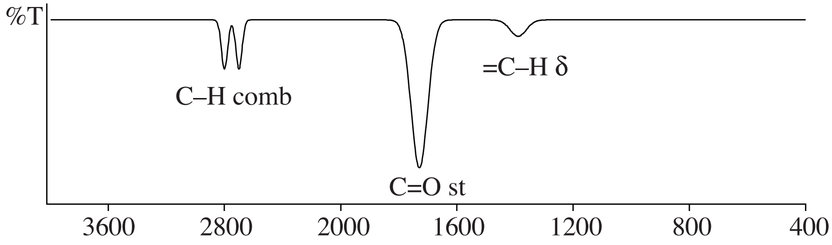
Examples ($\tilde{\nu}$ in cm^{-1})

	in CCl_4: 1719		in CCl_4: 1653		in CCl_4: 1757
	in CCl_4: 1718 1677 1640		neat: 1076		solid: 1058 in CCl_4: 1083 1079
	in CCl_4: 1662				gas: 2593 2548 neat: 2470
	in CS_2: 2562 2522		solid: 1212		solid: 1234
	solid: 1400		solid: 1130		solid: 1131

S

7.11 Carbonyl Compounds

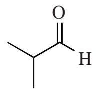
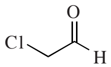
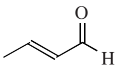
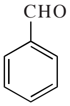
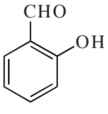
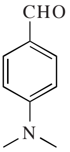
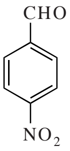
7.11.1 Aldehydes



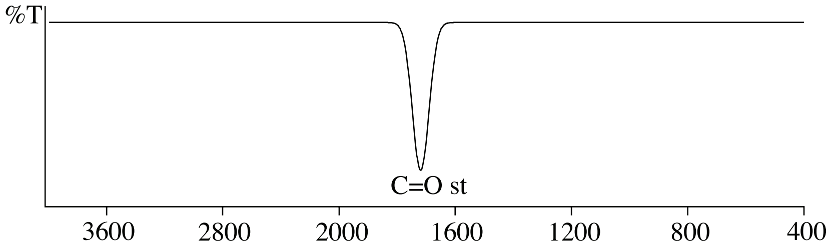
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C–H comb	2900–2800	Weak, Fermi resonance with C–H δ at ≈1390
	2780–2680	(for extreme position of C–H δ only one band)
	Subranges:	
	2830–2810, 2720–2690	Aliphatic
	2830–2810, 2750–2720	Aromatic, with *o*-substitution often higher
	In the same range: cyclohexanes at ≈2700, weak	
C=O st	1765–1645	Strong; in Raman, weak to medium
	Subranges:	
	1740–1720	Aliphatic
	1765–1730	α-Halogenated aliphatic
	1710–1685	Aromatic
	1695–1660	α,β-Unsaturated aromatic
	1670–1645	With intramolecular H bonds
C–H δ	1390	Weak, of no practical significance

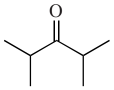
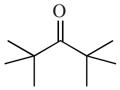
Examples ($\tilde{\nu}$ in cm^{-1})

CH_3–CHO	1748		1725		1742
CCl_3–CHO	1760		1687		1700
	1670		1696		in CCl_4: 1717 in $CHCl_3$: 1710

7.11.2 Ketones

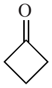
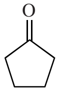
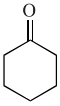
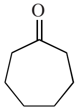
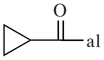
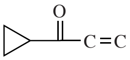
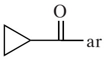


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C=O st	1775–1650	Strong; in Raman, weak to medium
	Subranges:	
	≈1715	Aliphatic, branching at α position causes shift to lower wavenumbers:  ≈1695  ≈1685
	≈1775–1705	Cyclic, $\tilde{\nu}$ decreases with increasing ring size

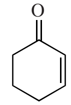
[contd.]

C = X

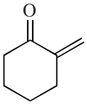
Assignment	Range	Comments
		 ≈1775  ≈1750
		 ≈1715  ≈1705
Conjugated:	≈1675	α,β-Unsaturated, often 2 bands (rotational isomers)
	1650–1600	C=C st
	≈1695	
	≈1665	α,β,γ,δ-Diunsaturated; α,β;α',β'-diunsaturated
	≈1670	
	≈1690	Aryl ketones
	≈1675	
	≈1665	Diaryl ketones, with N or O in *p*-position: down to ≈1600
α-Halogenated ketones:	Shifted toward higher wavenumbers depending on dihedral angle ϕ between C=O and C-hal; largest effect for ϕ = 0°, no effect for ϕ = 90°	
	Maximal shifts:	α-chloro ≈25; α-bromo ≈20
		α,α-dichloro ≈45; α-iodo ≈0
		α,α'-dichloro ≈45; α,α-difluoro ≈60
		perfluoro ≈90
α-Diketones:	≈1720	Aliphatic
	≈1775, ≈1760	Aliphatic 5-ring
	≈1760, ≈1730	Aliphatic 6-ring
	≈1675	Aliphatic enolized, C=C st: ≈1650
	≈1680	Aromatic
	≈1675	*o*-Quinones, with *peri*-OH: ≈1675, ≈1630
β-Diketones:	≈1720	Keto form, sometimes doublet
	≈1650	Enol form

Assignment	Range	Comments
	≈1615	Enol with intramolecular H bonds, C=C st: ≈1600, strong
γ-Diketones:		As monoketones
	≈1675	*p*-Quinones; with *peri*-OH: ≈1675, ≈1630 C=C st: ≈1600
C=C=O st as	2155–2130	Very strong

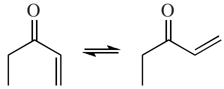
Examples ($\tilde{\nu}$ in cm⁻¹)



1691

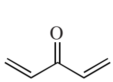


1697

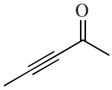


s-*trans* s-*cis*

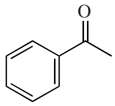
s-*trans* 1690
s-*cis* 1707



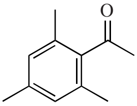
1672
1660



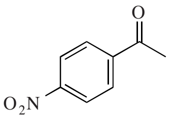
1678
(2222)



1692



1701



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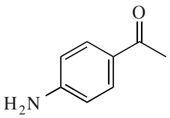

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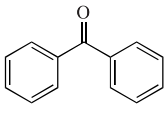

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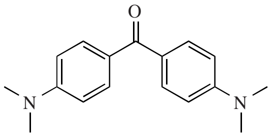
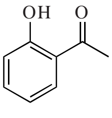
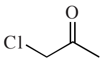


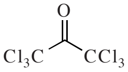
1639



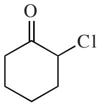
1648



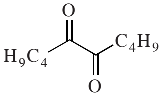
1752
1726
(rotamers)



1780
1751



1722



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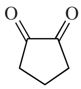
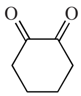
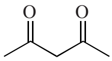


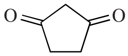
1700
1655



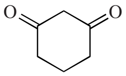
1735



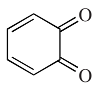
1724 (keto form)
1608 (enol form)



1755
1725
1635
1590



1630
1607



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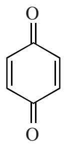

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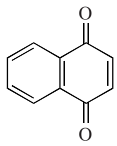

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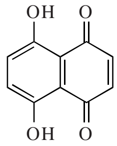
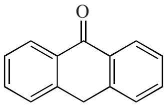
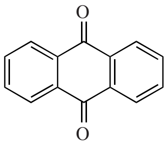


1623



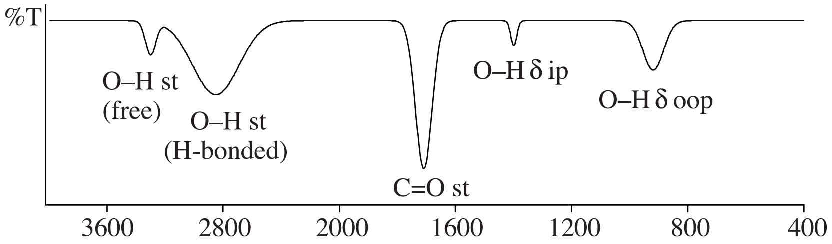
1662



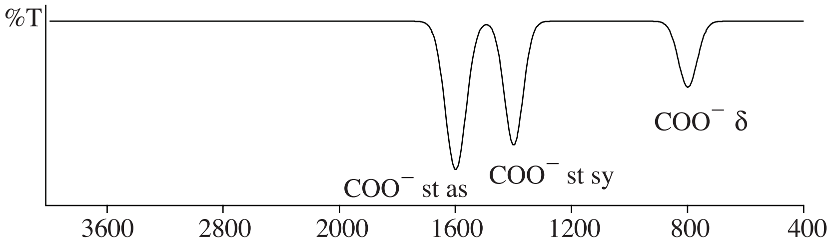
1678

7.11.3 Carboxylic Acids

Carboxylic Acids



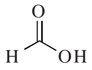
Carboxylate Anions



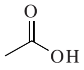
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
COO–H st	3550–2500	Intensity variable
	Subranges:	
	3550–3500	Free, sharp, only in highly diluted solutions
	3300–2500	H-bonded, broad, often more than one band
	In the same range: OH st, NH st, CH st, SiH st, SH st, PH st	
C=O st	1800–1650	Strong; in Raman, weak to medium
	1800–1740	Free (also in dicarboxylic acids)
	1740–1650	H-bonded (dimer, also in dicarboxylic acids)
	Subranges for H-bonded C=O:	
	1725–1700	al–COOH
	1715–1690	C=C–COOH
	1700–1680	ar–COOH
	1740–1720	hal–C–COOH
	1670–1650	Intramolecular H bond
OC–OH st, C–OH δ	1440–1210	Of no practical significance
OC–OH δ oop	960–880	Medium, generally broad (only in dimers); in the same range: =CH δ, ar CH δ, NH δ
$(COO)^-$ st as	1610–1550	Very strong; in α-halogen carboxylates near the higher value, with more than one α-hal beyond the normal range; in polypeptides at ≈1575
$(COO)^-$ st sy	1450–1400	Strong, of no practical significance, in polypeptides at ≈1470
$(COO)^-$ δ	≈775	Formates, weak
	≈925	Acetates
	≈680	Benzoates
	≈600	CF_3COO^-

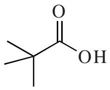
Examples ($\tilde{\nu}$ in cm^{-1})



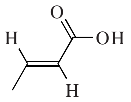
neat: 1727
in CCl_4: 1756, 1724



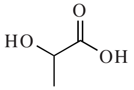
neat: 1759, 1718
in CCl_4: 1768, 1717



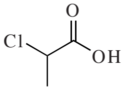
in CCl_4: 1704
solid: 1686



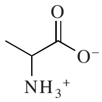
neat: 1700
in CCl_4: 1694



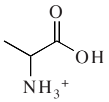
neat: 1725



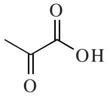
neat: 1730



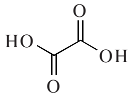
solid: 1605



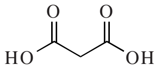
solid: 1740



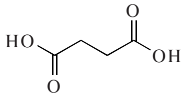
in CCl_4: 1788, 1725



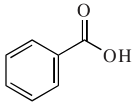
solid: 1650



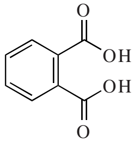
solid: 1735, 1703



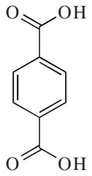
solid: 1724



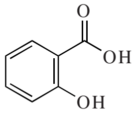
solid: 1690
in CCl_4: 1730
in $CHCl_3$: 1706



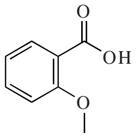
solid: 1690



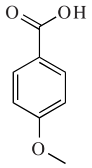
solid: 1690



solid: 1667
in CCl_4: 1696
in $CHCl_3$: 1661

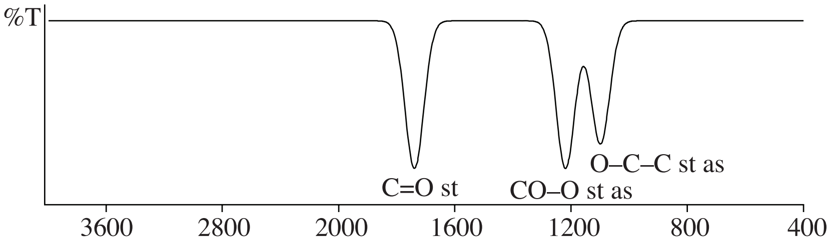


in CCl_4: 1750

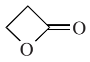
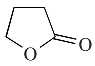
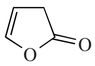
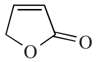
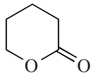
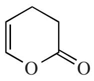
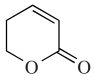
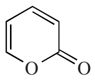


solid: 1693
in CCl_4: 1696

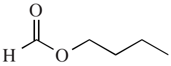
7.11.4 Esters and Lactones



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C=O st	1790–1650	Strong. In Raman, weak to medium
	Subranges:	
	1750–1735	Aliphatic esters
Conjugated esters:	1730–1710	α,β-Unsaturated esters
	1730–1715	Aromatic esters
	1690–1670	With intramolecular H bonds
	1790–1740	α-Halogenated esters
	≈1760	Vinyl esters, C=C st: 1690–1650, strong
	≈1760	Phenol esters
	≈1735	Phenol esters of aromatic acids
Diesters:		As the corresponding monoesters
Keto esters:	1755–1725	α-Keto esters, generally one band
	≈1750 (ketone) ≈1735 (ester)	β-Keto esters, keto form
	≈1650	β-Keto esters, enol form, C=C st: ≈1630, strong
	≈1740, ≈1715	γ-Keto esters, pseudoesters: ≈1770
Lactones:	 ≈1840	 ≈1770  ≈1800
		≈1750 (additional band at ≈1780 if α position is not substituted)  ≈1735
	 ≈1760	 ≈1720  ≈1730 (often doublet)
C–O st	1330–1050	2 bands: st as, very strong, at higher wavenumbers; st sy, strong, at lower wavenumbers
C–O st as:	*Subranges:*	
	≈1185	Formates, propionates, higher aliphatic esters
	≈1240	Acetates
	≈1210	Vinyl esters, phenol esters
	≈1180	γ-Lactones, δ-lactones
	≈1165	Methyl esters of aliphatic carboxylic acids
	In the same range:	Strong bands for C–F st, C–N st, N–O st, P–O st, C=S st, S=O st, P=O st, Si–O st, Si–H δ

Examples ($\tilde{\nu}$ in cm^{-1})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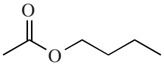
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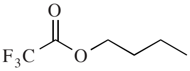
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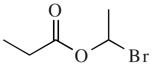
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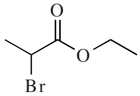
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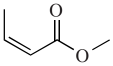
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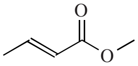
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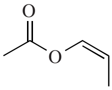
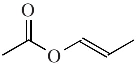
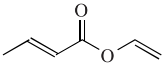
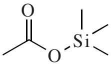

1758
(1690)


1752
(1675)


1730
(1658)
(1638)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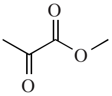
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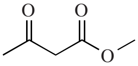
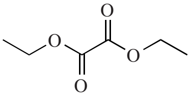

ester: 1704
ketone: 1690
enol: 1645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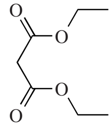
1754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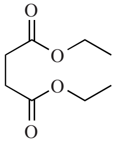
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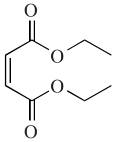
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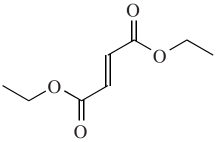
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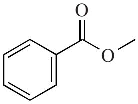
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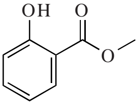
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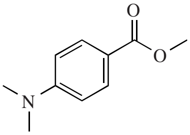
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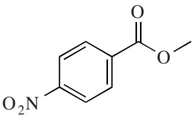
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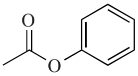
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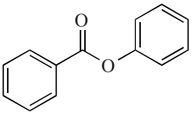
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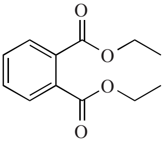
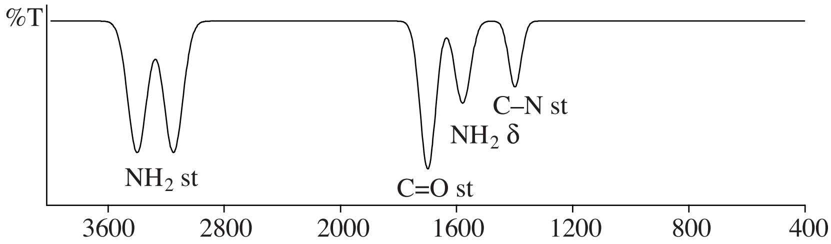
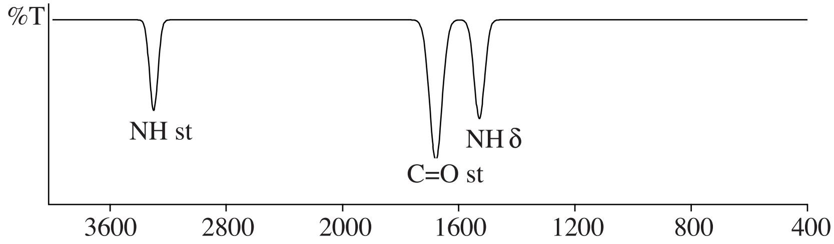

1746

7.11.5 Amides and Lactams

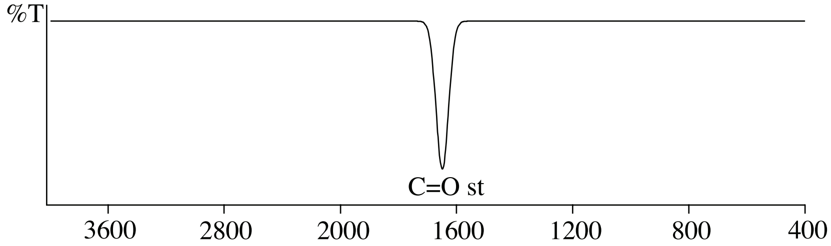
Primary Amides



Secondary Amides



Tertiary Amides

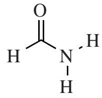


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

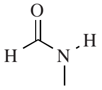
Assignment	Range	Comments
N–H st	3500–3100	Medium, in primary amides two bands, in proteins multiplet
	Subranges:	
	3500–3400	Free
	3350–3100	H-bonded
	≈3350, ≈3180	In primary amides generally two bands
	≈3200, ≈3100	In lactams generally two bands
	≈3200	Monohydrazides

Assignment	Range	Comments
	≈3100	Dihydrazides
	≈3250	Imides
	In the same range: OH st, ≡CH st (≈3300, sharp), H_2O	
C=O st	1740–1630	Generally strong. In Raman, weak to medium
(amide I)	*Subranges:*	
	≈1690	$NH_2C{=}O$ free amides, H-bonded: ≈1650
	≈1685	NHC=O free amides, H-bonded: ≈1660
	≈1650	NC=O free amides, H-bonded: ≈1650
	≈1745	4-Ring lactams
	≈1700	5-Ring lactams
	≈1650	6-, 7-Ring lactams
	≈1670	Monohydrazides
	≈1600	Dihydrazides
	1740–1670	Imides
	≈1750, 1700	5-Ring imides, 2 bands
	1655–1630	Polypeptides
	≈1690	Isocyanurates; with aromatic substitution: ≈1770
	≈1720, 1755 sh	Trifluoroacetamides
NH δ and N–C=O st sy	1630–1510	Generally strong, absent in lactams
(amide II)	*Subranges:*	
	≈1610	$NH_2C{=}O$ free, H-bonded: ≈1630
	≈1530	NHC=O free, H-bonded: ≈1540
	1560–1510	Polypeptides
	≈1555	Trifluoroacetamides
C–N st (?)	≈1400	$NH_2C{=}O$
	≈1250	NHC=O
	≈1330	Lactams
NH δ ip	≈1150	$NH_2C{=}O$
	≈1465	Lactams
NH δ oop	750–600	$NH_2C{=}O$
	≈700	NHC=O
	≈800	Lactams

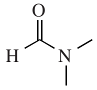
Examples ($\tilde{\nu}$ in cm^{-1})



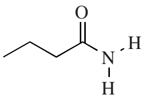
neat: 1672
in $CHCl_3$: 1709



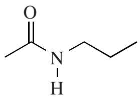
neat: 1672



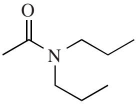
neat: 1670
in $CHCl_3$: 1673



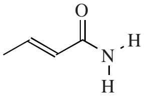
solid: 1631
in $CHCl_3$: 1679



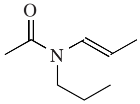
in CCl_4: 1690



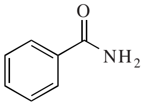
in CCl_4: 1647



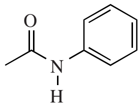
solid: 1677



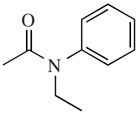
in CS_2: 1675
1650



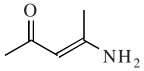
solid:1656
in $CHCl_3$: 1678



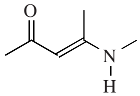
solid:1658
in $CHCl_3$: 1691
in CCl_4: 1705



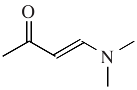
in CCl_4: 1667



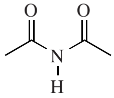
neat: 1700
1625
1540



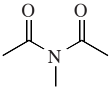
solid:1628
1595



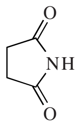
solid:1631
1584



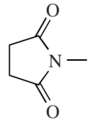
solid: 1734
1505



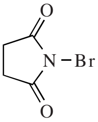
solid:1736
1706
1689



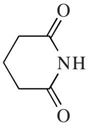
solid:1771
1698
in CCl_4: 1753
1727



solid: 1760
1690
in CCl_4: 1721
1705

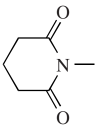


in $CHCl_3$: 1783
1733

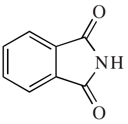


in CCl_4: 1742
1730
1718

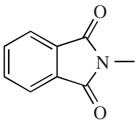
C = X



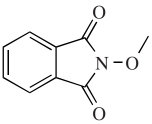
solid: 1718
1670
in CCl_4: 1729
1686



solid: 1774
1749
1724
in $CHCl_3$: 1778
1735

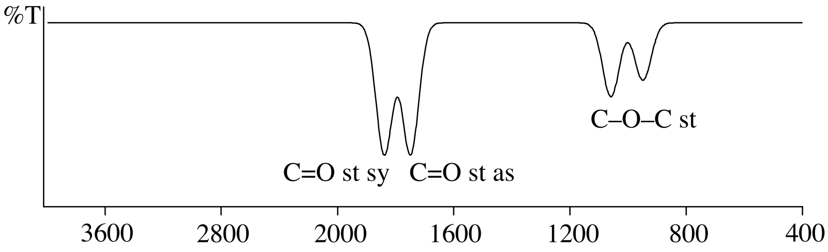


in $CHCl_3$: 1772
1712



solid:1790
1735

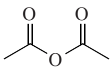
7.11.6 Acid Anhydrides



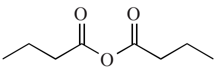
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
C=O st sy	1870–1770	Strong. In Raman, weak to medium
C=O st as	1800–1720	Strong. In Raman, weak to medium
	Subranges:	
	≈1820, ≈1760	Linear anhydrides, higher band stronger
	≈1850, ≈1775	5-Ring, lower band stronger
	≈1800, ≈1760	6-Ring, lower band stronger
C–O–C st	1300–900	Strong, several bands
	≈1040	Linear anhydrides
	≈920	Cyclic anhydrides

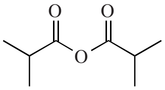
Examples ($\tilde{\nu}$ in cm^{-1})



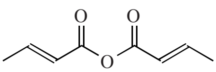
1825
1748



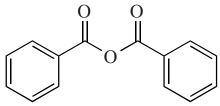
1810
1740
1045
1040



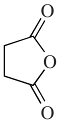
1803
1743



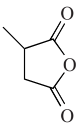
1780
1725

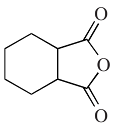


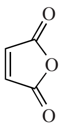
1790
1727
1035
1015
995

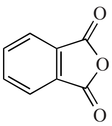


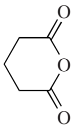
1865
1782
920

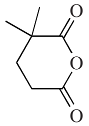

1859
1789


1845
1780

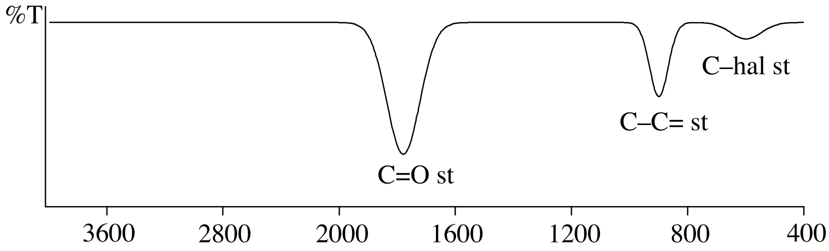

1850
1800
900


1840
1810
1760
912


1802
1761


1802
1761

7.11.7 Acid Halides

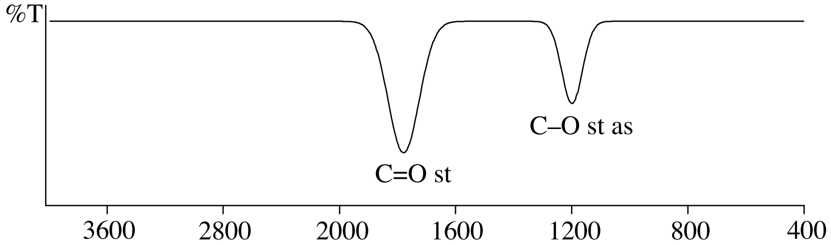


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

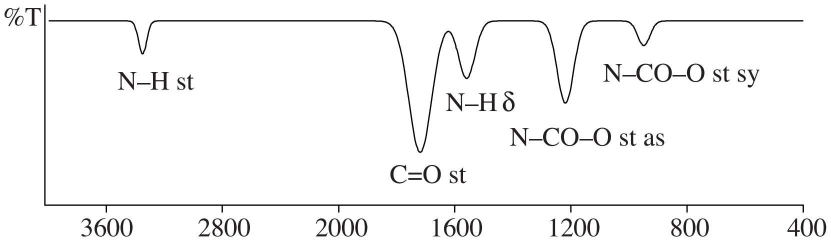
Assignment	Range	Comments
C=O st	1820–1750	Chlorides, strong; in Raman, weak to medium. Of narrow or medium width, for bromides and iodides at lower wavenumber
	1900–1870	Fluorides, strong, of narrow or medium width, additional band at ≈1725 in aromatic acid chlorides and bromides
C–CO st	1000–800	1000–900 aliphatic, assignment uncertain 900–800 aromatic, assignment uncertain
C–hal st	1200–500	1200–800 F 750–550 Cl 700–500 Br 600–500 I

7.11.8 Carbonic Acid Derivatives

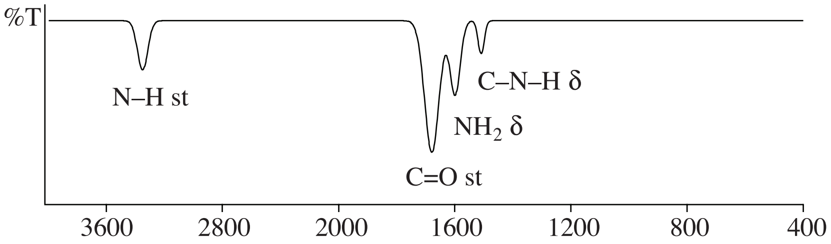
Carbonic Acid Esters



Carbamates



Ureas

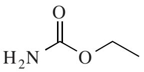


Typical Ranges ($\tilde{\nu}$ in cm^{-1})

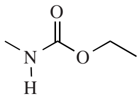
Assignment	Range	Comments	
C=O st	1820–1740	Strong. In Raman, weak to medium	Carbonic acid esters
	1750–1680	Strong. In Raman, weak to medium	Carbamates
	1690–1620	Strong. In Raman, weak to medium	Ureas
C–O st as	1260–1150	Strong	Carbonic acid esters

Assignment	Range	Comments	
N–H st	3500–3250	Medium, two bands for NH_2, one for NH	Carbamates
	3500–3200	Medium, two bands for NH_2	Ureas
N–H δ	1650–1500	Medium	Carbamates
NH_2 δ	1650–1600	Medium	Ureas
N–CO–O st as	1270–1210	Medium	Carbamates
N–CO–O st sy	1050–850	Weak	Carbamates
C–N–H δ	1600–1500	Weak	Ureas

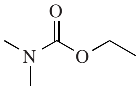
Examples ($\tilde{\nu}$ in cm^{-1})



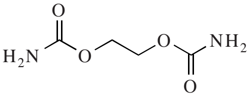
in $CHCl_3$: 1725



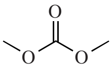
in CCl_4: 1727



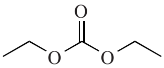
in $CHCl_3$: 1684



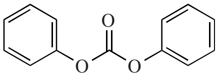
solid: 1690



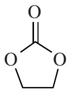
in CCl_4: 1758
in $CHCl_3$: 1751



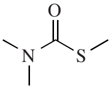
in CCl_4: 1748



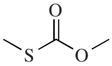
in CCl_4: 1786



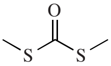
in CCl_4: 1822
1748



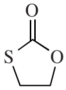
in CCl_4: 1662



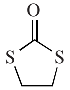
in CCl_4: 1719



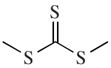
in CCl_4: 1653



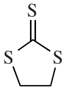
in CCl_4: 1757



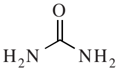
in CCl_4: 1718
1677
1640



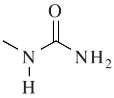
neat: 1076



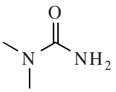
solid: 1058
in CCl_4: 1083
1079



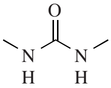
solid: 1679
1627



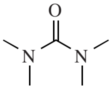
solid: 1645
1567
1418



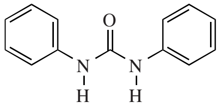
solid: 1656
1610
1511



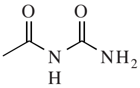
solid: 1622, 1580, 1530
in $CHCl_3$: 1663, 1548



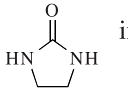
solid: 1645, 1560, 1497
in $CHCl_3$: 1675



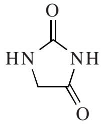
solid: 1650



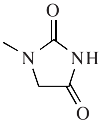
solid: 1667, 1634



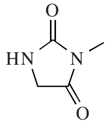
in CCl_4: 1735, 1718



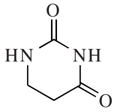
solid: 1776, 1697



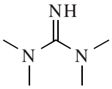
solid: 1712, 1676



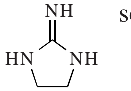
solid: 1748, 1706



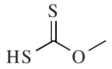
solid: 1767, 1695



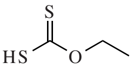
neat: 1600



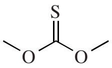
solid: 1767, 1681, 1621



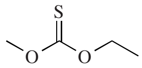
gas: 2593, 2548
neat: 2470



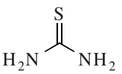
in CS_2: 2562, 2522



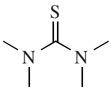
solid: 1212



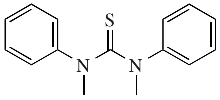
solid: 1234



solid: 1400



solid: 1130

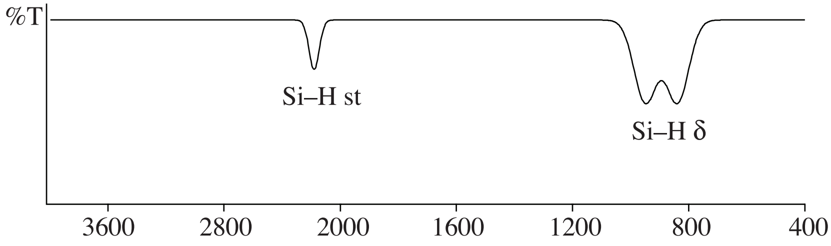


solid: 1131

C = X

7.12 Miscellaneous Compounds

7.12.1 Silicon Compounds



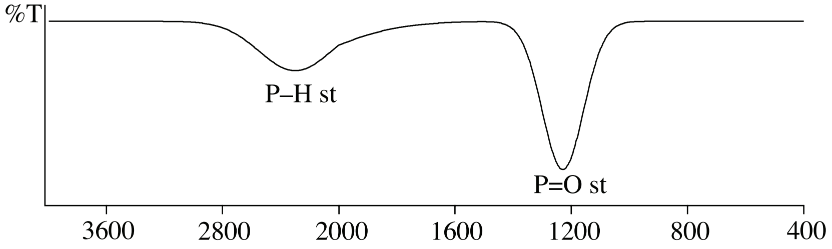
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
Si–H st	2250–2090	Medium. In Raman, medium to strong
	Subranges:	
	2160–2090	$R_3Si–H$; also for R as H; for SiH_3: 2 bands
	≈2250	hal–Si–H
	2220–2120	(Si–O)Si–H
Si–H δ	1010–700	Strong, broad, generally 2 bands
(Si–)CH_3 δ as	≈1410	Weak
(Si–)CH_3 δ sy	1275–1260	Very strong, sharp, typical for $SiCH_3$, not split for $Si(CH_3)_2$
(Si–)CH_3 γ	860–760	
	≈765	$SiCH_3$
	≈855, ≈800	$Si(CH_3)_2$
	≈840, ≈765	$Si(CH_3)_3$
Si–O st	1110–1000, 900– <600	
	1110–1000, 850–800	Si–O–C
	1090–1030, < 650	Si–O–Si
	900–800	Si–OH
	3700–3200	Si–OH st
	≈1030	Si–OH δ
Si–C st	850–650	
Si–N st	1250–830	
	Subranges:	
	950–830	Si–N–Si
	≈3400	Si_2NH st

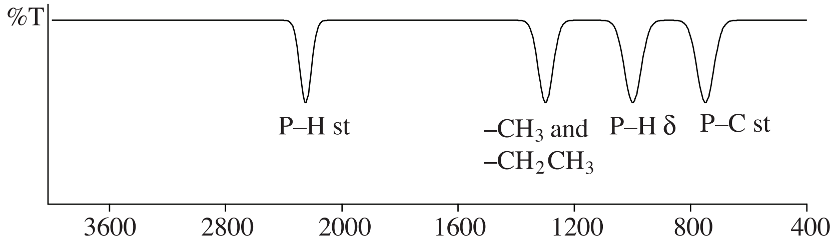
Assignment	Range	Comments
	950–830	N–Si–N
	1250–1100	Si–NH_2
	≈3570, ≈3390	SiN–H_2 st
	≈1540	Si–NH_2 δ
Si–F st	980–820	
	Subranges:	
	920–820	Si–F
	945–870	SiF_2, 2 bands
	980–860	SiF_3, 2 bands
Si–Cl st	< 625	

7.12.2 Phosphorus Compounds

Phosphorus Compounds

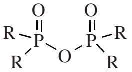
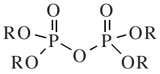
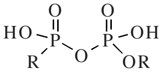
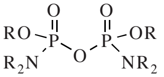
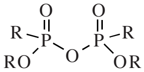
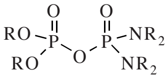
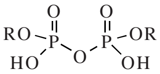
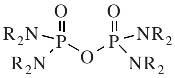


Phosphines



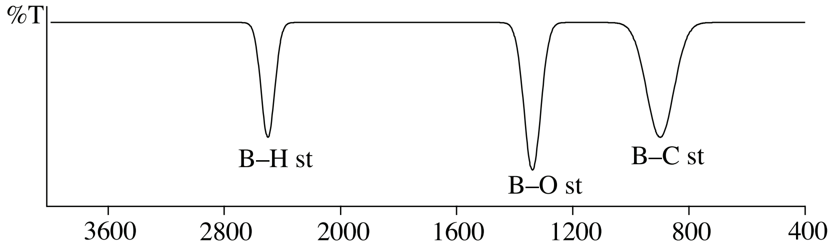
Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
P–H st	2440–2275	Weak to medium, generally one band, in R_3PH^+ very broad. In Raman, weak to medium
PO–H st	2700–2650	Weak, very broad
POH comb	2300–2250	Weak, very broad
	1740–1600	Additional band in O=P–OH (dimer?)
P–O st	1260–855	
	Subranges:	
	1050–970, 830–740	P–O–C al st; strong for upper band, often weak for lower band
	1260–1160	P–O–C ar st
	995–915	P(V)
	875–855	P(III)
	1100–940	P–OH st, broad, for $P(OH)_2$ often two bands
	980–900	P–O–P st
P=O st	1300–960	Strong. In Raman, weak to medium
	Subranges:	
	1190–1150	$R_3P{=}O$, also for R: H
	1265–1200	$R_2(R'O)P{=}O$, also for R: H
	1280–1240	$R(R'O)_2P{=}O$, also for R: H
	1300–1260	$(RO)_3P{=}O$
	1220–1150	$R(HO)_2P{=}O$
	1250–990	$R(HO)PO_2^-$, more than one band
	1125–970, 1000–960	RPO_3^{2-}
	1205–1090	$R_2(HO)P{=}O$
	1200–1090, 1090–995	$R_2PO_2^-$
	≈1250	$RO(HO)_2P{=}O$
	1230–1210, 1030–1020	$RO(HO)PO_2^-$
	1140–1050, 1010–970	$ROPO_3^{2-}$
	1250–1210	$(RO)_2(HO)P{=}O$
	1285–1120, 1120–1050	$(RO)_2PO_2^-$
	1220–1170	$R(RO)(HO)P{=}O$
	1245–1150, 1110–1050	$R(RO)PO_2^-$

Assignment	Range	Comments
	1240–1205	
	1310–1260	
	≈1195	
	≈1275	
	1265–1250	
	≈1300, ≈1240	
	≈1250	
	≈1235	
	1265–1240	R_2(X)P=O, X: F, Cl, Br
	1365–1260	R$(X)_2$P=O, X: F, Cl, Br
	1330–1280	$(RO)_2$(X)P=O, X: F, Cl, Br
	1365–1260	RO$(X)_2$P=O, X: F, Cl, Br
P=N	1500–1170	
P–OH δ	≈1280	Weak, of no practical significance
P–C st	800–700	Intensity varies widely, of no practical significance
P–H δ	1090–910	Strong, for $(RO)_2$HP=O very strong
P–N–C st	1110–930, 770–680	
P=N–al st	1500–1230	
P=N–ar st	1390–1300	
P=N–C=O st	1370–1310	
P=N–PR$_2$ st	1295–1170	
P=S st	750–580	Intensity varies widely
P–S st	<600	
(P–)CH$_3$ δ sy	1310–1280	
P–F st	905–760	

Assignment	Range	Comments
PF_2	1110–800	More than one band
P–Cl st	<600	

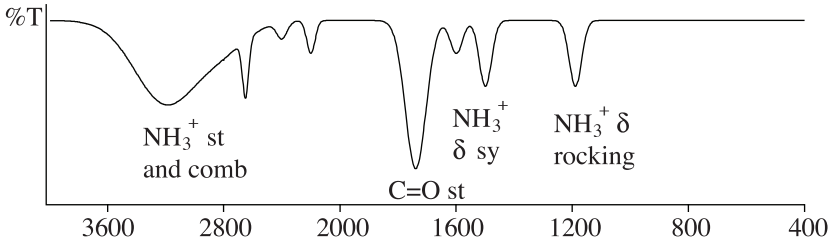
7.12.3 Boron Compounds



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

Assignment	Range	Comments
B–H st	2640–2200	Strong, in Raman weak to medium
	2200–1540	B–H···B, more than one band
B–O st	1380–1310	Very strong
	≈1500	Haloboroxines
BO–H st	3300–3200	Very broad
B–N st	1550–1330	Very strong
B–C st	1240–620	Strong, 2 bands if substitution highly asymmetric
B–F st	1500–800	
B–Cl st	1100–650	

7.13 Amino Acids



Typical Ranges ($\tilde{\nu}$ in cm^{-1})

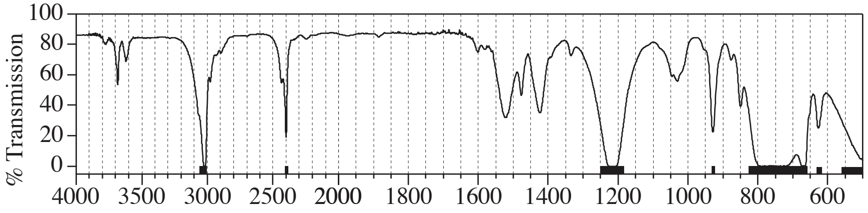
Assignment	Range	Comments
N–H st **O–H st**	3400–2000	Generally strong, broad, very structured
	Subranges:	
	3100–2000	Zwitterions, distinct side band at 2200–2000
	3350–2000	Hydrochlorides
	3400–3200	Na^+ salts
NH_3^+ δ as	1660–1590	Weak, for hydrochlorides near the lower limit
NH_3^+ δ sy	1550–1480	Medium
COO^- st as	1760–1595	Strong
	Subranges:	
	≈1595	Zwitterions
	1755–1700	Hydrochlorides; in α-amino acids: 1760–1730
	≈1595	Na^+ salts

7.14 Solvents, Suspension Media, and Interferences

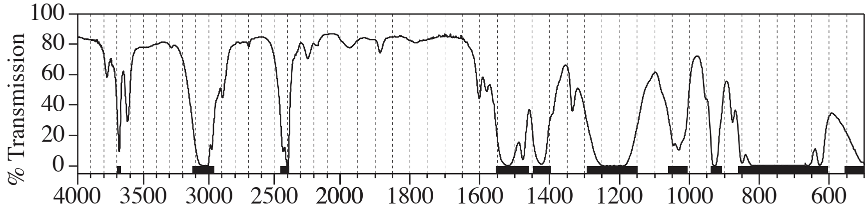
7.14.1 Infrared Spectra of Common Solvents

The low transmission in regions where the solvent absorbs may lead to artifacts. For the interpretation of spectra, these regions should be disregarded. In the following, they are indicated by bars.

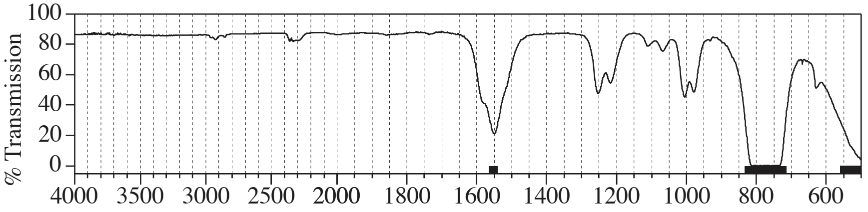
Chloroform: 0.2 mm cell



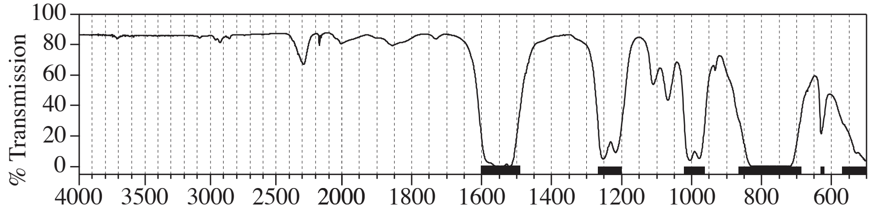
Chloroform: 1 mm cell



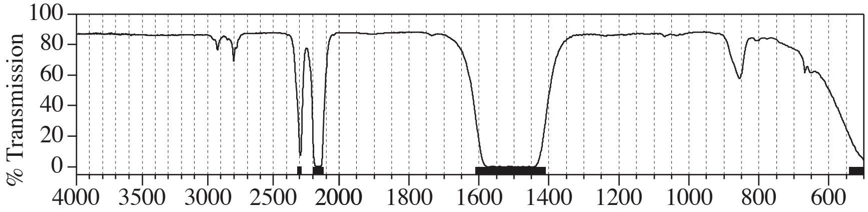
Carbon tetrachloride: 0.2 mm cell



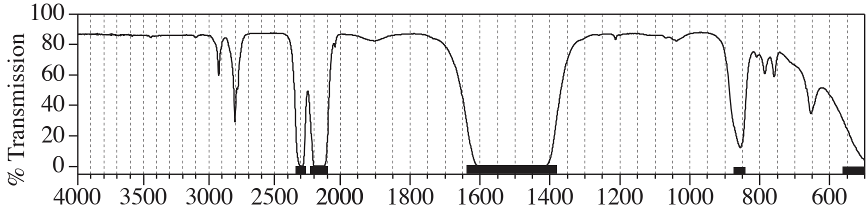
Carbon tetrachloride: 1 mm cell



Carbon disulfide: 0.2 mm cell



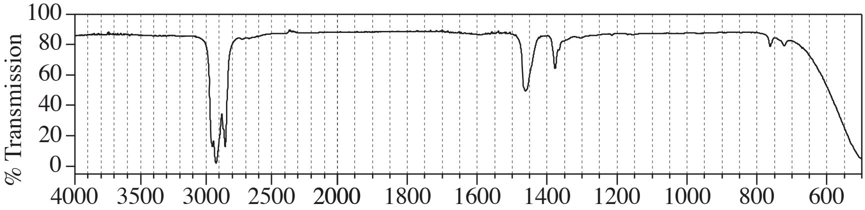
Carbon disulfide: 1 mm cell



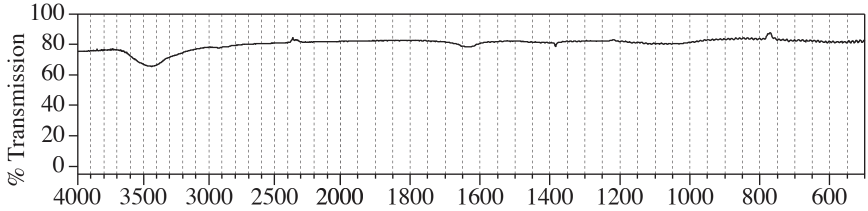
7.14.2 Infrared Spectra of Suspension Media

As it is difficult to prepare pellets and thin mineral oil films of reproducible thickness, the bands of these suspension matrixes are always found superimposed on the sample spectra.

Mineral oil (nujol): 10 μm thickness



Potassium bromide: pellet

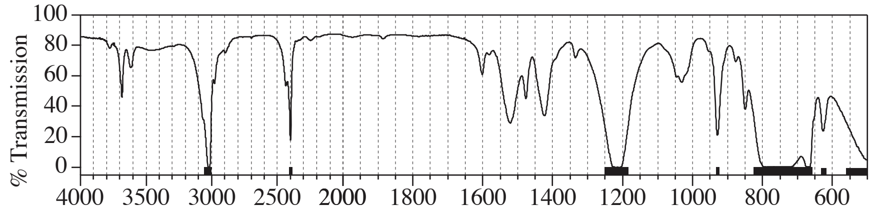


7.14.3 Interferences in Infrared Spectra

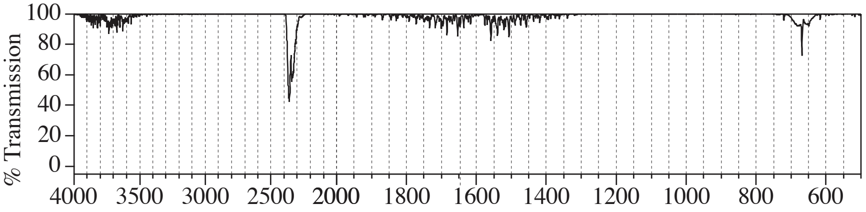
Traces of water in carbon tetrachloride or chloroform may give rise to two bands in the vicinity of 3700 and 3600 cm^{-1} as well as one around 1600 cm^{-1}. At higher concentrations, a broad band at 3450 cm^{-1} is found. Water in the vapor phase exhibits many sharp bands between 2000 and 1280 cm^{-1}. If present in high concentration, they may temporarily block the detector and appear as shoulders when occurring at a steep side of a strong signal.

Dissolved carbon dioxide shows an absorption band at 2325 cm^{-1}. In solutions that contain amines and traces of water, CO_2 can form carbonates, which lead to the appearance of unexpected bands of protonated N-containing groups. In improperly balanced double beam instruments, gaseous CO_2 can give rise to two signals at approximately 2360 and 2335 cm^{-1} as well as a signal at 667 cm^{-1}.

Chloroform, saturated with water: 0.2 mm cell



Water vapor with carbon dioxide



Commercially available polymers often contain phthalates as plasticizers, which can be found in apparently pure samples and give rise to a band at 1725 cm^{-1}. The presence of such phthalates can be confirmed by MS (m/z 149). In the course of chemical reactions, phthalates may be transformed into phthalic anhydride, which shows a band at 1755 cm^{-1}.

Other frequently encountered contaminants are silicones, which generally exhibit a band at 1625 cm^{-1}, together with a broad signal in the region from 1100 to 1000 cm^{-1}.

8 Mass Spectrometry

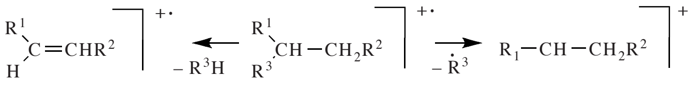
8.1 Alkanes [1]

Unbranched Alkanes [2,3]

Fragmentation: Larger alkyl fragments (with $C_{n>4}$) are chiefly formed by direct cleavage. They dehydrogenate and undergo substantial H and skeleton rearrangements. Smaller alkyl fragments (C_2 to C_4) are mainly formed by secondary decomposition of higher alkyl fragments. Eliminations of groups from within the chain (and recombination of its ends) also occur.
Ion series: Consecutive peaks corresponding to C_nH_{2n+1} (m/z 29, 43, 57, 71, ...), accompanied by C_nH_{2n-1} (m/z 27, 41, 55, 69, ...) and C_nH_{2n} (m/z 28, 42, 56, 70, ...) of lower intensity.
Intensities: Maximum intensity at m/z 43 or 57; with increasing masses, intensity of local maxima smoothly decreasing to a minimum at $[M-15]^+$.
Molecular ion: Medium intensity.

Branched Alkanes [2,3]

Fragmentation: In most cases, apparently simple bond cleavages, preferably at branched C atoms. The positive charge remains mainly on the branched C atom. Mechanistically, many H and skeleton rearrangements take place. This is reflected by the fact that no specific localization of heavy isotopes is possible.



Ion series: Consecutive peaks corresponding to C_nH_{2n+1} (m/z 29, 43, 57, 71, ...), accompanied by C_nH_{2n-1} (m/z 27, 41, 55, 69, ...) and C_nH_{2n} (m/z 28, 42, 56, 70, ...) of lower intensity.
Intensities: Local intensity maxima at those masses that result from cleavage at branched C atoms if the charge is localized there. Both C_nH_{2n+1} and (often more characteristically) C_nH_{2n} show this tendency.
Molecular ion: Intensity decreasing with increasing degree of branching. No $M^{+\cdot}$ is observed in highly branched systems.

References

[1] J.T. Bursey, M.M. Bursey, D.G. Kingston, Intramolecular hydrogen transfer in mass spectra. 1. Rearrangements in aliphatic hydrocarbons and aromatic compounds, *Chem. Rev.* **1973**, *73*, 191.
[2] K. Levsen, H. Heimbach, G.J. Shaw, G.W.A. Milne, Isomerization of hydrocarbon ions. VIII. The electron impact induced decomposition of *n*-dodecane, *Org. Mass Spectrom.* **1977**, *12*, 663.
[3] A. Lavanchy, R. Houriet, T. Gäumann, The mass spectrometric fragmentation of *n*-alkanes, *Org. Mass Spectrom.* **1979**, *14*, 79.

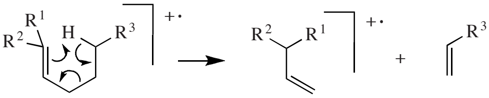
8.2 Alkenes [1–4]

Unbranched Alkenes

Fragmentation: Dominant loss of alkyl residues and neutral alkenes. The position of highly substituted double bonds can be localized because in this case alkene eliminations are specific McLafferty-type reactions. Otherwise, double bonds can be localized in derivatives, such as epoxides and glycols, or by means of low energy ionization techniques. Branching effects are less characteristic than in isoalkanes. Alicyclic compounds exhibit very similar spectra.
Ion series: Consecutive peaks corresponding to C_nH_{2n-1} (m/z 41, 55, 69, 83, ...), accompanied by alkyl and alkene ions, C_nH_{2n+1} (m/z 43, 57, 71, 85, ...) and C_nH_{2n} (m/z 42, 56, 70, 84, ...), mostly of lower intensity.
Intensities: Dominant maxima in the lower mass range, peaking around C_4. Local even-mass maxima due to alkene eliminations if the double bond is highly substituted.
Molecular ion: Significant, but not necessarily strong.

Branched Alkenes

Fragmentation: Highly substituted double bonds are less easily displaced than the unsubstituted ones and give rise to specific alkene eliminations of the McLafferty type, resulting in significant local maxima corresponding to C_nH_{2n} (see scheme). The latter may allow to localize the double bond. With unsubstituted double bonds, no reliable localization is possible and only moderately useful branching effects can be observed. The branching position is more easily determined after reduction to an alkane (in situ in GC/MS with H_2 as carrier gas and heated Pt wool as catalyst).



Ion series: Maxima of the alkene type (C_nH_{2n-1}; m/z 41, 55, 69, 83, ...), accompanied by weaker alkyl fragments, C_nH_{2n+1} (m/z 43, 57, 71, 85, ...), in the low mass range and more significant alkene ions, C_nH_{2n} (m/z 42, 56, 70, 84, ...).
Intensities: Intensive peaks in the lower mass range. Diagnostically important local maxima of even mass, frequently also in the higher mass range.
Molecular ion: Usually significant.

Polyenes and Polyynes

Fragmentation: The spectra of aliphatic compounds with several triple and/or double bonds are similar to those of aromatic hydrocarbons. A characteristic difference in the case of polyenes and polyynes is the presence of a signal at m/z 27, which is absent from spectra of purely aromatic compounds.
Ion series: Very similar to those of aromatic hydrocarbons, but fragments with higher hydrogen contents than in aromatics (m/z 54, 55; 66, 67; 79, 80) are usually found in polyenes and polyynes.

Intensities: Very similar distribution of peak intensities as for aromatic hydrocarbons.
Molecular ion: Usually strong, as in aromatic hydrocarbons.

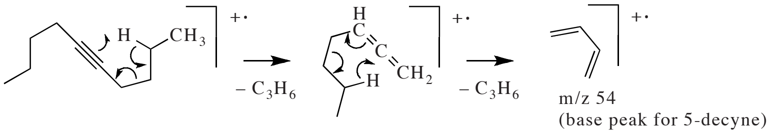
References

[1] A.G. Loudon, A. Maccoll, The mass spectrometry of the double bond. In: *The Chemistry of Alkenes, Vol. 2*; J. Zabicky, Ed.; Interscience: London, 1970; p 327.
[2] J.T. Bursey, M.M. Bursey, D.G. Kingston, Intramolecular hydrogen transfer in mass spectra. 1. Rearrangements in aliphatic hydrocarbons and aromatic compounds, *Chem. Rev.* **1973**, *73*, 191.
[3] N.J. Jensen, M.L. Gross, Localization of double bonds. *Mass Spectrom. Rev.* **1987**, *6*, 497.
[4] C. Dass, Ion–molecule reactions of $[\text{ketene}]^{+\cdot}$ as a diagnostic probe for distinguishing isomeric alkenes, alkynes, and dienes: A study of the C_4H_8 and C_5H_8 isomeric hydrocarbons, *Org. Mass Spectrom.* **1993**, *28*, 940.

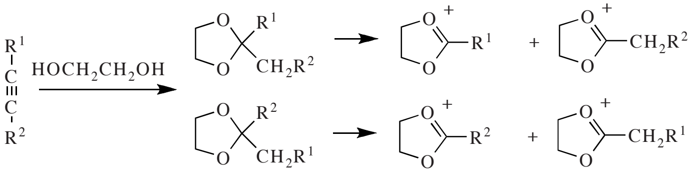
8.3 Alkynes [1]

Aliphatic Alkynes

Fragmentation: Tendency to lose a non-acetylenic $H^{\cdot}$ from $M^{+\cdot}$. Extensive rearrangements (including consecutive McLafferty rearrangements to the triple bond) result in uncharacteristic degradation:



In nonbranched alkynes with $C_{n>8}$, the rearrangement products at m/z 82 and 96 are dominant. Consecutive loss of methyl radical occurs. In general, no reliable localization of the triple bond is possible except in derivatives (as in ethylene glycol adducts [1], see scheme).



Ion series: Prominent peaks for C_nH_{2n-3} (m/z 25, 39, 53, 67, 81, …), accompanied by C_nH_{2n-1} (m/z 41, 55, 69, 83, …) and alkyl ions C_nH_{2n+1} (m/z 43, 57, 71, 85, …). Occasionally, even-mass maxima for C_nH_{2n-2} (m/z 26, 40, 54, 68, 82, …).
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Weak or missing in spectra of smaller molecules, significant in those of larger ones. Generally, $[M-1]^+$ is present. In terminal acetylenes, it is normally more abundant than $M^{+\cdot}$.

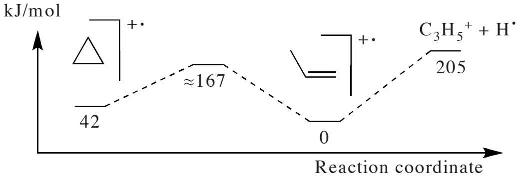
References

[1] C. Lifshitz, A. Mandelbaum, Mass spectrometry of acetylenes. In: *The Chemistry of the Carbon-Carbon Triple Bond, Part 1*; S. Patai, Ed.; Wiley: Chichester, 1978; p 157.

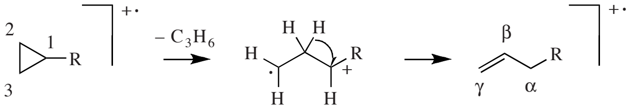
8.4 Alicyclics [1]

Cyclopropanes [2,3]

Fragmentation: Generally, spectra of cyclopropanes and alkenes are very similar because at 70 eV ionization, the ring readily isomerizes to the corresponding alkene radical cations.



Preferred primary fragmentation by bond cleavage at branched C atoms. Loss of alkyl residues and of neutral alkenes dominates. The ring of monosubstituted cyclopropanes is opened exclusively at the 1,2- and not at the 2,3-bond. The primarily formed double bond is predominantly (for R: OCH_3) or exclusively (for R: H, alk, $COOCH_3$) found in the β,γ-position (even for $COOCH_3$, where the α,β-unsaturation is thermodynamically more stable).



Molecular ions of cyclopropyl cyanide, allyl cyanide, methacrylonitrile, and pyrrole rearrange to one common radical cation, most likely that of pyrrole [4].
Ion series: Consecutive maxima corresponding to C_nH_{2n-1} (m/z 41, 55, 69, 83, ...), accompanied by alkyl and alkenyl ions of the type C_nH_{2n+1} (m/z 43, 57, 71, 85, ...) and C_nH_{2n} (m/z 42, 56, 70, 84, ...), mostly of lower intensity.
Intensities: Dominant peaks in the low mass range, peaking around C_4. Local even-mass maxima due to alkene eliminations if the resulting double bond is highly substituted.
Molecular ion: Significant, but not necessarily strong.

Saturated Monocyclic Alicyclics [5]

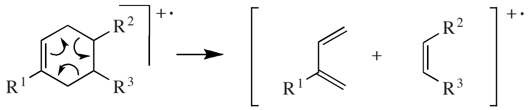
Fragmentation: Preferred primary fragmentation by bond cleavage at branched C atoms, followed by loss of alkyl residues and alkenes.
Ion series: Consecutive maxima corresponding to C_nH_{2n-1} (m/z 41, 55, 69, 83, ...), accompanied by C_nH_{2n+1} (m/z 43, 57, 71, 85, ...) and C_nH_{2n} (m/z 42, 56, 70, 84, ...) of lower intensities. In general, the maxima are so similar to those of alkenes that no clear distinction is possible.
Intensities: Overall distribution of peaks maximizing in the lower mass range, around C_4 or C_5. Local maxima can result from branching effects.
Molecular ion: Significant, mostly of medium intensity.

Polycyclic Alicyclics

Fragmentation: Most important primary cleavage at highly branched carbon atoms, followed by H rearrangements and complex fragmentations.
Ion series: With increasing number of rings, the position of unsaturated hydrocarbon fragments in the upper m/z range shifts from C_nH_{2n-1} (m/z 41, 55, 69, 83, ...) to C_nH_{2n-3} (m/z 39, 53, 67, 81, ...) and to C_nH_{2n-5} (m/z 51, 65, 79, 93, ...). Typically, maxima in the lower m/z range have a lower degree of unsaturation than those in the upper m/z range.
Intensities: Major maxima evenly distributed, somewhat more intensive in the high mass or $M^{+\cdot}$ range.
Molecular ion: Strong.

Cyclohexenes

Fragmentation: Loss of larger ring substituents as well as retro-Diels–Alder reaction, yielding fragments of even-mass maxima with one or two double-bond equivalents, C_nH_{2n} (m/z 42, 56, 70, 84, ...) and C_nH_{2n-2} (m/z 40, 54, 68, 82, ...), unless the retro-Diels–Alder product corresponds to ethylene. Somewhat unexpectedly, the base peak of cyclohexene is at $[M-15]^+$.
The retro-Diels–Alder reaction often accounts for prominent fragments of cyclohexenes and 1,4-cyclohexadienes:



However, double-bond migration may or may not occur beforehand. Also, other fragmentation pathways may dominate. Therefore, a reliable localization of the double bond in cyclohexene derivatives of unknown structure is not necessarily possible. For example, the base peak of 1,2-dimethylcyclohexene is at m/z 68 rather than at the expected m/z 82.
Ion series: Unsaturated hydrocarbon fragments in the upper m/z range are shifted, relative to cyclohexane fragments, by two mass units to C_nH_{2n-3} (m/z 39, 53, 67, 81, ...). Typically, maxima in the lower m/z range correspond to a lower degree of unsaturation than those in the upper m/z range.
Intensities: Intensive peaks evenly distributed over whole mass range.
Molecular ion: Medium intensity (ca. 40% in cyclohexene).

References

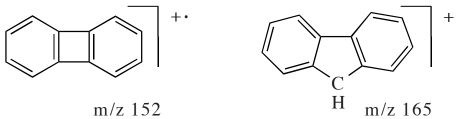
[1] J.T. Bursey, M.M. Bursey, D.G. Kingston, Intramolecular hydrogen transfer in mass spectra. 1. Rearrangements in aliphatic hydrocarbons and aromatic compounds, *Chem. Rev.* **1973**, *73*, 191.
[2] H. Schwarz, The chemistry of ionized cyclopropanes in the gas phase. In: *The Chemistry of the Cyclopropyl Group, Part 1*; Z. Rappoport, Ed.; Wiley: Chichester, 1987; p 173.

[3] J.R. Collins, G.A. Gallup, Energy surfaces in the cyclopropane radical ion and the photoelectron spectrum of cyclopropane, *J. Am. Chem. Soc.* **1982**, 104, 1530.
[4] G.D. Willet, T. Baer, Thermochemistry and dissociation dynamics of state-selected C_4H_4X ions. 3. $C_4H_5N^+$, *J. Am. Chem. Soc.* **1980**, *102*, 6774.
[5] E.F.H. Brittain, C.H.J. Wells, H.M. Paisley, Mass spectra of cyclobutanes and cyclohexanes of molecular formula $C_{10}H_{16}$, *J. Chem. Soc. B* **1968**, 304.

8.5 Aromatic Hydrocarbons [1–4]

Aromatic Hydrocarbons

Fragmentation: Weak tendency of fragmentation. Elimination of H˙ and successive H_2 eliminations, yielding $[M\text{-}1]^+$, $[M\text{-}3]^+$, and $[M\text{-}5]^+$ of decreasing intensities. In condensed aromatics, $[M\text{-}2]^{+\cdot}$ can be a dominating fragment. Further typical fragmentation reactions are the eliminations of acetylene (Δm 26) and C_3H_3 (Δm 39). Some CH_3 elimination frequently occurs in pure aromatic compounds. In the case of diphenyl compounds, biphenylene (m/z 152) and, if a CH_2 group is available, fluorene (m/z 165) ions are typically observed.

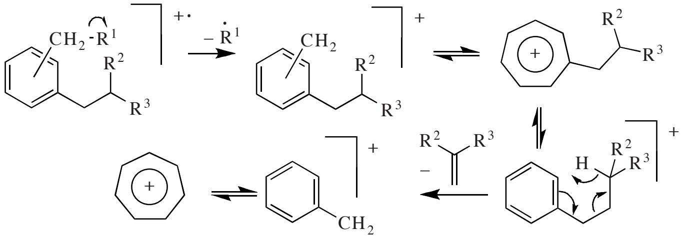


m/z 152 m/z 165

Ion series: C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...), for polycyclic aromatics gradually changing to more highly unsaturated ions. Doubly charged ions occur frequently, in particular as the size of the π-electron system increases.
Intensities: Weak fragments. The intensity pattern of doubly charged ions does not follow that of the corresponding singly charged ions.
Molecular ion: Strong.

Alkylsubstituted Aromatic Hydrocarbons

Fragmentation: Dominant loss of alkyl residues by benzylic cleavage, followed by elimination of alkenes.



At low resolution, methylbenzyl and β-phenylethyl have the same mass as benzoyl (m/z 105). In contrast to benzoyl, dehydrogenation products (m/z 104, 103) as well as protonated benzene (m/z 79) are also present if m/z 105 is a hydrocarbon rest.
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...), in the lower mass range.
Intensities: Intensive peaks mainly in the higher mass range. Maxima by benzylic cleavage.
Molecular ion: Strong or medium.

References

[1] J.T. Bursey, M.M. Bursey, D.G. Kingston, Intramolecular hydrogen transfer in mass spectra. 1. Rearrangements in aliphatic hydrocarbons and aromatic compounds, *Chem. Rev.* **1973**, *73*, 191.

[2] W. Schönfeld, Fragmentation diagrams for elucidation of decomposition reactions of organic compounds. 1. Aromatic hydrocarbons (in German), *Org. Mass Spectrom.* **1975**, *10*, 321.

[3] C. Lifshitz, Tropylium ion formation from toluene: Solution of an old problem in organic mass spectrometry. *Acc. Chem. Res.* **1994**, *27*, 138.

[4] M.V. Buchanan, B. Olerich, Differentiation of polycyclic aromatic hydrocarbons using electron-capture negative chemical ionization, *Org. Mass Spectrom.* **1984**, *19*, 486.

8.6 Heteroaromatic Compounds [1,2]

General Characteristics

Fragmentation: Mostly fragments of aromatic character with specific eliminations including heteroatoms, e.g., elimination of HCN, CO, CHO, CS, and CHS from $M^{+\cdot}$, and of HCN, CO, and CS from fragments. In the case of alkyl-substituted heteroaromatics, occurrence of benzylic-type cleavage and McLafferty rearrangements of substituents with $C_{n>1}$ as well as specific rearrangements including heteroatoms, especially in N aromatics.
Ion series: Aromatic fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, …), in the lower mass range if the necessary number of C atoms is present (no such fragments, e.g., in pyrazine). Ions including heteroatoms like $HCN^{+\cdot}$ (m/z 27), CH_3CNH^+ (m/z 42), and $CS^{+\cdot}$ (m/z 44).
Intensities: Intensive peaks mainly in the higher mass range.
Molecular ion: Generally strong. $[M-1]^+$ is often relevant in alkyl-substituted heteroaromatics.

Furans [3]

Fragmentation: Oxygen can be lost from $M^{+\cdot}$ together with the neighboring C as CHO (Δm 29). In 2- or 6-methylfurans, CH_3CO^+ (m/z 43) can be seen (base peak in 2,5-dimethylfuran). As in aromatic methyl ethers, $[M-43]^+$ is a product of a two-step reaction: ($M^{+\cdot}$-$CH_3^{\cdot}$-CO). Furans substituted with an alkyl group ($C_{n>1}$): benzylic-type cleavage (to pyrylium ion, $C_5H_5O^+$, m/z 81), followed by loss of CO.
Ion series: Mainly aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, …).
Intensities: Intensive peaks mainly in the higher mass range. The fragments are usually more important than in purely aromatic hydrocarbons.
Molecular ion: Strong. No pronounced tendency to protonate. Usually, $[M-1]^+$ is very strong in methylfurans.

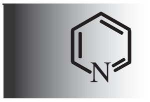
Thiophenes [4]

Fragmentation: Sulfur can be lost from $M^{+\cdot}$ together with the neighboring C as CHS (Δm 45) or CS (Δm 44). Typical for thiophenes substituted with an alkyl group ($C_{n>1}$) is benzylic-type cleavage followed by loss of CS (Δm 44). Protonated thiophene (m/z 85) is a characteristic product of monoalkylated thiophenes.
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, …). Besides the isotope peak at $[M+2]^{+\cdot}$, the signals at m/z 44 and 45 ($CS^{+\cdot}$ and CHS^+) are indicators of sulfur.
Intensities: Dominant peaks for $M^{+\cdot}$ and products of benzylic-type cleavage.
Molecular ion: Strong. Characteristic S isotope signal ($[M+2]^{+\cdot}$ corresponds to 4.5% of $M^{+\cdot}$). No pronounced tendency of protonation. Usually, $[M-1]^+$ is very strong in methylthiophenes.

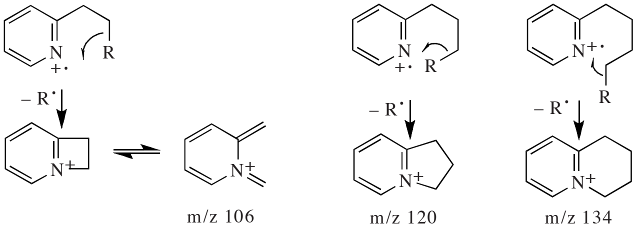
Pyrroles [5]

Fragmentation: HCN elimination (Δm 27) from $M^{+\cdot}$ and from fragments. In methylpyrroles, $[M\text{-}1]^+$ is dominant. Benzylic-type cleavage in *C*- and *N*-alkylpyrroles with or without (nonspecific) H rearrangements.
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, ...).
Intensities: Dominant peaks for $M^{+\cdot}$ and products of benzylic-type cleavage.
Molecular ion: Strong (odd mass for odd number of N in the molecule). No tendency to protonate. In methyl-substituted pyrroles, $[M\text{-}1]^+$ is dominant.

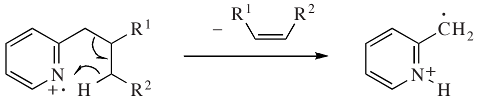
Pyridines



Fragmentation: HCN elimination (Δm 27) from fragments and the ion H_2CN^+ (m/z 28) are characteristic. Additional reactions in 2- or 6-methylpyridines are CH_3CN elimination (Δm 41) and the formation of CH_3CNH^+ (m/z 42). Benzylic cleavage is dominant for 3-alkyl-, strong for 4-alkyl-, and weak for 2-alkylpyridines. Typical rearrangements with participation of the N atom in 2- and 6-alkylpyridine derivatives.
Intramolecular *N*-alkylation in 2-alkyl derivatives:



McLafferty rearrangements are important in 2- and 4-alkylpyridines:



Ion series: Aromatic hydrocarbon fragments, C_nH_n, $C_nH_{n\pm1}$ and $C_nH_{n\pm1}N$ (m/z 39–41, 51–54, 63–67, 75–80, ...).
Intensities: Dominant peaks for $M^{+\cdot}$ or, if possible, for products of benzylic-type cleavage.
Molecular ion: Strong, except when benzylic-type cleavage is possible. Odd mass for an odd number of N in the molecule. No tendency to protonate. $[M\text{-}1]^+$ is usually present and is strong in alkyl-substituted pyridines.

N-*Oxides of Pyridines and Quinolines*

Fragmentation: The $[M\text{-}O]^{+\cdot}$ radical ion, of variable intensity, is probably due to thermal decomposition. The fragments $[M\text{-}CO]^{+\cdot}$ and, if an alkyl group is present on the neighboring C atom, $[M\text{-}OH]^+$ are relevant for quinoline *N*-oxides. Rearrangements with ring formation including the N–O moiety if alkyl or aryl groups

are present in the neighboring positions.
Ion series: As for the corresponding heteroaromatics, aromatic hydrocarbon fragments, C_nH_n, $C_nH_{n\pm1}$ and $C_nH_{n\pm1}N$ (m/z 39–41, 51–54, 63–67, 75–80, ...), are observed.
Intensities: Dominant peaks for $M^{+\cdot}$ and products of benzylic-type cleavage.
Molecular ion: Strong, except when $[M\text{-}O]^{+\cdot}$ dominates due to experimental conditions or when benzylic-type cleavage is possible. Odd mass for odd number of N atoms in the molecule. No tendency to protonate.

Pyridazines and Pyrimidines

Fragmentation: Loss of N_2 or $CH_2N^{\cdot}$ (Δm 28) from pyridazines. Also, loss of $N_2H^{\cdot}$ (especially important in methylpyridazines) to give $[M\text{-}29]^+$. In pyridazine *N*-oxides, consecutive loss of $NO^{\cdot}$ and HCN. Consecutive losses of two HCN ($2 \times \Delta$m 27) molecules from pyrimidines. From 2-, 4-, and 6-methylpyrimidines, CH_3CN (Δm 41) is eliminated and the ion CH_3CNH^+ (m/z 42) occurs.
Ion series: Aromatic hydrocarbon fragments (C_nH_n, $C_nH_{n\pm1}$) and, for pyrimidines, $C_nH_{n\pm1}N$, at low masses (m/z 39, 51–53).
Intensities: Dominant peak for $M^{+\cdot}$.
Molecular ion: Strong. No tendency to protonate. For pyrimidines, $[M\text{-}1]^+$ is usually observable.

Pyrazines

Fragmentation: Consecutive losses of two HCN ($2\times\Delta$m 27) molecules. For methylpyrazines, elimination of CH_3CN (Δm 41) and formation of CH_3CNH^+ (m/z 42).
Ion series: No aromatic character of the spectra.
Intensities: Dominant peak for $M^{+\cdot}$.
Molecular ion: Strong. No tendency to protonate. Usually, $[M\text{-}1]^+$ is observable; it can be stronger than $M^{+\cdot}$ in alkyl-substituted ($C_{n>1}$) pyrazines.

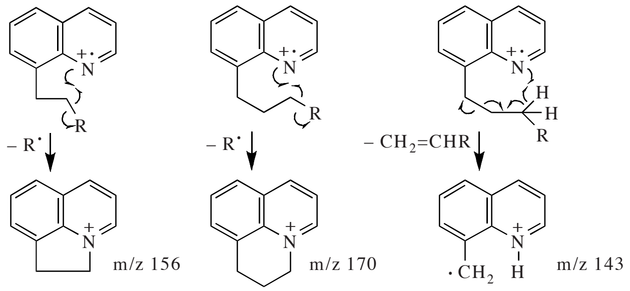
Indoles

Fragmentation: Analogous to pyrrole; HCN elimination (Δm 27) from $M^{+\cdot}$ and from fragments. From $M^{+\cdot}$ also $CH_2N^{\cdot}$ (Δm 28) elimination (in one or two steps). In methyl-substituted indoles, $[M\text{-}1]^+$ is dominant. In *N*-methylindoles, $[M\text{-}15]^+$ is significant. Benzylic-type cleavage in *C*- and *N*-alkylindoles with or without (nonspecific) H rearrangements.
Ion series: Aromatic ion series.
Intensities: Dominant maxima in the higher mass range.
Molecular ion: Strong. No tendency to protonate. In methyl-substituted indoles, strong signal for $[M\text{-}1]^+$.

Quinolines and Isoquinolines

Fragmentation: Similar to pyridine: HCN elimination (Δm 27) from $M^{+\cdot}$, $[M-1]^+$, and fragments. In methylquinolines and methylisoquinolines also CH_3CN elimination (Δm 41). In alkyl-substituted ($C_{n>1}$) quinolines, benzylic cleavage dominates except when neighboring effects of N play a role. For 2- and 8-alkylquinolines as well as 1- and 3-alkylisoquinolines, see rearrangements in pyridines.
Ion series: Aromatic hydrocarbon fragments, C_nH_n, $C_nH_{n\pm1}$, and $C_nH_{n\pm1}N$ (m/z 39–41, 51–54, 63–67, 75–80, …).
Intensities: Dominant peak for $M^{+\cdot}$ or, if possible, for products of benzylic-type cleavage.
Molecular ion: Strong, except when benzylic-type cleavage is possible. Odd mass for odd number of N atoms in the molecule. No tendency to protonate. $[M-1]^+$ is usually present and is strong in alkyl-substituted quinolines.

Rearrangements in 8-alkylquinolines:



Cinnoline, Phthalazine, Quinazoline, Quinoxaline

Fragmentation: Same as for the corresponding monocyclic heteroaromatics pyridazine, pyrimidine, and pyrazine. Characteristic for pyridazine, cinnoline, and phthalazine is the elimination of N_2 (Δm 28) and $N_2H^\cdot$ (Δm 29) from their alkyl derivatives. Phthalazine loses HCN (Δm 27) twice.
Ion series: Aromatic hydrocarbon fragments, (C_nH_n, $C_nH_{n\pm1}$) and $C_nH_{n\pm1}N$ (m/z 39–41, 51–54, 63–67, 75–80, …).
Intensities: Dominant maximum for $M^{+\cdot}$ or, if possible, for products of benzylic-type cleavage.
Molecular ion: Strong, except when benzylic-type cleavage is possible. Odd mass for odd number of N atoms in the molecule. No tendency to protonate. $[M-1]^+$ is usually present and is strong in alkyl-substituted compounds.

References

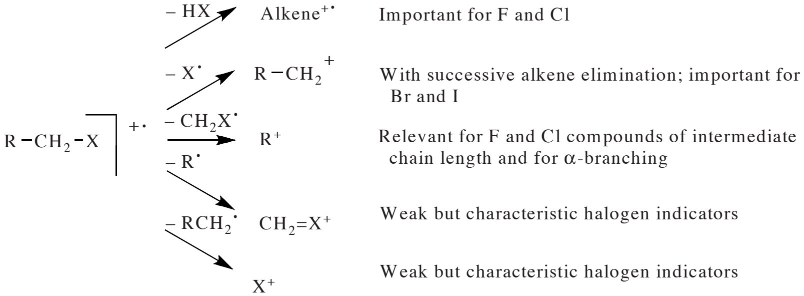
[1] Q.N. Porter, *Mass Spectrometry of Heterocyclic Compounds*, 2nd ed.; Wiley: New York, 1985.
[2] D.G.I. Kingston, B.W. Hobrock, M.M. Bursey, J.T. Bursey, Intramolecular hydrogen transfer in mass spectra. III. Rearrangements involving the loss of small neutral molecules, *Chem. Rev.* **1975**, *75*, 693.

[3] R. Spilker, H.-F. Grützmacher, Isomerization and fragmentation of methylfuran ions and pyran ions in the gas phase, *Org. Mass Spectrom.* **1986**, *21*, 459.
[4] W. Riepe, M. Zander, Mass-spectrometric fragmentation behavior of thiophene benzologs. *Org. Mass Spectrom.* **1979**, *14*, 455.
[5] H. Budzikiewicz, C. Djerassi, A.H. Jackson, G.W. Kenner, D.J. Newmann, J.M. Wilson, Mass spectra of monocyclic derivatives of pyrrole, *J. Chem. Soc.* **1964**, 1949.

8.7 Halogen Compounds [1–3]

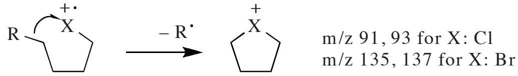
Saturated Aliphatic Halides

Fragmentation: Loss of halogen radical (I > Br > Cl > F) followed by elimination of alkenes. Loss of alkyl radical followed by elimination of acid HX. Loss of acid HX to give an alkene radical cation.



Ion series: The dominant hydrocarbon fragments are mainly alkenyl fragments (C_nH_{2n-1}) for F and Cl, mixed alkyl (C_nH_{2n+1}) and alkenyl fragments (C_nH_{2n-1}) for Br, and mainly alkyl fragments (C_nH_{2n+1}) for I.

Intensities: Intensive peaks mainly in the lower mass range. Characteristic maxima for Cl and Br at $C_4H_8X^+$ (m/z 91, 93 and 135, 137, respectively), which has a cyclic structure:



Alkyl substituents on the chain reduce the intensity of this fragment. If it is strong, $[M\text{-}X]^+$ is weak. In the case of iodoalkanes, some I^+ and $HI^{+\cdot}$ at m/z 127, 128 is usually detectable.

Molecular ion: Strong for the smallest alkanes, with increasing intensity in the sequence F, Cl, Br, I. Decreases rapidly with increasing mass and with increasing branching. It is negligible for F and Cl if the *n*-alkyl chains are longer than pentyl, and for Br and I if they are longer than heptyl and nonyl, respectively. Low tendency to protonate. Characteristic isotope patterns for Cl and Br. Iodine can be detected because of its high mass; the ^{13}C signals of $M^{+\cdot}$ and its fragments are conspicuously weak.

Polyhaloalkanes

Fragmentation: Preferred fragmentation of the C–C bond if several halogen atoms are bonded to one of these carbon atoms. CF_3 (m/z 69) is often the base peak in terminally perfluorated alkanes, and so is $CHCl_2$ (m/z 83, 85, 87) in terminally dichlorinated compounds. Often, X_2 is eliminated besides the usual fragmentation of $X^\cdot$ and HX. Interchange of halogens may occur. For example, m/z 85 (CF_2Cl) is a dominant signal (ca. 60%) for CF_3CFCl_2.

Ion series: Most fragments are halogenated alkyl and alkenyl groups, easily detectable on the basis of the isotope signals in the cases of Cl and Br.

Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Weak, decreasing with increasing number of halogen atoms. Absent from the spectra of many polyhalogenated compounds.

Aromatic Halides

Fragmentation: Consecutive losses of halogen radicals and/or acid HX. In perhalogenated aromatics, decomposition down to C_x^+, with x from 1 to 6 (m/z 12, 24, 36, 48, 60, 72). If alkyl-substituted ($C_{n>1}$), the base peak is mostly the result of benzylic cleavage. In an otherwise aromatic environment, m/z 57 is a F indicator ($C_3H_2F^+$). Elimination of CF_2 (Δm 50) from CF_3 groups attached to the aromatic ring (from $M^{+\cdot}$ or fragments).
Ion series: Aromatic fragments, C_nH_n, C_nH_{n-1}, and C_nH_{n-2} (m/z 39, 51–53, 63–65, 75–77, ...). In the higher mass range: $C_n(H,X)_n$.
Intensities: Dominant peaks in the $M^{+\cdot}$ region.
Molecular ion: Usually very strong. Characteristic isotope signals for Cl and Br.

References

[1] A.G. Loudon, Mass spectrometry and the carbon-halogen bond. In: *The Chemistry of the Carbon-Halogen Bond, Part 1*; S. Patai, Ed.; Wiley: London, 1973; p 223.
[2] D.G.I. Kingston, B.W. Hobrock, M.M. Bursey, J.T. Bursey, Intramolecular hydrogen transfer in mass spectra. III. Rearrangements involving the loss of small neutral molecules, *Chem. Rev.* **1975**, *75*, 693.
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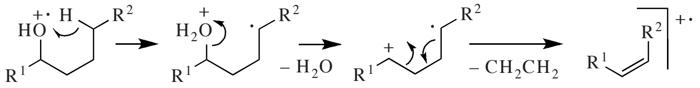
8.8 Alcohols, Ethers, and Related Compounds [1,2]

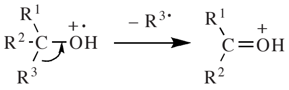
8.8.1 Alcohols and Phenols

Aliphatic Alcohols [3]

Fragmentation: Elimination of water from $M^{+\cdot}$ and from fragments. Strong for primary alcohols. If an aliphatic H atom can be transferred in a 6-ring process, it is involved in the water elimination in 90% of the investigated cases. If a CH_2CH_2 group is attached to the O-bearing C atom, water elimination is often followed by loss of ethylene. Water elimination is dominant for long-chain alcohols, rendering their spectra similar to those of alkenes.



Cleavage of bonds next to the OH-bearing C atom to form oxonium ions, then elimination of water and of alkenes. The α-cleavage is often dominant. Usually, its importance increases with increasing branching at the α-carbon atom. The larger substituent is lost most readily.

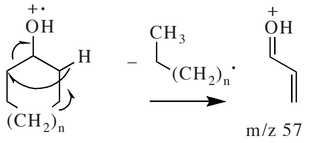


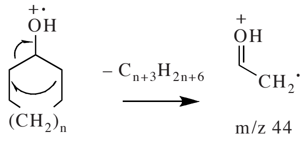
m/z 31 for primary alcohols (R^1, R^2: H)
m/z 30 + R^1 (45, 59, 73, ...) for secondary alcohols (R^2: H)
m/z 29 + R^1 + R^2 (59, 73, 87, ...) for tertiary alcohols

Consecutive H_2O and alkene eliminations in longer-chain primary alcohols lead to $[M-46]^{+\cdot}$, $[M-74]^{+\cdot}$, $[M-102]^{+\cdot}$, ... In particular, branched alcohols frequently show a typical series of fragments at $[M-15]^+$, $[M-18]^{+\cdot}$, and $[M-33]^+$.
Ion series: Dominant alkene ions corresponding to C_nH_{2n-1} (m/z 41, 55, 69, ...), C_nH_{2n} (m/z 42, 56, 70, ...), accompanied by weaker fragments, $C_nH_{2n+1}O$ (m/z 31, 45, 59, ...), with one or more local maxima in the latter series (m/z 31 dominates in primary alcohols).
Intensities: Intensive peaks in the lower mass range, local maxima among alkene-type fragment ions of the type $C_nH_{2n+1}O^+$.
Molecular ion: Mostly weak, often missing, especially in tertiary and long-chain alcohols. Indirect determination of $M^{+\cdot}$ is often possible from the fragments at $[M-15]^+$, $[M-18]^{+\cdot}$ and $[M-33]^+$. $[M+1]^+$ is often significant. In primary and secondary alcohols also $[M-1]^+$ can usually be seen. Sometimes, $[M-2]^{+\cdot}$ is formed because of oxidation to carbonyl compounds during sample introduction.

Alicyclic Alcohols

Fragmentation: Elimination of water from $M^{+\cdot}$, followed by loss of alkyl or alkenyl residues. Ring cleavage at the O-bearing C atom, followed by loss of alkyl residues after H rearrangement (see scheme).
Ion series: Alkene hydrocarbon fragments C_nH_{2n-1} (m/z 41, 55, 69, ...), C_nH_{2n-3} (m/z 39, 53, 67, 81, ...), and unsaturated O fragments, $C_nH_{2n-1}O$ (m/z 43, 57, 71, ...), as well as acetaldehyde and its homologues (m/z 44, 58, 72, ...).

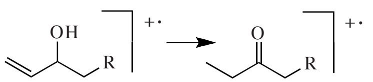




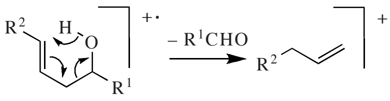
Intensities: Local maxima evenly distributed over the whole mass range.
Molecular ion: Usually weak but in contrast to aliphatic alcohols practically never missing. $[M+1]^+$ typically contains a significant amount of $[M+H]^+$.

Unsaturated Aliphatic Alcohols [3]

Allyl alcohols: The spectra are similar to those of the corresponding carbonyl compounds, which are (partly) formed by double H rearrangement of $M^{+\cdot}$.



γ,δ-Unsaturated alcohols: Aldehyde elimination through a McLafferty-type rearrangement.



Vicinal Glycols

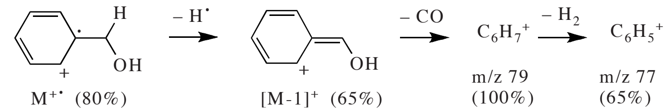
Fragmentation: Cleavage of bonds next to the OH-bearing C atom (α-cleavage) dominates. Preferable fragmentation of the C–C bond between the two oxygens, the charge remaining predominantly on the larger fragment. Water elimination from these fragments, but scarcely from $M^{+\cdot}$.
Ion series: Saturated and unsaturated aliphatic ions (m/z 43, 57, 71, … and 41, 55, 69, …) and intensive peaks from O-containing saturated rests (m/z 45, 59, 73, …).
Intensities: Dominant peaks for the products of α-cleavages and their dehydrated derivatives.
Molecular ion: Weak.

Phenols

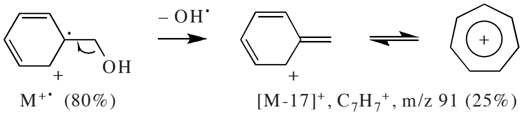
Fragmentation: Decarbonylation (Δm 28) and loss of $CHO^\cdot$ (Δm 29) followed by elimination of acetylene. An important fragment of alkyl derivatives is $[M-1]^+$, as is $[M-15]^+$ if at least two alkyl carbons are present (dimethyl or ethyl). Elimination of CO from the primary fragments. $[M-18]^{+\cdot}$ mainly with *ortho*-alkylphenols. In derivatives with a longer alkyl chain, benzylic cleavage and alkene elimination (McLafferty rearrangement) are the dominant primary fragmentation processes. The fragments then lose CO (Δm 28).
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …). The presence of some m/z 55 (C_3H_3O) is common. A peak at m/z 69 (O≡CCH=C=O) is characteristic of 1,3-dihydroxy substitution.
Intensities: Dominant peaks in the higher mass range.
Molecular ion: Dominant, no tendency to form $[M+H]^+$; $[M-1]^+$ is weak.

Benzyl Alcohols

Fragmentation: Loss of H˙ and consecutive elimination of CO (Δm 28) to give a protonated benzene molecule, which further loses H_2.



Elimination of OH˙ (Δm 17) to yield the tropylium cation is the second important fragmentation path:



Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Dominant peaks for the products described under *Fragmentation*. For benzyl alcohol decreasing in the sequence of $[M\text{-}29]^+$, $M^{+\cdot}$, $[M\text{-}1]^+$, $[M\text{-}31]^+$, $[M\text{-}17]^+$.
Molecular ion: Strong.

8.8.2 Hydroperoxides

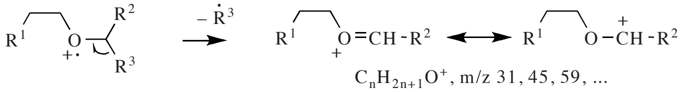
Aliphatic Hydroperoxides [4]

Fragmentation: Most pronounced is the loss of the hydroperoxy radical $HO_2^{\cdot}$ (Δm 33), especially when a tertiary alkyl cation is formed. Important, in decreasing order, is loss of H_2O_2 (Δm 34), H_2O (Δm 18), HO˙ (Δm 17), and O (Δm 16).
Ion series: Mainly saturated and unsaturated alkyl fragments, C_nH_{2n+1} (m/z 43, 57, 71, ...) and C_nH_{2n-1} (m/z 41, 55, 69, ...). The oxygen-indicating fragment at m/z 31 and its homologues are always present.
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Weak.

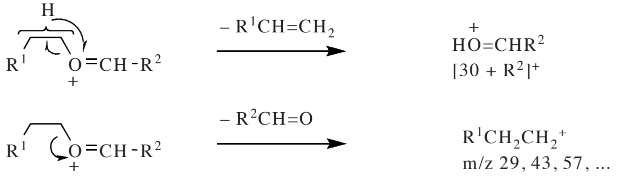
8.8.3 Ethers

Aliphatic Ethers [5,6]

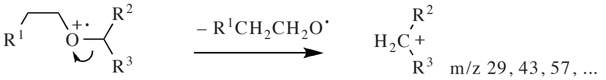
Fragmentation: Homolysis of the C–C bond next to the O atom to yield oxygen-containing fragments. Preferably, the bond at the highest substituted C atom breaks and the larger alkyl group is lost.



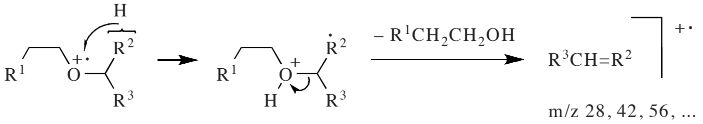
This homolysis is followed by the elimination of alkenes, aldehydes, or, less importantly, of water.



As a competing process, especially with increasing molecular weight, heterolysis at the O atom takes place to yield strong alkyl ion signals. The larger as well as the branched alkyl rests are fragmented preferably. The base peak often arises from heterolysis of the C–O bond.



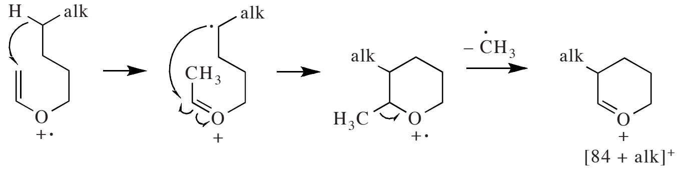
In contrast to the H_2O elimination from alcohols, the H transfer involved in the elimination of RCH_2CH_2OH from ethers is nonspecific.



Ion series: Alkyl fragments, C_nH_{2n+1} (m/z 29, 43, 57, …), with maxima due to cleavage of the C–O bond. Alkene ion series, C_nH_{2n} (m/z 28, 42, 56, …), due to elimination of alcohol. Oxygen-containing fragments, $C_nH_{2n+1}O$ (m/z 31, 45, 59, …), with maxima due to cleavage of the C–C bond next to the oxygen. O
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Significant or weak. Decreasing with increasing chain length and branching.

Unsaturated Ethers [7]

Fragmentation of vinylic and acetylenic alkyl ethers: Dominant homolysis of the alkyl C–C bond next to the O atom on the saturated side, leading to $C_3H_5O^+$ (m/z 57) for vinylic and $C_3H_3O^+$ (m/z 55) for acetylenic ethers of primary aliphatic alcohols. For alkyl ($C_{n>5}$) vinyl ethers, ethanol elimination after triple H transfer. $[M-15]^+$ in vinyl ethers predominantly by elimination of the vinyl CH_2 after H rearrangement.



Fragmentation of allylic ethers: Heterolysis of both C–O bonds, leading to strong $C_3H_5^+$ (m/z 41) and alkyl (m/z 29, 43, 57, …) cations. Formation of ionized allyl alcohol ($C_3H_6O^{+\bullet}$, m/z 58) by nonspecific H transfer from the alkyl rest. In allylic

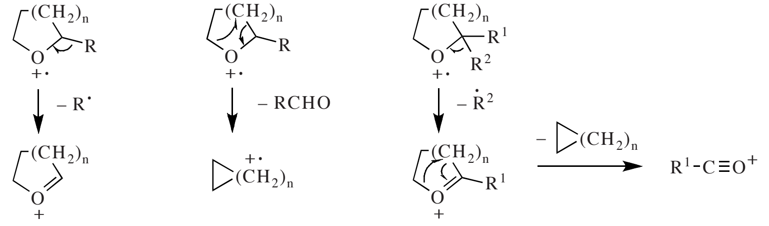
and propargylic ethers, no cleavage of the C–C bond next to the O atom of the alkenyl group occurs. Hence, loss of vinyl or acetylenyl cannot be observed.
Ion series: $C_nH_{2n}O$ (m/z 44, 58, 72, ...) for alkenyl alkyl ethers and $C_nH_{2n-2}O$ (m/z 42, 56, 70, ...) for dialkenyl ethers. Unsaturated aliphatic (C_nH_{2n-1}; m/z 41, 55, 69, ...) as well as saturated aliphatic and unsaturated oxygen-containing fragments (C_nH_{2n+1} and $C_nH_{2n-1}O$; m/z 43, 57, 71, ...).
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Weak to medium, very weak for acetylenic ethers.

Alkyl Cycloalkyl Ethers

Fragmentation of methyl ethers of cycloalkanols with > 3 C atoms: After primary cleavage of the ring C–C bond next to the O atom, the prominent fragments formed are $CH_3OCH{=}CH_2^{+\cdot}$ (m/z 58) and, for alicyclics with > 4 C atoms, $CH_3O{=}CHCH{=}CH_2^+$ (m/z 71, rearrangement in analogy to that observed for cycloalkanols). Loss of methanol to give hydrocarbon fragments, C_nH_{2n-2} (m/z 54, 68, 82, ...).
Fragmentation of ethyl and higher alkyl ethers of cycloalkanols with > 3 C atoms: Alkene elimination to yield the protonated cycloalkanol (m/z 72, 86, 100, ...) and heterolytic cleavage of the C–O bond to give dominating cycloalkyl ions (m/z 69, 83, ...).
Ion series: Besides the fragments already mentioned, mainly unsaturated hydrocarbon fragments (C_nH_{2n-1}, m/z 27, 41, 55, 69, ...).
Intensities: The above mentioned fragments dominate the spectrum.
Molecular ion: Weak or intermediate.

Cyclic Ethers

Fragmentation: Primary ring cleavage at C–C bonds next to the O atom, followed by loss of CH_2O (Δm 30), H_2O (Δm 18), or alkyl (Δm 15, 29, ...). Elimination of $H^\cdot$ to give $[M-1]^+$, followed by CO elimination (Δm 28) to $[M-29]^+$. When α-substituted, dominant loss of substituents, followed by water elimination. Formation of acyl cation if two α-substituents are present.



Ion series: Mainly ions of the alkene type. Weak saturated, oxygen-containing fragments (m/z 31, 45, ...).
Intensities: Intensive peaks evenly distributed over the whole mass range.
Molecular ion: Often significant but sometimes weak, especially when α-substituted. Intensity of $[M-1]^+$ usually comparable to that of $M^{+\cdot}$ if no α substituent is present.

Methoxybenzenes

Fragmentation: Loss of methyl radical, followed by decarbonylation to $[M\text{-}43]^+$; elimination of formaldehyde (Δm 30) from $M^{+\cdot}$ or from primary fragments.
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …).
Intensities: Intensive peaks in the $M^{+\cdot}$ region.
Molecular ion: Strong.

Alkyl Aryl Ethers [8]

Fragmentation: Commonly dominating alkene elimination to give the corresponding phenol ion (nonspecific hydrogen migration), followed by decarbonylation. In the case of aryl methyl ethers, loss of CH_2O from $M^{+\cdot}$ or from primary fragments as well as $CH_3^{\cdot}$ elimination followed by decarbonylation.
Ion series: Mostly aromatic fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …).
Intensities: Usually maximum at the mass of the corresponding phenol. Otherwise, intensive peaks mainly concentrated in the high and medium mass range.
Molecular ion: Strong.

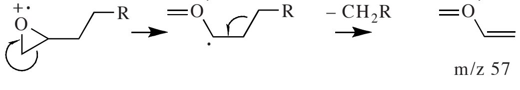
Aromatic Ethers

Fragmentation: Loss of $H^{\cdot}$ (Δm 1), CO (Δm 28), and $CHO^{\cdot}$ (Δm 29) from $M^{+\cdot}$. Cleavage at the C–O bond and decarbonylation of the resulting product, followed by dehydrogenation.
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …).
Intensities: Intensive peaks mainly in the $M^{+\cdot}$ region.
Molecular ion: Strong.

8.8.4 Aliphatic Epoxides [9]

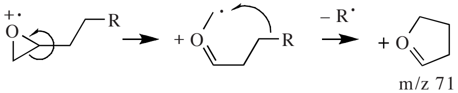
Fragmentation: The most important primary fragmentation is the cleavage of C–C bonds next to the O atom (α-cleavage), resulting in complex degradation due to the related multiple choice and extensive secondary rearrangements. The products allow mass-spectrometric localization of double bonds after epoxidation.

Due to ring opening prior to fragmentation, β-cleavage is as relevant as the α-cleavage.



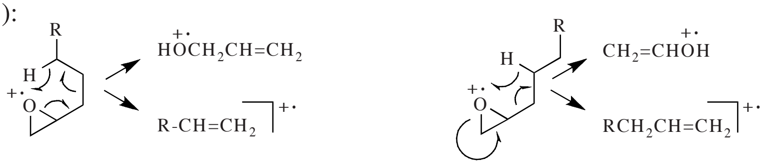
m/z 57

γ-Cleavage is the most important fragmentation mechanism, especially in terminal epoxides:

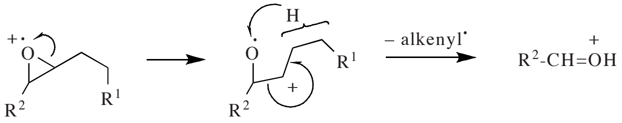


m/z 71

Mainly in terminal epoxides, rearrangement with alkene elimination, formally leading to alkene-$OH^{+\cdot}$ ($C_nH_{2n}O$, m/z 44, 58, 72, ...) and alkene$^{+\cdot}$ (C_nH_{2n}, m/z 28, 42, 56, ...):



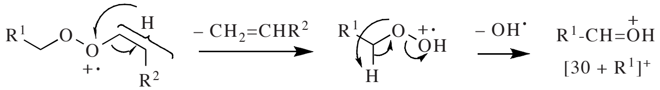
Mainly in nonterminal epoxides, transannular cleavage with H transfer and elimination of an alkenyl radical, leading to $C_nH_{2n+1}O$ fragments (m/z 45, 59, 73, ...):

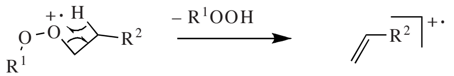


Ion series: Mixed, not characteristic.
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Usually weak.

8.8.5 Aliphatic Peroxides [4]

Fragmentation: Alkene elimination to give hydroperoxide radical cations and hydroperoxide elimination to yield alkene radical cations (dominating if larger alkyl groups are present). Alkene elimination can be followed by loss of OH·, resulting in products that formally correspond to those obtained by O–O cleavage, which probably is not a one-step process:





Elimination of O· or O_2 may occur in cyclic peroxides. *tert*-Butyl peroxides predominantly eliminate *tert*-butyl-OO· to give $[M-89]^+$.
Ion series: Saturated or unsaturated alkyl groups (C_nH_{2n+1}, m/z 29, 43, 57, ...; C_nH_{2n-1}, m/z 27, 41, 55, ...) and alkenyl ions (C_nH_{2n}, m/z 28, 42, 56, ...) dominate. The fragment at m/z 31 and sometimes its homologues indicate the presence of oxygen.
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Weak to moderate.

8.8.6 References

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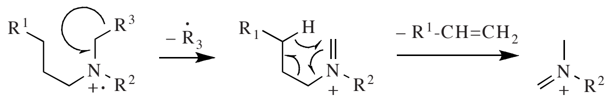
8.9 Nitrogen Compounds [1,2]

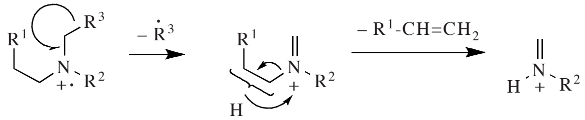
8.9.1 Amines

Saturated Aliphatic Amines [3]

Fragmentation: Dominating loss of alkyl residues by cleavage of the C–C bond next to the N atom ("N-cleavage"). Larger substituents are eliminated preferably. When a γ-H is available, subsequent elimination of alkenes by McLafferty-type reactions:



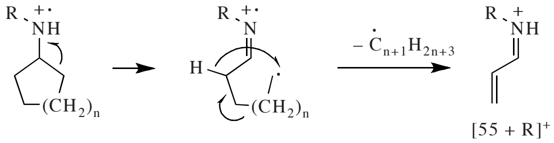
Otherwise, unspecific H transfer onto the N atom:

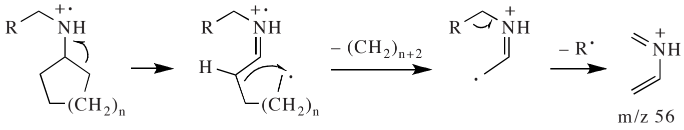


NH_3, RNH_2, and RR'NH eliminations from primary, secondary, and tertiary amines, respectively, are negligible except from some multifunctional compounds (e.g., diamines and phenyl-phenoxy-substituted amines).
Ion series: Even-mass fragments of the type $C_nH_{2n+2}N$ (m/z 30, 44, 58, 72, 86, …).
Intensities: Mainly peaks in the low mass range. Dominating base peak from "N-cleavage" at $[28 + m(R^1) + m(R^2) + m(R^4) + m(R^5)]^+$ for $R^1R^2R^3CNR^4R^5$ (e.g., m/z 30 for RCH_2NH_2, m/z 44 for RCH_2NHCH_3, m/z 58 for $RCH_2N(CH_3)_2$, and m/z 86 for $RCH_2N(CH_2CH_3)_2$). Local maximum at m/z 86 ($C_5H_{12}N^+$) for *n*-alk–NH_2 (protonated piperidine, 6-membered ring).
Molecular ion: Usually weak or absent, especially if the α-C atom is substituted. Decreasing intensity with increasing molecular weight. Tendency to protonate to $[M+H]^+$. Odd mass for odd number of N atoms in the molecule.

Cycloalkylamines

Fragmentation: The most important primary reaction is the ring cleavage next to the N atom, followed by H rearrangement and loss of an alkyl residue. Some elimination of amine, R^1R^2NH.

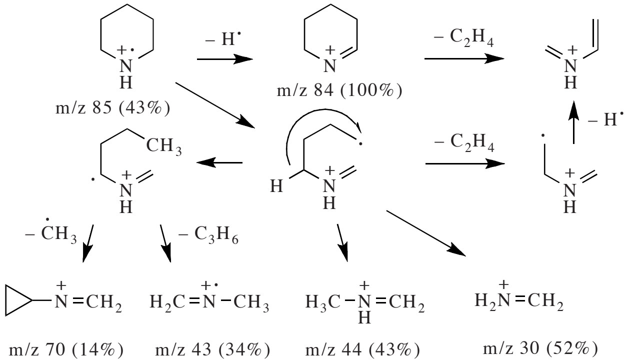




Ion series: Even-mass fragments of the type $C_nH_{2n}N$ (m/z 42, 56, 70, 84, ...).
Intensities: Intensive local maxima evenly distributed over the whole mass range.
Molecular ion: Usually significant. Odd mass for odd number of N atoms in the molecule.

Cyclic Amines

Fragmentation: Dominating primary reaction is the cleavage of C–C bonds next to N, resulting in the loss of substituents next to N or in primary ring cleavage. Primary ring cleavage is followed by H rearrangement and loss of alkenes or alkyl groups. The most important primary fragmentation for substituted cyclic amines is the loss of substituents at C atoms next to N.
Piperidine:



Ion series: Even-mass fragments of the type $C_nH_{2n}N$ (m/z 42, 56, 70, 84, ...) and $C_nH_{2n+2}N$ (m/z 30, 44, 58, ...) as well as odd-mass fragments of the type $C_nH_{2n+1}N$ (m/z 43, 57, 71, 85, ...).
Intensities: Intensive local maxima evenly distributed over the whole mass range if no substituent is bonded to the C atom next to N. Otherwise, dominating maxima by loss of such substituents.
Molecular ion: Significant or strong if no substituent is bonded to the C atom next to N; otherwise, weak. Tendency to form $[M\text{-}H]^+$. Odd mass for odd number of N atoms in the molecule.

N

Piperazines

Fragmentation: As for cyclic amines, enhanced primary ring cleavage at C–C bonds next to the N atom.
Ion series: Even-mass fragments of the type $C_nH_{2n}N$ (m/z 42, 56, 70, 84, ...) and $C_nH_{2n+2}N$ (m/z 30, 44, 58, ...) as well as odd-mass series of the type $C_nH_{2n+1}N$ (m/z 43, 57, 71, 85, ...).
Intensities: Intensive local maxima evenly distributed over the whole mass range if no substituent is bonded to the C atom next to N. Otherwise, dominating maxima by loss of such substituents.
Molecular ion: Significant or strong if no substituent is bonded to the C atom next to N; otherwise, weak. Tendency to form $[M\text{-}H]^+$. Odd mass for odd number of N atoms in the molecule.

Aromatic Amines

Fragmentation: Dominating cleavage of alkyl bond at N-bearing C atom ("N-cleavage") followed by alkene elimination if aliphatic substituents with $C_{n\geq2}$ are present. Otherwise, loss of $H^{\cdot}$ from primary and secondary anilines and benzylic amines. Loss of HCN from $M^{+\cdot}$ or from fragments. A local maximum at m/z 42 is typical of an aromatically bonded dimethylamino group.
Ion series: Aromatic hydrocarbon fragments (C_nH_n and $C_nH_{n\pm1}$; m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Dominating maxima by "N-cleavage" and following alkene loss if aliphatic substituents with $C_{n>1}$ are present.
Molecular ion: Abundant if no aliphatic substituents with more than one C atom are present, otherwise, medium or weak. No tendency to protonate. In primary and secondary aromatic and benzylic amines, $[M-H]^+$ is important. Odd mass for odd number of N atoms in the molecule.

8.9.2 Nitro Compounds

Aliphatic Nitro Compounds

Fragmentation: Loss of $NO^{\cdot}$ (Δm 30), $NO_2^{\cdot}$ (Δm 46), and HNO_2 (Δm 47) as well as the formation of some m/z 30 as N indicator. Spectra with only few characteristic features.
Ion series: Mixed alkyl and alkenyl fragments, C_nH_{2n+1} (m/z 43, 57, 71, ...) and C_nH_{2n-1} (m/z 41, 55, 69, ...).
Intensities: Dominant peaks in the lower mass range.
Molecular ion: Weak or missing. Odd mass for odd number of N atoms in the molecule.

N

Aromatic Nitro Compounds

Fragmentation: Loss of O (Δm 16), $NO^{\cdot}$ (Δm 30, followed by elimination of CO, Δm 28), and $NO_2^{\cdot}$ (Δm 46) from $M^{+\cdot}$ or from a major primary cleavage product. Extensive rearrangement of the functional group to a nitroso ester.
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks mainly in the upper mass range.
Molecular ion: Strong. Odd mass for odd number of N atoms in the molecule.

8.9.3 Diazo Compounds and Azobenzenes

Diazo Compounds [4,5]

Diazonium: Because of the low volatility of diazo compounds, their electron impact mass spectra show thermal decomposition products. These are formed by loss of N_2

(e.g., a diazonium chloride gives rise to the corresponding aromatic chloro compound). From a phenyl diazonium *ortho*-carboxylate zwitterion, biphenylene is formed as dimerization product.
Diazomethane and derivatives: $M^{+\cdot}$ is strong except when catalytic decomposition occurs on metal surfaces of the inlet system. Loss of N_2 is a dominant reaction of diazomethane and diazoketones.

Azobenzenes

Fragmentation: Cleavage at the azo group followed by loss of N_2, giving rise to the dominant base peak.
Ion series: Aromatic fragments (C_nH_n, $C_nH_{n\pm1}$; m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Dominant $M^{+\cdot}$ and azo cleavage products.
Molecular ion: Strong. Odd mass for odd number of N atoms in the molecule.

8.9.4 Azides

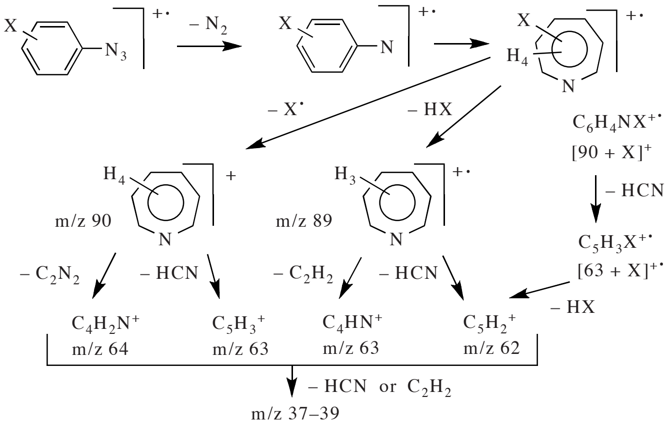
Aliphatic Azides [6]

Fragmentation: $[M\text{-}42]^+$ ($N_3^\cdot$ elimination) or $[M\text{-}28]^{+\cdot}$ (N_2 elimination) dominant in most cases. The spectra are similar to those of the corresponding aliphatic compounds.
Ion series: Aliphatic hydrocarbon series.
Intensities: Dominant peaks in the lower mass range, as in aliphatic compounds.
Molecular ion: Absent or weak. Odd mass for odd number of N atoms in the molecule.

Aromatic Azides [7]

N

Fragmentation: In most cases, $[M\text{-}28]^{+\cdot}$ (N_2 elimination) is the base peak. The next step is the elimination of HCN (Δm 27) or acetylene (Δm 26), or, if there is a substituent X on the ring, of $X^\cdot$ or HX.



Ion series: Aromatic hydrocarbon fragments (C_nH_n and $C_nH_{n\pm1}$; m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Dominant peaks in the higher mass range; $[M-28]^{+\cdot}$ (N_2 elimination) and $[M-55]^{+\cdot}$ (N_2 and HCN elimination) are the most intensive peaks.
Molecular ion: Weak. Odd mass for odd number of N atoms in the molecule.

8.9.5 Nitriles and Isonitriles

Aliphatic Nitriles (R–CN) [4]

Fragmentation: Elimination of alkyl radicals to give $(CH_2)_nCN^+$ (m/z 40, 54, 68, ...). McLafferty rearrangement yielding $CR_2{=}C{=}NH^{+\cdot}$ (m/z 41 for R: H). In most cases, C–CN cleavage and HCN elimination are not significant reactions. Complex rearrangements in unsaturated nitriles if other functional groups are present.
Ion series: Saturated and unsaturated alkyl ions mainly in the lower mass range (C_nH_{2n+1} and C_nH_{2n-1}; m/z 29, 43, 57, ... and 27, 41, 55, ...). Rearrangement products corresponding to $C_nH_{2n-1}N$ contribute, to a significant extent, to the ion series m/z 41, 55, 69, ... For alkyl chains with $C_{n>5}$, dominating $(CH_2)_nCN^+$ (i.e., $C_nH_{2n-2}N$, m/z 82, 96, 110, ..., probably with a cyclic structure).
Intensities: Intensive peaks due to the above mentioned ions.
Molecular ion: Weak or missing. Both $[M+H]^+$ and $[M-H]^+$ are usually more intensive than $M^{+\cdot}$. In some aliphatic nitriles, $[M+2H]^{+\cdot}$ is as intensive as $M^{+\cdot}$. Odd mass for odd number of N atoms in the molecule.

Aromatic Nitriles (R–CN)

N

Fragmentation: Consecutive elimination of HCN and acetylene.
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks in the $M^{+\cdot}$ region.
Molecular ion: Dominant intensity, often base peak. In contrast to aliphatic and benzylic nitriles, $[M-1]^+$ is usually not important. Odd mass for odd number of N atoms in the molecule.

Aliphatic Isonitriles (R–NC)

Fragmentation: In general, the spectra are similar to those of the corresponding nitriles. The most important difference lies in the loss of $CN^\cdot$ (Δm 26) and the higher probability of losing HCN (Δm 27). Further important fragmentations are the elimination of alkyl radicals to give $(CH_2)_nCN^+$ ions and the McLafferty rearrangement to yield $CR_2{=}N{=}CH^{+\cdot}$ (m/z 41 for R: H).
Ion series: Saturated and unsaturated alkyl ions mainly in the lower mass range (C_nH_{2n+1}, m/z 29, 43, 57, ... and C_nH_{2n-1}, m/z 27, 41, 55, ...). Rearrangement products corresponding to $C_nH_{2n-1}N$ contribute, to a significant extent, to the ion series of m/z 41, 55, 69, ...
Intensities: Intensive peaks in the lower mass range.

Molecular ion: Weak, decreasing with increasing chain length and degree of branching. Both $[M+H]^+$ and $[M-H]^+$ can be stronger than $M^{+\cdot}$. Odd mass for odd number of N atoms in the molecule.

Aromatic Isonitriles (R–NC) [4]

Fragmentation: Dominant loss of HCN ($[M-27]^{+\cdot}$). In methylphenyl and benzyl isocyanides also formation of isocyanotropylium ion, $[M-1]^+$, followed by loss of HCN to $[M-28]^+$.
Ion series: Aromatic (C_nH_n and $C_nH_{n\pm1}$; m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks in the higher mass range.
Molecular ion: Dominant; base peak for phenyl isocyanide. Odd mass for odd number of N atoms in the molecule.

8.9.6 Cyanates, Isocyanates, Thiocyanates, and Isothiocyanates

Aliphatic Cyanates (R–OCN) [8]

Fragmentation: Spectra often very similar to those of the corresponding isocyanates (see below). Cleavage of the C–C bond next to O, with the charge remaining on CH_2OCN (m/z 56) for short-chain cyanates and preferably on the alkyl substituent if it has a $C_{n>2}$ chain (m/z 29, 43, 57, ...). Cleavage of the C–O bond with H rearrangement to give $HCNO^{+\cdot}$ (m/z 43) or alkene$^{+\cdot}$ (m/z 42, 56, 70, ...). For cyanates with $C_{n>5}$ substituents, alkene elimination yields m/z 99.
Ion series: Saturated and unsaturated alkyl cations (C_nH_{2n+1}, m/z 29, 43, 57, ... and C_nH_{2n-1}, m/z 27, 41, 55, ...). Alkene radical cations (C_nH_{2n}, m/z 42, 56, 70, ...) together with isobaric ions of the composition $C_nH_{2n}NCO$.
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Usually weak or absent. $[M-H]^+$ is often more intensive. Odd mass for odd number of N atoms in the molecule.

N

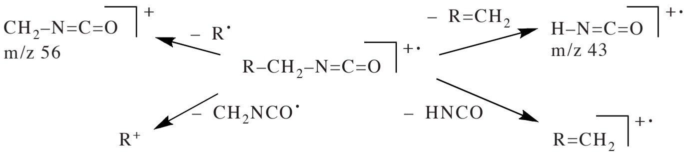
Aromatic Cyanates (R–OCN) [8]

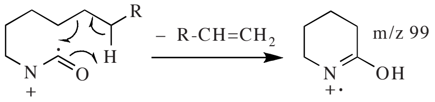
Fragmentation: Loss of $OCN^{\cdot}$ (Δm 42) or, to a lesser extent, of CO (Δm 28), with subsequent HCN elimination (Δm 27).
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks in the higher mass range.
Molecular ion: Strong. Odd mass for odd number of N atoms in the molecule.

Aliphatic Isocyanates (R–NCO) [8]

Fragmentation: Spectra often very similar to those of the corresponding cyanates. Cleavage of the C–C bond next to N, the charge remaining on the CH_2NCO (m/z 56) for short-chain isocyanates and preferably on the alkyl substituent for compounds with a $C_{n>2}$ chain (m/z 29, 43, 57, ...). Cleavage of the C–N bond with H

rearrangement to give $HNCO^{+\cdot}$ (m/z 43) or alkene$^{+\cdot}$ (m/z 42, 56, 70, …) ions. For isocyanates with $C_{n>5}$ alkyl chains, alkene elimination, yielding m/z 99.





Ion series: Saturated and unsaturated alkyl cations (C_nH_{2n+1}, m/z 29, 43, 57, … and C_nH_{2n-1}, m/z 27, 41, 55, …). Alkene radical cations (C_nH_{2n}, m/z 42, 56, 70, …) together with isobaric ions of the composition $C_nH_{2n}OCN$.
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Usually weak or absent. $[M-H]^+$ is often more intensive. Odd mass for odd number of N atoms in the molecule.

Aromatic Isocyanates (R–NCO) [8]

Fragmentation: Consecutive elimination of CO (Δm 28) and HCN (Δm 27). In contrast to aromatic cyanates, practically no elimination of $NCO^{\cdot}$ (Δm 42).
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …).
Intensities: Intensive peaks in the higher mass range.
Molecular ion: Dominating; base peak for phenyl isocyanate. Odd mass for odd number of N atoms in the molecule.

N

Aliphatic Thiocyanates (R–SCN) [8]

Fragmentation: Elimination of HCN (Δm 27) followed by loss of an alkyl group. The cleavage of the C–C bond next to SCN is unimportant except in short-chain thiocyanates.
Ion series: Saturated and unsaturated alkyl cations (C_nH_{2n+1}, m/z 29, 43, 57, … and C_nH_{2n-1}, m/z 27, 41, 55, …).
Intensities: Intensive peaks in the lower mass range.
Molecular ion: Weak. Decreasing with increasing chain length and degree of branching; absent from the spectrum of hexyl thiocyanate. Odd mass for odd number of N atoms in the molecule. Both $[M+H]^+$ and $[M-H]^+$ are detectable. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Aromatic Thiocyanates (R–SCN) [8]

Fragmentation: The most important fragmentation is the elimination of $SCN^{\cdot}$ (Δm 58). Further elimination reactions are loss of $CN^{\cdot}$ (Δm 26), HCN (Δm 27),

and CS (Δm 44).
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...). Weak signal at m/z 45 (CHS^+) indicates sulfur.
Intensities: Intensive peaks in the higher mass range.
Molecular ion: Dominant; base peak in phenyl thiocyanate. Odd mass for odd number of N atoms in the molecule. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Aliphatic Isothiocyanates (R–NCS) [8]

Fragmentation: Cleavage of the C–C bond next to NCS, leading to m/z 72 (CH_2NCS) or to its homologues if the α-C atom is substituted. Loss of the alkyl residue with concomitant double hydrogen rearrangement to yield H_2NCS^+ (m/z 60). With a $C_{n>4}$ alkyl chain, loss of $SH^\cdot$ (Δm 33). With $C_{n>5}$ alkyl chains, loss of alkene leading to m/z 115, probably according to the mechanism shown for aliphatic isocyanates.
Ion series: Mainly saturated and unsaturated alkyl cations (C_nH_{2n+1}, m/z 29, 43, 57, ... and C_nH_{2n-1}, m/z 27, 41, 55, ...). Signal for CH_2NCS^+ (m/z 72) or its homologues (m/z 86, 100, 114, ...) if the α-C atom is substituted.
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Medium to weak, decreasing with increasing chain length and degree of branching. More intensive than in the corresponding thiocyanates; 1% for hexadecyl isothiocyanate. Both $[M+H]^+$ and $[M-H]^+$ are relevant. Odd mass for odd number of N atoms in the molecule. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Aromatic Isothiocyanates (Ar-NCS) [8]

N

Fragmentation: Dominant loss of $NCS^\cdot$ (Δm 58). In contrast to aromatic thiocyanates, the loss of HCN (Δm 27) or CS (Δm 44) leads to very weak fragments only.
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...). Weak signal at m/z 45 (CHS^+) indicates sulfur.
Intensities: Intensive peaks in the higher mass range.
Molecular ion: Dominant; base peak in phenyl isothiocyanate. Odd mass for odd number of N atoms in the molecule. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

8.9.7 References

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8.10 Sulfur Compounds [1]

8.10.1 Thiols

Aliphatic Thiols [2]

Fragmentation: Elimination of H_2S (Δm 34; or SH, Δm 33, from secondary thiols) followed by loss of alkenes; consecutive losses of ethylene from unbranched thiols. Cleavage of the α,β-C–C bond (next to the SH group) leads to CH_2SH^+ (m/z 47). Note that this fragment also occurs in secondary and tertiary thiols. The S atom is poorer than N, but better than O, at stabilizing such a fragment. Cleavage at the next C–C bonds leads to signals at m/z 61, 75, and 89. In secondary and tertiary thiols, prominent fragments are formed by loss of the largest α-alkyl group.
Ion series: Dominant alkenyl fragments (C_nH_{2n-1}, m/z 41, 55, 69, …) and smaller aliphatic fragments (C_nH_{2n+1}, m/z 43, 57, 71, …). Sulfur-containing aliphatic fragments: $C_nH_{2n+1}S$ (m/z 47, 61, 75, 89, …). Often significant sulfur-indicating fragments: HS^+, $H_2S^{+\cdot}$, H_3S^+, and CHS^+ (m/z 33, 34, 35, and 45).
Intensities: More intensive peaks in the lower mass range, mostly of the alkene type. Characteristic local maxima from S-containing fragments, $C_nH_{2n+1}S$ (m/z 47, 61, 75, 89, …). In *n*-alkyl thiols, the intensity of the signal at m/z 61 is roughly half that of m/z 47; the signal at m/z 89 is more intensive than that at m/z 75, presumably because it is stabilized by cyclization.
Molecular ion: Relatively strong except for higher tertiary thiols. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Aromatic Thiols [2]

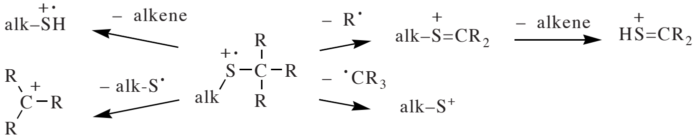
Fragmentation: CS elimination from $M^{+\cdot}$ and $[M-1]^+$, yielding $[M-44]^{+\cdot}$ and $[M-45]^+$. $HS^\cdot$ elimination from $M^{+\cdot}$ to give $[M-33]^+$.
Ion series: HCS^+ (m/z 45) is characteristic besides the aromatic fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …).
Intensities: Intensive peaks in the higher mass range.
Molecular ion: Usually dominating; base peak in thiophenol. $[M-1]^+$ is usually strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

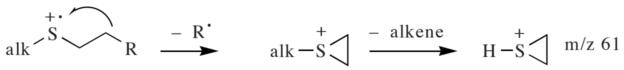
S

8.10.2 Sulfides and Disulfides

Aliphatic Sulfides [1]

Fragmentation: Loss of alkyl radicals by cleavage of the C–C bond next to S (the largest group being lost preferably) and of the C–S bond, followed by alkene and H_2S elimination. Alkene elimination from $M^{+\cdot}$ to form the corresponding thiol ions. In contrast to thiols and cyclic sulfides, no H_2S or $HS^\cdot$ elimination from $M^{+\cdot}$.





In general, the H rearrangements are nonspecific. The transfer of secondary H predominates over that of primary H.
Ion series: Sulfur-containing aliphatic fragments, $C_nH_{2n+1}S$ (m/z 47, 61, 75, 89, ...). The hydrocarbon fragments may dominate in long-chain sulfides.
Intensities: Intensive peaks in the lower mass range. Characteristic local maxima from S-containing fragments, $C_nH_{2n+1}S$ (m/z 47, 61, 75, 89, ...).
Molecular ion: Usually strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Alkyl Vinyl Sulfides

Fragmentation: Loss of alkyl radicals (Δm 15, 29, 43, ...). Elimination of thioethanol (Δm 62) after triple H rearrangement. Dominant m/z 60 ($CH_3CH{=}S^{+\cdot}$) accompanied by m/z 61 ($CH_3CH_2S^+$).
Ion series: Sulfur-containing unsaturated aliphatic fragments, $C_nH_{2n-1}S$ (m/z 45, 59, 73, ...). Unsaturated hydrocarbon ions, C_nH_{2n} (m/z 42, 56, 70, ...) and C_nH_{2n-2} (m/z 40, 54, 68, ...)
Intensities: Intensive peaks evenly distributed over the whole mass range.
Molecular ion: Of medium intensity. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Cyclic Sulfides

Fragmentation: Primary cleavage of the C–C bond next to S, followed by rearrangements and elimination of $CH_3^{\cdot}$ (base peak for tetrahydrothiapyrane) and $C_2H_5^{\cdot}$. In tetrahydrothiophene, $[M-1]^+$ is also significant. $HS^{\cdot}$, H_2S, and C_2H_4 elimination from $M^{+\cdot}$.
Ion series: Sulfur-containing aliphatic fragments with one degree of unsaturation, $C_nH_{2n-1}S$ (m/z 45, 59, 73, 87, 101, ...), m/z 87 being of special dominance.
Intensities: Overall distribution of peaks maximizing in the low mass range due to S-containing fragments, $C_nH_{2n-1}S$ (m/z 45, 59, 73, 87, ...).
Molecular ion: Very strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Aromatic Sulfides [2]

Fragmentation: Loss of CS (Δm 44) and of $HS^{\cdot}$ (Δm 33) from $M^{+\cdot}$.
Ion series: HCS^+ (m/z 45) is characteristic besides the aromatic fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).

Intensities: Intensive peaks mainly in the higher mass range.
Molecular ion: Strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ (4.5% relative to $M^{+\cdot}$ per S atom) and [frag+2] for S-containing fragments.

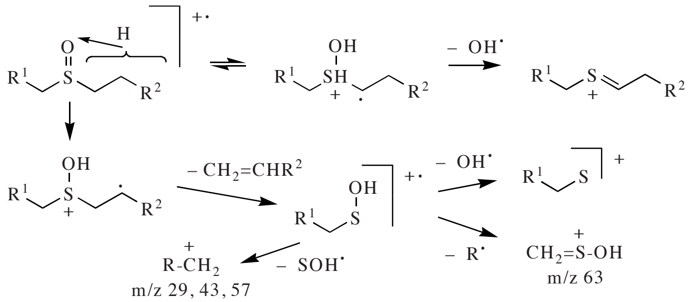
Disulfides

Fragmentation: Loss of $RSS^{\cdot}$, leading to alkyl cations and alkene elimination to give $RSSH^{+\cdot}$. Cleavage of the S–S bond with or without H rearrangements, leading to RS^{+}, $[RS\text{-}H]^{+\cdot}$, and $[RS\text{-}2H]^{+}$. Loss of one or two S with or without H atoms is a common process in cyclic, unsaturated, and aromatic disulfides.
Ion series: In saturated aliphatic disulfides, H_2S_2 and its alkyl homologues are characteristic (m/z 66, 80, 94, ...).
Intensities: Variable.
Molecular ion: Usually strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

8.10.3 Sulfoxides and Sulfones

Aliphatic Sulfoxides [4,5]

Fragmentation: Most fragments are produced after rearrangement with non-specific H transfer to the O atom and subsequent $OH^{\cdot}$ elimination to yield $[M\text{-}17]^{+}$ or alkene elimination to $[M\text{-alkene}]^{+\cdot}$, followed by $OH^{\cdot}$, $SOH^{\cdot}$ (giving alk^{+} ions), or $alk^{\cdot}$ elimination (yielding CH_2=S–OH^{+}, m/z 63).

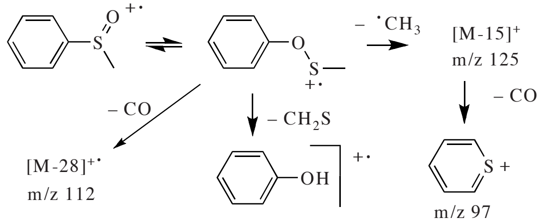


Ion series: Characteristic ion at m/z 63 (CH_2=S–OH^{+}) as well as alkyl and alkenyl fragments, C_nH_{2n+1} (29, 43, 57, 71, ...) and C_nH_{2n-1} (27, 41, 55, 69, ...).
Intensities: Intensive peaks evenly distributed over the whole mass range.
Molecular ion: Of medium intensity. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

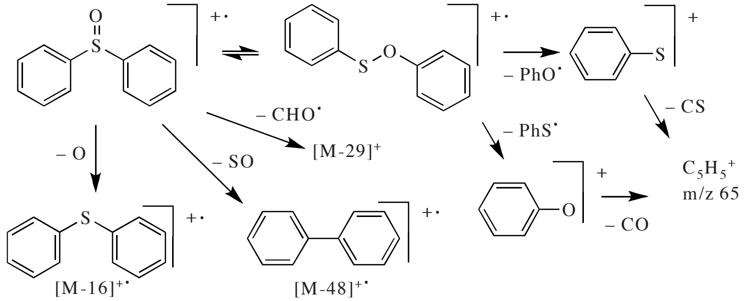
S

Alkyl Aryl and Diaryl Sulfoxides [4,5]

Fragmentation: Most fragments of methyl aryl sulfoxides are produced, after rearrangement to $CH_3S–O–ar^{+\cdot}$, by elimination of CH_2S (yielding $[M-46]^{+\cdot}$, a phenol), of CO (to $[M-28]^{+\cdot}$), and of $CH_3^{\cdot}$ (to $[M-15]^{+}$). The latter ion loses CO to give the thiapyranyl cation (m/z 97 if ar is phenyl).



The skeletal rearrangement is not relevant for the fragmentation of higher alkyl aryl sulfoxides. Here, direct cleavage of the C–S bonds and McLafferty rearrangements dominate.
For diaryl sulfoxides, elimination of SO (to give $[M-48]^{+\cdot}$) as well as of O, $OH^{\cdot}$, and $CHO^{\cdot}$ (yielding $[M-16]^{+\cdot}$, $[M-17]^{+}$, and $[M-29]^{+}$, respectively). After rearrangement to sulfenates, cleavage of the S–O bond to produce $ar–S^{+}$ and $ar–O^{+}$ ions, which further lose CS and CO, respectively, to give $C_5H_5^{+}$ (m/z 65).



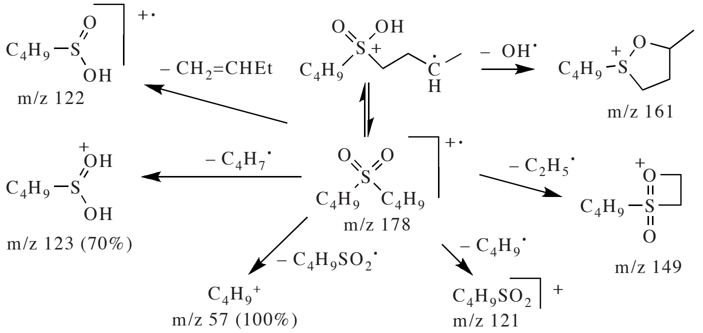
Ion series: Besides the ions described under *Fragmentation*, mainly fragments of the aromatic type, i.e., C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …), as well as O- and S-containing ions.
Intensities: Intensive peaks mainly in the high mass range.
Molecular ion: Very strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Aliphatic Sulfones [4,5]

Fragmentation: Fragmentation of the S–C bond with the charge remaining on either side. Single and double H rearrangements to give $RS(O)OH^{+\cdot}$ and $RS(OH)_2^{+}$. The probability of the double H rearrangement increases with increasing chain length. If one of the substituents is unsaturated, rearrangement to RS(O)O–alkene followed by cleavage of the S–O bond yields the ion RSO^{+}.

S

Ion series: Dominating aliphatic fragments, C_nH_{2n+1} (m/z 29, 43, 57, ...) and C_nH_{2n-1} (m/z 27, 41, 55, ...). Usually, one significant fragment corresponding to alk–S(O)OH$^{+\cdot}$ (from the series of m/z 80, 94, 108, ...) or alk–S(OH)$_2^+$ (from the series of m/z 81, 95, 109, ...) can be observed.
Intensities: Intensive peaks mainly of aliphatic fragments in the lower mass range.
Molecular ion: Weak. Characteristic ^{34}S isotope peak at [M+2]$^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).



Cyclic Sulfones [4]

Fragmentation: Dominant elimination of SO_2 (Δm 64, followed by loss of $CH_3^\cdot$), $HSO_2^\cdot$ (Δm 65, followed by loss of C_2H_4), or CH_2SO_2 (Δm 78). Weak signal at [M-17]$^+$ due to OH$^\cdot$ elimination.
Ion series: Mainly unsaturated hydrocarbon fragments, C_nH_{2n-1} (m/z 27, 41, 55, ...).
Intensities: Intensive peaks in the lower mass range.
Molecular ion: Moderate. Characteristic ^{34}S isotope peak at [M+2]$^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Alkyl Aryl Sulfones [4]

Fragmentation: Isomerization of M$^{+\cdot}$ to ar–OS(=O)alk and formation of the phenoxy ion or the phenol radical cation with H rearrangement. The migration of the aryl group depends on the type of substituents. It is facilitated by electron donators and hindered by acceptors. Mainly in substituted or unsaturated alkyl derivatives also isomerization to ar–S(=O)O–alk(ene) and formation of ar–S=O$^+$ (m/z 125 if ar is phenyl). Single and double H rearrangements to give ar–S(=O)OH$^{+\cdot}$ and ar–S(OH)$_2^+$. The probability of the double H rearrangement increases with increasing chain length. In some derivatives, SO_2 elimination from M$^{+\cdot}$ dominates. Substituents X of the alkyl group may migrate to the aryl group to yield X–ar–S=O$^+$ ions.
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...), as well as S- and O-containing aromatic fragments at higher masses.
Intensities: Intensive peaks mainly in the higher mass range.
Molecular ion: Strong. Characteristic ^{34}S isotope peak at [M+2]$^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Diaryl Sulfones [4,5]

Fragmentation: Predominant aromatic fragments of the type ar–O^+ and ar–SO^+ (m/z 125 if ar is phenyl), formed after migration of one of the aryl groups. The ar–SO_2^+ ion is unimportant; ar^+ is intensive. Small signals due to SO_2, $SO_2H^{\cdot}$, and SO_2H_2 eliminations (Δm 64, 65, and 66, respectively). With alkyl substituents in *ortho* position, $[M\text{-}OH]^+$ and $[M\text{-}H_2O]^{+\cdot}$ are formed, upon which SO elimination follows.
Ion series: Aromatic fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …) and the S- and O-containing aromatic fragments at higher masses. Usually, ar–SO^+ (m/z 125 if ar is phenyl) is very strong.
Intensities: Intensive peaks mainly in the higher mass range.
Molecular ion: Strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

8.10.4 Sulfonic Acids and Their Esters and Amides

Aromatic Sulfonic Acids [6]

Fragmentation: The most prominent fragment, $[M\text{-}HSO_3]^+$ (Δm 81), is formed in a two-step process. In the first step, $OH^{\cdot}$ elimination leads to a weak fragment ion $[M\text{-}OH]^+$ (Δm 17). If an alkyl group is present in *ortho* position, $[M\text{-}H_2SO_3]^{+\cdot}$ (Δm 82) is formed instead of $[M\text{-}81]^+$. Other important fragments are $[M\text{-}SO_2]^{+\cdot}$ (Δm 64), $[M\text{-}HSO_2]^+$ (Δm 65), and $[M\text{-}SO_3]^{+\cdot}$ (Δm 80).
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...), and O-containing aromatic fragments at higher masses.
Intensities: Intensive peaks mainly in the higher mass range.
Molecular ion: Very strong. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

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Alkylsulfonic Acid Esters [6]

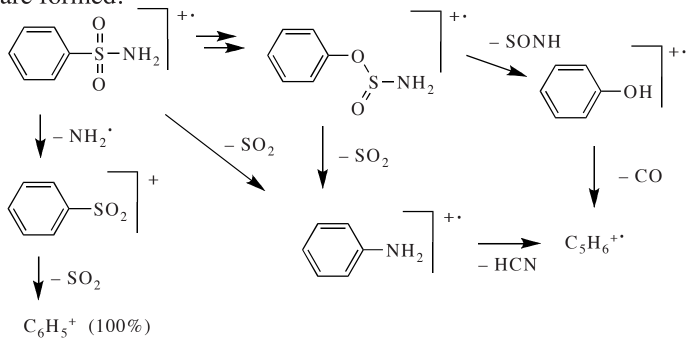
Fragmentation: Loss of alkyl by fragmentation of the C–O bond with concomitant double H rearrangement to form the protonated sulfonic acid ion (m/z 97 for methanesulfonates), which then loses water. Loss of the alkoxyl residue (fragmentation of the S–O bond). Formation of an alkene ion from the alkyl ester group by a McLafferty-type rearrangement. In aryl esters, the phenoxy ion and the phenol radical cations dominate the spectrum.
Ion series: Besides $RSO_3H_2^+$ and RSO_2^+ (m/z 97 and 79 for methanesulfonates), for aliphatic esters mainly alkene fragments. In aryl esters, aromatic fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, …), as well as O-containing aromatic fragments at higher masses.
Intensities: Intensive peaks in the lower mass range.
Molecular ion: Small or negligible signal for alkyl esters; intensive for aryl esters. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Arylsulfonic Acid Esters [6]

Fragmentation: Dominating fragments resulting from cleavage of the S–O bond (leading to the ar–SO_2^+ ion), which loses SO_2 (m/z 155 and 91 for *p*-toluenesulfonates). In arylsulfonates with longer chains, double H rearrangement to give the protonated acid (m/z 173 for *p*-toluenesulfonates).
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks mainly in the higher mass range.
Molecular ion: Medium or weak. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

Aromatic Sulfonamides [6]

Fragmentation: In *N*-alkylamides, the C–C bond next to N is split preferably. In *N*-arylsulfonamides, besides $[M\text{-}SO_2]^{+\cdot}$ and $[M\text{-}HSO_2]^+$, the ions ar–SO_2^+ and ar'–NH^+ are formed.



Ion series: Typical for the tosyl group are ions at m/z 155, 91, and 65.
Molecular ion: In arylsulfonamides, $M^{+\cdot}$ is dominant. Characteristic ^{34}S isotope peak at $[M+2]^{+\cdot}$ and [frag+2] for S-containing fragments (4.5% per S atom).

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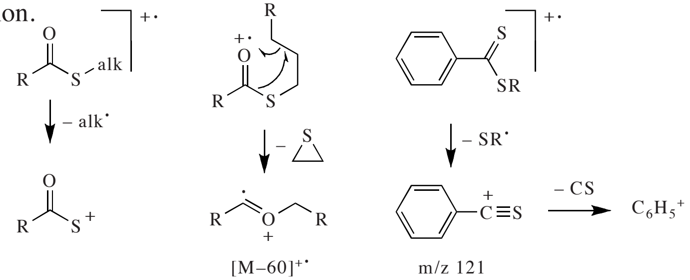
8.10.5 Thiocarboxylic Acid Esters [7]

In contrast to esters, the major fragmentation process is elimination of the alkyl radical from the thiol site. Ethylene sulfide is eliminated from thioesters with longer alkyl chains. Aromatic dithiocarboxylic acid esters usually fragment in two steps to the aryl cation.



8.10.6 References

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8.11 Carbonyl Compounds [1–4]

8.11.1 Aldehydes

Aliphatic Aldehydes [5]

Fragmentation: Cleavage of the bond next to CO. The fragmentation of the hydrocarbon chain is similar to that in corresponding alkanes. McLafferty rearrangement with localization of the charge on either side, giving rise to $C_nH_{2n}^{+\cdot}$ (m/z 28, 42, 56, ...) and, often less important, to $C_nH_{2n}O^{+\cdot}$ ions (m/z 44, 58, 72, ...). At least one product (often both) is significant. Elimination of water from the molecular ion to give $[M\text{-}18]^{+\cdot}$, occasionally very pronounced.
Ion series: Dominating fragments of the series of C_nH_{2n+1} and $C_nH_{2n-1}O$ (in both cases: m/z 29, 43, 57, ...). Weaker signals of the series C_nH_{2n-1} (m/z 41, 55, 69, ...) and rearrangement products, C_nH_{2n} (m/z 28, 42, 56, ...).
Intensities: Intensive peaks concentrated in the lower mass range. Local even-mass maxima from McLafferty-type reactions ($[M\text{-}44]^{+\cdot}$ when the aldehyde is not substituted in α-position).
Molecular ion: Only strong for molecules of low molecular weight; very weak for $C_{n>9}$. $[M\text{-}1]^+$ may be more relevant than $M^{+\cdot}$.

Unsaturated Aliphatic Aldehydes

Fragmentation: Cleavage of the bond next to CO, leading to $[M\text{-}1]^+$ (more significant than in saturated aldehydes), $[M\text{-}29]^+$, and m/z 29. No McLafferty rearrangement occurs if the γ-hydrogen atom is attached to a double-bonded carbon or if there is a double bond in α,β-position.
Ion series: Fragments of the series of C_nH_{2n-1} and $C_nH_{2n-3}O$ (in both cases, m/z 41, 55, 69, ...).
Molecular ion: Stronger than in saturated aldehydes. Usually, $[M\text{-}1]^+$ is relevant.

Aromatic Aldehydes

Fragmentation: Characteristic $H^\cdot$ loss to yield the corresponding benzoyl ion, $[M\text{-}1]^+$, followed by decarbonylation to a phenyl ion, $[M\text{-}1\text{-}28]^+$, of lower intensity. To a small extent also decarbonylation of the molecular ion, leading to $[M\text{-}28]^{+\cdot}$. Weak signal at m/z 29 (CHO^+).
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks predominantly in the molecular ion region.
Molecular ion: Usually prominent. $[M\text{-}1]^+$ is strong.

8.11.2 Ketones

Aliphatic Ketones

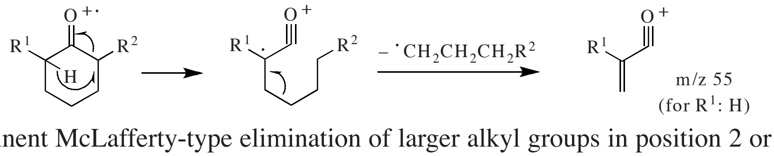
Fragmentation: Cleavage of the bond next to CO is the most important primary fragmentation. The charge can remain on either side. The acyl ions then lose CO. McLafferty rearrangement giving rise to $C_nH_{2n}O^{+\cdot}$ ions (m/z 58, 72, 86, ...). Consecutive rearrangements occur if both alkyl chains contain a γ-H atom. Keto-enol tautomerism of the first rearrangement product is not a prerequisite for the second rearrangement to occur. Oxygen is sometimes indicated by weak signals at $[M-18]^{+\cdot}$ and m/z 31, 45, 59. Fragmentation of the hydrocarbon chain similar to that in the corresponding alkanes.
Ion series: Dominating fragments of the series C_nH_{2n+1} and $C_nH_{2n-1}O$ (in both cases m/z 29, 43, 57, ..., but often distinguishable by the intensity of the ^{13}C isotope signal), with maxima due to cleavage at the CO group to give acyl ions and their decarbonylation products. Weaker signals in the series C_nH_{2n-1} (m/z 41, 55, 69, ...). Even-mass maxima, $C_nH_{2n}O$ (m/z 58, 72, 86, ...), due to alkene elimination (McLafferty rearrangement). Usually, m/z 43 (CH_3CO^+) is strong if an unsubstituted α-CH_2 group is present.
Intensities: Intensive peaks mainly in the lower mass range.
Molecular ion: Relatively abundant, weak in long-chain and branched ketones.

Unsaturated Ketones

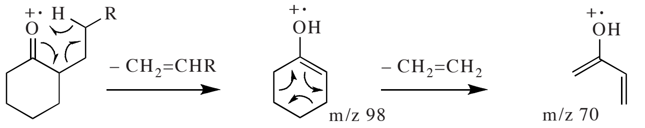
Fragmentation: Cleavage of the bond next to CO, more favorably on the saturated side, is the most important primary fragmentation. The acyl ion then loses CO. The McLafferty rearrangement occurs neither when the unsaturated substituents are in α,β-position nor when the only available γ-hydrogen atom is attached to a double-bonded carbon.
Molecular ion: Relatively abundant.

Cyclic Ketones

Fragmentation: Major primary fragmentation by bond cleavage next to carbonyl, followed by loss of alkyl residue.



Prominent McLafferty-type elimination of larger alkyl groups in position 2 or 6 as alkenes. This rearrangement is very favorable; even aromatically bonded H atoms can rearrange. For cyclohexanones, a consecutive retro-Diels–Alder reaction can occur:



Oxygen is sometimes indicated by a weak signal at $[M-18]^{+\cdot}$.
Ion series: Alkene fragments of the type of C_nH_{2n-1} or $C_nH_{2n-3}O$ (for both: m/z 41, 55, 69, ...) with maxima due to alkyl loss after ring opening next to the carbonyl group and H transfer. Prominent even-mass maxima by elimination of substituents at position 2 or 6 as alkenes via sterically favored McLafferty rearrangements.
Intensities: Overall more intensive peaks in the lower mass range or even distribution of major peaks over the whole mass range. Local maxima from major fragmentation pathway.
Molecular ion: Abundant.

Aromatic Ketones

Fragmentation: Dominant α-cleavage to give the benzoyl ion, followed by decarbonylation to a phenyl ion of lower intensity. α-Cleavage in acetophenone also produces the acetyl cation (m/z 43). Even-mass maxima due to alkene elimination via McLafferty rearrangement. CO elimination from diaryl ketones through skeletal rearrangements.
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks predominantly in the molecular ion region.
Molecular ion: Strong.

8.11.3 Carboxylic Acids

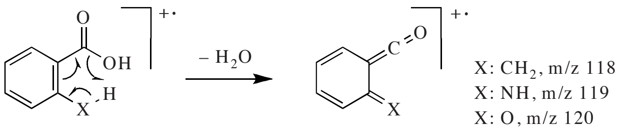
Aliphatic Carboxylic Acids

Fragmentation: Cleavage of the C–CO bond leading to m/z 45 and to $[M-45]^+$. Loss of $OH^\cdot$ leading to $[M-17]^+$; may be followed by decarbonylation. Cleavage of the γ-bond (relative to CO) leads to $^+CH_2CH_2COOH$ (m/z 73) if there is no branching on the α- and β-C atoms. Loss of $H^\cdot$ (not the carboxylic one) gives $[M-1]^+$. Water elimination to give $[M-18]^{+\cdot}$ if the alkyl group consists of at least 4 C atoms; may be followed by decarbonylation. McLafferty rearrangement to m/z 60 (acetic acid) if there is no α-substituent.
Ion series: Saturated and unsaturated alkyl ions mainly in the lower mass range (C_nH_{2n+1} and C_nH_{2n-1}, m/z 29, 43, 57, ... and 27, 41, 55, ...). With long-chain aliphatic acids, $C_nH_{2n-1}O_2$ series (m/z 59, 73, 87, ...), exhibiting maxima for n = 3, 7, 11, 15, ... (m/z 73, 129, 185, 241, ...). Even-mass maxima, $C_nH_{2n}O_2$ (m/z 60, 74, 88, ...), due to McLafferty rearrangements.
Intensities: Intensive peaks due to the above mentioned ions.
Molecular ion: Generally detectable. Easily protonated to $[M+H]^+$.

Aromatic Carboxylic Acids

Fragmentation: Pronounced loss of $OH^\cdot$, leading to $[M-17]^+$ and followed by decarbonylation (Δm 28) to a phenyl ion of lower intensity. Water elimination to

$[M-18]^{+\cdot}$ if a H-bearing *ortho*-substituent is present. Some acids decarboxylate (Δm 44). Loss of CO (Δm 28) from $M^{+\cdot}$.



Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks predominantly in the molecular ion region.
Molecular ion: Strong.

8.11.4 Carboxylic Acid Anhydrides

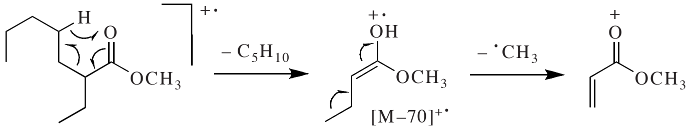
Fragmentation: In the case of linear anhydrides, abundant acyl ions due to cleavage next to carbonyl group. For cyclic anhydrides, maxima due to decarboxylation (Δm 44), followed by decarbonylation.
Molecular ion: Weak or absent (especially in linear aliphatic anhydrides), easily protonated to $[M+H]^+$. Relatively strong for phthalic anhydrides.

8.11.5 Esters and Lactones

Esters of Aliphatic Carboxylic Acids

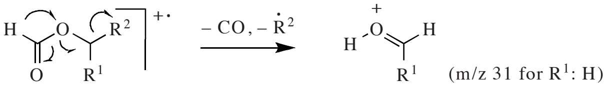
Fragmentation: Dominant fragmentation of the bonds next to the carbonyl C, leading to alk-CO^+ (m/z 43, 57, 71, ...; decreasing intensity with increasing length of the alkyl chain) and followed by decarbonylation, as well as fragmentation to $COOR^+$ (m/z 59, 73, 87, ...) and to alk$^+$ (m/z 15, 29, 43, ...). Alcohol elimination to give $C_nH_{2n-2}O$ (m/z 42, 56, 70, ...), followed by decarbonylation (Δm 28) or ketene elimination (Δm 42). Alkene elimination from the acid side via McLafferty rearrangements, leading to $C_nH_{2n}O_2$ (m/z 60, 74, 88, ...). The larger alkyl group participates in the rearrangement if several γ-H atoms are available. In the following example, the alternative process leading to $[M-C_2H_4]^{+\cdot}$ is negligible:

C = X



Nonspecific H rearrangements on the alcohol side (from $M^{+\cdot}$ or the McLafferty product) lead to $C_nH_{2n}O_2$ and to the corresponding alkene, C_nH_{2n} (m/z 28, 42, 56, ...). In methyl esters of long chain acids, the ions $[(CH_2)_{2+4n}COOCH_3]^+$ (m/z 87, 143, 199, ...) correspond to maxima. For esters of higher alcohols ($C_{n\geq3}$), double H rearrangement to the protonated acid, $C_nH_{2n+1}CO_2H_2^+$ (m/z 61, 75, 89, ...). α-Substituted esters may lose the substituent and then CO (Δm 28) via alkoxyl rearrangement. Analogously, β-substituted esters may eliminate ketene (Δm 42).

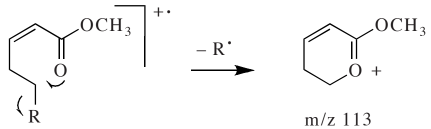
Besides usual ester reactions, specific rearrangements can be observed in formates.



Ion series: C_nH_{2n+1} (m/z 29, 43, 57, …) for the alkyl groups at the ester oxygen (except for methyl esters). C_nH_{2n-1} (m/z 27, 41, 55, …). $C_nH_{2n-1}O_2$ (m/z 59, 73, 87, ...), exhibiting maxima for n = 4, 8, 12, … (m/z 87, 143, 199, …) in the case of methyl esters of long-chain acids. Even-mass maxima for $C_nH_{2n}O_2$ (m/z 60, 74, 88, …) due to alkene elimination via McLafferty rearrangements on both sides of the carboxyl group. C_nH_{2n} (m/z 28, 42, 56, …) as H rearrangement product from the alcohol side.
Intensities: Intensive peaks due to the above mentioned ions in the lower mass range.
Molecular ion: Often of low abundance. Easily protonated to $[M+H]^+$.

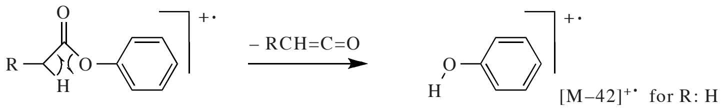
Esters of Unsaturated Carboxylic Acids

α,β-Unsaturated esters: Loss of alk–O˙ followed by CO elimination is the dominant fragmentation path. Also, loss of the δ-substituent yields a 6-membered oxonium ring:



Significant difference between *Z* and *E* isomers of long-chain α,β-unsaturated esters: Single H rearrangement occurs with *Z* esters, and double H rearrangements (leading to protonated acids) have been found for *E* esters.
β,γ-Unsaturated esters: Only slight qualitative, but significant quantitative differences have been observed as compared to α,β-unsaturated esters (e.g., less intensive signals for $M^{+\cdot}$ of *β,γ-* than of α,β-unsaturated esters).
γ,δ-Unsaturated esters: Loss of the alcohol chain as a radical, R˙, followed by ketene elimination.
Aliphatic enol esters and aryl esters: Formation of alk–CO$^+$ (m/z 43, 57, 71, …). Elimination of a ketene to give the enol or phenol radical cation. The rearrangement occurs prodominantly, but not exclusively, through a 4-membered transition state:

C=X



Esters of Aromatic Acids

Fragmentation: Dominant loss of RO˙ to form the benzoyl ion, followed by decarbonylation (Δm 28) and further loss of acetylene (Δm 26). Ethyl esters also eliminate C_2H_4 (Δm 28) to give the acid radical cation, which then loses OH˙ to yield the benzoyl ion. In higher alkyl esters, besides the acid, the protonated

acid is formed (double H rearrangement). In *ortho*-substituted aryl esters with an α-hydrogen atom on the substituent, an alcohol is eliminated from $M^{+\cdot}$. In the case of alkyl phthalates (other than dimethyl phthalate), alkenyl elimination from one ester group to give the protonated ester acid, followed by alkene elimination from the other ester group, and subsequent water elimination to the protonated anhydride ion, which forms the base peak at m/z 149.
Ion series: Aromatic hydrocarbon fragments, C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Prominent maximum at the mass of the related benzoyl ion and its decarbonylation product.
Molecular ion: Usually strong.

Lactones

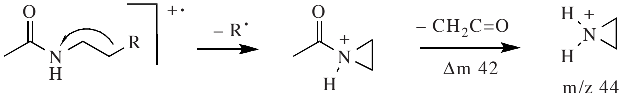
Fragmentation: The most prominent reaction is the loss of substituents (or $H^\cdot$) at the O-bearing C atom, followed by decarbonylation (Δm 28), decarboxylation (Δm 44, mainly in smaller molecules), and ketene elimination (Δm 42). Decarboxylation of $M^{+\cdot}$ is rarely significant. Competing reactions are several kinds of primary ring cleavages. Aromatic lactones show maxima due to two consecutive decarbonylations.
Ion series: No specific ion series. The acetyl ion (m/z 43) is often an important fragment.
Intensities: Maxima at the mass resulting from loss of substituents at the C atom next to oxygen. Otherwise, intensive peaks evenly distributed over the whole mass range.
Molecular ion: Usually of low intensity and easily protonated to $[M+H]^+$ in aliphatic lactones; abundant in the case of aromatic lactones.

8.11.6 Amides and Lactams

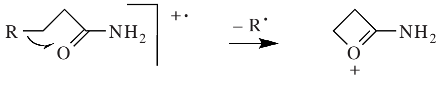
Amides of Aliphatic Carboxylic Acids

C = X

Fragmentation: Alkene elimination on the acid side via McLafferty reaction to yield the corresponding acetamide radical cation. Loss of alkenes on the amine side to give the ion of the desalkyl amide, often via double H rearrangement to the protonated desalkyl amide ion. Bond cleavage on both sides of the carbonyl group. Cleavage of the C–C bond attached to N, and the β,γ-C–C bond (relative to N):



Cleavage of the bonds to the β-C (see scheme) and to the γ-C on the acid side.



Ion series: Even-mass fragments corresponding to $C_nH_{2n}NO$ (m/z 44, 58, 72, ...)

produced by cleavage of the bond next to CO on the acid side. Odd-mass fragments (in secondary and tertiary amides), $C_nH_{2n-1}O$ (m/z 43, 57, 71, ...), produced by cleavage of the bond next to CO on the amine side.
Intensities: Overall peak distribution maximizing in the low mass range. Local maxima from McLafferty and from γ-cleavage products.
Molecular ion: Significant. Strong tendency to protonate to $[M+H]^+$.

Amides of Aromatic Carboxylic Acids

Fragmentation: Maxima due to amide bond cleavage yielding the benzoyl ion, followed by decarbonylation (Δm 28).
Ion series: Aromatic fragments corresponding to C_nH_n and $C_nH_{n\pm1}$ (m/z 39, 51–53, 63–65, 75–77, ...).
Intensities: Intensive peaks predominantly in the molecular ion region.
Molecular ion: Abundant. $[M-H]^+$ is significant in *N,N*-disubstituted anilides, weaker in monosubstituted derivatives, and absent from the spectrum of benzamide. It is formed exclusively by loss of *ortho*-hydrogens of the aromatic ring.

Anilides

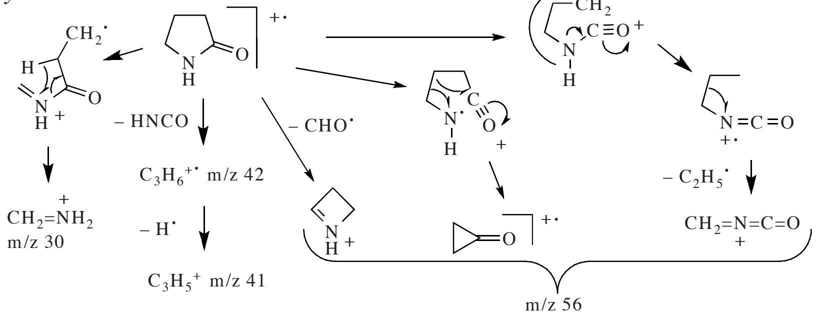
Formanilides: Loss of CO (Δm 28) to give the aniline radical cation and consecutive HCN elimination (Δm 27).
Acetanilides: Ketene elimination gives the aniline radical cation (often base peak), which can eliminate HCN (Δm 27), and formation of the acetyl cation (m/z 43).
Trichloroacetanilides: Dominant loss of $CCl_3^{\cdot}$ (Δm 117).
Pivalanilides: Besides reactions analogous to those of acetanilides (Δm 84, formation of the aniline radical cation), also formation of the *tert*-butylbenzene radical cation through elimination of HNCO (Δm 43).

Lactams

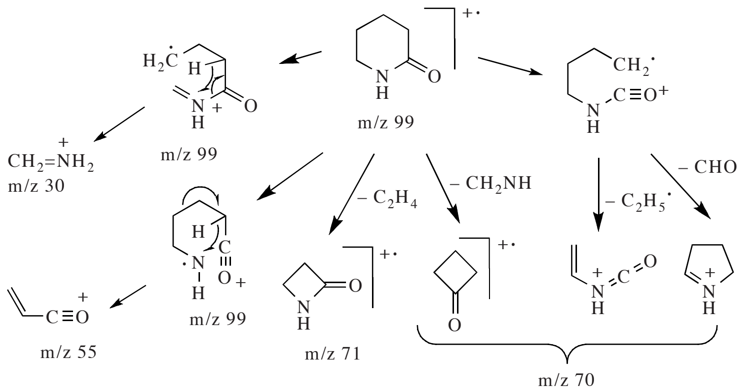
Fragmentation: Cleavage of the C–C bond at the N-bearing C atom. Cleavage of the CO–N bond, followed by loss of CO (Δm 28) or by further cleavage of the C–C bond next to N, giving an iminium ion. In 2-pyrrolidone and 2-piperidone, the signal at m/z 30 ($[CH_2{=}NH_2]^+$) is strong. The base peak of 2-pyridone is formed by CO elimination (Δm 28).

C = X

2-Pyrrolidone:



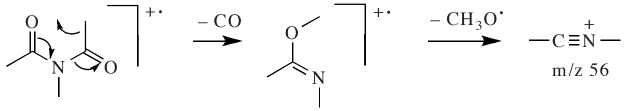
2-Piperidone:



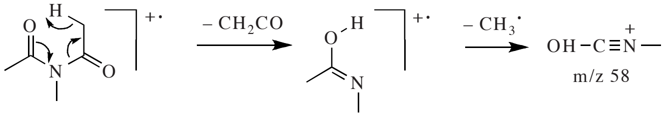
Molecular ion: Often observable; more abundant than for the corresponding lactones.

8.11.7 Imides

Saturated acyclic imides: Consecutive CO (Δm 28) and alkoxy elimination:



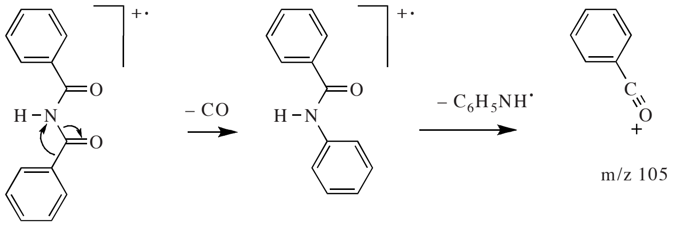
Ketene elimination:



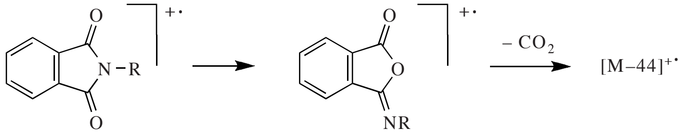
If the *N*-substituent chain is sufficiently long, cleavage of the C–C bonds next to N, with or without H rearrangement.

C = X

Dibenzoylamine: Loss of CO to *N*-phenylbenzamide:



Cyclic imides: The spectra of saturated cyclic imides are almost identical to those of the corresponding diketones. Loss of HNCO (Δm 43) from succinimide, followed by CO elimination (Δm 28). Aroyl migration and loss of CO_2 from aromatic cyclic imides.



8.11.8 References

[1] J.H. Bowie, Mass spectrometry of carbonyl compounds. In: T*he Chemistry of the Carbonyl Group, Vol. 2*; J. Zabicky, Ed.; Interscience: London, 1970; p 277.

[2] S.W. Tam, Mass spectra of acid derivatives. In: *Suppl. B, The Chemistry of Acid Derivatives, Part 1*; S. Patai, Ed.; Wiley: Chichester, 1979; p 121.

[3] D.G.I. Kingston, J.T. Bursey, M.M. Bursey, Intramolecular hydrogen transfer in mass spectra. II. The McLafferty rearrangement and related reactions, *Chem. Rev.* **1974**, *74*, 215.

[4] D.G.I. Kingston, B.W. Hobrock, M.M.Bursey, J.T. Bursey, Intramolecular hydrogen transfer in mass spectra. III. Rearrangements involving the loss of small neutral molecules, *Chem. Rev.* **1975**, *75*, 693.

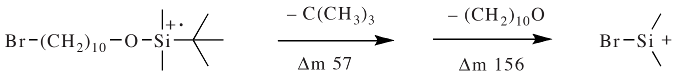
[5] A.G. Harrison, High-resolution mass spectra of aliphatic aldehydes, *Org. Mass Spectrom.* **1970**, *3*, 549.

C = X

8.12 Miscellaneous Compounds

8.12.1 Trialkylsilyl Ethers [1,2]

Fragmentation: Loss of alkyl attached to Si (preferential loss of larger groups). Cleavage of the C–C bond adjacent to O, followed by alkene elimination. Loss of alkoxyl, followed by alkene eliminations. Elimination of trialkylsilanol. The R_2Si–OR' cation has the tendency to attack, in an electrophilic manner and even over long distances, free electron pairs and π-electron centers, causing the expulsion of neutral fragments from the interior of the molecule via a rearrangement:



Ion series: $[C_nH_{2n+3}OSi]^+$ (m/z 75, 89, 103, 117, ...). $[C_nH_{2n+3}Si]^+$ (m/z 45, 59, 73, 87, ...). Occasionally, maxima at even mass due to elimination of trialkylsilanol.
Molecular ion: $M^{+\cdot}$ often of low abundance or absent, easily protonated to $[M+H]^+$. Typical isotope patterns owing to ^{28}Si, ^{29}Si, and ^{30}Si (see Chapter 2.5.5).

8.12.2 Phosphorus Compounds

Alkyl Phosphates [3]

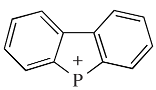
Fragmentation: Maxima due to alkenyl loss from $M^{+\cdot}$ via double H rearrangement, followed by successive alkene eliminations down to protonated phosphoric acid (m/z 99).
Ion series: PO^+ (m/z 47), $H_2PO_2^+$ (m/z 65), $H_2PO_3^+$ (m/z 81), often as nonspecific P indicators.
Molecular ion: $M^{+\cdot}$ observable.

Aliphatic Phosphines

Ion series: Maxima of the ion series of $[C_nH_{2n+3}P]^+$ (m/z 48, 62, 76, 90, ...) due to alkene eliminations.
Molecular ion: $M^{+\cdot}$ observable.

Aromatic Phosphines and Phosphine Oxides

Fragmentation: Maxima due to loss of an aryl group, followed by H_2 elimination to yield the 9-phosphafluorenyl ion (m/z 183).
Molecular ion: $M^{+\cdot}$ abundant, easily losing $H^\cdot$ to give $[M-1]^+$.



m/z 183

8.12.3 References

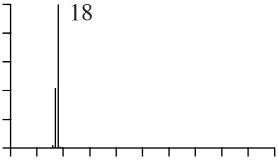
[1] D.G.I. Kingston, B.W. Hobrock, M.M. Bursey, J.T. Bursey, Intramolecular hydrogen transfer in mass spectra. III. Rearrangements involving the loss of small neutral molecules, *Chem. Rev.* **1975**, *75*, 693.
[2] H. Schwarz, Positive and negative ion chemistry of silicon-containing molecules in the gas phase. In: *The Chemistry of Organic Silicon Compounds, Part 1*; S. Patai, Z. Rappoport, Eds.; Wiley: Chichester, 1989; p 445.
[3] D.G.I. Kingston, J.T. Bursey, M.M. Bursey, Intramolecular hydrogen transfer in mass spectra. II. The McLafferty rearrangement and related reactions, *Chem. Rev.* **1974**, *74*, 215.

8.13 Mass Spectra of Common Solvents and Matrix Compounds

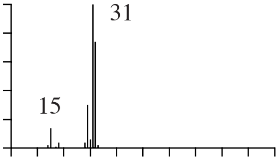
8.13.1 Electron Impact Ionization Mass Spectra of Common Solvents

The label {50} indicates that the intensity scale ends at 50% relative intensity and is subdivided in 10% steps. In these cases, the height of the base peak has to be doubled to bring it to 100%. All spectra represent positive ions only.

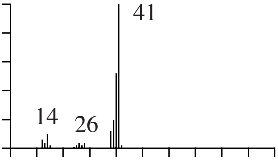
Water {50}



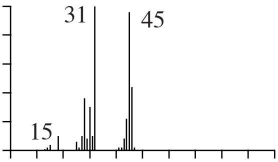
Methanol



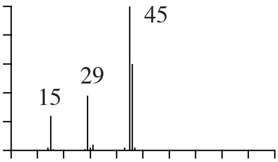
Acetonitrile



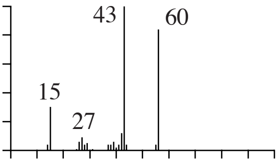
Ethanol {50}



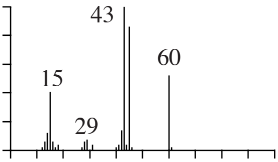
Dimethyl ether



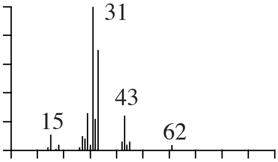
Acetone {50}



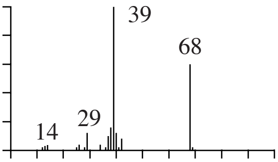
Acetic acid



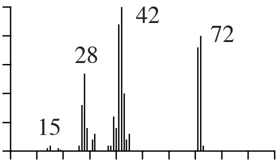
Ethylene glycol {50}



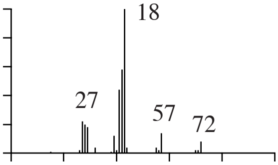
Furan



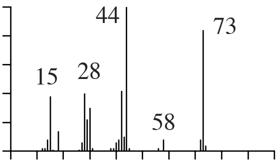
Tetrahydrofuran {50}



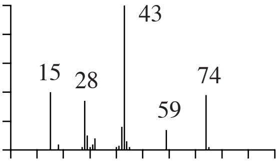
Pentane



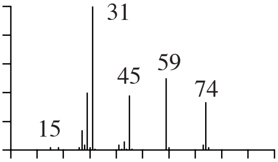
N,N-Dimethylformamide



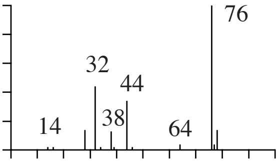
Methyl acetate {50}



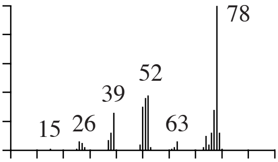
Diethyl ether



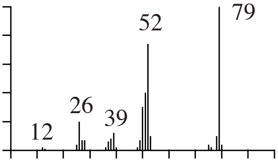
Carbon disulfide {50}



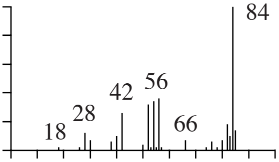
Benzene {50}



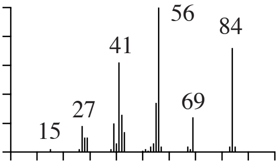
Pyridine



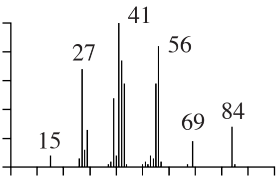
Benzene-d_6 {50}



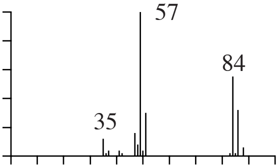
Cyclohexane



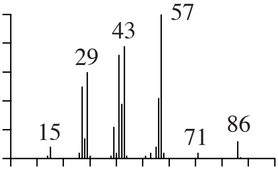
1-Hexene



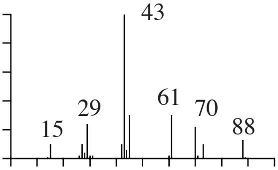
Methylene chloride



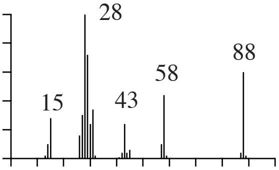
Hexane



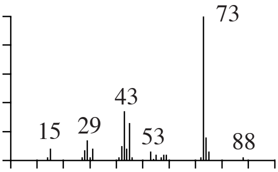
Ethyl acetate {50}



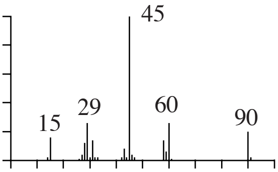
1,4-Dioxane {50}



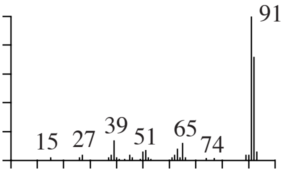
Tetramethylsilane {50}



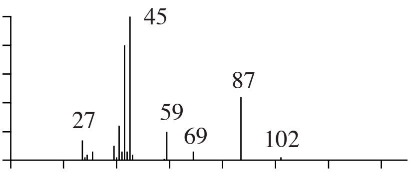
1,2-Dimethoxyethane {50}



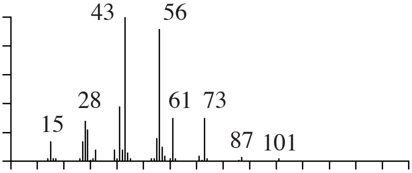
Toluene



Diisopropyl ether {50}

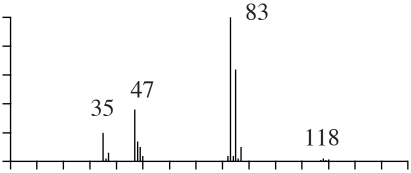


Butyl acetate {50}

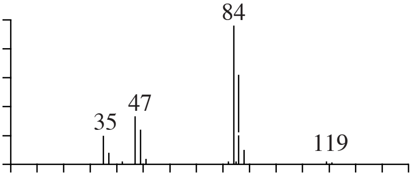


Solvents

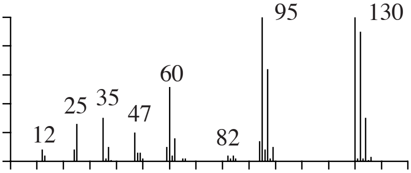
Chloroform



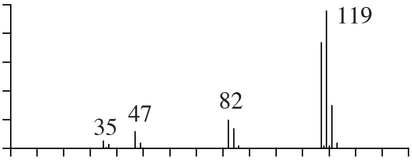
Chloroform-*d*



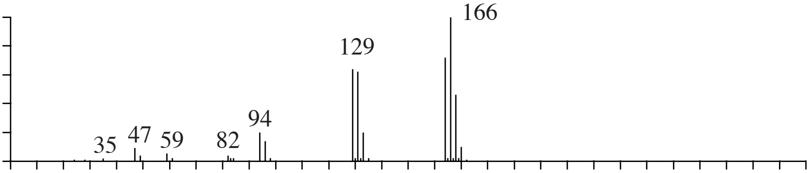
Trichloroethylene



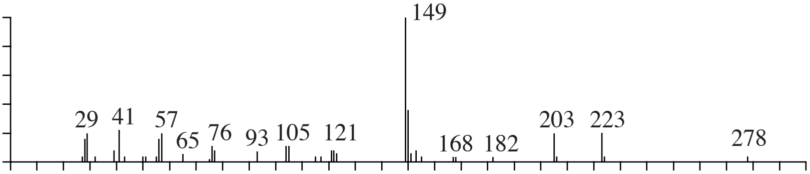
Carbon tetrachloride



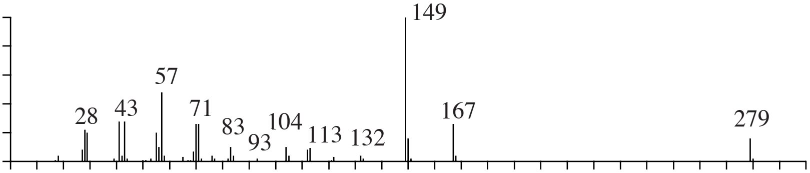
Tetrachloroethylene



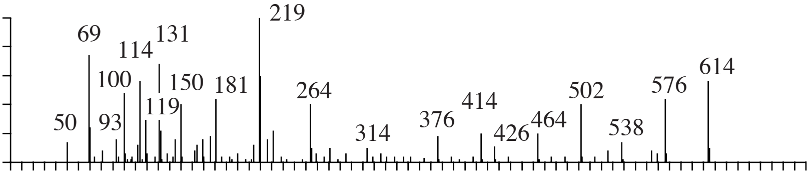
Dibutyl phthalate [25] (frequent impurity due to its use as polymer plasticizer)



Dioctyl phthalate (frequent impurity due to its use as polymer plasticizer)



Heptacosafluorotributylamine (calibration reagent)

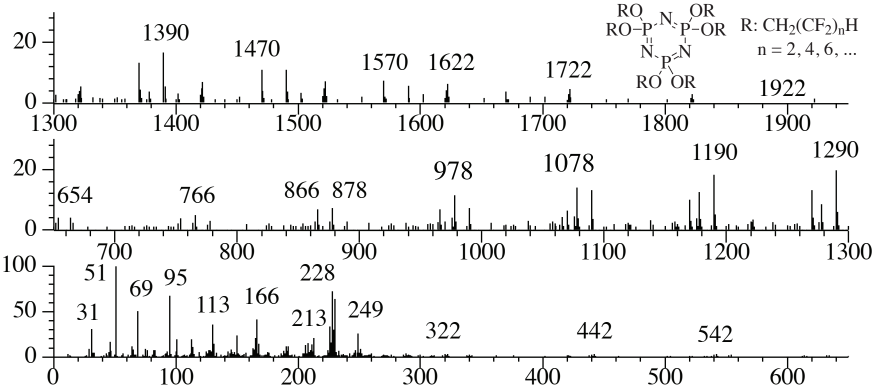


8.13.2 Spectra of Common FAB MS Matrix and Calibration Compounds

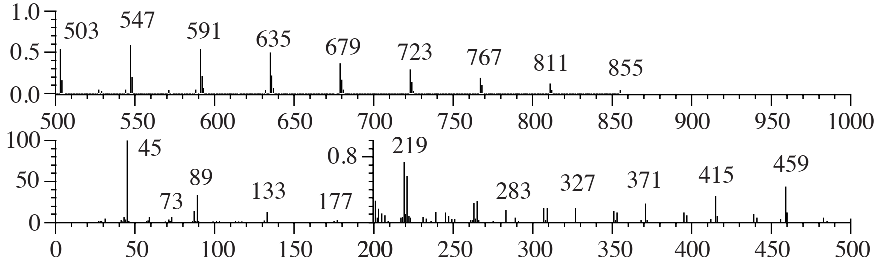
Fast atom bombardment (FAB) mass spectra (MS) usually exhibit signals for the protonated or deprotonated molecular ions, $[M{\pm}H]^{\pm}$, and protonated clusters, $[M_n{+}X_m{\pm}H]^{\pm}$ (n, m = 0, 1, 2, …), of the sample and matrix molecules, X. Even traces of metal salts in the sample give rise to clusters of the type $[M_n{+}X_m$+metal cation$]^{+}$. Na^+ (23 u) and K^+ (39 u) adducts are often found. The nature of the clusters is often revealed by the regular intervals at which their peaks occur in the spectra.

Calibration Compounds in Positive Ionization FAB Mass Spectra

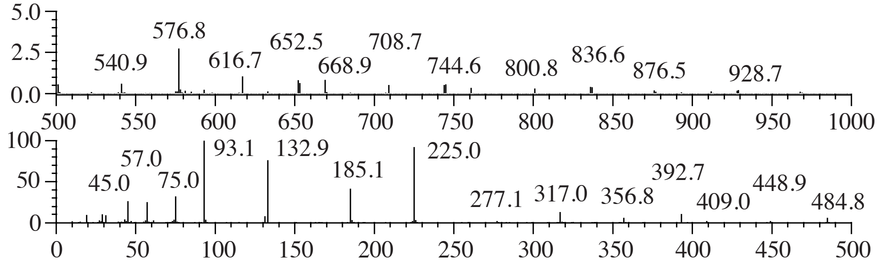
Ultramark 1621 (erroneously also referred to as perfluoroalkyl phosphazine)



Polyethylene glycol 600 (often used as internal reference for high resolution m/z determinations)

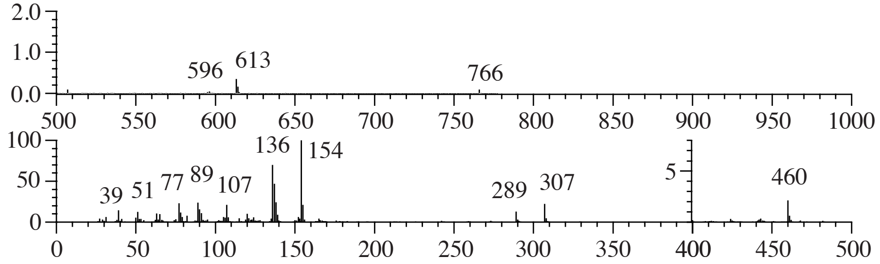


CsI (Cs^+, 132.9; I^-, 126.9) in glycerol (formation of $[\text{glycerol}_m\text{-}H_n{+}Cs_p{+}I_q]^+$)

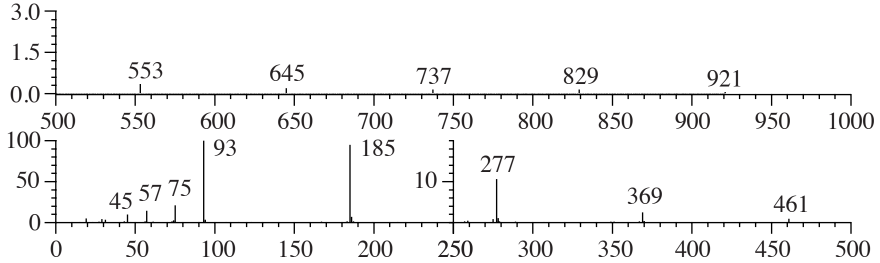


Matrix Compounds in Positive Ionization FAB Mass Spectra

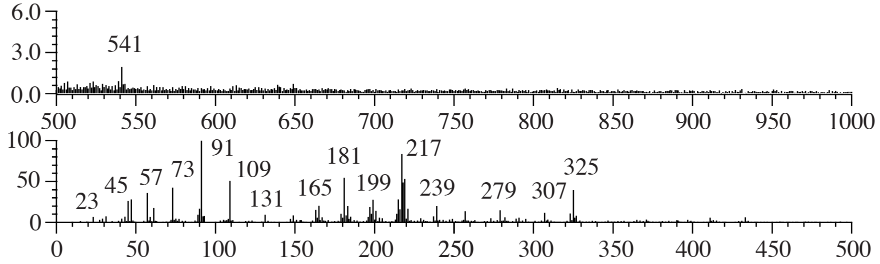
3-Nitrobenzyl alcohol (M_r 153.1)



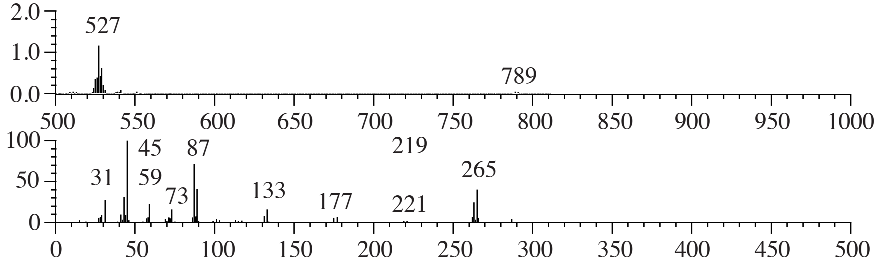
Glycerol (M_r 92.1)



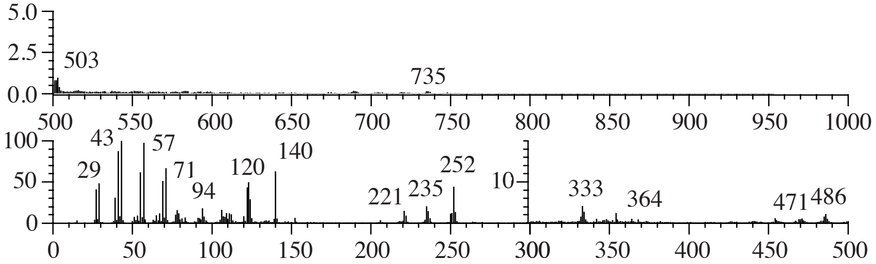
1-Thioglycerol (M_r 108.2. Note m/z 23, Na^+; 131, $[M+Na]^+$; 239, $[2M+Na]^+$. Similarly, small K^+ impurities give signals at m/z 39, 147, 255)



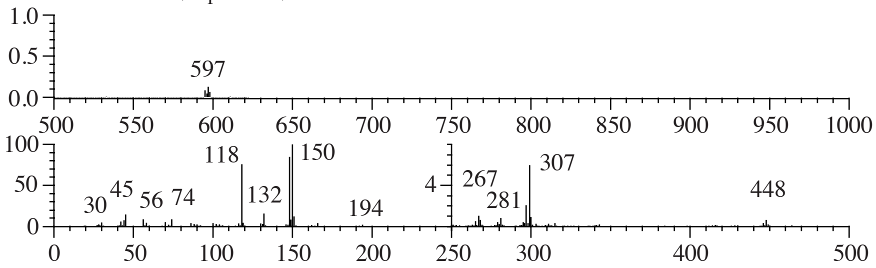
1,4,7,10,13,16-Hexaoxacyclooctadecane (18-crown-6, M_r 264.3. Also used as an additive; binds metal ions and reduces $[M+\text{metal ion}]^+$ in favor of $[M+H]^+$, which can be important for samples with exchangeable H^+, such as for peptides [1])



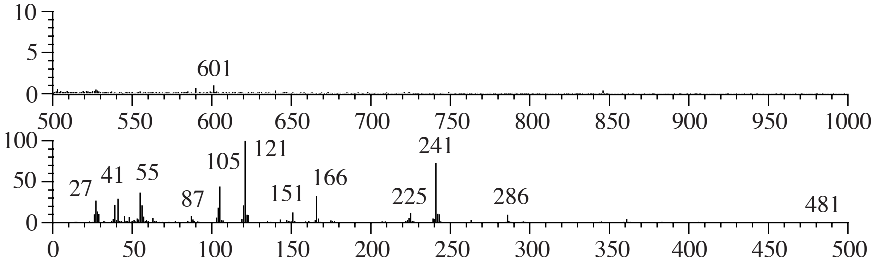
2-Nitrophenyl octyl ether (M_r 251.3)



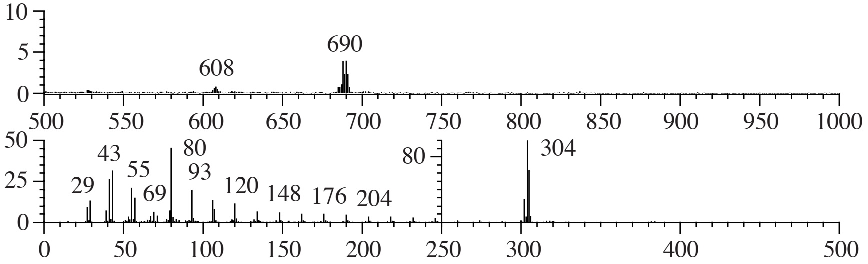
Triethanolamine (M_r 149.2)



Sulfolane (M_r 120.2) [2]

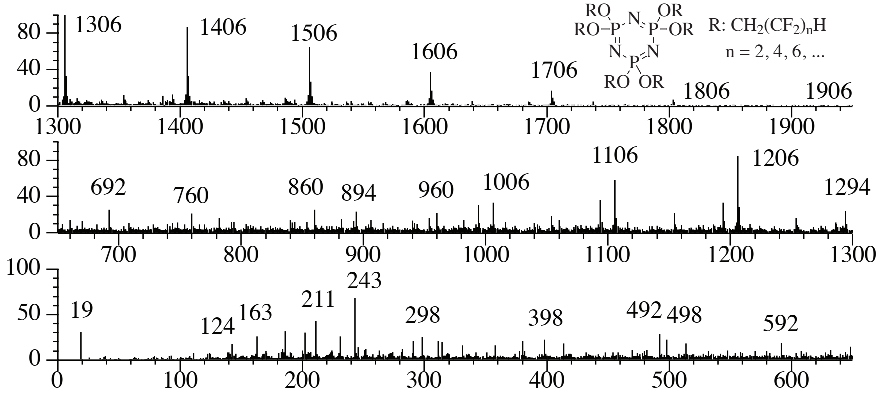


Hexadecylpyridinium bromide (M_r 384.4; for [hexadecylpyridinium]$^+$ m/z 304.3) in 2-nitrobenzyl alcohol

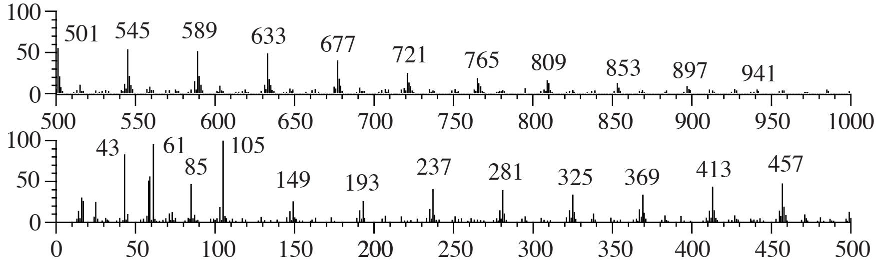


Calibration Compounds in Negative Ionization FAB Mass Spectra

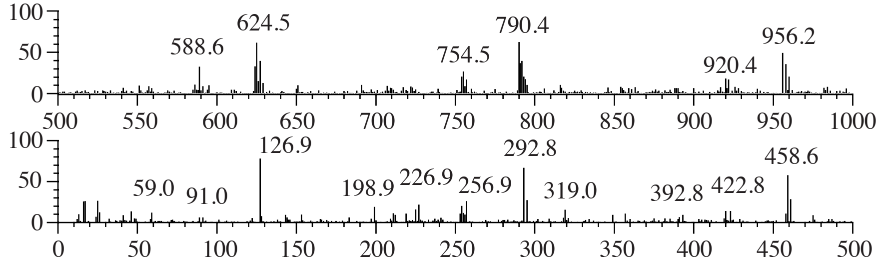
Ultramark 1621 (erroneously also referred to as perfluoroalkyl phosphazine)



Polyethylene glycol 600 (often used as internal reference for high resolution m/z determinations)

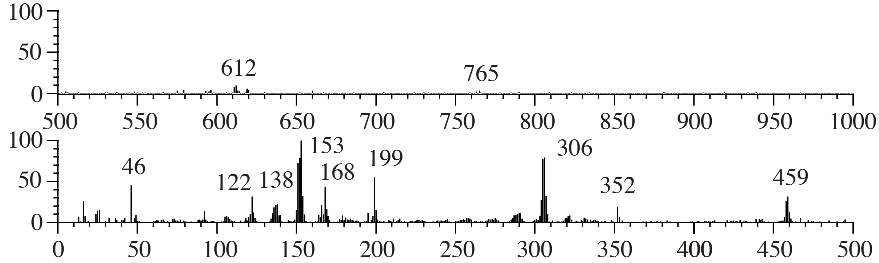


KI (K^+, 39.1; I^-, 126.9) in glycerol (formation of $[\text{glycerol}_m\text{-H}_n\text{+K}_p\text{+I}_q]^-$)

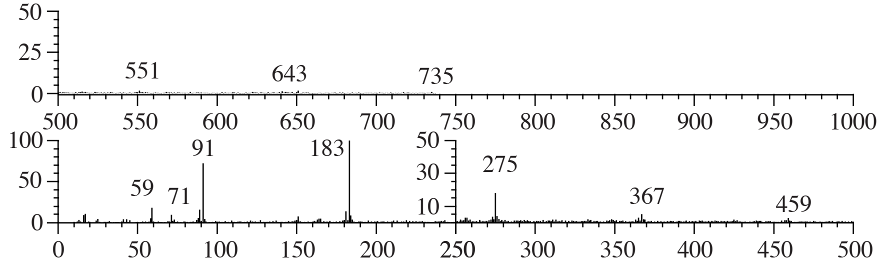


Matrix Compounds in Negative Ionization FAB Mass Spectra

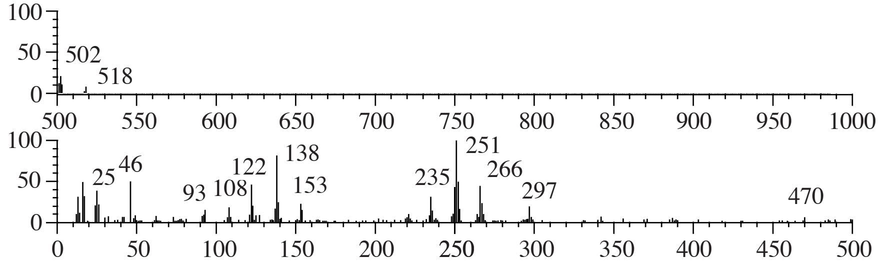
3-Nitrobenzyl alcohol (M_r 153.1)



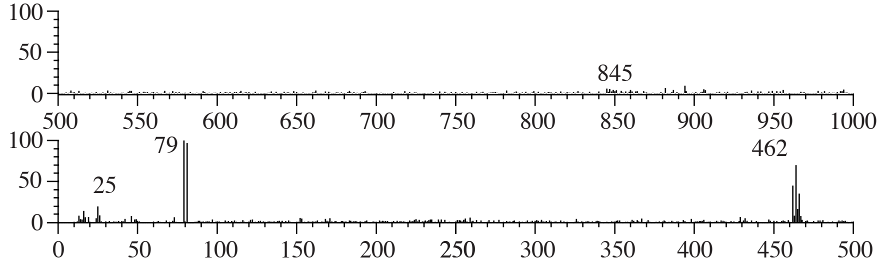
Glycerol (M_r 92.1)



2-Nitrophenyl octyl ether (M_r 251.3)



2-Nitrobenzyl alcohol solution of hexadecylpyridinium bromide (M_r 384.4; enhances detectability and reduces metal ion adducts of sample [3])

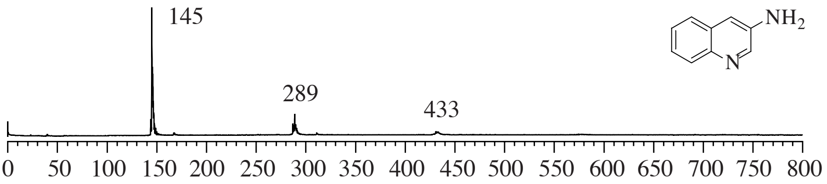


8.13.3 Spectra of Common MALDI MS Matrix Compounds

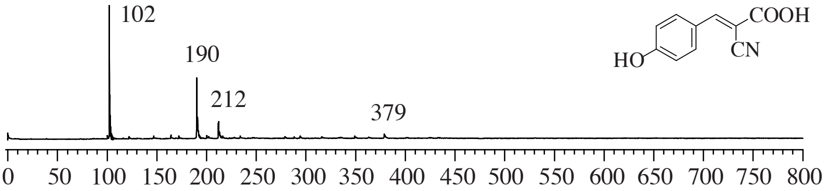
Matrix-assisted laser desorption ionization (MALDI) mass spectra (MS) usually show signals for protonated or deprotonated molecular ions, $[M\pm H]^{\pm}$, and protonated clusters, $[M_n+X_m\pm H]^{\pm}$ (n, m = 0, 1, 2, …), of the sample and matrix molecules, X. In positive ionization mass spectra, clusters of the type $[M_n+X_m+\text{metal cation}]^+$ occur even if there are only traces of metal salts in the sample. Sodium (23 u) and potassium (39 u) ion adducts are often encountered. The nature of the clusters is revealed by the regular intervals at which their signals occur in the spectra [4].

Matrix Compounds in Positive Ionization MALDI Mass Spectra

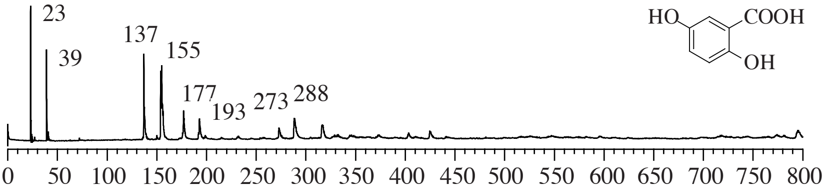
3-Aminoquinoline (M_r 144.2)



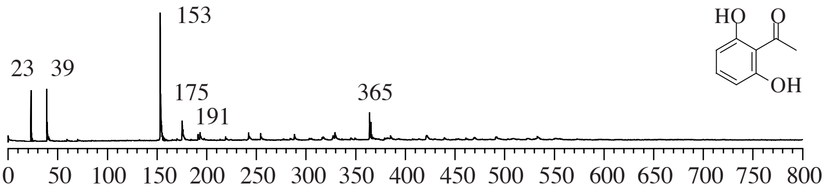
α-Cyano-4-hydroxycinnamic acid (M_r 189.2; m/z 212, $[M+Na]^+$)



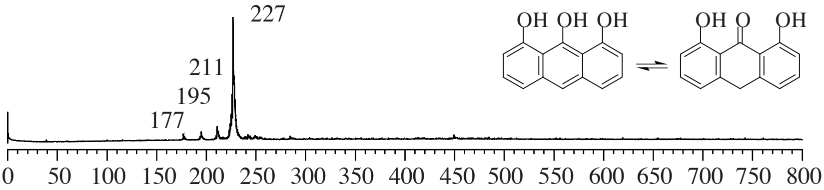
2,5-Dihydroxybenzoic acid (M_r 154.1; m/z 177, $[M+Na]^+$; m/z 193, $[M+K]^+$)



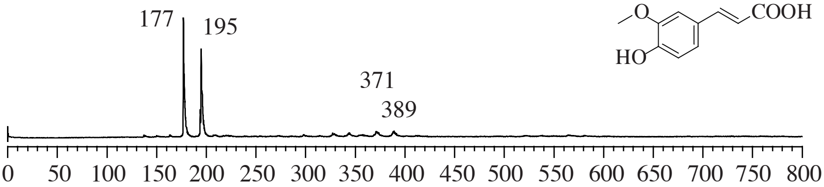
2,6-Dihydroxyacetophenone (M_r 152.1; m/z 175, $[M+Na]^+$; m/z 191, $[M+K]^+$; m/z 365, $[2M+Na+K-H]^+$?)



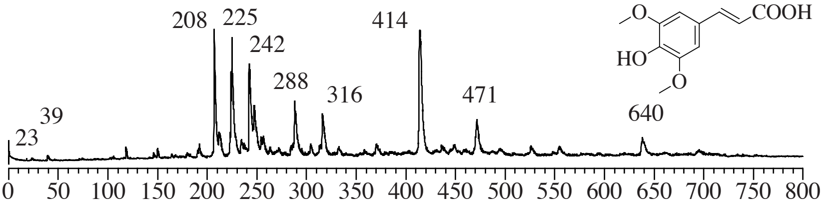
Dithranol (M_r 226.2)



Ferulic acid (4-hydroxy-4-methoxycinnamic acid; M_r 194.2)

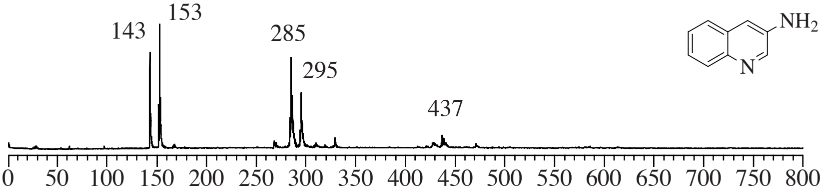


Sinapinic acid (3,5-dimethoxy-4-hydroxycinnamic acid; M_r 224.2; m/z 471, $[2M+Na]^+$)

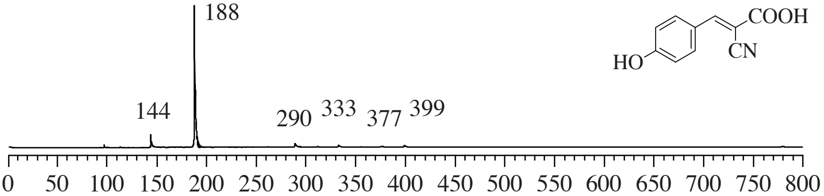


Matrix Compounds in Negative Ionization MALDI Mass Spectra

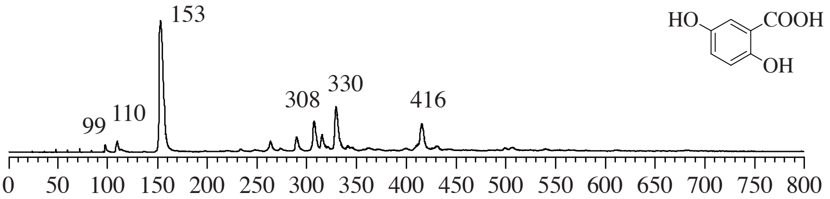
3-Aminoquinoline (M_r 144.2)



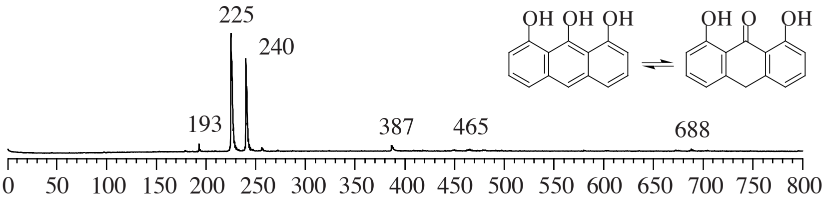
α-Cyano-4-hydroxycinnamic acid (M_r 189.2; m/z 399, $[2M+Na-2H]^-$)



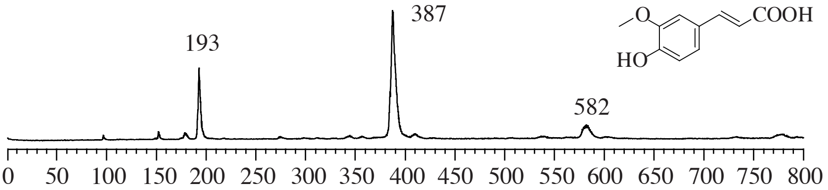
2,5-Dihydroxybenzoic acid (M_r 154.1)



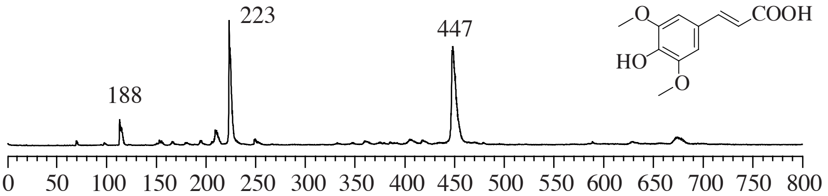
Dithranol (M_r 226.2)



Ferulic acid (4-hydroxy-4-methoxycinnamic acid; M_r 194.2)



Sinapinic acid (3,5-dimethoxy-4-hydroxycinnamic acid; M_r 224.2)



8.13.4 References

[1] R. Orlando, Analysis of peptides contaminated with alkali-metal salts by fast atom bombardment mass spectrometry using crown ethers, *Anal. Chem.* **1992**, *64*, 332.

[2] P.K. Singh, L. Field, B.J. Sweetman, Organic disulfides and related substances, *J. Org. Chem.* **1988**, *53*, 2608.

[3] Z.-H. Huang, B.-J. Shyong, D.A. Gage, K.R. Noon, J. Allison, N Alkylnicotinium halides: A class of cationic matrix additives for enhancing the sensitivity in negative ion fast-atom bombardment mass spectrometry of polyanionic analytes, *J. Am. Soc. Mass Spectrom.* **1994**, *5*, 935.

[4] A.E. Ashcroft, *Ionization Methods in Organic Mass Spectrometry*, The Royal Society of Chemistry: Cambridge, 1997.

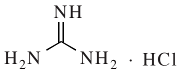
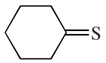
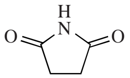
9 UV/Vis Spectroscopy

9.1 Correlation between Wavelength of Absorbed Radiation and Observed Color

Absorbed light		Observed (transmitted) color
Wavelength [nm]	Corresponding color	
400	violet	yellow-green
425	indigo blue	yellow
450	blue	orange
490	blue-green	red
510	green	purple
530	yellow-green	violet
550	yellow	indigo blue
590	orange	blue
640	red	blue-green
730	purple	green

9.2 Simple Chromophores

Chromophore	Compound	Transition	λ_{max} [nm]	ε_{max}	Solvent
C–H	CH_4	$\sigma \rightarrow \sigma^*$	122	strong	gas
C–C	CH_3–CH_3	$\sigma \rightarrow \sigma^*$	135	strong	gas
C=C	CH_2=CH_2	$\pi \rightarrow \pi^*$	162	15000	heptane
	$(CH_3)_2C$=$C(CH_3)_2$	$\pi \rightarrow \pi^*$	196	11500	heptane
C=C=C	CH_2=C=CH_2		170	4000	
			227	630	
C≡C	HC≡CH		173	6000	gas
	n-C_5H_{11}–C≡C–CH_3		178	10000	hexane
			196	2000	
			222	160	
C–Cl	CH_3Cl	$n \rightarrow \sigma^*$	173	200	hexane
C–Br	*n*-C_3H_7Br	$n \rightarrow \sigma^*$	208	300	hexane

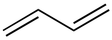
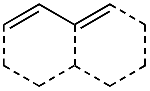
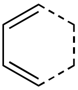
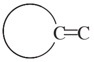
Chromophore	Compound	Transition	λ_{max} [nm]	ε_{max}	Solvent
C–I	CH_3I	$n \rightarrow \sigma^*$	259	400	hexane
C–O	CH_3OH	$n \rightarrow \sigma^*$	177	200	hexane
	CH_3OCH_3	$n \rightarrow \sigma^*$	184	2500	gas
C–N	$(C_2H_5)_2NH$	$n \rightarrow \sigma^*$	193	2500	hexane
	$(CH_3)_3N$	$n \rightarrow \sigma^*$	199	4000	hexane
C=N			265	15	water
	$(CH_3)_2C{=}NOH$		193	2000	ethanol
	$(CH_3)_2C{=}NONa$		265	200	ethanol
N=N	$CH_3{-}N{=}N{-}CH_3$		340	16	ethanol
N=O	$(CH_3)_3C{-}NO$		300	100	ether
			665	20	
	$(CH_3)_3C{-}NO_2$		276	27	ethanol
	$n\text{-}C_4H_9{-}O{-}NO$		218	1050	ethanol
			313–384	20–40	ethanol
	$C_2H_5{-}O{-}NO_2$		260	15	ethanol
C≡N	$CH_3C{\equiv}N$		<190		
X=Y=Z	$C_2H_5{-}N{=}C{=}S$		250	1200	hexane
	$C_2H_5{-}N{=}C{=}N{-}C_2H_5$		230	4000	
			270	25	
C–S	CH_3SH	$n \rightarrow \sigma^*$	195	1800	gas
		$n \rightarrow \sigma^*$	235	180	
	$C_2H_5{-}S{-}C_2H_5$	$n \rightarrow \sigma^*$	194	4500	gas
		$n \rightarrow \sigma^*$	225	1800	
	$C_2H_5{-}S{-}S{-}C_2H_5$	$n \rightarrow \sigma^*$	194	5500	hexane
		$n \rightarrow \sigma^*$	250	380	
C=S	$(CH_3)_2C{=}S$		460	weak	
			495	weak	ethanol
C=O	$(CH_3)_2C{=}O$	$n \rightarrow \sigma^*$	166	16000	gas
		$\pi \rightarrow \pi^*$	189	900	hexane
		$n \rightarrow \pi^*$	279	15	hexane
	CH_3COOH	$n \rightarrow \pi^*$	200	50	gas
	CH_3COONa	$n \rightarrow \pi^*$	210	150	water
	$CH_3COOC_2H_5$	$n \rightarrow \pi^*$	210	50	gas
	CH_3CONH_2	$n \rightarrow \pi^*$	220	63	water
			191	15200	CH_3CN
C=C=O	$(C_2H_5)_2C{=}C{=}O$		227	360	
			375	20	

9.3 Conjugated Alkenes

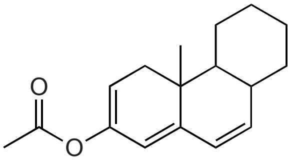
9.3.1 Dienes and Polyenes

The $\pi \rightarrow \pi^*$ transition of conjugated double bonds is above ≈200 nm with typical intensities of the order of log $\varepsilon \approx 4$. Its position can be estimated with the Woodward–Fieser rule. For cross-conjugated systems, the value for the chromophore absorbing at the longest wavelength has to be calculated.

Woodward–Fieser rule for estimating the position of the $\pi \rightarrow \pi^*$ transition (λ_{max} in nm)

Parent system		acyclic	217
		heteroannular	214
		homoannular	253
Increments	for each additional conjugated double bond		+30
	for each exocyclic double bond 		+5
	for each substituent	C-substituent	+5
		Cl	+5
		Br	+5
		O–alkyl	+6
		$OCOCH_3$	0
		N(alkyl)$_2$	+60
		S–alkyl	+30
Solvent correction			≈0

Example: Estimation of the absorption maximum for

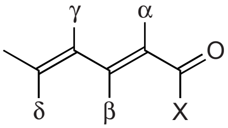


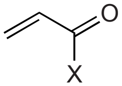
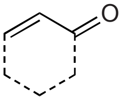
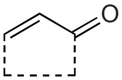
base value (homoannular)	253
1 additional conjugated double bond	30
1 exocyclic double bond	5
3 C-substituents	15
1 $OCOCH_3$	0
estimated	303
experimental	306

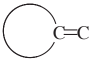
9.3.2 α,β-Unsaturated Carbonyl Compounds

The π→π* transition of α,β-unsaturated carbonyl compounds is above ≈200 nm with typical intensities of the order of log ε ≈ 4. Its position can be estimated with the extended Woodward rule. For cross-conjugated systems, the value for the chromophore absorbing at the longest wavelength must be calculated.

Extended Woodward rule for estimating the position of the π→π* transition (λ_{max} in nm)



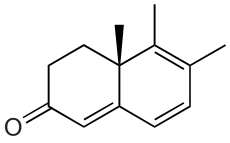
Parent system		X: alkyl X: H X: OH X: O–alkyl	215 207 193 193
			215
			202

Increments		
	for each additional conjugated double bond	+30
	for each exocyclic double bond 	+5
	for each homoannular diene system	+39

For each substituent on double bond system	Increment α	β	γ	δ and beyond
C-substituent	10	12	18	18
Cl	15	12		
Br	25	30		
OH	35	30		50
O–alkyl	35	30	17	31
O–$COCH_3$	6	6	6	6
S–alkyl		85		
N(alkyl)$_2$		95		

Solvent corrections	Solvent	Correction term
	water	-8
	hexane	11
	cyclohexane	11
	chloroform	1
	methanol	0
	ethanol	0
	diethyl ether	7
	dioxane	5

Example: Estimation of the absorption maximum in ethanol for



base value	215
2 additional conjugated double bonds	60
exocyclic double bond	5
homoannular diene system	39
1 β-C-substituent	12
3 additional C-substituents	54
solvent correction	0
estimated	385
experimental	388

9.4 Aromatic Hydrocarbons

9.4.1 Monosubstituted Benzenes

Typical Ranges for Monosubstituted Benzenes (λ_{max} in nm)

Transition	λ_{max}	ε
π → π* (allowed)	180–230	2000–10000
π → π* (forbidden)	250–290	100–2000
π → π* (substituent delocalized by aryl; K band)	220–250	10000–30000
n → π* (substituent with lone pair; R band)	275–350	10–100

Specific Examples of Monosubstituted Benzenes (λ_{max} in nm)

	π → π* (allowed)		π → π* (forbidden)		π → π* (K band)		n → π* (R band)	
Substituent R (solvent)	λ_{max}	ε	λ_{max}	ε	λ_{max}	ε	λ_{max}	ε
–H (cyclohexane)	198	8000	255	230				
$-CH_3$ (hexane)	208	7900	262	230				
$-CH{=}CH_2$ (ethanol)			282	450	244	12000		
–C≡CH (hexane)			278	650	236	12500		
–Cl (ethanol)	210	7500	257	170				
–OH (water)	211	6200	270	1450				
$-O^-$ (water)	235	9400	287	2600				
$-NH_2$ (water)	230	8600	280	1430				
$-NH_3^+$ (water)	203	7500	254	160				
$-NO_2$ (Hexan)	208	9800	270	800	251	9000	322	150
	213	8100						
–C≡N (water)			271	1000	224	13000		
–CHO (hexane)			280	1400	242	14000	≈330	≈60
$-COCH_3$ (ethanol)			278	1100	243	13000	319	50
–COOH (water)	202	8000	270	800	230	10000		

9.4.2 Polysubstituted Benzenes

Estimation of the position of the allowed $\pi \rightarrow \pi^*$ transition in multiply substituted benzenes (λ_{max} in nm, log $\varepsilon \approx 4$)

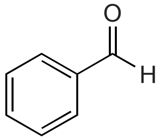
Base value: 203.5

Substituent	Increment [nm]
$-CH_3$	3.0
–Cl	6.0
–Br	6.5
–OH	7.0
$-O^-$	31.5
$-OCH_3$	13.5
$-NH_2$	26.5
$-NHCOCH_3$	38.5
$-NO_2$	65.0
–C≡N	20.5
–CHO	46.0
$-COCH_3$	42.0
–COOH	25.5

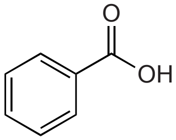
9.4.3 Aromatic Carbonyl Compounds

Scott rules for estimating the position of the K band (solvent: ethanol; λ_{max} in nm, ε = 10000–30000)

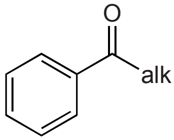
Parent systems



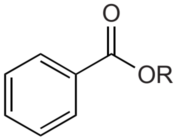
250



230



246

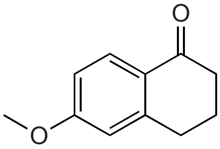


230

Increments

Substituent	*ortho*	*meta*	*para*
–alkyl	3	3	10
–cycloalkyl	3	3	10
–Cl	0	0	
–Br	2	2	15
–OH	7	7	25
–O–alkyl	7	7	25
$–O^{-}$	11	20	78
$–NH_2$	13	13	58
$–N(CH_3)_2$	20	20	85
$–NHCOCH_3$	20	20	45

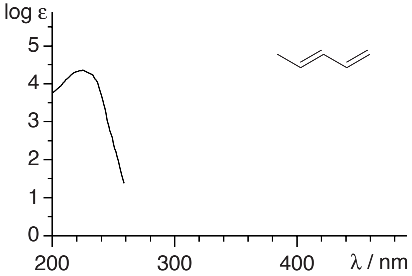
Example: Estimation of the absorption maximum (K band) for

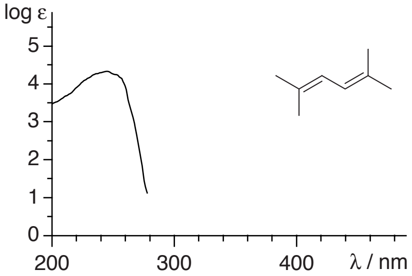


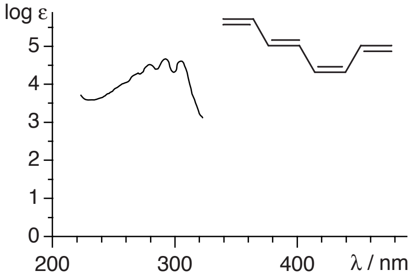
base value	246
ortho-cycloalkyl	3
para-O–alkyl	25
estimated	274
experimental	276

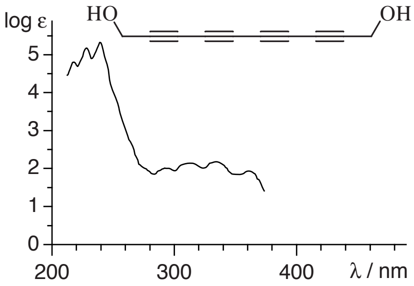
9.5 Reference Spectra

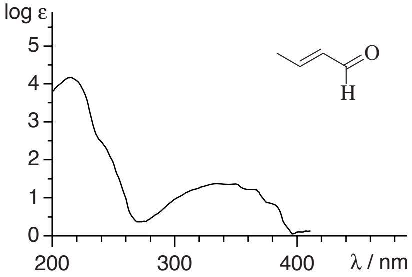
9.5.1 Alkenes and Alkynes

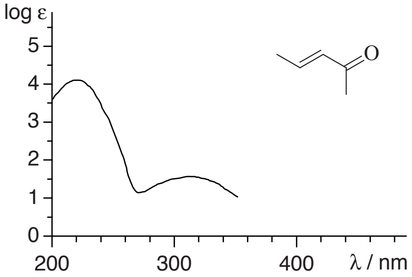


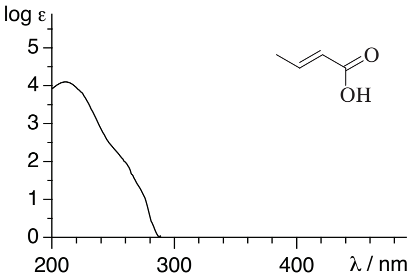


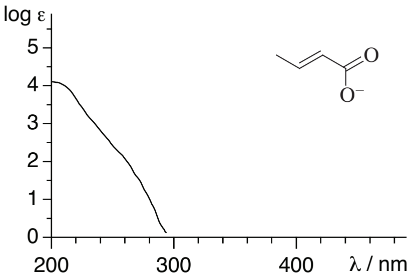




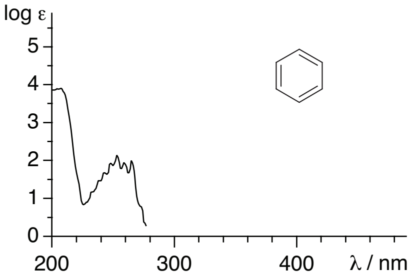


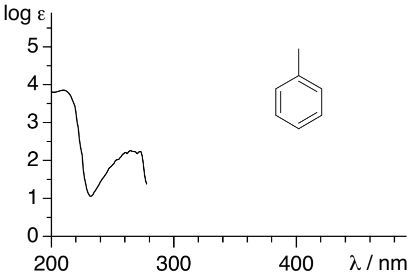


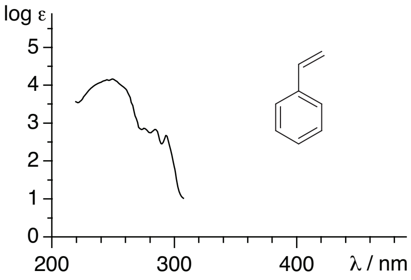


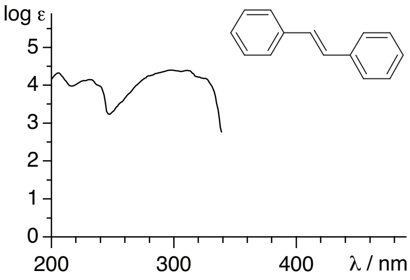


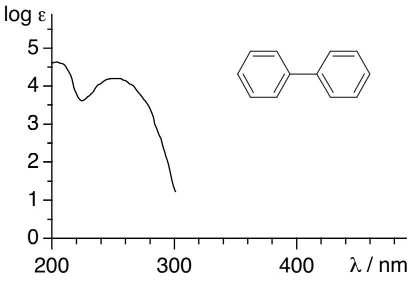
9.5.2 Aromatic Compounds

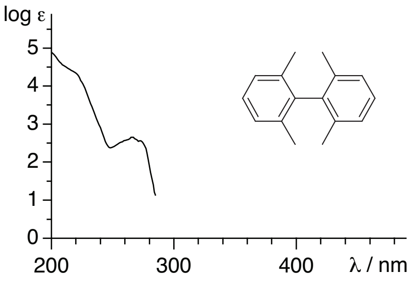


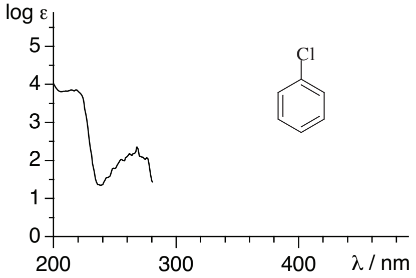








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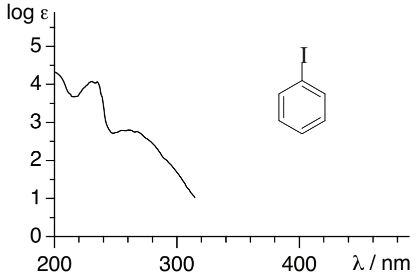
4
3
2
1
0
200
300
400
λ / nm
Cl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λ / nm
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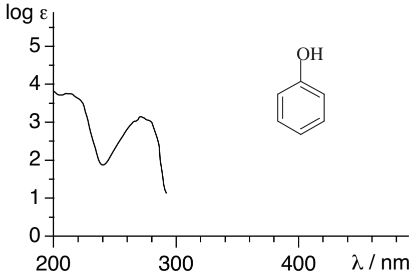
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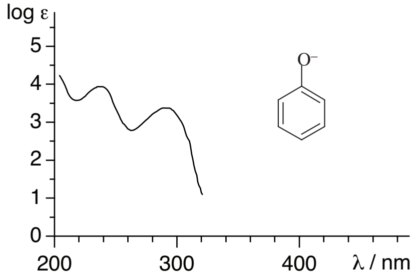
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λ / nm
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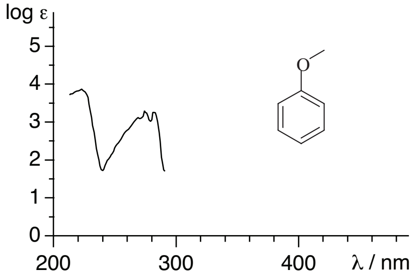
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300
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λ / nm
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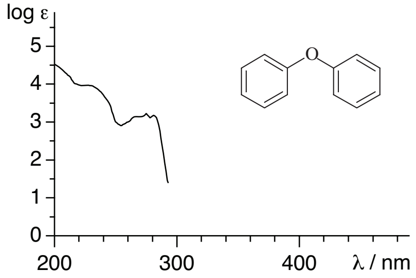
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400
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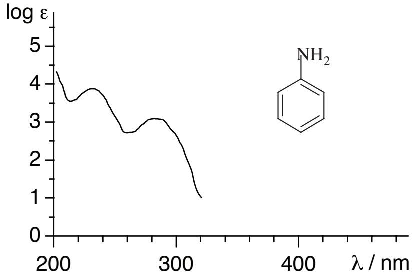
O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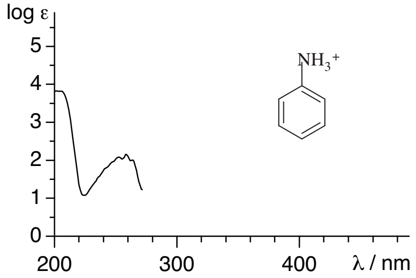
NH_2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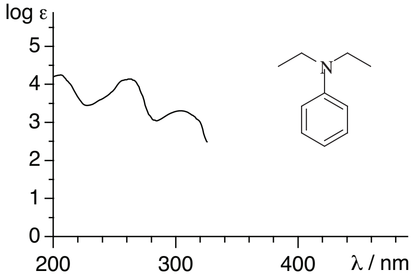
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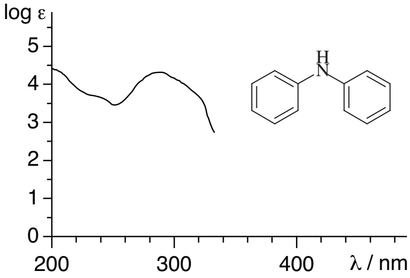
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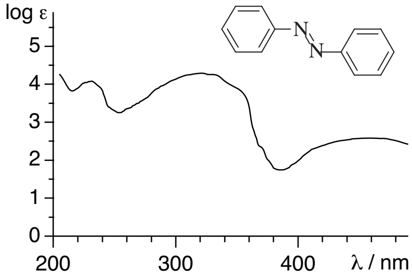
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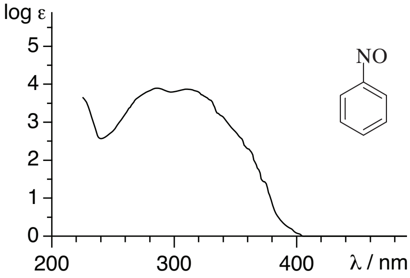
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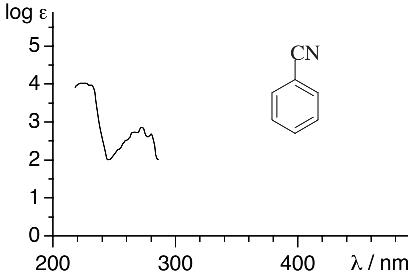
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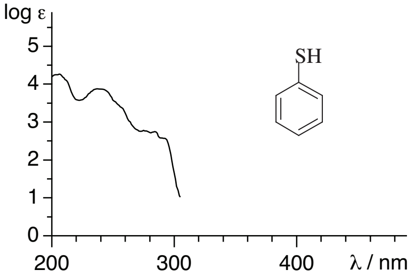
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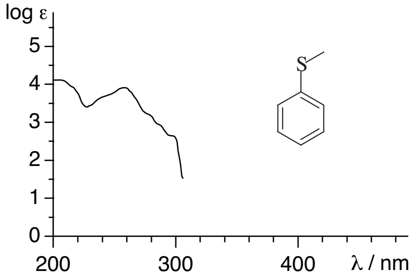
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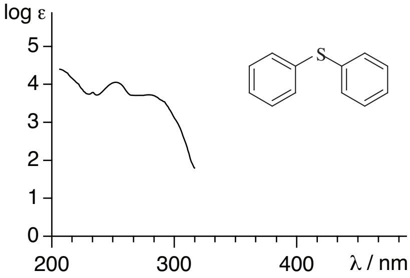
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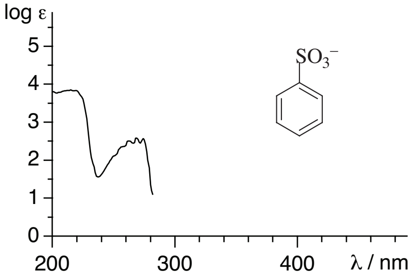
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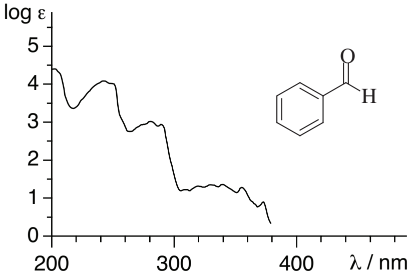
SO3−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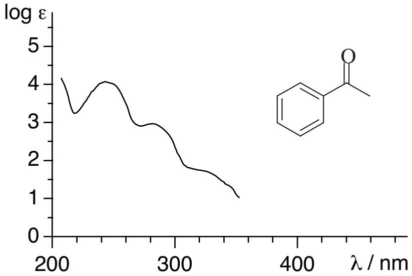
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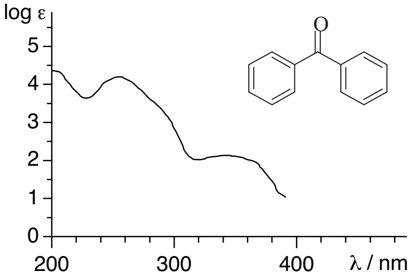
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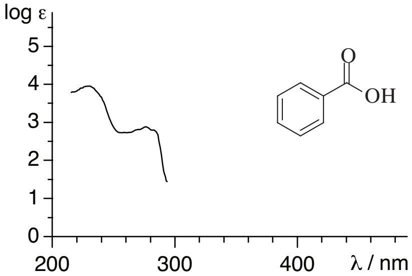
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λ / nm
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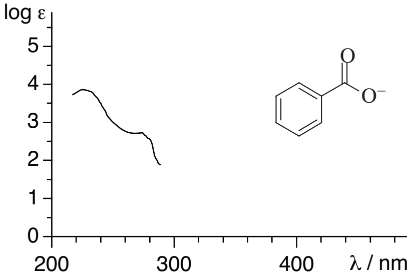
OH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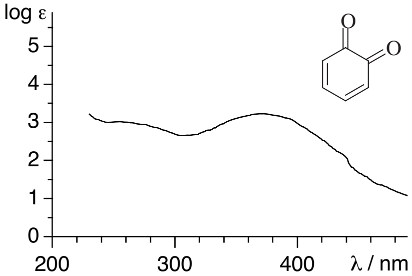
O−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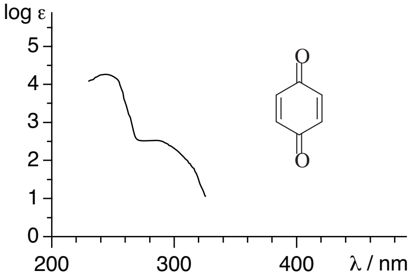
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O
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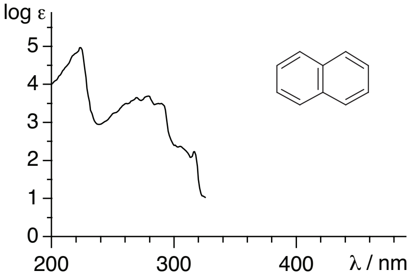
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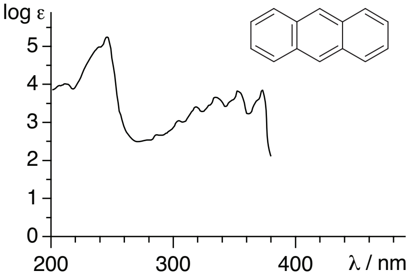
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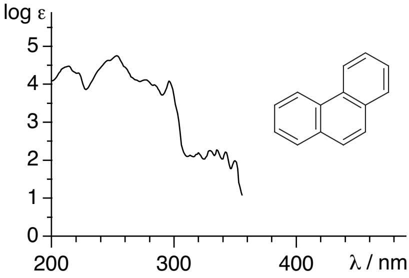
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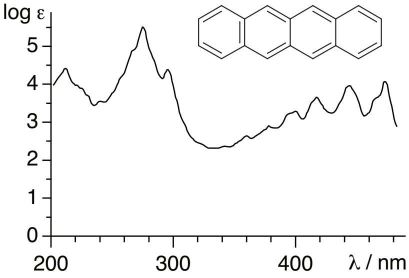
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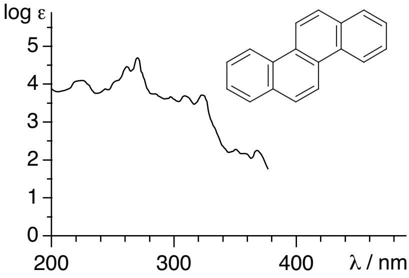
λ / nm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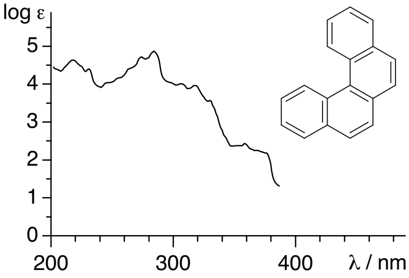
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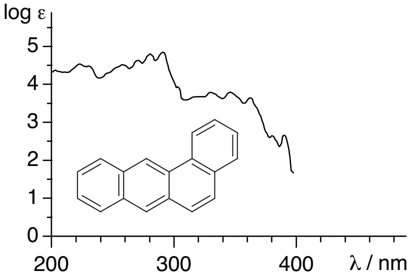
λ / nm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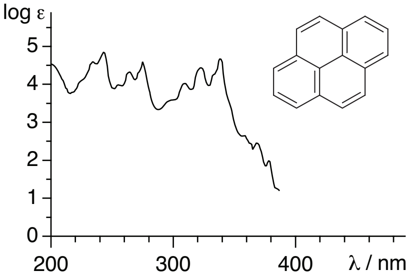
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300
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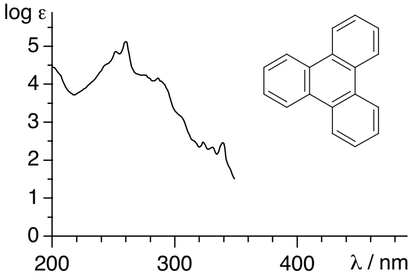
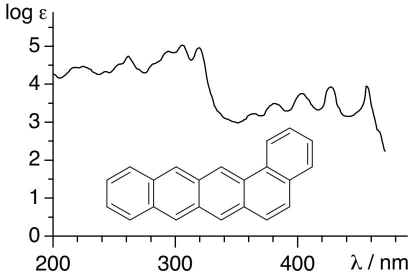
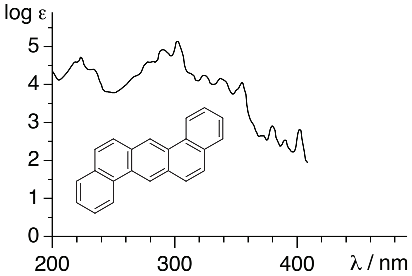
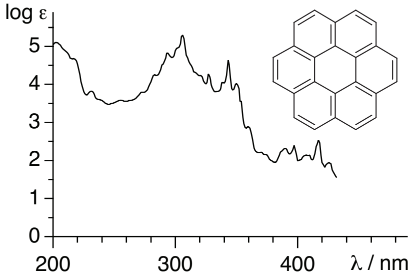
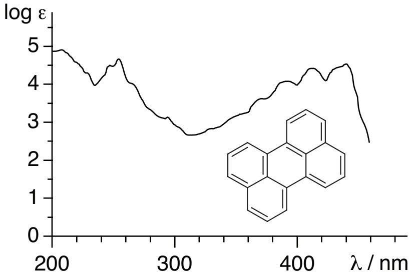
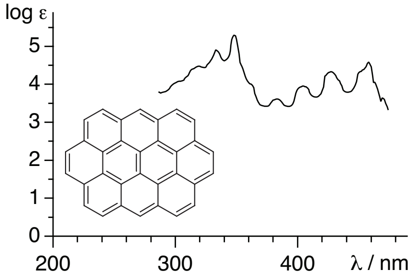




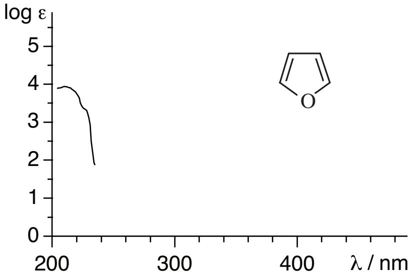


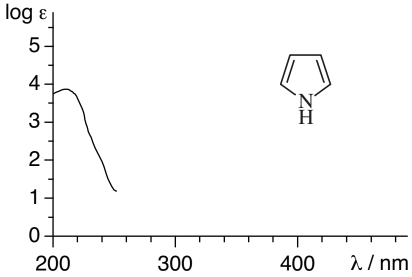


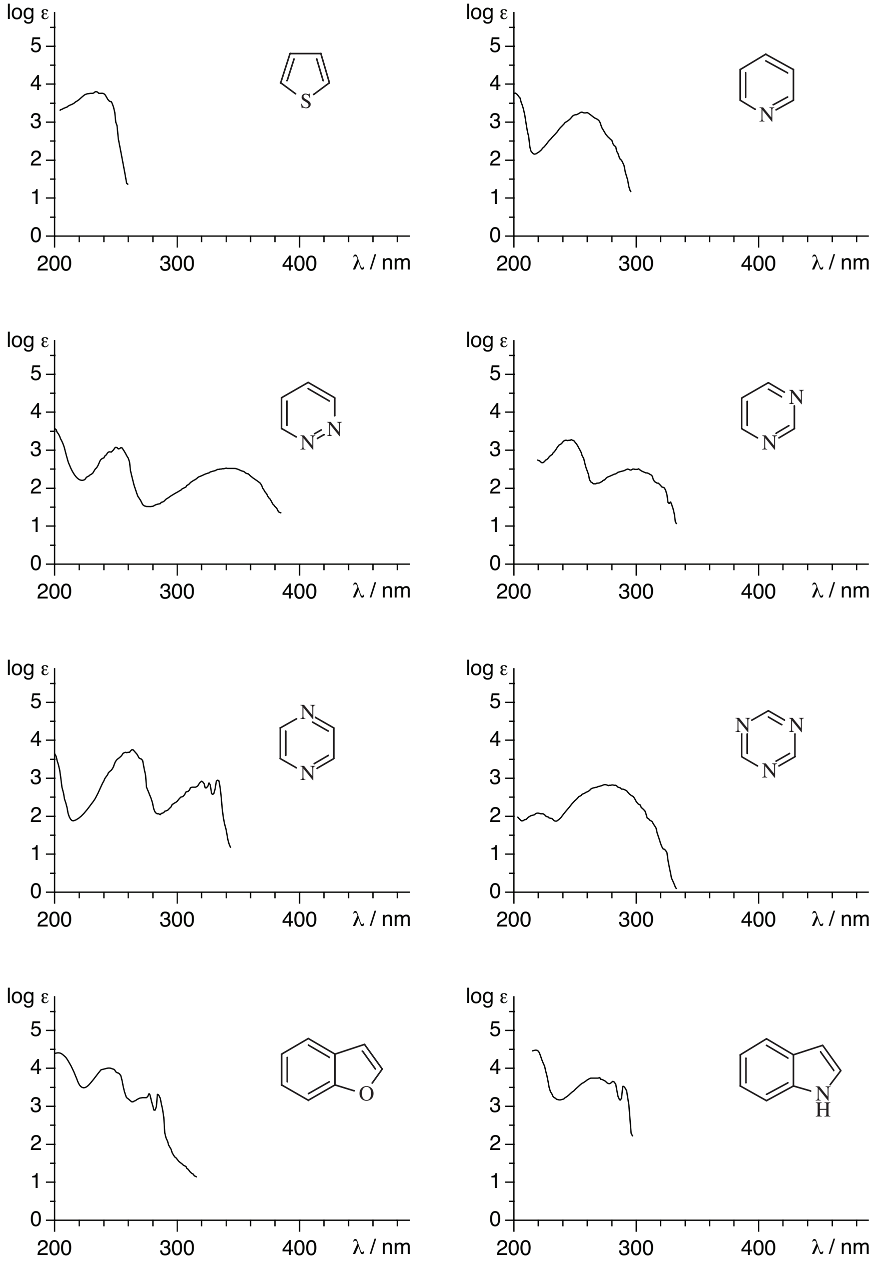




9.5.3 Heteroaromatic Compounds



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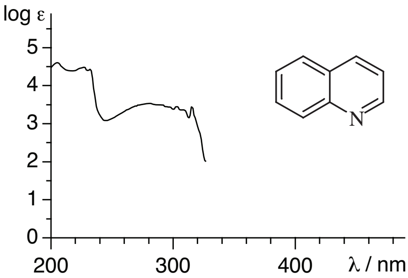
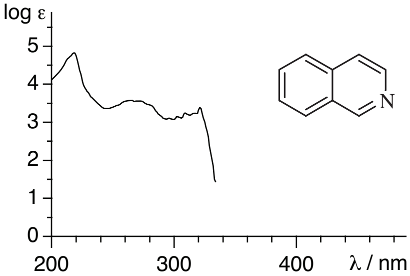
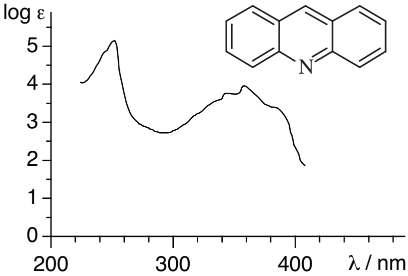
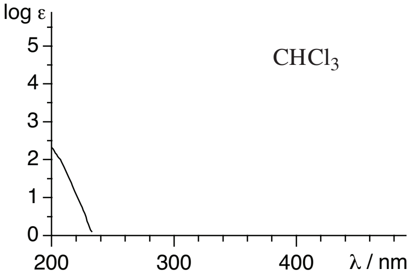
4
3
2
1
0
200
300
400
λ / nm
S
N
N
N
N
N
N
N
N
N
N
O
N
H

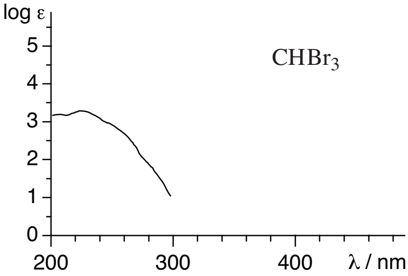


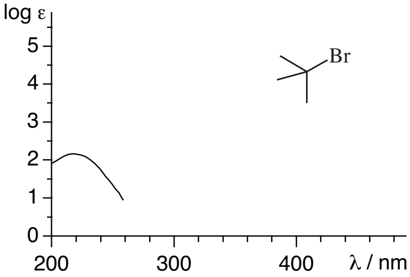


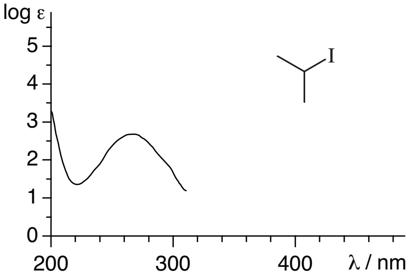


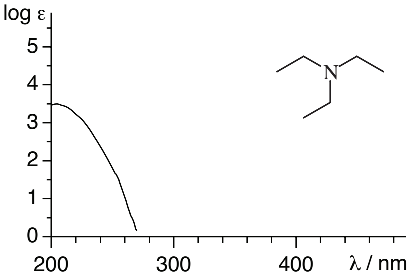
9.5.4 Miscellaneous Compounds







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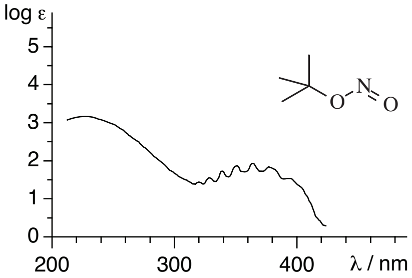
1
2
3
4
5
200
300
400
λ / nm
N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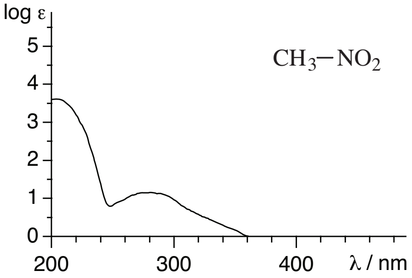
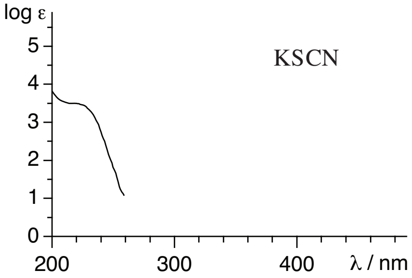
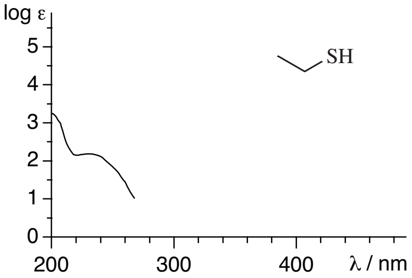
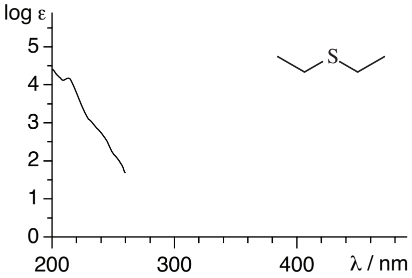
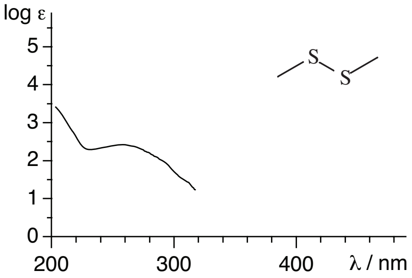
0
1
2
3
4
5
200
300
400
λ / nm
O
N
O

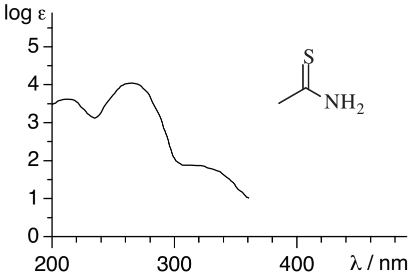

log ε
0
1
2
3
4
5
200
300
400
λ / nm
CH3—NO2

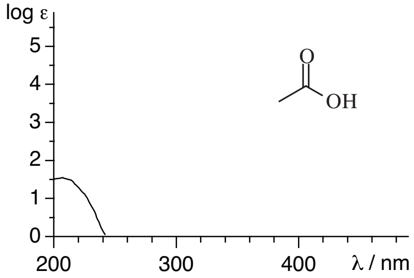

log ε
0
1
2
3
4
5
200
300
400
λ / nm
KSCN

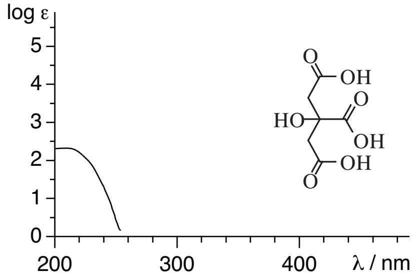

log ε
0
1
2
3
4
5
200
300
400
λ / nm
SH

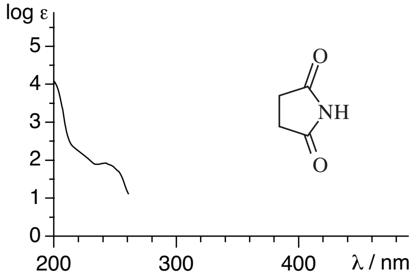

log ε
0
1
2
3
4
5
200
300
400
λ / nm
S

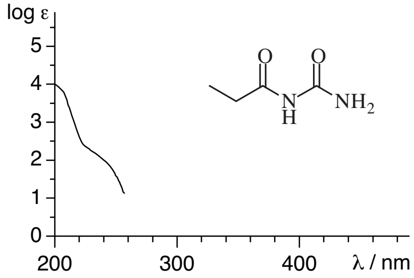

log ε
0
1
2
3
4
5
200
300
400
λ / nm
S
S


log ε
0
1
2
3
4
5
200
300
400
λ / nm
S
NH2

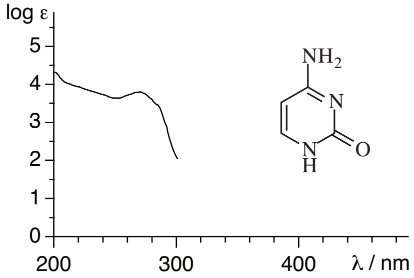


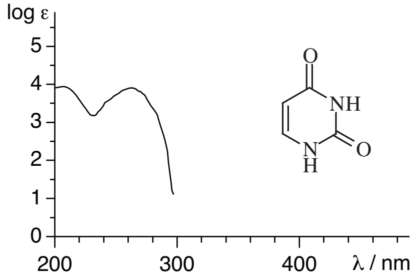


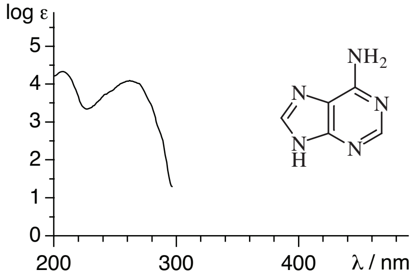


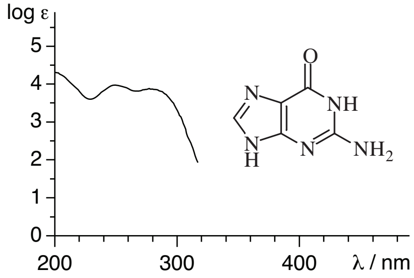


9.5.5 Nucleotides









9.6 Common Solvents

The end absorption, λ_{end}, of several common solvents is given here as the wavelength at which the solvents absorb 80% of the irradiated light (λ_{end} in nm; cell length, 1 cm; reference, water).

Solvent	λ_{end}
acetone	335
acetonitrile	190
benzene	285
carbon disulfide	380
carbon tetrachloride	265
chloroform	245
cyclohexane	210
dichloromethane	230
diethyl ether	210
1,4-dioxane	215
ethanol	205

Solvent	λ_{end}
ethyl acetate	205
heptane	195
hexane	195
methanol	205
pentane	200
2-propanol	205
pyridine	305
tetrahydrofuran	230
toluene	285
2,2,4-trimethylpentane	210
xylene	290